

## Article

# Investigation of Cubic and Spherical IONPs' Rheological Characteristics and Aggregation Patterns from the Perspective of Magnetic Targeting

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**Abstract:** Based on our prior research, we synthesized iron-oxide nanoparticles (IONPs) in two shapes (spherical and cubic) and sized them for the current inquiry. This research examined the magnetic, rheological, and sedimentation properties of the suspensions containing PEG-coated IONPs, considering that both produced particles are intended to be employed for magnetic targeting applications. The saturation magnetization of both IONPs was lower than the magnetite domain magnetization of 92 emu/g due to the surfactant and the dead surface magnetic layer. Under each investigated magnetic field (0, 34 and 183 mT), the shear viscosity behaviour of the MNP suspensions of both kinds was comparable. Shear thinning behaviour was seen for both particle suspensions in the low shear rate area ( $0.1 \text{ s}^{-1}$  to  $1 \text{ s}^{-1}$ ). The rheological curves from this paper show that the suspensions present a higher viscosity at lower shear rates for spherical and cubic PEG-coated nanoparticles when a magnetic field is applied. The aggregation behaviour demonstrates that cubic-shaped IONPs are more stable throughout time, with hydrodynamic diameter measurements showing a relatively slow variation of the DLS size distribution from 250 nm to 210 nm in the first 600 s; contrarily, the hydrodynamic diameter of spherical IONPs fluctuated significantly, from 855 nm to 460 nm. Another key finding relates to the sedimentation profile, specifically that PEG-coated IONPs with spherical shapes have a stronger tendency to sediment than those with cubic forms, which are more stable.

**Keywords:** magnetic nanoparticles; particle synthesis; particle targeting; nanoparticle morphology; magnetic characteristics; biomedical applications



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## 1. Introduction

To improve size, shape, magnetism, stability, toxicity, biocompatibility, and biodistribution, various magnetic nanoparticles (MNPs) have been synthesized using multiple synthetic routes and coated with different materials [1–3]. However, several challenges still need to be resolved. For example, the size of the nanoparticle must be large enough that it has a strong magnetism for effective magnetic targeting and can transport drugs deep into the body. Critical considerations for in vivo medical applications include [4]: (i) Particles smaller than 10 nm are quickly eliminated by renal clearance; (ii) Particles larger than 5  $\mu\text{m}$  cause capillary blockade; (iii) Particles more prominent than 200 nm are excreted by the liver and spleen; and (iv) The upper size limit for the appearance of superparamagnetic properties of MNPs is 120 nm. Therefore, MNPs should be synthesized with 10–100 nm diameters for magnetically controlled drug release systems [4].

MNPs must be manufactured from a non-toxic substance with small particle sizes to stay in circulation after injection and pass through the capillary systems of organs and

tissues, preventing vessel embolism, in order to be used for drug targeting applications. Additionally, they must be highly magnetized to be fixed near the targeted diseased tissue and to allow magnetic field control of their mobility in the blood [5,6].

The drug-targeting process using a magnetic carrier is based on the competition between forces exerted on the particles by the blood and magnetic forces generated from the external magnet.

By far the most prevalent type of iron oxide nanoparticles (IONPs) utilized in medicine is magnetite ( $\text{Fe}_3\text{O}_4$ ). When this kind of particle is being used for biomedical applications, it is vital that their size is smaller than 20 nm (especially if this utilization of the particles is for diagnostic magnetic resonance imaging) to quickly enter and migrate throughout the human body [7].

The magnetic particles can carry medications to a particular location of interest when the magnetic forces are more significant than the drag force generated by the linear blood flow in arteries (flow velocity 10 cm/s) or capillaries (flow velocity 0.05 cm/s) [6,8,9]. Therefore, the loading of IONPs, the strength of the external magnetic field, and the size and magnetism of the nanoparticles are all directly correlated with the targeting performance of magnetic drug delivery systems [10].

In general, the saturation magnetization decreases as the size of IONPs increases [11]. Different saturation magnetizations are produced depending on the  $\text{Fe}^{2+}/\text{Fe}^{3+}$  ratios in the tetrahedral or octahedral cationic sublattices affected by the synthesis process [11]. The two basic chemical processes for making MNPs are coprecipitation in an aqueous solution employing polymers or charged molecules as surfactants, and thermal decomposition of the iron organometallic compounds in a higher boiling point organic solvent [11–14].

Based on our prior research [8], we have synthesized IONP particles in two shapes (spherical vs. cubic) and sized them for the current inquiry. This research examined the magnetic, rheological, and sedimentation properties of the suspensions containing PEG-coated IONPs. In addition, investigations were carried out considering that both produced particles are intended to be employed for magnetic targeting applications.

This article differs in the following ways from results that have already been published:

- Compared to the PEG-coated nanocomposite MNCs used in earlier work [8,9], the PEG-coated IONPs particles generated using the thermal decomposition synthesis process were used as carriers;
- This study investigated the PEG-coated spherical and cubic-shaped IONPs against the spherical nanocomposite previously employed;
- The employed magnetic nanoparticle size ranged from 10 to 15 nm (ideal for particle targeting), as opposed to the 40 to 150 nm PEG-coated magneto-responsive nanocomposite in our earlier work.

## 2. Materials and Methods

### 2.1. Chemicals

For the preparation of magnetic iron-oxide nanoparticles (IONPs'), the following chemicals have been used: ferric oleate, 1-octadecene, oleic acid (for the spherical nanoparticles), sodium oleate (for the cubic nanoparticles), hexane, acetone, isopropanol, and toluene. The ferric oleate precursor was prepared from ferric chloride hexahydrate.

For the phase transfer of the MNP, the synthesis required using  $\alpha,\omega$ -Bis(2-[(3-carboxy-1-oxopropyl)amino]ethyl)polyethylene glycol (PEG diacid, MW: 3000 g/mol), N-hydroxysuccinimide (NHS), dicyclohexylcarbodiimide (DCC), dopamine hydrochloride, chloroform ( $\text{CHCl}_3$ ), N, N-dimethylformamide (DMF), anhydrous sodium bicarbonate ( $\text{Na}_2\text{CO}_3$ ), and hexane.

All the chemicals presented were obtained from Sigma-Aldrich in analytical or pure-grade quality.

## 2.2. Syntheses

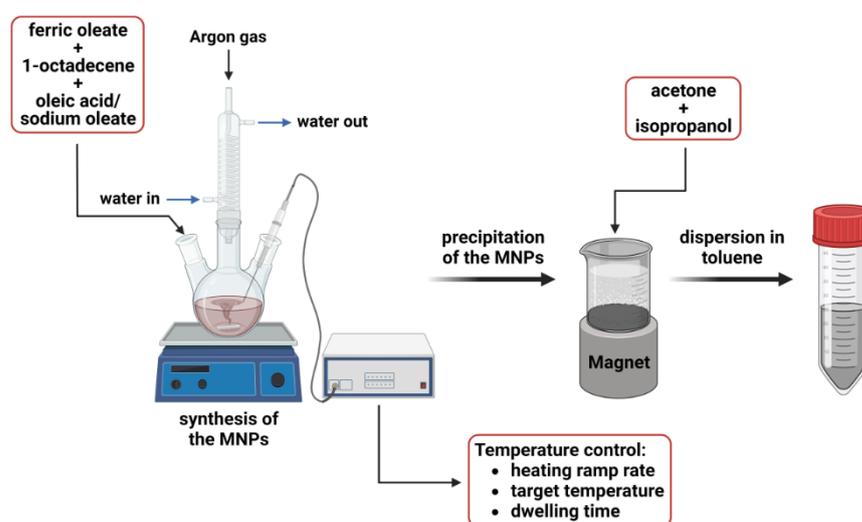
### 2.2.1. Ferric Oleate Synthesis

A mixture of ferric chloride hexahydrate (5.395 g, 0.02 mol), sodium oleate (18.25 g, 0.06 mol), MQ Water (water obtained using Milli-Q Direct Water Purification System) (25 mL, 1.38 mol), hexane (70 mL, 0.56 mol), and ethanol (40 mL, 0.69 mol) was refluxed in an oil bath with magnetic stirring for 4 h at 70 °C. After that, the solution was left to cool down before washing it three times with MQ water using a separatory funnel. Finally, the excess hexane and ethanol were removed using a rotavapor until a waxy solution was obtained. The ferric oleate and the magnetic nanoparticle synthesis were based on the protocol published in the previous work [15].

### 2.2.2. Magnetic Nanoparticles Synthesis

Hydrophobic monodispersed magnetic nanoparticles were synthesized through thermal decomposition. The synthesis of the MNPs relies on reducing the iron oleate at high temperatures. Iron oleate (1.6 g, 1.78 mmol), oleic acid (600 µL, 1.78 mmol) and 1-octadecene (25 mL, 1.9 mmol) were added to a two-necked round bottom flask (RBF), which was then connected to a reflux-condenser and a thermocouple. The mixture was magnetically stirred under Argon atmosphere while the temperature was gradually increased to 320 °C. The inert (Argon) atmosphere from the synthesis process was used to avoid possible oxygen contamination and to reduce the particles' size [16]. The heating ramp rate of the process was 3 °C/min. After the first hour of the reaction, the mixture changed colour from dark orange to black, indicating the nucleation of crystallites. The reaction mixture was kept at the target temperature of 320 °C for 45 min, which had the purpose of growing the nanoparticles. Next, the RBF was cooled down to room temperature. Next, 25 mL of hexane was added to the mixture, transferred to a beaker, and placed on a magnet. Hexane is used to prevent flocculation and facilitate MNPs' dispersion.

The magnetic nanoparticles were precipitated by adding a mixture of a 1:1 molar ratio of acetone and isopropanol. When all the particles were wholly precipitated, the supernatant was discarded entirely. Since magnetite oxidizes to maghemite while exposed to air [17], consequently, the obtained nanoparticles were washed thrice with acetone and then redispersed in 30 mL toluene. Figure 1 shows a generalized perspective of the synthesis strategy.



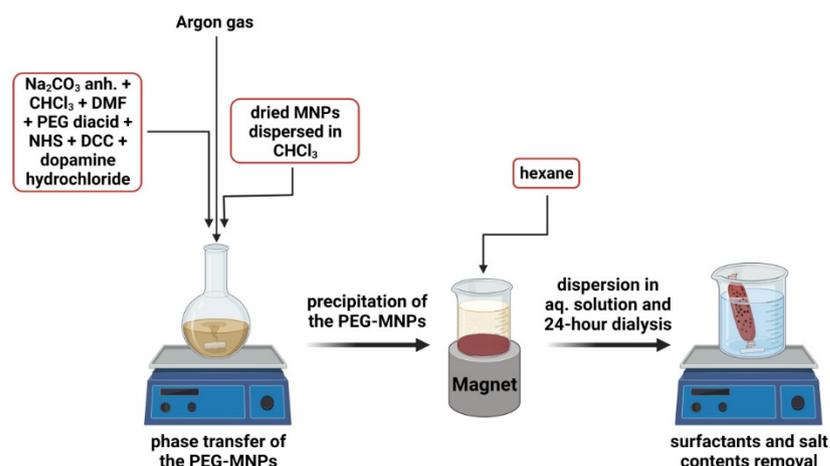
**Figure 1.** Schematic illustration of the magnetic nanoparticles synthesis strategies using the thermal decomposition method. Image "Created with BioRender".

The presented method was used to synthesize spherical iron oxide nanoparticles. For the cubic nanoparticles, the same procedure was applied but with the following adjustments:

- The precursors and the corresponding quantities were iron oleate (0.833 g, 0.96 mmol), sodium oleate (0.213 g, 0.7 mmol), and 1-octadecene (14 mL, 1.06 mmol);
- The target temperature was 325 °C;
- The heating ramp rate of the process was 2.8 °C/min.

### 2.2.3. Phase Transfer of the Magnetic Nanoparticles

The protocol for the phase transfer process corresponds to the work performed by [18]. In a 25 mL RBF, anhydrous Na<sub>2</sub>CO<sub>3</sub> (10 mg, 0.094 mmol) was dissolved in CHCl<sub>3</sub> (2 mL, 0.025 mmol) and DMF (1 mL, 0.013 mmol), after which PEG diacid (20 mg, 0.007 mmol), NHS (2 mg, 0.017 mmol), DCC (3 mg, 0.015 mmol), and dopamine hydrochloride (1.27 mg, 0.0083 mmol) were added to the mixture and magnetically stirred for two hours at room temperature. After that, 5 mg of dried MNPs dispersed in 1 mL CHCl<sub>3</sub> were added slowly. The solution was left overnight while stirring under an inert atmosphere. The solution was transferred to a beaker, where the PEG-MNPs were precipitated with hexane and magnetically separated. The functionalized nanoparticles were left to dry for at least 24 h and then redispersed in a 5 mL aqueous phase (MQ Water). Dispersed nanoparticles were left for 24 h dialysis to certify the removal of surfactants or salt contents. While in the dialysis bag, the water was changed every two hours thrice. Figure 2 shows a generalized perspective of the phase transfer process.



**Figure 2.** Schematic illustration describing the principle of the MNPs polymer coating process. Image “Created with BioRender”.

### 2.3. Characterization

The shape and core diameter of the magnetic nanoparticles were examined using a bright field scanning transmission electron microscope (BFSTEM; Hitachi S-5500, Hitachi High-Tech Corporation, Tokyo, Japan). The samples were prepared by drying a small amount of the diluted MNP solution on carbon-coated copper grids. The particles' hydrodynamic size and zeta potential were measured using an Anton Paar Litesizer (Anton Paar GmbH, Graz, Austria) and a Malvern Zetasizer Nano-ZS device (Malvern Panalytical Ltd., Malvern, UK). Before the measurements, the samples were diluted to a 0.1 mg/mL concentration. For the PEG-MNP, disposable cuvettes (for size measurements) and Omega cuvettes (for zeta potential measurements) were employed, whereas glass cuvettes were used for the bare nanoparticles. An Alternating Gradient Magnetometer (AGM), Princeton PMC Model 2900 (Lake Shore Cryotronics, Inc., Westerville, OH 43082, USA), and a vibrating sample magnetometer (VSM 880-ADE Technologies, Westwood, MA, USA) were used to measure the MNPs' saturation magnetization (*M<sub>s</sub>*) curves. The instrument's

sensitivity was adjusted to 500 memu, with a  $10^4$  Oe maximum magnetic field applied. The average duration was 300 ms, and the applied field increment was 100 Oe. A rotating rheometer (Anton Paar MCR 300 Physica, Anton Paar GmbH, Graz, Austria) with a 20 mm diameter plate–plate magnetorheological cell (MRD 170/1T-SN80730989) was used to test the magneto-viscous properties of the pegylated nanoparticles at 25 °C in both the presence and absence of a magnetic field.

### 3. Results

#### 3.1. MNPs' Synthesis and Morphology

The final shape of the magnetic nanoparticles is influenced by the solution content, temperature, pressure, and synthesis technique [19].

Magnetic nanoparticles synthesized using the thermal decomposition method can have a consistent particle dispersion [20]. This technique's key benefit is that it produces monodisperse, small-sized, and highly crystalline magnetic nanoparticles [21].

Iron (III)-oleate was thermally decomposed to produce the nanoparticles, using oleic acid as a stabilizing ligand.

The purpose of precipitating the nanoparticles with acetone, followed by rigorous stirring, was to remove the leftover oleate precursor and surplus oleic acid. After discharging the supernatant, the nanoparticles were redispersed in toluene and stored at 4 °C. Table 1 provides an overview of the synthesis procedures for cubic- and spherical-shaped nanoparticles.

**Table 1.** The synthesis protocol for spherical- and cubic-shaped MNPs.

Particle Shape	Particle Core Diameter/ Edge Length [nm]	Heating Rate [°C/min]	Target Temperature [°C]	Dwelling Time [min]
Spherical	8–11	3.0	320	45
Cubic	14–16	2.8	325	45

##### 3.1.1. Oleic Acid

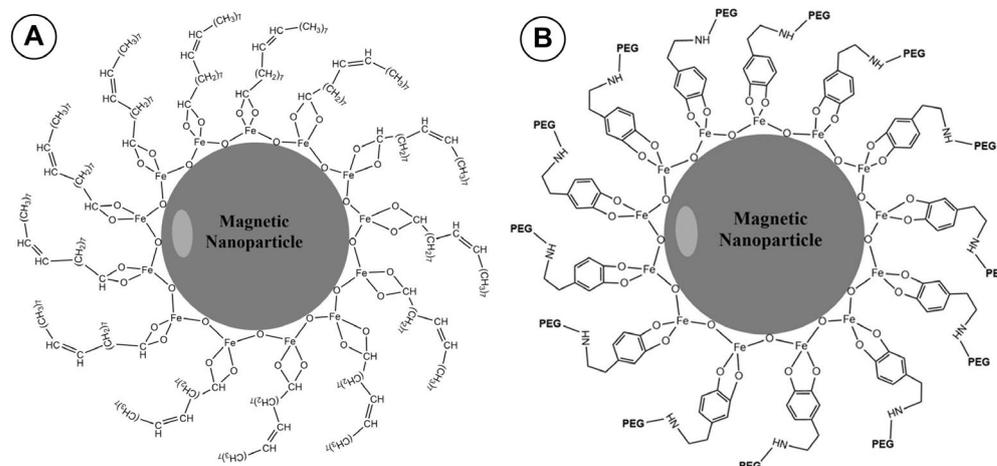
The MNPs' coating affects stability, upcoming functionalization, or hydrophobic or hydrophilic behaviour [22].

Oleic acid is widely used as a capping agent in ferrite nanoparticle synthesis. It can form a dense protective monolayer film, strongly bonded to the surface of the particle. This capping agent leads to the formation of highly uniform and monodispersed particles [23,24]. However, coating MNPs with OA makes them dispersible only in organic solvents, consequently limiting their use for biomedical applications [25]. Therefore, this hydrophobic coating must be replaced with a hydrophilic coating for biomedical applications in aqueous environments.

Oleic acid was employed as the coating surfactant in our work because it can stop particle oxidation and aggregation. The MNPs aggregate and become unstable in solution if not coated with surfactants [26].

Paper [26] shows that the strong bond between oleic acid and the surface of the nanoparticles is caused by the terminal carboxylic group, with a negative charge from the acid's chain interacting with the positive charge from the surface (Figure 3A). Additionally, the cis-double bond from the oleic acid chain generates a kink required for the MNP to stabilize well [27]. In the phase transfer process mentioned in Section 2.2.2, when bonding the PEG to the surface of the magnetic nanoparticle, the following interactions took place [28]: first, the dopamine was linked to one of the carboxylic groups of the PEG diacid through the NHS reagent, which is used to trigger the reaction. Second, the dopamine-linked PEG bound covalently to the surface of the nanoparticle due to dopamine's high affinity to iron oxide. In the process, the oleate was displaced from nanoparticle's surface in

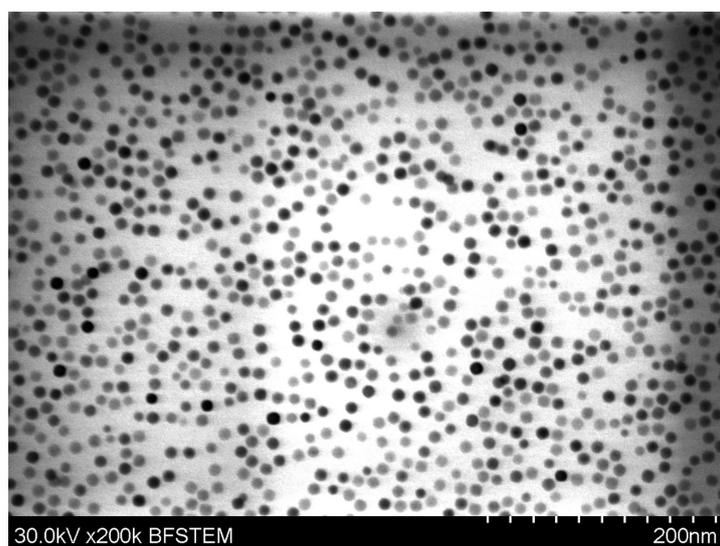
the organic solution of  $\text{CHCl}_3/\text{DMF}$ , which enabled the phase transfer of the PEG-covered MNP to aqueous solution (Figure 3B).



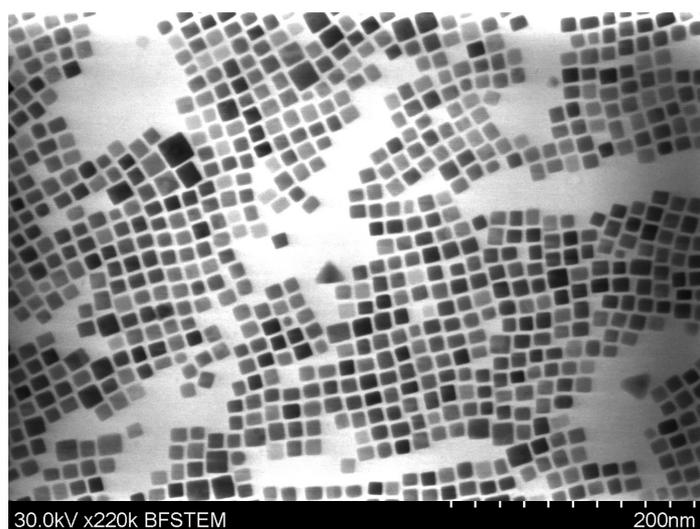
**Figure 3.** Schematic illustrations of (A) the structure of the oleic acid-coated MNP and (B) the PEG-covered MNP after the displacement of the oleate in the phase transfer process.

### 3.1.2. MNPs' Morphology

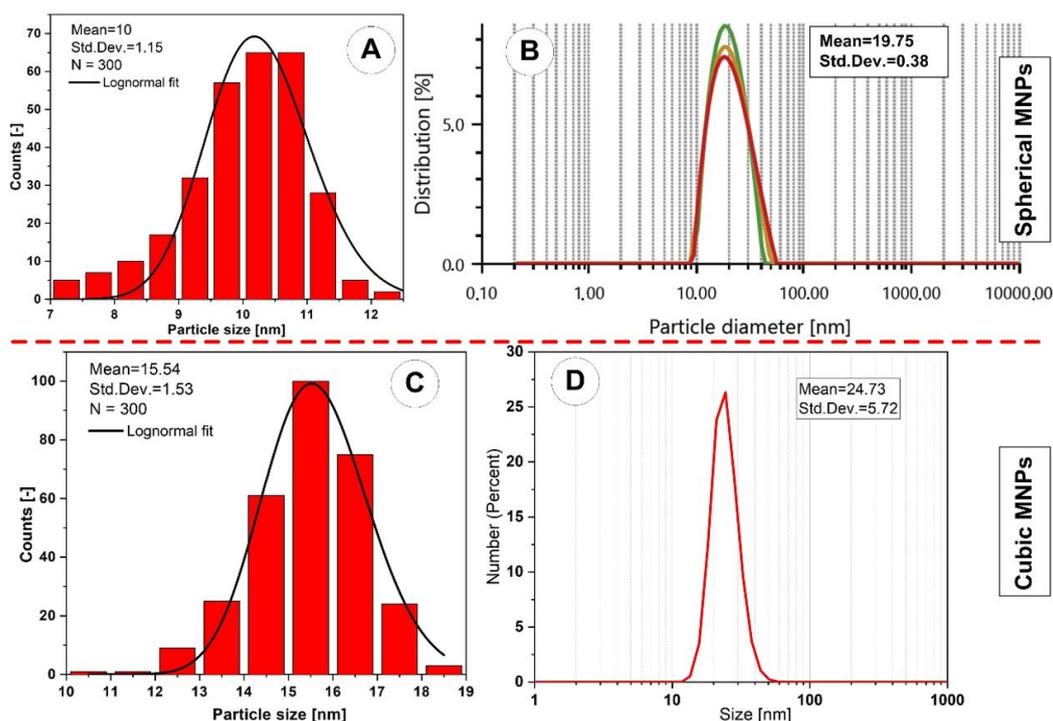
Bright-field scanning transmission electron microscopy (BFSTEM; Hitachi S-5500) was used to conduct a primary analysis of the size and form of both spherical (Figure 4) and cubic (Figure 5) individual core particles. The diameters of 300 cores were extracted from the TEM images using the program ImageJ [<https://imagej.nih.gov/ij/>] (accessed on 10 February 2023)] to produce a frequency distribution of the MNPs' core diameter. The mean core sizes of the spherical and cubic shapes were, respectively, 10.2 nm (Figure 6A) and 15.6 nm (Figure 6C), according to the fit of the log-normal [29] distribution of the core diameters.



**Figure 4.** The morphologies of the magnetic nanoparticles for the synthesized hydrophobic systems. Representative TEM image of the iron oxide nanospheres. The samples have a regular spherical shape distribution. Figures show a round uniform structure of magnetic nanoparticles, well dispersed, with an average diameter of 10 nm.



**Figure 5.** Representative TEM image of the core-shell iron oxide hydrophobic magnetic nanoparticles with the cubic structure of edge length of  $15 \pm 2$  nm. TEM image of the cubic shape MNPs. The nanocubes have a very regular cubic shape without any edge truncations.



**Figure 6.** IONPs' core size distribution by TEM images for spherical (A,C) cubic MNPs. The mean hydrodynamic diameter was measured using the Malvern Instrument Zetasizer Nano series for (B) spherical MNPs, and (D) cubic MNPs. Subfigure (B) shows three measurements for the same spherical MNPs sample.

Positive skewness can be seen in the diameter distribution for both spherical and cubic IONPs, with a mean diameter of 10 nm for spherical IONPs (Figure 6A) and a mean edge length of 15.5 nm for cubic IONPs (Figure 6C).

### 3.1.3. MNPs' Colloidal Stability

DLS measurements were used to confirm the different MNPs' aqueous colloidal stability. In addition, a Zetasizer Nano series was used to assess the size distribution and

surface charge of the nanoparticles at room temperature of 25 °C using dynamic light scattering (DLS) and zeta potential (Malvern Instruments, Malvern, UK).

The effectiveness of hydrodynamic diameter measurements was utilized to test the colloidal stability of IONPs [30]. Aggregation and larger effective hydrodynamic diameters indicate a decrease in colloidal stability. The zeta descriptor determines NP's colloidal stability, quantifying the electrostatic potential between the NP's shear plane and the solvent. Charge differences over ( $>\pm 10$  mV) cause higher interparticle repulsion.

A Malvern Instrument Zetasizer Nano series with a He-Ne laser ( $\lambda = 633$  nm, max 5 mW) and a scattering angle of 173° was used for all DLS experiments. A quartz cuvette measuring the dimensions of 10 mm by 10 mm was filled with 1 mL of particle suspension for each measurement.

It was discovered that the hydrodynamic diameter of the original IONPs in toluene before polymer coating was around  $20 \pm 0.4$  nm for spherical MNPs (Figure 6B) and  $25 \pm 5.7$  nm for cubic MNPs (Figure 6D), with polydispersity index values less than 0.14 for spherical MNPs and, respectively, 0.11 for cubic MNPs (Table 2).

**Table 2.** MNP size is determined by magnetization, TEM and DLS measurements.

Sample	$D_m$ [nm]	$D_{TEM}$ [nm]	$D_H$ [nm]	$D_H - D_{TEM}$ [nm]	The Polydispersity Index (PDI)	Zeta Potential (mV)
Spherical MNPs	$7.3 \pm 1.6$	$10 \pm 1.2$	$19.8 \pm 0.4$	$9.8 \pm 0.8$	0.14	−14
Cubic MNPs	$9.0 \pm 1.6$	$15.6 \pm 1.6$	$24.7 \pm 5.7$	$9.2 \pm 4.2$	0.11	−23.9

Where  $D_m$  is the magnetic diameter (see Section 3.3),  $D_{TEM}$  [nm] is the particle diameter measured using TEM images, and  $D_H$  [nm] is the particle hydrodynamic diameter measured by DLS (z-average).

Moreover, a Zetasizer Nano Series device (Malvern Instruments Ltd., Malvern, UK) was used to measure the coated IONPs' zeta potential (ZP). ZP values typically fall within the  $-100$  to  $+100$  mV range. According to [31], colloidal stability can be predicted by the magnitude of ZP. Lower dispersions ZP values (where attraction defeats repulsion) will cause aggregation, coagulation, or flocculation due to van der Waals interparticle attraction [31]. In contrast, the ZP of MNPs with values of  $-25$  mV or  $+25$  mV usually have a high degree of stability (which prevents the particles from aggregating due to electric repulsion) [31].

Table 2 shows that the cubic MNPs have a ZP of  $-23.9$  mV rather than the  $-14$  mV found in spherical shaped MNPs, indicating that the cubic MNPs are more stable. The samples were homogenized for 10 min in an ultrasonic bath before each measurement, followed by 2 min of repose.

### 3.2. Synthesis of the PEG-coated IONPs

If a biomedical use is envisioned, addressing the colloidal stability of the MNPs under physiological settings is crucial. These nanoparticle aggregates would represent an immediate hazard to health if the prepared nanoparticles were administered topically to a living being. Additionally, unprotected inorganic nanoparticles are vulnerable to opsonization, an innate immune system defensive mechanism in which specific serum proteins bind to the particle surface and trigger phagocytosis. As a result, it is essential to swap out the hydrophobic protective shell for hydrophilic ligands, which provide excellent colloidal stability in an aqueous environment and shield serum proteins from opsonizing the material under physiological circumstances.

It has been demonstrated that surface alterations or polymer coatings enhance the stability, dispersity, and biocompatibility of MNPs [32].

To stop NP aggregation or agglomeration, polyethylene glycol (PEGylation) is frequently used [33]. Although it often slows down the process, the PEGylation of NPs does not ensure the eradication of NP aggregation or agglomeration. In addition, even PEGylated NPs

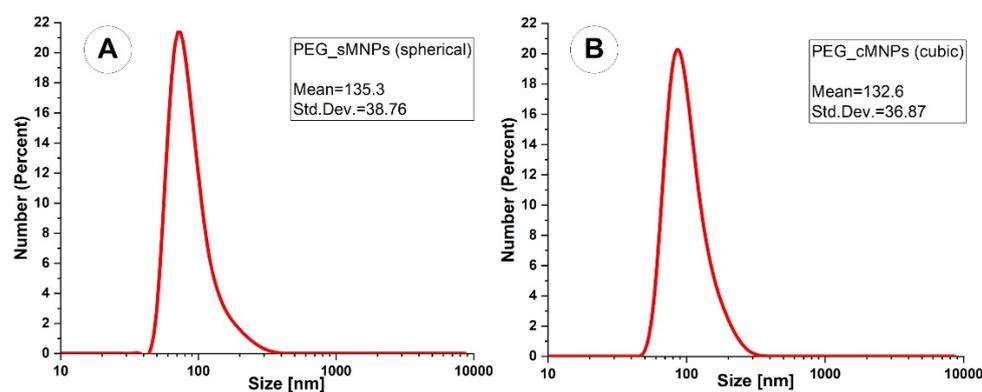
have the potential to self-associate over time, either due to the PEG layer degrading or the introduction of instability factors into the colloidal system [34]. Therefore, PEGylated and non-PEGylated samples should be concerned with the distribution of these self-associated structures.

Before the polymer coating, it was discovered that the original spherical IONPs in toluene had a hydrodynamic diameter of 19.8 nm. Following the polymer coating, the equivalent hydrodynamic diameters for the PEG-coated IONPs were  $135.3 \pm 38.8$  nm for spherical particles and  $132.6 \pm 36.9$  nm for cubic particles; see Table 3 and Figure 7. Note that we used a 220 nm filter to filter the particles because the hydrodynamic particle size after PEG coating was more than 250 nm (the literature states that the size is recommended to be lower than 250 to 300 nm). For instance, the initial hydrodynamic diameter of coated cubic IONPs was between 260 and 320 nm, whereas coated spherically shaped particles ranged between 300 and 900 nm. Afterwards, in Section 4.1.2, this outcome will be described.

**Table 3.** DLS measurements of the PEG-coated MNPs.

Sample	$D_H$ before Coating [nm]	$D_H$ after Coating [nm]	Zeta Potential (mV)
Spherical MNPs	$19.8 \pm 0.4$	$135.3 \pm 38.8$	$-26.9$
Cubic MNPs	$24.7 \pm 5.8$	$132.6 \pm 36.9$	$-28.3$

Where  $D_H$  [nm] is the particle hydrodynamic diameter measured by DLS (z-average).



**Figure 7.** The hydrodynamic diameters of the spherical PEG-coated MNPs (A) and cubic MNPs (B) were obtained from DLS measurement in the figure: PEG-sMNPs—PEG-coated spherical MNPs, and PEG-cMNPs—PEG-coated cubic MNPs.

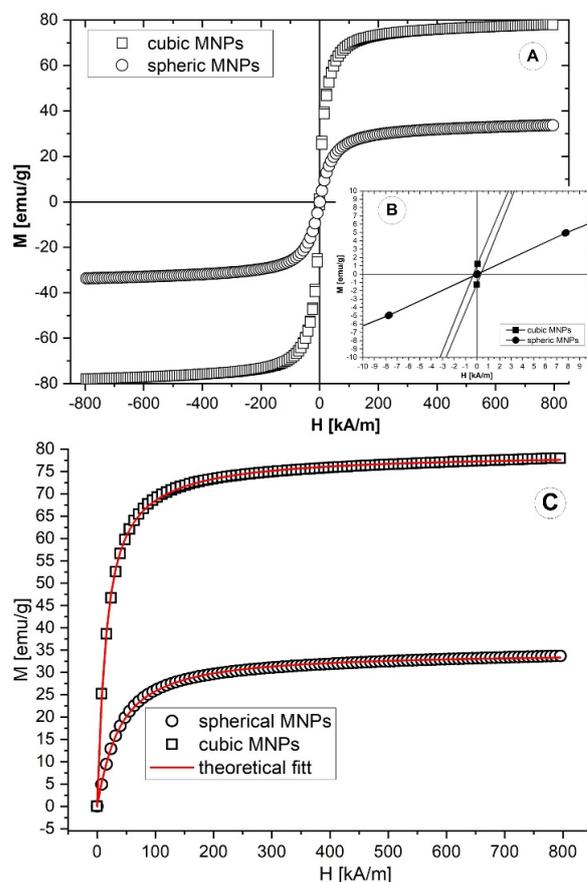
Zeta potential studies revealed that the average zeta potential for the cubic nanoparticles was  $-28.3$  mV. For the spherical nanoparticles, it was  $-26.9$  mV, indicating better biocompatibility and successful surface passivation for both coated IONPs.

### 3.3. IONPs' Magnetic Property

The DC magnetization of the two IONPs under investigation was measured at room temperature using a magnetometer (Alternating Gradient Magnetometer (AGM), Princeton PMC Model 2900, Westerville, OH, USA; VSM 880-ADE Technologies, Westwood, MA, USA) with a field range of 0 kA/m to 900 kA/m.

The hysteresis (M-H) loops for both samples are shown in Figure 8A,B at 294 K. The magnetic characteristics of the IONPs are summarized in Table 4. The spherical IONPs have zero remanence, while the cubic IONPs have 1.1 emu/g remanent magnetization due to their shape anisotropy. The reduced values of the squareness ratio,  $M_r/M_s$ , indicate that the particles are superparamagnetic and that their magnetization is dominated by thermal fluctuations [35]. The saturation magnetization ( $M_s$ ) at 900 kA/m is 34 emu/g for spheres and 78 emu/g for cubes. The saturation magnetization of both IONPs is lower than

the magnetite domain magnetization, 92 emu/g [36], due to the surfactant and the dead surface magnetic layer [37]. The latter is confirmed by the discrepancy between SPIONs' TEM and magnetic diameter (see Table 2). The SPIONs' magnetic diameter was determined by fitting the first magnetization curves (Figure 8C—symbols) with a theoretical model for highly concentrated magnetic colloids [38,39].



**Figure 8.** (A) Sphere- and cubic-shaped nanoparticle magnetization curves at room temperature (300 K). (B) Inset to (A): detail of the hysteresis loop. (C) Initial magnetization curve. Magnetic diameter is calculated from these curves.

**Table 4.** Magnetization saturation of the synthesized MNPs and comparison with published values.

Particles	Synthesis Procedure	Size, $D_{TEM}$ [nm]	Magnetic Diameter, $D_M$ [nm]	Saturation Magnetization, $M_s$ [emu/g]	References
Spherical	Thermal decomposition	$10 \pm 1.2$	$7.3 \pm 1.6$	34	Our results
Cubic	Thermal decomposition	$15.6 \pm 1.6$	$9.0 \pm 1.6$	78	Our results
Spherical	Thermal decomposition	14.5	-	$75 \pm 1$	[17]
Cubic	Thermal decomposition	12	-	$75 \pm 1$	[17]
Spherical	Thermal decomposition	$18.5 \pm 1.7$	$10.7 \pm 5.6$	74	[20]
Spherical	Controlled co-precipitation	$7.2 \pm 0.02$	5.6	42.1	[16]
Spherical	Co-precipitation	12	8.7	52.03	[40]
Spherical	Co-precipitation	10	-	34 ( $CM_{OA} = 16$ mM) 45 ( $CM_{OA} = 64$ mM)	[26]

Where  $CM_{OA}$ —molar concentration (CM) of the oleic acid (OA) (for  $Fe_3O_4$  nanoparticles colloids functionalized with oleic acid).

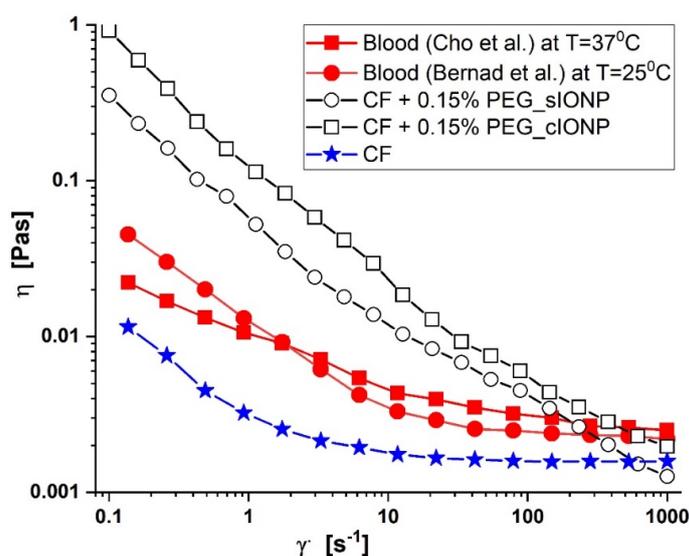
Table 4 compares the saturation magnetization of the MNPs synthesized in this paper and the different values presented in the literature.

### 3.4. Rheological Properties of the IONP's Aqueous Dispersion

#### 3.4.1. Model Suspension Preparation and Characteristics

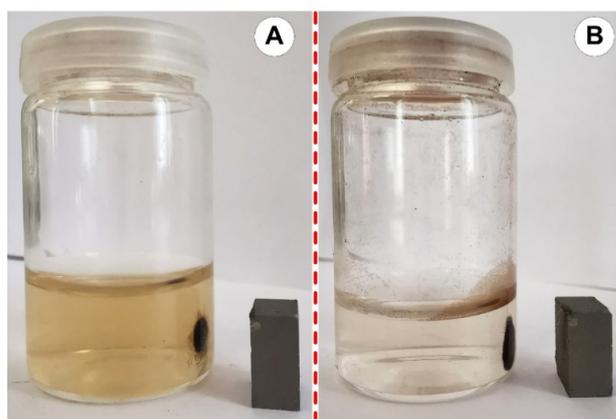
Glycerol–water solutions with a density ( $1060 \text{ kg/m}^3$ ) identical to that of blood served as the carrier fluid (CF) in our research [8,9]. These solutions were created by combining calculated weights of glycerol and distilled water. This CF facilitated the experimental studies by adequately reproducing the rheological behaviour of the fluid flow at the implanted stent site. In addition, the model suspension of magnetic carriers used in experiments was made by mixing PEG-coated IONPs with a carrier fluid (CF) at a mass concentration of 0.15 per cent.

Figure 9 displays the comparison of the blood sample [8] (obtained from a 38-year-old female volunteer in good health), the model suspension's viscosity curve (CF + 0.15% PEG\_IONPs) at  $T = 25^\circ\text{C}$ , and the findings reported in the literature [41].



**Figure 9.** Viscosity curves for blood (values from the literature and a healthy volunteer) and model suspension fluid (carrier fluid + 0.15% PEG-coated IONPs), (Cho et al. [41], Bernad et al. [42]).

The generated magnitude of the applied field determines the amount of change in the model suspension. The PEG-coated IONPs utilized in the model suspensions demonstrated the necessary responsiveness to a magnetic field's action, as seen in Figure 10. When a magnet is placed next to the bottle, the magnetic IONPs are drawn to it (Figure 10A,B), but the functionalized IONPs can quickly disperse in the suspensions when there is no magnetic field.



**Figure 10.** PEG-coated IONPs separation in the presence of the magnetic field. (A) cubic IONPs, (B) spherical IONPs.

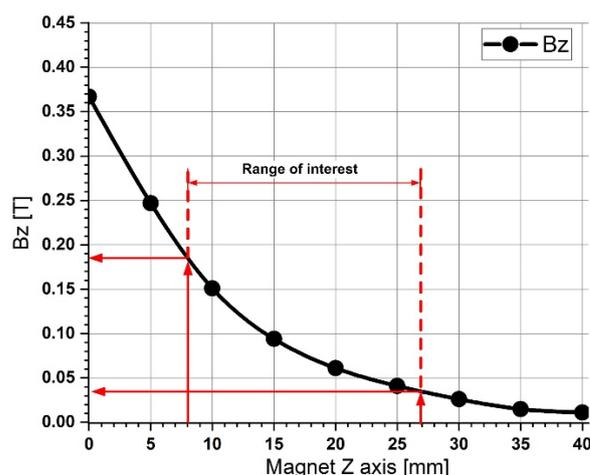
### 3.4.2. Magnetic Field

This paper aims to examine the potential of manufactured nanoparticles in drug-targeting applications; hence, it is crucial to understand the rheological behaviour of both simple and functionalized (with PEG) particles in the presence or absence of an external magnetic field that influences them.

Because the magnetic field's force must outweigh the hydrodynamic (drag) force and the shear effects produced by the running blood cells, the magnetic capture of nanoparticles at flow conditions typical of larger arteries may be more challenging. The magnetic field gradients and the flow dynamics consequently govern the behaviour of magnetic particles in circulation and the effectiveness of their aggregation in addition to the nanoparticle features [43].

External magnetic stimuli can affect magnetic suspensions and cause them to react. Whether a static magnetic field is homogenous or inhomogeneous affects how an object reacts [42]. If the field is homogenous (uniform in direction and amplitude), the material will undergo microstructural changes that will significantly affect its rheological and macroscopic features, such as the agglomeration of the particles into chains or columns [44].

In this paper, we used a Neodymium-type 52 magnet (NdFeB52) to produce the magnetic field with a maximum energy product ( $B \times H$ ) of 52 MGOe. In addition, we investigated the magnetic particle targeting in our research for magnetic fields generated between 0.18 T and 0.04 T, which corresponded to the magnet position between 8 mm and 27 mm from the magnet surface (along with the magnet z-axis—Figure 11).



**Figure 11.** A magnetic field generated by the NbFeB52 permanent magnet, and field distribution at various Z-axis positions from the magnet surface, measured using an F.W. Bell Gaussmeter, model 5080.

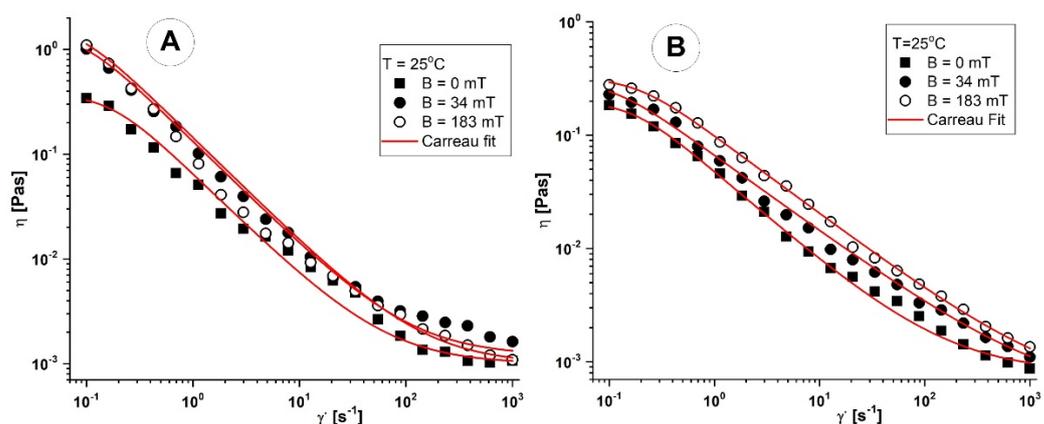
### 3.4.3. Viscosity Curves of the Spherical and Cubic MNPs Aqueous dispersions

In both the presence and absence of an external magnetic field, MNPs particle suspensions' steady-state shear flow behaviour was assessed. Because shear flow is the predominant flow element in most applications, the study was restricted to it. The viscosity curves were used to define the behaviour of the MNPs particle suspensions. The shear rate regime was selected to encompass the range of shear rates necessary for the drug targeting applications.

Parallel plate geometry was used to apply steady shear strain to the suspensions. A shear rate ramp-up in the  $0.1\text{--}1000\text{ s}^{-1}$  range was used to obtain the viscosity curves of apparent viscosity ( $\eta$ ) vs. shear rate ( $\dot{\gamma}$ ). The magnetorheological (MR) accessory produced a perpendicular external magnetic field to the shear flow. The temperature for all rheological testing was  $25.0\text{ }^{\circ}\text{C}$ . For these tests, the materials were placed into the rheometer immediately following ultrasonic homogenization (for all the samples, the applied energy

was 2000 J with 0.5 cycles and an amplitude of 80% in three consecutive steps), and we used freshly made suspensions for these tests.

Figure 12 depicts the variations in the shear viscosity of the cubic and spherical IONPs mixture suspensions as a function of shear rate for three distinct magnetic field intensities ranging from 0 to 183 mT. Stepwise, as the magnetic field intensity rose, the shear viscosity increased. However, the shear viscosity behaviour of the MNP suspensions of both kinds was comparable under each magnetic field. In addition, shear thinning behaviour was seen for both particle suspensions in the low shear rate area ( $0.1 \text{ s}^{-1}$  to  $1 \text{ s}^{-1}$ ), demonstrating that the structures of the dispersed magnetic particles formed under the external magnetic field persisted until the shear rate reached  $400 \text{ s}^{-1}$ .



**Figure 12.** Viscosity curves in the absence and presence of the magnetic field for investigated IONPs. (A) Spherical shape MNPs. (B) Cubic shape MNPs.

The Carreau model (Equation (1)) and the data from the  $\eta = f(\dot{\gamma})$  were associated [45]:

$$\eta(\dot{\gamma}) = \eta_{\infty} + (\eta_0 - \eta_{\infty}) \left[ 1 + (C\dot{\gamma})^2 \right]^{-p}, \quad (1)$$

in which  $C$  [s] represents the Carreau constant (the value of the slope of the viscosity curve in the log–log scale at high values of the shear rate  $\dot{\gamma}$ ),  $p$  [–] is the Carreau exponent,  $\eta_0$  is the viscosity at infinitely low shear rates, and  $\eta_{\infty}$  [Pas] is viscosity at infinitely high shear rates.

## 4. Discussion

### 4.1. Perspective of Biomedical Applications for the Cubic and Spherical Shape MNPs

The functionalized MNPs enter the bloodstream by intravenous injection and pass through the cardiovascular system. At this stage, protein adsorption might cause a considerable change in the surface chemistry, highlighting the need for effective shielding of the MNPs [46].

The injected MNPs will gradually diffuse outside the blood vessels and into the surrounding extracellular matrix (ECM). As a result, the local MNPs concentration in the blood is relatively high after injection, proportionate to the extravascular area. In this condition, some NPs are absorbed by tissue cells once they reach the extravascular space. In contrast, others are maintained in the interstitial fluid and eventually enter the lymphatic system [47].

It has been demonstrated that surface alterations, particularly polymer coatings, enhance the biocompatibility, dispersity, and stability of MNPs [32].

The steric repulsion prevents particle aggregation after the PEG surface coating since the nanoparticles are always expected to stay enclosed by their protective ligand shell.

The manufactured spherical and cubic PEG-coated MNPs were dispersed in an aqueous carrier (blood analogue fluid) to form the model suspensions, which precisely reflected the rheological behaviour of the fluid flow during the magnetic drug targeting application.

The model suspension results from the carrier fluid (CF) combination and 0.15 per cent mass concentration of the PEG-coated MNPs and the preparation of the blood analogue fluid (the carrier fluid) were presented in our previous chapter.

#### 4.1.1. Rheological Aspects of the PEG-Coated Cubic and Spherical IONPs Suspension from the Application Point of View

External magnetic stimuli have an impact on magnetic suspensions and can control them. For example, suppose the external magnetic field is homogenous (uniform in direction and strength). In that case, the material will undergo microstructural changes (such as the aggregation of the particles into chains), which will significantly affect the rheological characteristics of the suspensions.

So, from the perspective of magnetic targeting, altering the rheological behaviour of the injected therapeutical suspension can significantly impact the targeting processes' effectiveness. Therefore, in the following chapter, we will investigate the rheological behaviour of the model suspension both with and without an external magnetic field.

As mentioned in the previous chapter, a coil integrated into the plate–plate magnetorheological cell and set perpendicular to the plane of the plates created a homogeneous magnetic field that was used to measure the viscosity curves of the two model suspensions both with and without it.

First, we discuss how viscosity changes with the shear rate when the magnetic field strength is zero. Figure 13A,B display the logarithmically scaled viscosity curves of the suspensions with PEG-coated spherical and cubic magnetite that are not in a magnetic field. The two PEG-coated MNP suspensions behaved similarly and did not significantly differ. The aggregation of the particles caused the shear thinning of both samples (reduction in viscosity with increasing shear rate). When the shear rate was low, the particles attracted one another and stayed together in the suspension's volume.

The Van der Waals attraction force between the coated MNPs drives the agglomeration without a magnetic field. The higher viscosity ( $\eta \approx 1\text{--}10$  Pas) of the MNPs suspensions at low shear is caused by this network of agglomerated particles (both for PEG-coated spherical and cubic MNPs) (Figure 13A,B).

Shear forces destroy the particle network as the suspension starts to flow and the shear rate rises. As a result, the size of the agglomerates continues to decrease with increasing  $\dot{\gamma}$ , which lowers the viscosity. The viscosity may attain a constant value at high shear rates, at which point shear forces alone cannot decrease the size of the clusters.

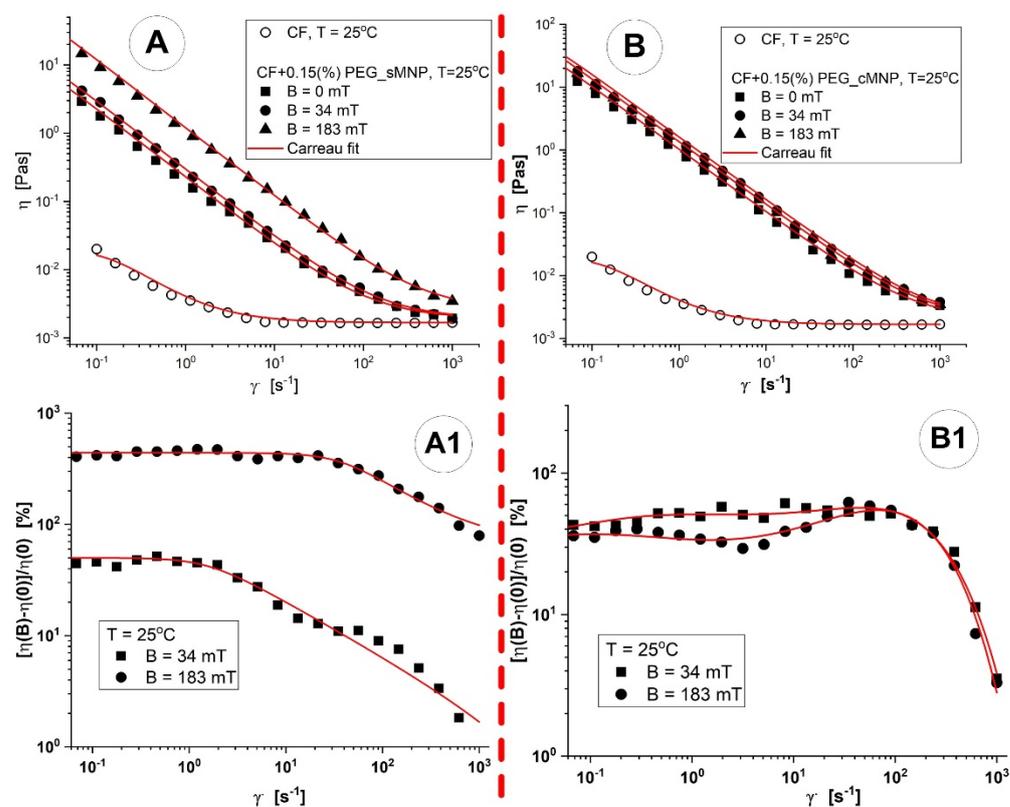
All suspensions showed altered flow behaviour in a magnetic field. Figure 13A,B compare the viscosity curves for zero magnetic fields and two different field strengths. Both suspensions displayed the previously noted shear-thinning property in a magnetic field.

The values obtained for the fit parameters are listed in Table 5. As expected, in the presence of the magnetic field, the values of the fit parameters  $\eta_0$ ,  $C$ , and  $p$  increase with the intensification of the field.

**Table 5.** The values of the fit parameters obtained by fitting the viscosity curves with the Carreau model for different magnetic flux density values for spherical and cubic nanoparticles at  $T = 25\text{ }^{\circ}\text{C}$ .

Particle	B [mT]	$\eta_{\infty}$ [Pas]	$\eta_0$ [Pas]	C [s]	$p$ [–]	$r^2$
Spherical	0	0.0019	15.4	67.6	0.4325	0.971
	34	0.0019	20.2	67.9	0.4335	0.971
	183	0.0026	140.0	116.6	0.474	0.989
Cubic	0	0.002	90	87.4	0.454	0.982
	34	0.002	130	95.4	0.460	0.984
	183	0.002	100	61.9	0.395	0.941

Where  $T$  [ $^{\circ}\text{C}$ ] is the carrier fluid temperature,  $B$  [T] is magnetic flux density,  $\eta_{\infty}$  [Pas] is the viscosity at infinite shear rates,  $\eta_0$  [Pas] is zero shear viscosity,  $C$  [s] is the characteristic time constant, and  $p$  [–] flows behaviour index. The  $r^2$  values for all fits are close to unity, indicating an excellent fit ( $r^2$  is the coefficient of determination used to evaluate the quality of the Carreau fits).



**Figure 13.** The model suspension viscosity curve in the presence and absence of the magnetic field: (A) suspension with spherical MNPs; (B) suspension with cubic MNPs. Magneto-viscous effects (MVE) as a function of shear rate at different magnetic flux densities: (A1) for suspension with spherical MNPs and (B1) suspension with cubic MNPs. CF—carrier fluid. All measurements were made at temperature  $T = 25\text{ }^{\circ}\text{C}$ .

Additionally, to better compare the characteristics of the two nanoparticles under different conditions, Table 6 summarizes the rheological properties of the MNPs under an applied magnetic field.

The single-domain IONPs provide the particle aggregates with a persistent magnetic dipole moment. The difference in magnetic permeability between the particles and the dispersion medium also induces a magnetic dipole moment in an external magnetic field. In a magnetic field, the attraction force is amplified due to the appearance of dipole-dipole interactions, which causes chain-like structures to develop parallel to the field. The suspensions in a magnetic field are more viscous due to this attractive structure perpendicular to the shear flow.

**Table 6.** Viscosity evolution of the spherical and cubic MNPs for different values of the shear rate for the investigated magnetic field.

B [mT]	Viscosity Values [Pa·s]									
	$\dot{\gamma} = 1000 \text{ s}^{-1}$		$\dot{\gamma} = 100 \text{ s}^{-1}$		$\dot{\gamma} = 10 \text{ s}^{-1}$		$\dot{\gamma} = 1 \text{ s}^{-1}$		$\dot{\gamma} = 0.1 \text{ s}^{-1}$	
	sMNPs	cMNPs	sMNPs	cMNPs	sMNPs	cMNPs	sMNPs	cMNPs	sMNPs	cMNPs
0	0.0013	0.0011	0.0041	0.0034	0.0103	0.0079	0.0525	0.0289	0.3531	0.1341
34	0.0017	0.00127	0.0045	0.0038	0.0159	0.0098	0.0696	0.0318	0.8124	0.1494
183	0.0029	0.00166	0.0076	0.0039	0.0175	0.0122	0.1069	0.1003	1.2086	0.699

Where sMNPs—spherical magnetic nanoparticles, and cMNPs—cubic magnetic nanoparticles.

When the applied shear rate rises, the nanoparticles organize themselves in the shearing direction. Additionally, as the shear rate increases, aggregate formations with a chain-like structure are destroyed, which causes the viscosity to decrease.

The magneto-viscous effect (MVE) is the term used to describe the relative increase in viscosity caused by a magnetic field. This effect is most pronounced when using weak magnetic fields; as the magnetic field increases, it usually decreases.

It can be observed in Figure 13(A1,B1) that the effect of MVE on the range of low shear rates is almost independent of the shear rate but significantly decreases at high speeds. At this point, the particle agglomerations are destroyed.

The main finding of this study is that, compared to cubic IONPs, spherical IONPs exhibit a significantly higher shift in flow behaviour due to the external magnetic field (Figure 13). This finding is crucial for magnetic targeting because it can impact the injectability and workability of the active suspension [48] when changes in the rheological properties of the model suspensions occur when an external magnetic field is present.

#### 4.1.2. PEG-Coated IONPs Sedimentation

This investigation also looked at the sedimentation of the PEG-coated nanoclusters (Figure 14). The sedimentation was recorded throughout 24 h as a function of time. The current findings demonstrate that the PEG-coated IONPs dispersion was stable enough to be used in the experimental study, given that the injection period of the model suspension used for experimental investigations in our prior work [8] was roughly 30 s.

Moreover, as can be seen from Figure 14, the cubic nanoparticles are more stable compared to the spherical ones. This could be due to the zeta potential value, which is two times larger for the cubic nanoparticles (Table 2), and which is known to influence the stability of nanoparticles in suspension [49]. A large positive or negative value for zeta potential indicates good stability as a result of the electrostatic repulsion of the particles. On the other hand, small values for zeta potential may lead to particle aggregation and flocculation. This is the result of the van der Waals attractive forces that act upon the particles [50].

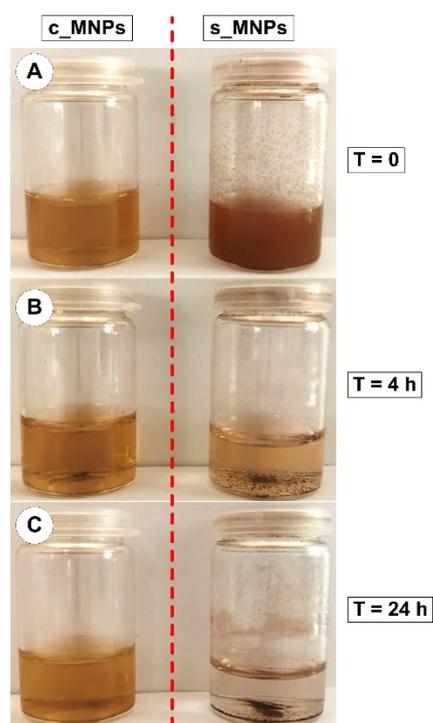
#### 4.1.3. PEG-Coated IONPs Aggregation

Monitoring the particles' colloidal stability is another crucial application of DLS [51]. Thus, it is essential to employ DLS to monitor the aggregation kinetic of MNPs to give precise input regarding the duration of this process [52].

In our investigations, the IONPs particles considerably agglomerate during storage in the suspensions. Therefore, the practical solutions were bath sonicated before each experiment to disperse these agglomerates and achieve a consistent initial particle size.

In this chapter, we assessed how the two synthesized IONPs aggregated. The increase in hydrodynamic radius caused by particle–particle aggregation was measured using DLS.

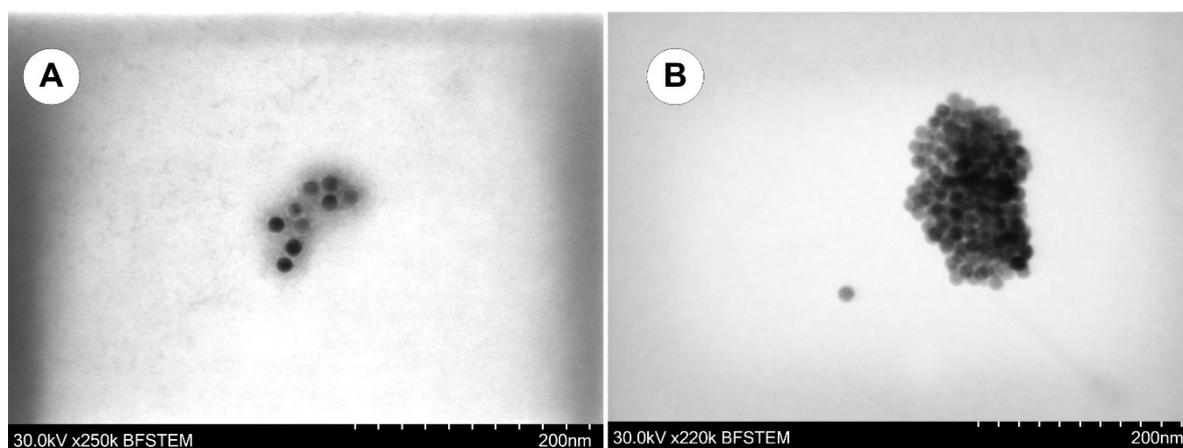
Several published sources detailing nanoparticle synthesis mention cluster sizes far more prominent than the initial size [53].



**Figure 14.** Sedimentation as a function of time for the PEG-coated IONPs suspensions.

Following synthesis, the measured hydrodynamic radius for spherical and cubic MNPs was 19.8 nm and 25 nm, respectively, 2 and 1.7 times bigger than the individual particle radius reported by TEM investigations. In addition, the PEG coating of the particles raised their measured radius by  $\approx 13$  times, to 153 nm for spherical IONPs and 154 nm for cubic IONPs, respectively.

Figure 15 shows that the IONPs particles strongly agglomerate after being stored in suspensions for 24 h. So, for an hour, we looked at the aggregation behaviour of both IONPs. As previously noted, the suspensions underwent a 10 min bath sonication procedure before the measurements began. When particle clusters more prominent than 200 nm in diameter are found during DLS measurements, accumulation is seen. It is essential to mention that the hydrodynamic diameter of both IONPs after PEGylation was around 135 nm (Figure 7).

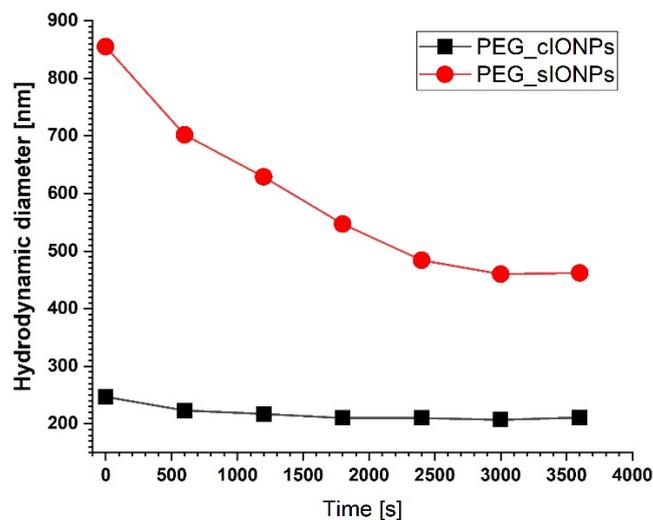


**Figure 15.** Iron oxide spherical nanoparticles coated with the PEG ligand according to the presented protocol. The 2D clusters were obtained after PEGylation of the IONPs. (A,B) A sample of the clustering after PEGylation, resulting from TEM investigation.

These clusters, which most likely developed during the production of nanomaterials, were very hard to break apart by bath sonication or solution agitation (Figures 14 and 15).

The aggregation behaviour depicted in Figure 15 aligns with earlier studies [54] looking at the aggregation of nanoparticles and nanoparticle clusters.

Figure 16 demonstrates that cubic-shaped IONPs are more stable throughout time. Hydrodynamic diameter measurements show a relatively slow variation of the DLS size distribution from 250 nm to 210 nm in the first 600 s. The PEG\_cIONPs continue to have a diameter of about 210 nm after that. Contrarily, over the 60 min investigation period, the hydrodynamic diameter of PEG\_sIONPs fluctuated significantly, from 855 nm to 460 nm.



**Figure 16.** Aggregation behaviour of the PEG-coated IONPs. Notation in the figure: PEG\_cIONPs—PEG-coated cubic IONPs, PEG\_sIONPs—PEG-coated spherically shaped IONPs.

The findings showed that most MNCs are soldered in substantial permanent clusters because of bridge contacts between polymer shells.

Another key finding relates to the sedimentation profile, specifically that PEG-coated IONPs with spherical shapes have a stronger tendency to sediment than those with cubic shapes, which are more stable.

Given that the average sizes of the IONPs particles (as determined by DLS) are  $19.8 \pm 0.4$  nm for spherical IONPs and  $24.7 \pm 5.7$  nm for cubic IONPs, it is assumed that the significant structures produced in the range of several hundred nanometres might result from the MVE's characteristics discussed in the previous chapter.

## 5. Conclusions

Different synthesis routes lead to nanoparticles with contrasting dimensions, morphologies, and magnetization. Moreover, an essential aspect of synthesis is the type of reagent used and its concentration. In our work, MNPs were synthesized through thermal decomposition. Two morphologies were obtained, spherical and cubic, by altering the reagent, the heating ramp rate, and the target temperature. The spherical and cubic nanoparticles had similar dimensions but significant magnetization differences. Even though the particles had superparamagnetic characteristics, the magnetic saturation obtained for the cubic particles ( $d = 15$  nm) was more than twice the one for the spherical particles ( $d = 10$  nm), a difference that may be influenced by the 5 nm distinction between the two core types. Additionally, the low magnetic saturation values may result as an effect of the oleic acid coating, which is considered to influence the magnetic moment from the particle's surface.

After PEGylation, the MNPs aggregated with 130–140 nm hydrodynamic diameters. These diameters exceeded the average values of 10–100 nm, which they must integrate to

be used for medical purposes [55]. In addition, by investigating the particles for 24 h, it was observed that the cubic ones were more stable in suspension than the spherical ones.

In the case of using PEGylated nanoparticles for medical applications such as drug targeting, several aspects need to be known and understood, including the use of a suitable magnetic field and the rheological characteristics of the nanoparticles. Furthermore, the rheological behaviour of the active suspension containing PEG-coated IONPs should be as similar as possible to that of the blood. The rheological curves from this paper showed that the suspensions present a higher viscosity at lower shear rates for spherical and cubic PEG-coated nanoparticles when a magnetic field is applied. In the case of the PEG-coated spherical nanoparticles, when a higher magnetic field is used ( $B = 183$  mT), the corresponding values show a sharp increase. In contrast, the PEG-coated cubic nanoparticles' rheological curves are almost superimposed regardless of the magnetic field's intensity. This suggests that the rheological behaviour of the PEG-coated cubic nanoparticles is weakly influenced if different values for the magnetic field are applied.

Considering all the above, both spherical and cubic magnetic nanoparticles can be suitable for medical applications due to their appropriate dimensions and high magnetization. However, a different method for coating the MNPs with PEG needs to be investigated to reduce the formation of aggregates and, consequently, the hydrodynamic diameters.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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