

*Communication*

# Co-Precipitated $Mn_{0.15}Ce_{0.85}O_{2-\delta}$ Catalysts for NO Oxidation: Manganese Precursors and Mn-Ce Interactions

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**Abstract:** Two  $Mn_{0.15}Ce_{0.85}O_{2-\delta}$  mixed oxides were synthesized by a co-precipitation method using  $Mn(NO_3)_2$  and  $KMnO_4$  as the manganese precursors, respectively. Structural analyses by X-ray powder diffraction and Raman spectroscopy reveal the formation of  $MnO_x$ - $CeO_2$  solid solutions. The  $Mn_{0.15}Ce_{0.85}O_{2-\delta}$  catalyst prepared from the high-valent manganese precursor exhibits higher activity for the catalytic oxidation of NO. The advantage of  $KMnO_4$  is related to the improved redox property of the catalyst as supported by  $H_2$  temperature-programmed reduction (TPR) and  $O_2$  temperature-programmed desorption (TPD). The Mn-Ce interactions create more  $Mn^{4+}$ ,  $Ce^{3+}$  and oxygen vacancies on the  $KMnO_4$ -synthesized mixed oxides based on the Raman and X-ray photoelectron spectra (XPS).

**Keywords:** solid solutions; precursor; oxidation state; lattice defects; NO oxidation

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## 1. Introduction

NO oxidation is an important reaction for environmental protection, such as NO oxidative adsorption [1], fast selective catalytic reduction (SCR) of NO [2,3], lean  $NO_x$  adsorption (LNT) [4], and  $NO_2$ -assisted soot oxidation [5,6]. Pt catalysts are widely considered as the most efficient catalysts despite their high costs. Ceria is one of the cheap catalytic materials and plays important roles in many catalytic reactions owing to easy redox cycle between  $Ce^{4+}$  and  $Ce^{3+}$ .

The redox property of ceria can be improved by interaction with transition metals, e.g.,  $Fe^{x+}$  [7],  $Mn^{x+}$  [8],  $Co^{x+}$  [9], and  $Cu^{x+}$  [10]. Among the mixed oxides,  $MnO_x$ - $CeO_2$  catalysts are the most favorable ones due to a favorable synergistic effect between these two metal oxides. Nevertheless, their catalytic activities for NO oxidation have not been fully explored. Machida et al. [11] found that the enhanced NO oxidative adsorption of the mixed oxides is facilitated by both the high NO oxidation activity of  $MnO_x$  and the strong  $NO_x$  adsorption ability of  $CeO_2$ . They pointed out that the active sites for NO oxidation were Mn species accompanied with reversible sorption/desorption of lattice oxygens based on XPS and  $O_2$ -TPD studies [12]. Li et al. [13] evidenced the existence of a synergistic mechanism between the oxides in the  $MnO_x$ - $CeO_2$ . They found that Mn-Ce-O<sub>x</sub> improved the activities of NO oxidation due to the increased surface area and the enhanced dispersion of  $MnO_x$  with the addition of Ce. In our previous study [14], this synergistic mechanism was further verified in a tight-contact mixture of  $MnO$  and  $CeO_2$  via a so-called long-ranged electronic interaction, which created more  $Mn^{4+}$  and oxygen vacancies compared with the loose-contact mixture of the oxides.

It is acceptable that such an electronic interaction would be much more significant in the solid solutions of  $\text{MnO}_x$  and  $\text{CeO}_2$  synthesized by chemical methods. In this work, two  $\text{Mn}_{0.15}\text{Ce}_{0.85}\text{O}_{2-\delta}$  mixed oxides were synthesized by co-precipitation using different manganese precursors. Detailed microstructural and surface property characterizations were carried out to establish structure-performance relationships of the catalysts for NO oxidation.

## 2. Experimental

### 2.1. Catalyst Preparation

Two  $\text{Mn}_{0.15}\text{Ce}_{0.85}\text{O}_{2-\delta}$  (corresponding to a Mn/Ce molar ratio of 15/85) were synthesized by a co-precipitation method with different manganese precursors. KOH solution was added dropwise to a precursor solution of  $\text{Mn}(\text{NO}_3)_2$  (aqueous solution, 50 wt.%, Aladdin, China) and  $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$  (99.99 wt.%, Aladdin) to adjust the pH to 10.5 and stirred at 50 °C for 2 h. In order to ensure the complete removal of potassium ions, the precipitate was filtrated and washed at least four times. After drying the solid at 110 °C for 12 h and calcination at 500 °C for 6 h in air, the obtained catalyst was denoted as MnCe-L (the low-valent manganese sample). Another  $\text{Mn}_{0.15}\text{Ce}_{0.85}\text{O}_{2-\delta}$  catalyst was prepared by the same method except using  $\text{KMnO}_4$  (Aladdin) as the manganese precursor, and the obtained catalyst was denoted as MnCe-H (the high-valent manganese sample). Due to electronic interactions between  $\text{Mn}^{x+}$  and  $\text{Ce}^{x+}$  and thermal oxidation/reduction of metal ions during calcination, different manganese and cerium ions co-existed in the final products.

For reference,  $\text{Mn}_3\text{O}_4$  and  $\text{CeO}_2$  were synthesized by a similar precipitation method using  $\text{Mn}(\text{NO}_3)_2$  and  $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$  as the precursors, respectively. The physical mixture of  $\text{Mn}_3\text{O}_4$  and  $\text{CeO}_2$  at a Mn/Ce molar ratio of 15/85 was prepared by mixing  $\text{Mn}_3\text{O}_4$  and  $\text{CeO}_2$  powders using a spatula for two minutes.  $\text{MnO}_2$  (Aladdin) and  $\text{MnO}$  (Aladdin) were also used for reference.

### 2.2. Activity Measurement

Temperature-programmed oxidation (TPO) of NO was performed in a fixed-bed reactor connected to an infrared spectrometer Nicolet iS10 (Thermo Fisher, Waltham, MA, USA) with a gas mixture of 1000 ppm NO/10%  $\text{O}_2/\text{N}_2$  (500  $\text{mL} \cdot \text{min}^{-1}$ ) as the reaction gas. In order to smooth airflow and heat dissipation, the catalyst powders (100 mg) were mixed with silica pellets (300 mg) using a spoon. Then, the mixture was placed in a tubular quartz reactor, and was heated to 700 °C at a rate of 10 °C·min<sup>-1</sup>.

### 2.3. Catalyst Characterization

X-ray diffraction (XRD) patterns were measured on a D8 ADVANCE diffractometer (Bruker, Germany) using  $\text{Cu K}_\alpha$  radiation ( $\lambda = 0.15418 \text{ nm}$ ). The XRD patterns ranged between 10° and 80° at 0.02° intervals and recorded at a scanning speed of 4°·min<sup>-1</sup>.

Specific surface areas of the samples were determined on an automatic surface analyzer (F-Sorb 3400, Gold APP Instrument, Beijing, China) by the four-point Brunauer–Emmett–Teller (BET) method. Prior to the experiment, the samples were degassed at 200 °C for 2 h to remove the molecules adsorbed on the sample surface.

X-ray fluorescence analysis (XRF) was performed on a Shimadzu instrument (1800 Kyoto, Japan) to determine the elemental contents of the samples.

Raman spectra were collected on a confocal micro-Raman apparatus (Aurora J300, IDSpec, Hong Kong, China) using an  $\text{Ar}^+$  laser with a wavelength of 632.8 nm.

X-ray photoelectron spectra (XPS) were performed on a PHI-Quantera SXM system equipped with  $\text{Al K}_\alpha$  radiation. The data was corrected based on C 1s (284.8 eV).

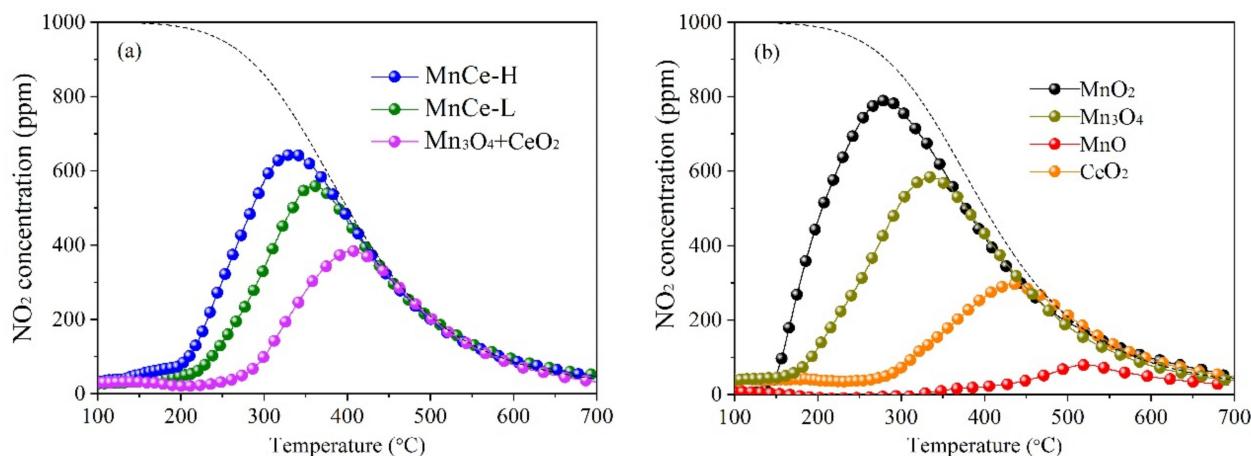
Temperature-programmed reduction (TPR) of  $\text{H}_2$  was conducted on an Auto Chem II Chemisorption Analyzer (Micromeritics, USA). A 50 mg sample was pretreated with a He flow (50  $\text{mL} \cdot \text{min}^{-1}$ ) at 450 °C for 30 min. After being cooled down to room temperature, the sample was heated in 10%  $\text{H}_2/\text{Ar}$  (50  $\text{mL} \cdot \text{min}^{-1}$ ) to 900 °C with a rate of 10 °C·min<sup>-1</sup>.

Temperature-programmed desorption (TPD) of O<sub>2</sub> was carried out on the Auto Chem II Chemisorption Analyzer. After purging in He at 300 °C for 1 h, 200 mg sample was exposed to 10% O<sub>2</sub>/He at 50 °C for 1 h. The reactor was heated to 900 °C at a rate of 10 °C·min<sup>-1</sup> after flushed by He for half an hour.

### 3. Results and Discussion

#### 3.1. NO Oxidation Activity

Figure 1 shows the NO oxidation activities of the mixed oxide catalysts and reference samples. The dotted NO<sub>2</sub> profile was drawn by the thermodynamic equilibrium of  $\text{NO} + \text{O}_2 \rightleftharpoons \text{NO}_2$ . Pure CeO<sub>2</sub> shows low NO conversions (<30%) within the whole temperature range, which is ascribed to its relatively poor redox behavior. The NO oxidation activity of manganese oxides increases with the valance of the metal ions, i.e., MnO<sub>2</sub> > Mn<sub>3</sub>O<sub>4</sub> > MnO. Specifically, Mn<sub>3</sub>O<sub>4</sub> exhibits a quite strong ability to oxidize NO to NO<sub>2</sub> and reaches the maximum NO conversion (59%) at 337 °C. The mixture of Mn<sub>3</sub>O<sub>4</sub> and CeO<sub>2</sub> shows NO conversions between the two monoxide components. Given the same composition of the mixed oxides, both the co-precipitated samples present much higher catalytic performance than the mechanical mixture. MnCe-H even shows slightly higher activity than Mn<sub>3</sub>O<sub>4</sub>, while MnCe-L has somewhat lower NO conversions at the temperatures lower than 360 °C. Considering the dominating species of Ce in the mixed oxides, these facts demonstrate strong synergistic effects between manganese oxide and ceria in the co-precipitated samples, which depend importantly on the preparation methods and the precursor adopted. Although MnO<sub>2</sub> exhibits higher NO oxidation activity than Mn<sub>0.15</sub>Ce<sub>0.85</sub>O<sub>2-δ</sub> mixed oxides, the poor thermal stability [15], poor sulfur resistance [16] and relatively high cost limit its industrial applications. It is well known that phase transformation of MnO<sub>2</sub> to Mn<sub>2</sub>O<sub>3</sub> occurs readily at 200–720 °C [17], while Mn<sub>0.15</sub>Ce<sub>0.85</sub>O<sub>2-δ</sub> catalysts maintain high NO oxidation activity after calcination at 650 °C [18].

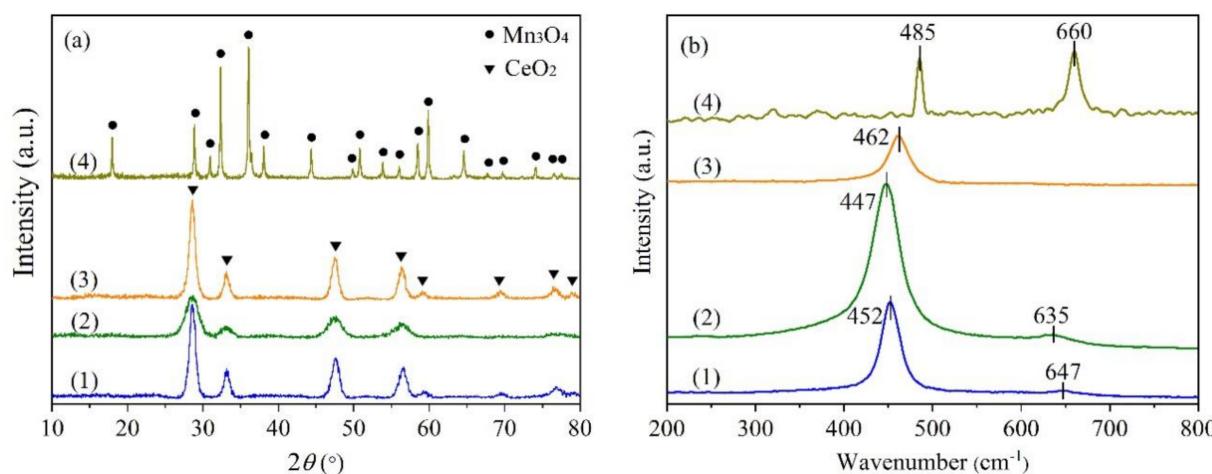


**Figure 1.** NO<sub>2</sub> evolution during NO-TPO tests over (a) Mn<sub>0.15</sub>Ce<sub>0.85</sub>O<sub>2-δ</sub> mixed oxides and (b) reference oxide catalysts. Reaction conditions: NO = 1000 ppm, O<sub>2</sub> = 10%, N<sub>2</sub> in balance, GHSV = 30,000 h<sup>-1</sup>.

#### 3.2. Structural Properties

The powder XRD patterns of the samples are shown in Figure 2a. The diffraction peaks of the precipitated manganese oxide and ceria correlate well with the characteristics of tetragonal Mn<sub>3</sub>O<sub>4</sub> and cubic CeO<sub>2</sub>, respectively [19]. Only diffraction peaks of ceria are observed in the patterns of two Mn<sub>0.15</sub>Ce<sub>0.85</sub>O<sub>2-δ</sub> mixed oxides. The absence of any MnO<sub>x</sub>-related peaks implies highly dispersed state of MnO<sub>x</sub> clusters in the matrix of ceria or the formation of MnO<sub>x</sub>-CeO<sub>2</sub> solid solutions [20]. Table 1 lists the lattice constants and mean crystallite sizes of the samples calculated according to Cohen's method and Williamson-Hall equation, respectively. Apparently, the calculated lattice constants of ceria

in the mixed oxides are smaller than that of pure ceria. The shrinkage of the ceria crystal cell is attributed to the incorporation of smaller manganese ions ( $Mn^{4+}$ : 0.053 nm;  $Mn^{3+}$ : 0.065 nm;  $Mn^{2+}$ : 0.083 nm) into the ceria ( $Ce^{3+}$ : 0.114 nm;  $Ce^{4+}$ : 0.097 nm) lattice to form metastable pseudo-solid solutions [11,21]. It herein indicates the substitution of cerium by manganese in the crystal cell of ceria in the co-precipitated mixed oxides. Furthermore, MnCe-L presents a larger lattice parameter than MnCe-H, which is associated with the incorporation of a larger amount of large  $Mn^{2+}$  cations in the former mixed oxides. As listed in Table 1,  $Mn_3O_4$  presents an exceptionally lower BET surface area owing to significant grain growth of manganese oxide [12]. It is also noted that MnCe-H shows a higher specific surface area than MnCe-L due to the smaller crystallites. No potassium was detected by XRF, and thus its influence can be excluded. The atomic ratio of  $Mn/(Mn + Ce)$  is close to 0.15 for both the mixed oxides, correlating with the theoretical value.



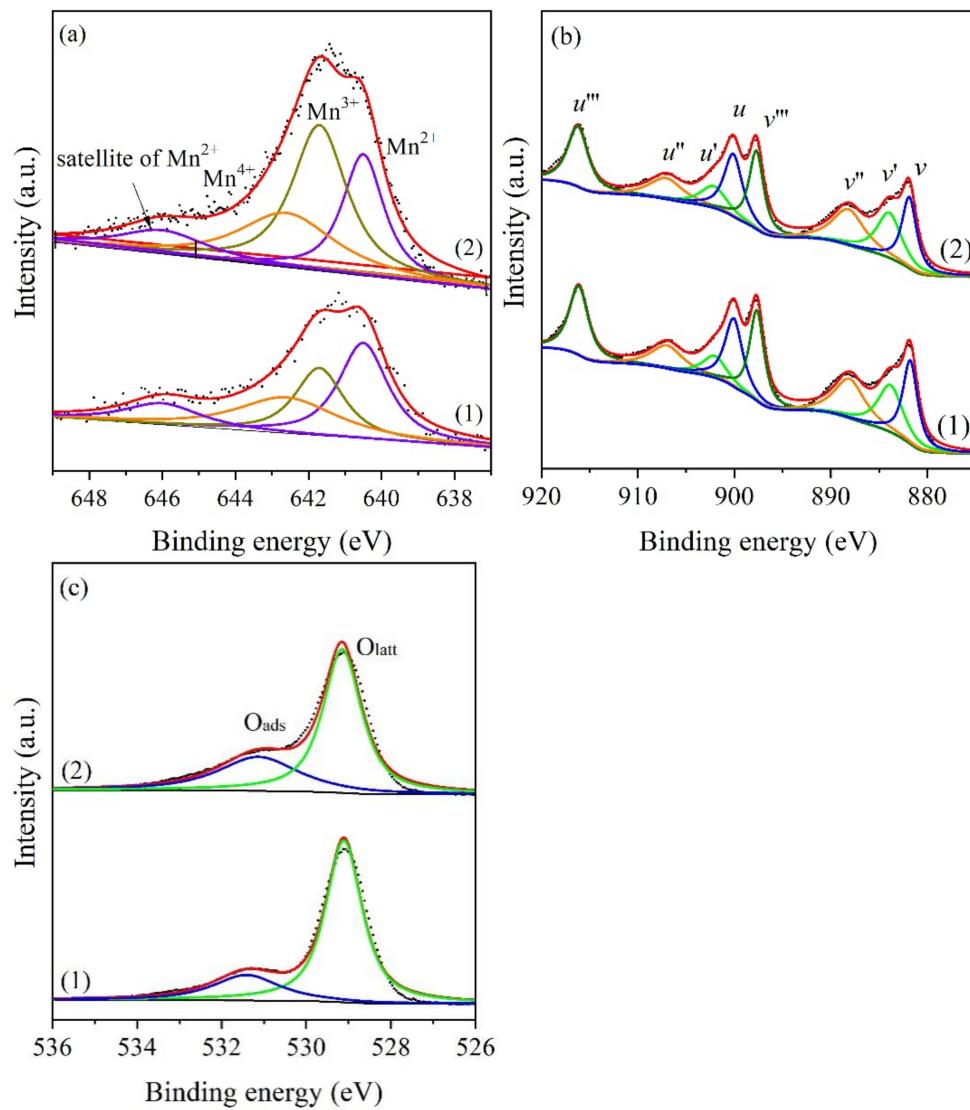
**Figure 2.** (a) XRD patterns and (b) Raman spectra of (1) MnCe-L, (2) MnCe-H, (3) CeO<sub>2</sub> and (4) Mn<sub>3</sub>O<sub>4</sub>.

**Table 1.** Structural properties of the catalysts.

Catalyst	Lattice Parameter (nm)	Crystallite Size (nm)	S <sub>BET</sub> (m <sup>2</sup> ·g <sup>-1</sup> )
MnCe-L	$a = b = c = 0.5387$	6.9	96
MnCe-H	$a = b = c = 0.5382$	4.2	126
CeO <sub>2</sub>	$a = b = c = 0.5394$	6.7	98
Mn <sub>3</sub> O <sub>4</sub>	$a = b = 0.5754; c = 0.9432$	36.3	23

Raman spectroscopy was applied to obtain additional information of both M-O band and lattice defects, and the results are shown in Figure 2b. The bands at 485 and 660 cm<sup>-1</sup> can be attributed to the stretching mode of the Mn-O lattice of Mn<sub>3</sub>O<sub>4</sub> [22]. CeO<sub>2</sub> exhibits a distinct at 462 cm<sup>-1</sup> assigned to a vibration mode of F<sub>2g</sub> symmetry. This band shifts towards lower wavenumbers to different degrees for the mixed oxides. Such red shifts are related to the creation of more oxygen vacancies, or the expansion of the crystal cell [23]. According to the XRD results, the latter possibility can be excluded. Thus, the red shifts correspond to nonstoichiometry of CeO<sub>2-δ</sub>. The greater shift degree of MnCe-H is attributed to a more significant change of CeO<sub>2</sub> environment interacted by the substituted high-valent Mn<sup>4+</sup> and Mn<sup>3+</sup> cations. The ceria band is symmetrical and remains undisturbed by the manganese oxide feature at 485 cm<sup>-1</sup>. Additionally, the spectrum of the Mn<sub>3</sub>O<sub>4</sub> + CeO<sub>2</sub> mixture (not shown) exhibits not only the characteristics of CeO<sub>2</sub> at 456 cm<sup>-1</sup> but also that of Mn<sub>3</sub>O<sub>4</sub> at 653 cm<sup>-1</sup>. Herein, the bands at 635 and 647 cm<sup>-1</sup> are associated with the oxygen vacancies induced by the generation of Ce<sup>3+</sup> [19,24], which show a similar shift trend with the main band. These oxygen vacancies are conducive to the diffusion of lattice oxygen, further promoting the oxidation reaction.

Figure 3 shows the deconvoluted XPS spectra of the two  $\text{Mn}_{0.15}\text{Ce}_{0.85}\text{O}_{2-\delta}$  mixed oxides. As shown in Figure 3a, the Mn  $2p_{3/2}$  spectra reveal three main peaks at 640.5, 641.7 and 642.6 eV attributed to the presence of  $\text{Mn}^{2+}$ ,  $\text{Mn}^{3+}$  and  $\text{Mn}^{4+}$ , respectively [25,26]. The peak at 646.0 eV is the shakeup satellite of divalent Mn bound to O as the charge-transfer compound [24]. The relative percentage of  $\text{Mn}^{2+}$  was calculated by the peak area ratio of  $\text{Mn}^{2+}/(\text{Mn}^{2+} + \text{Mn}^{3+} + \text{Mn}^{4+})$ . Obviously, the  $\text{Mn}(\text{NO}_3)_2$ -derived MnCe-L processes more  $\text{Mn}^{2+}$  (40.6%) than the  $\text{KMnO}_4$ -derived one (27.6%). That is, the choice of the Mn precursor affects the oxidation state of manganese ions in the obtained products.



**Figure 3.** XPS spectra of (a) Mn  $2p_{3/2}$ , (b) Ce  $3d$  and (c) O  $1s$  on (1) MnCe-L and (2) MnCe-H.

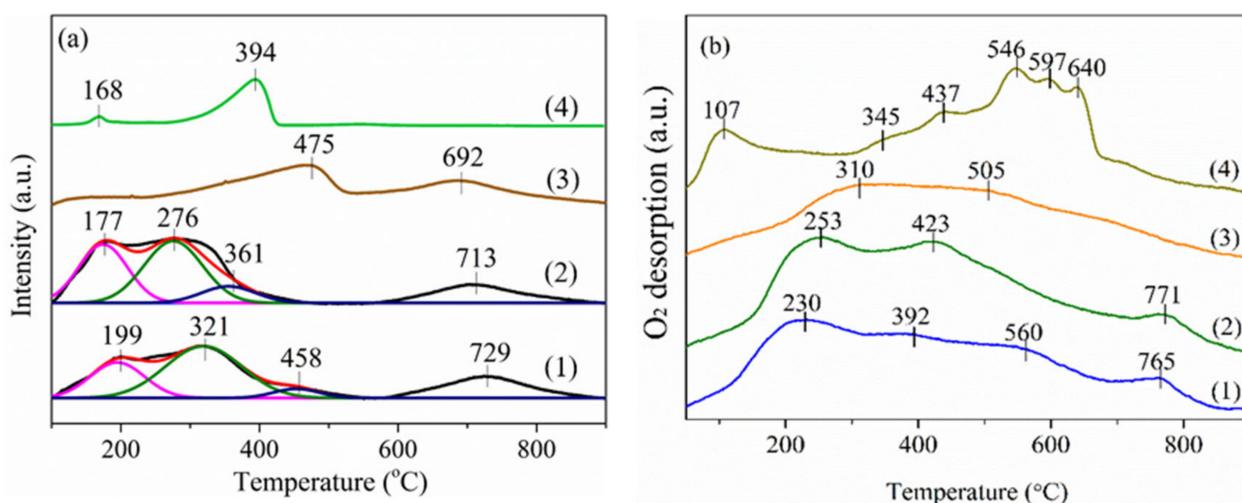
Figure 3b shows the XPS spectra of Ce  $3d$  fitted with eight peaks, including those of  $\text{Ce}^{4+} 3d_{3/2}$  ( $u$ ,  $u''$  and  $u'''$ ),  $\text{Ce}^{4+} 3d_{5/2}$  ( $v$ ,  $v''$  and  $v'''$ ),  $\text{Ce}^{3+} 3d_{3/2}$  ( $u'$ ), and  $\text{Ce}^{3+} 3d_{5/2}$  ( $v'$ ). The relative percentage of  $\text{Ce}^{3+}$  was calculated by the peak area ratio of  $v'/(v + v' + v'' + v''')$ . The obtained surface  $\text{Ce}^{3+}/\text{Ce}$  ratio on MnCe-H (25.7%) is somewhat higher than that on MnCe-L (23.8%). The changes are not so significant for the multiple splitting of the Ce  $3d$  signals, which may be related to the dominant content of Ce in the mixed oxides. According to Machida's work [11], it is plausible for  $\text{Mn}^{3+}$  to substitute  $\text{Ce}^{4+}$  in the fluorite structure when considering their structural similarity, although their ionic radius are quite different. In the present work, the introduction of more  $\text{Mn}^{3+}$  instead of  $\text{Mn}^{2+}$  in MnCe-H produces

more  $\text{Ce}^{3+}$  due to the electronic equilibrium and structural stability since  $\text{Ce}^{3+}$  has a lower oxidation state and a larger ionic radius than  $\text{Ce}^{4+}$ .

The formation of  $\text{Ce}^{3+}$  can also produce more oxygen vacancy and adsorbed oxygen species, which is confirmed by the XPS spectra of O 1s in Figure 3c. The peaks at 529.4 and 531.9 eV are characteristic of lattice oxygen ( $\text{O}_{\text{latt}}$ ) and surface adsorbed oxygen ( $\text{O}_{\text{ads}}$ ), respectively [27]. The relative percentage of  $\text{O}_{\text{ads}}$  was calculated by the peak area ratio of  $\text{O}_{\text{ads}}/(\text{O}_{\text{latt}} + \text{O}_{\text{ads}})$ . More adsorbed oxygen (33.1%) is detected on the surface of MnCe-H than on MnCe-L (22.3%), accompanied with a high ratio of  $\text{Ce}^{3+}$ .

### 3.3. Redox Properties

The reducibility of metal ions was investigated by  $\text{H}_2$ -TPR. Figure 4a shows the reduction profiles of  $\text{Mn}_3\text{O}_4$ ,  $\text{CeO}_2$  and the co-precipitated mixed oxides.  $\text{CeO}_2$  has two typical peaks at 475 and 692 °C assigned to the reduction of surface/subsurface  $\text{Ce}^{4+}$  and bulk  $\text{Ce}^{4+}$ , respectively.  $\text{Mn}_3\text{O}_4$  shows a broad peak at 394 °C due to reduction of  $\text{Mn}^{3+}$  to  $\text{Mn}^{2+}$ . The calculated  $\text{H}_2$  consumption ( $1720 \mu\text{mol}\cdot\text{g}^{-1}$ ) is much smaller than the theoretical value ( $4370 \mu\text{mol}\cdot\text{g}^{-1}$ ) for the reduction of  $\text{Mn}_3\text{O}_4$  to  $\text{MnO}$ , which is mainly caused by the difficult reduction of large manganese oxides [28]. A shoulder at 168 °C belongs to the reduction of surface  $\text{MnO}_2$  clusters to  $\text{Mn}_3\text{O}_4$  [29], although they cannot be detected by XRD.



**Figure 4.** (a)  $\text{H}_2$ -TPR and (b)  $\text{O}_2$ -TPD profiles of (1) MnCe-L, (2) MnCe-H, (3)  $\text{CeO}_2$  and (4)  $\text{Mn}_3\text{O}_4$ .

Compared with  $\text{Mn}_3\text{O}_4$ , MnCe-L and especially MnCe-H show a similar but much stronger peak at 177–199 °C, which implies the generation of more  $\text{Mn}^{4+}$  in the mixed oxides. The calculated  $\text{H}_2$  consumption for the first deconvoluted peak in MnCe-H ( $624 \mu\text{mol}\cdot\text{g}^{-1}$ ) is larger than that in MnCe-L ( $568 \mu\text{mol}\cdot\text{g}^{-1}$ ). Meanwhile, the second peak shifts towards lower temperatures (276–321 °C). The calculated  $\text{H}_2$  consumption for this peak is 804 and  $989 \mu\text{mol}\cdot\text{g}^{-1}$  for MnCe-H and MnCe-L, respectively. These data are larger than the theoretical value ( $321 \mu\text{mol}\cdot\text{g}^{-1}$ ) by assuming that all manganese ions in the mixed oxides exist in the form of  $\text{Mn}_3\text{O}_4$  and  $\text{MnO}$  is the final product. Herein, it is ascribed not only to the reduction of  $\text{Mn}_3\text{O}_4$  to  $\text{MnO}$  but also to that of  $\text{CeO}_2$  to  $\text{Ce}_2\text{O}_3$  promoted by  $\text{Mn}^{x+}$  in the solid solutions or at the interface. Similarly, the promoted reduction of  $\text{Ce}^{4+}$  may also contribute to the first peak. Those small peaks at 361–458 °C are suggested to be associated with the reduction of less affected  $\text{Mn}_3\text{O}_4$  and  $\text{CeO}_2$ . The high-temperature peaks at 713–729 °C are assigned to the reduction of bulk oxygen from ceria unpromoted with Mn.

The  $\text{O}_2$ -TPD technique was used to determine the reactivity of oxygen species of metal oxides. As shown in Figure 4b, the peak at 107 °C is related to desorption of superoxide ion  $\text{O}_2^-$  on the  $\text{Mn}_3\text{O}_4$  surface [30]. The peaks at 345 and 437 °C are assigned to the desorption

of chemisorbed oxygen species [31]. The peaks range from 450 to 650 °C are associated with the surface lattice oxygen species [25]. Overlapped broad peaks, which are attributed to the reduction of superoxide ions  $O_2^-$  and peroxide ions  $O_2^{2-}/O^-$  on the surface, are observed for  $CeO_2$  [32]. The desorption curves of the mixed oxides are more complicated. The distinct peaks at 230–253 °C and 392–423 °C are attributed to the desorption of surface-active oxygen species bound to oxygen vacancies (i.e.,  $O_2^-$  and  $O^-$ ) and the coordinately unsaturated surface lattice oxygen, respectively [33]. It can be seen by integrating the peak area that more surface-active oxygen species are produced on MnCe-H than on MnCe-L. It implies that the strong metal-metal interaction in the MnCe-H results in more lattice defects (such as oxygen vacancies) and facilitates the creation of active  $O_2^{2-}/O^-$ . The peak at around 560 °C is related to the coordinately saturated surface lattice oxygen species. The desorption peaks at 765–771 °C are attributed to lattice oxygen ion  $O^{2-}$  from  $CeO_2$  promoted by  $Mn^{x+}$  [34]. It suggests that the incorporation of Mn into ceria enhances the mobility of lattice oxygen. These results are consistent with the H<sub>2</sub>-TPR results.

#### 4. Conclusions

$Mn(NO_3)_2$  and  $KMnO_4$  were applied as the manganese precursors to synthesize  $Mn_{0.15}Ce_{0.85}O_{2-\delta}$  mixed oxides by coprecipitation, respectively. Manganese existed mainly in a pseudo-solid solution with a fluorite-type structure of  $CeO_2$ . The application of  $KMnO_4$  leads to a higher proportion of  $Mn^{4+}$  and  $Mn^{3+}$  in the mixed oxide than that in the  $Mn(NO_3)_2$ -dervied one. The incorporation of  $Mn^{4+}$  and  $Mn^{3+}$  into the ceria lattice results in the formation of more  $Ce^{3+}$  due to the electronic equilibrium and structural stability. Herein, more oxygen vacancies and active oxygen species are created on the former catalyst, resulting in a maximum NO conversion of 64% achieved at 331 °C for NO oxidation.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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