Supplementary material

Anion effects on the liquid-liquid equilibrium behavior of pluronic L64 + water + sodium salts at different pH: Determination of thermodynamic parameters

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Table S1. Coefficients from binodal curves for the salt and polymer rich-phase.

pH	Salt	Coefficient - (Equation 3 and 4)				R^2
		a	b	c	d	
	Sodium sulfate	-521.86	198.5	-23.916	0.9199	0.9768
5.0	Sodium citrate	-348.05	152	-21.812	1.105	0.9921
	Sodium tartrate	-109.68	67.048	-13.416	0.9285	0.9965
	Sodium sulfate	-438.27	161.87	-18.756	0.6906	0.9721
7.5	Sodium citrate	-295.71	131.98	-19.184	0.9426	0.9939
	Sodium tartrate	-246	114.01	-17.25	0.8922	0.9972
	Sodium sulfate[a]	n.a	n.a	n.a	n.a	n.a
10.0	Sodium citrate[b]	n.a	n.a	n.a	n.a	n.a
	Sodium tartrate	-176.17	92.821	-15.197	0.7833	0.9811

[a] The adjusted equation was $y = ex^f$, with e = 0.0064, f = -1.026, and R^2 = 0.9933.
[b] The adjusted equation was $y = ex^f$, with e = 0.0041, f = -1.27, and R^2 = 0.9732.

Table S2. Thickness of immobilized water spherical shell surrounding the ion and effective radius of anions.

Anion	V (cm³ mol⁻¹)	r (nm)	Δr (nm)	(r + Δr) (nm)
$C_6H_5O_7^{3-}$	72.0	0.3056	0.0301	0.3357
SO_4^{2-}	36.5	0.2437	0.0375	0.2812
$C_4H_4O_6^{2-}$	60.1	0.2878	0.0243	0.3121