

## Article

# BODIPY-Pyridylhydrazone Probe for Fluorescence Turn-On Detection of Fe<sup>3+</sup> and Its Bioimaging Application

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**Abstract:** A novel pyridylhydrazone-tethered BODIPY (BODIPY-PH) was synthesized, fully characterized via nuclear magnetic resonance (NMR), Fourier transform infrared spectroscopic (FTIR), and single-crystal X-ray diffraction (SC-XRD) techniques, and developed for the selective detection of Fe<sup>3+</sup> through fluorescent enhancement process. This derivative showed 1:1 binding with Fe<sup>3+</sup> in an acetonitrile-water mixture (1:9 v/v) with the binding constant (K) of 5.4 × 10<sup>4</sup> M<sup>-1</sup> and the limit of detection of 0.58 μM. The Fe<sup>3+</sup> complexation reaction has been proved to be a reversible process and could be effectively repeated up to three cycles. The electronic properties of BODIPY-PH and its Fe<sup>3+</sup> complex modeled by the density functional theory (DFT) method suggested the presence of chelation-enhanced fluorescence (CHEF) effect in the Fe<sup>3+</sup> binding reaction. The X-ray absorption spectroscopy (XAS) probed at Fe K-edge confirmed the complex formation between BODIPY-PH and the Fe<sup>3+</sup> in an octahedral geometry. Finally, bioimaging against human embryonic kidney (Hek293) cell, through confocal fluorescence microscopic technique indicated that the BODIPY-PH displayed good permeability and low toxicity toward the tested cell lines and showed enhanced fluorescent signal in the cells incubated with Fe<sup>3+</sup> proving its capability for Fe<sup>3+</sup> analysis in cellular matrix.

**Keywords:** BODIPY; ferric ions; fluorescence; bioimaging; pyridyl group

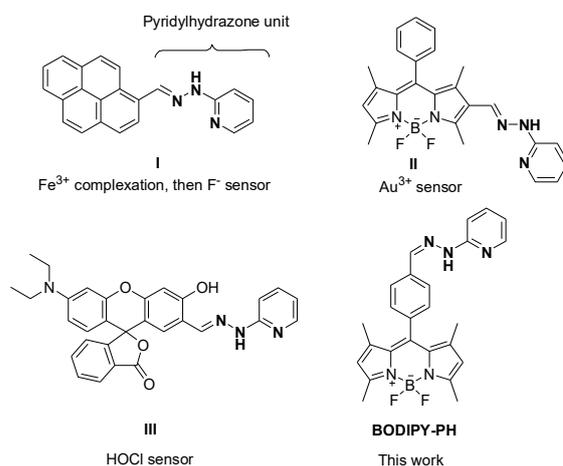
## 1. Introduction

Iron is one of the most important transition metals due to its vital roles in oxygen metabolism and electron transfer processes in numerous biological systems [1–4]. It catalyzes oxygen-related reactions in several enzymes and involves in DNA and RNA synthesis [5,6]. However, excessive iron in live cells can induce the generation of reactive oxygen species (ROS) leading to cellular toxicity. This phenomenon has been linked to serious neurological disorders, such as Alzheimer's, Wilson's, and Parkinson's disease [7–9]. In addition, immoderate iron uptake could cause detrimental diseases, including kidneys,

pancreas, and liver failures as well as certain cancers [10–13]. Therefore, the development of efficient sensors for iron detections in clinical, environmental, and industrial samples is highly needed for safety issues.

Nowadays, standard methods for ferric ion ( $\text{Fe}^{3+}$ ) detection rely on sophisticated instruments, such as atomic absorption spectrometry (AAS) [14], inductively coupled plasma mass spectroscopy (ICP-MS) [15], and inductively coupled plasma-atomic emission spectrometry (ICP-AES) [16]. These techniques are expensive, time-consuming, and not suitable for both on-site and cellular analysis. On the other hand, fluorescence technique cooperating with appropriate molecular sensors or carbon dots is a simple and cost-effective detection method and it can be applicable for the detections in field tests, cellular samples, as well as on paper substrates [17–21].

As part of our continuing interest in the development of boron-based imaging agents [22–24], we decided to synthesize a novel fluorescence compound based on boron difluoride dipyrromethene (BODIPY) and functionalize it with a pyridylhydrazone (PH) unit acting as a metal-binding site (Figure 1). Although the capability of PH unit in  $\text{Fe}^{3+}$  complexation has been demonstrated in pyrene-based compounds (I), the sensing properties of the PH-functionalized fluorescent probe toward ferric ions have not been systematically reported elsewhere [25]. Apart from metal binding, the PH moiety is involved in reaction-based fluorescence turn-on sensor for  $\text{Au}^{3+}$  ions and hypochlorous acid (HOCl) in BODIPY (II) and rhodamine (III) backbones, respectively (Figure 1) [26,27]. Herein, we reported the synthesis of pyridylhydrazone-tethered BODIPY compound (BODIPY-PH) and its systematic sensing properties toward ferric ions in aqueous solutions using UV-Vis/fluorescence spectroscopic, nuclear magnetic resonance (NMR), Fourier transform infrared spectroscopic (FTIR), and X-ray absorption spectroscopic (XAS) techniques. The latter technique was operated through high-energy synchrotron radiation which can give in-depth details on metal coordination environment. Furthermore, fluorescence imaging experiments have been carried out in living Hek293 cells to prove the biocompatibility and capability in mapping  $\text{Fe}^{3+}$  concentrations in the cellular matrix of the synthesized compound.



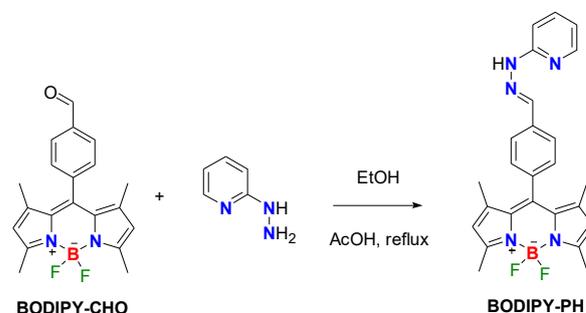
**Figure 1.** Structures of pyridylhydrazone-based fluorescent sensors in the literature and the BODIPY-PH in this work.

## 2. Experimental Section

### 2.1. Materials and Instruments

The aldehyde-functionalized BODIPY (BODIPY-CHO) shown in Scheme 1 was synthesized according to the reported procedures [28,29]. 2-Hydrazinopyridine, Dess–Martin periodinane, 2,3-dichloro-5,6-dicyanoquinone (DDQ), 2,4-dimethylpyrrole, and glacial acetic acid were purchased from Tokyo Chemical Industry (TCI), Tokyo, Japan; boron trifluoride diethyl etherate purified by redistillation, and 4-(hydroxymethyl)benzaldehyde

dimethyl acetal were purchased from Sigma-Aldrich, St. Louis, MO, USA; ethanol (ACS grade) was purchased from Honeywell, Charlotte, NC, USA. Standard buffer solutions pH 1 to 12 were purchased from Merck, Darmstadt, Germany (glycine/NaCl/HCl for pH 1, Citric acid/NaOH/HCl for pH 2–6, Na<sub>2</sub>HPO<sub>4</sub>/KH<sub>2</sub>PO<sub>4</sub> for pH 7, boric acid/KCl/NaOH for pH 8–11, Na<sub>3</sub>PO<sub>4</sub>/NaOH for pH 12). All chemicals were used without further purification. Electrospray mass spectra were obtained from a Bruker micrOTOF spectrometer (Bruker, Billerica, MA, USA). NMR spectra were recorded on a Bruker Avance 500 MHz NMR spectrometer (Bruker, Billerica, MA, USA) at ambient temperature. Chemical shifts are given in ppm and are referenced to residual <sup>1</sup>H and <sup>13</sup>C solvent signals. Fourier transform infrared spectroscopy (FTIR) spectra were acquired from an IRTracer-100 spectrophotometer (SHIMADZU, Kyoto, Japan). The crystallographic measurement was performed using a Bruker D8 Venture diffractometer (Bruker, Billerica, MA, USA). UV-VIS absorption and fluorescence spectra were acquired from a Cary Series UV-Vis-NIR spectrophotometer (Agilent Tech, Santa Clara, CA, USA) and a Perkin Elmer LS55 fluorescence spectrometer (PerkinElmer, Waltham, MA, USA), respectively.



**Scheme 1.** The synthesis of BODIPY-PH.

## 2.2. Synthetic Procedures

Synthesis of (E)-5,5-difluoro-1,3,7,9-tetramethyl-10-((2-(pyridin-2-yl)hydrazono)methyl)phenyl)-5H-dipyrrolo [1,2-c:2',1'-f][1-3]diazaborinin-4-ium-5-uide (BODIPY-PH). BODIPY-CHO (438 mg, 1.24 mmol) and 2-hydrazinopyridine (160 mg, 1.49 mmol) were dissolved in ethanol (10 mL). A few drops of glacial acetic acid were then added to the mixture. The reaction mixture was refluxed at 90 °C for 4 h. After solvent evaporation, the red solid was obtained. The solid was then washed with DI water (2 × 30 mL), filtered, and dried under reduced pressure. The BODIPY-PH was obtained in 58.2% yield (75.5 mg). <sup>1</sup>H NMR (500 MHz, DMSO) δ 10.98 (s, 1H), 8.08 (d, *J* = 4.2 Hz, 1H), 8.06 (s, 1H), 7.80 (d, *J* = 8.0 Hz, 2H), 7.61 (t, *J* = 7.6 Hz, 1H), 7.35 (d, *J* = 8.0 Hz, 2H), 7.25 (d, *J* = 8.4 Hz, 1H), 6.74 (t, 1H), 6.15 (s, 2H), 2.42 (s, 6H), 1.37 (s, 6H). <sup>13</sup>C NMR (126 MHz, DMSO) δ 157.38, 155.38, 148.23, 143.14, 142.09, 138.40, 138.20, 136.71, 134.23, 131.08, 128.73, 127.06, 121.86, 115.67, 106.89. HRMS (ESI), *m/z* calcd for C<sub>25</sub>H<sub>24</sub>BF<sub>2</sub>N<sub>5</sub> [M + H]<sup>+</sup>, 444.2170; found, 444.2165.

## 2.3. Crystallographic Measurements

A single crystal of BODIPY-PH was grown by slow diffusion of pentane vapor into its solution in tetrahydrofuran for 5 days. The diffraction spots were collected on a Bruker D8 Venture diffractometer controlled by APEX3 software [30]. Absorption correction was carried out by the multi-scan method using SADABS [31]. Cell refinement and data reduction were carried out by SAINT [30]. The structure solution was performed on Olex2 software [32] using the intrinsic phasing method of SHELXT package [33]. The structure after structure solution process was then refined on *F*<sup>2</sup> by a full-matrix least-squares method using the SHELXL program package [34]. All non-hydrogen atoms were treated anisotropically, while the N-bound H (N4) atom was located in a difference map, and their positions were freely refined. The C-bound H atoms were placed in idealized positions (C—H = 0.95–1.00 Å depending on hybridization) and refined as riding atoms. The solvent mask command was employed in order to remove the contribution of the electron scattering

from the disordered solvents [32]. The constraint  $U_{iso}(H) = 1.2U_{eq}(\text{carrier})$  was applied in all cases. The software package used to prepare molecular graphics and materials for publication was Mercury [35].

#### 2.4. General Details for UV-Vis and Fluorescence Measurements

**Preparation of the stock solutions:** The stock solutions of BODIPY-PH compounds were prepared by dissolving 20 mg of BODIPY-PH with 50 mL of acetonitrile in standard volumetric flasks (0.9 mM). The stock solution of  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$  was prepared by dissolving 50 mg of  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$  with 25 mL water in standard volumetric flasks (5 mM). UV-Vis and fluorescence measurements were performed by taking appropriate amount of these stock solutions.

**UV-Vis absorption measurement:** The suitable amount of stock solution of BODIPY-PH was added to acetonitrile:water (1:9) solution (3 mL) in a 3.5 mL quartz cuvette (final concentration = 10  $\mu\text{M}$ ). The UV-Vis absorption spectra were recorded before and after 15 min incubation time upon addition of appropriate amounts of  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ .

**Fluorescence measurement:** The suitable amount of stock solution of BODIPY-PH was added to acetonitrile:water (1:9) solution (3 mL) in a 3.5 mL quartz cuvette (final concentration = 10  $\mu\text{M}$ ). The fluorescence spectra were recorded before and after 15 min incubation time upon addition of appropriate amounts of  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ , using the following parameters: excitation wavelength = 475 nm, excitation slit = 10 nm, and emission slit = 10 nm.

The solvent system, acetonitrile:water = 1:9  $v/v$ , was used for UV-Vis and fluorescence measurements throughout the manuscript as we aim to compare the  $\text{Fe}^{3+}$ -sensing performance of our probe with the reported BODIPY-based sensor in the literature [36] using the same solvent system.

#### 2.5. Fluorescence Quantum Yields Calculations

Fluorescence quantum yields of the BODIPY-PH and its iron complex were measured in acetonitrile-water mixture (1:9  $v/v$ ) using fluorescein in 0.1M NaOH solution as a standard ( $Q_{std} = 0.95$ ) and were calculated based on the Equation (1) [37]:

$$Q = Q_{std} \times \left( \frac{A_{sample}}{A_{std}} \right) \times \left( \frac{I_{std}}{I_{sample}} \right) \times \left( \frac{\blacksquare_{sample}}{\blacksquare_{std}} \right)^2 \quad (1)$$

where  $Q$  denotes the fluorescence quantum yields,  $A$  is the peak area of emission spectra,  $I$  stands for absorption intensities at the excitation wavelength, and  $\blacksquare$  is the solvent reflective index.

#### 2.6. Binding Stoichiometry and Binding Constant Determination

The binding stoichiometry of the complexes between BODIPY-PH and ferric ions ( $\text{Fe}^{3+}$ ) was determined by Job's plot experiments [38]. The samples for Job's plot were prepared by mixing BODIPY-PH with  $\text{Fe}^{3+}$  at different ratios of BODIPY-PH over total concentration of BODIPY-PH and  $\text{Fe}^{3+}$  ( $[\text{BODIPY-PH}]/([\text{BODIPY-PH}] + [\text{Fe}^{3+}])$ ) in acetonitrile-water solution (1:9  $v/v$ ), while keeping the overall concentration of BODIPY-PH and  $\text{Fe}^{3+}$  ( $[\text{BODIPY-PH}] + [\text{Fe}^{3+}]$ ) at 50  $\mu\text{M}$ . The emission intensities of the samples were recorded at 518 nm (excitation wavelength = 475 nm) and plotted against molar fractions of BODIPY-PH ( $[\text{BODIPY-PH}]/([\text{BODIPY-PH}] + [\text{Fe}^{3+}])$ ).

The binding constant values were determined from emission intensities of BODIPY-PH (10  $\mu\text{M}$ ) at 518 nm upon the constant addition of  $\text{Fe}^{3+}$  from 1 molar equivalent (10  $\mu\text{M}$ ) to 16 molar equivalent (160  $\mu\text{M}$ ) (10  $\mu\text{M}$  increment). The data were fitted to the modified Benesi-Hildebrand Equation (2) [39]:

$$\frac{1}{F - F_{min}} = \frac{1}{K(F_{max} - F_{min})[\text{Fe}^{3+}]} + \frac{1}{F_{max} - F_{min}} \quad (2)$$

where  $F_{min}$  denotes the emission intensities of BODIPY-based dyes in the absence of ferric ions,  $F$  is the emission intensities of the dyes at intermediate  $Fe^{3+}$  concentrations,  $F_{max}$  stands for the emission intensities of the dyes at the complete interaction with ferric ions, and  $K$  is the binding constant values.

### 2.7. Theoretical Calculations

To understand the geometry of BODIPY-PH and the formation of a complex system of BODIPY-PH and  $Fe^{3+}$ , BH and HLYP function [40] was performed in implicit aqueous solution using conductor-like polarized continuum model (C-PCM) framework [41,42]. The basis set of 6-311G\* was applied for all atoms, except for the Fe atom which the LANL2DZ effective core potential was applied to describe the Fe core electrons to reduce computational cost. The electronic properties based on optimized structure of this system were then explored. The frontier molecular orbitals (FMOs) and HOMO-LUMO energy diagrams of BODIPY-PH and its  $Fe^{3+}$  complex were also calculated. All calculations were performed by the Gaussian 16 program package [43].

### 2.8. X-ray Absorption Spectroscopy (XAS)

To investigate the binding and complexation between the BODIPY-PH ligand and the  $Fe^{3+}$  ions, the mixture in acetonitrile-water (1:9 *v/v*) was characterized by employing X-ray absorption spectroscopy (XAS) in both X-ray absorption near edge structure (XANES) and extended X-ray absorption fine structure (EXAFS) regions. The XANES and EXAFS at Fe K-edge were collected at Beamline 1.1W: Multiple X-ray Techniques, Synchrotron Light Research Institute (SLRI), Thailand. The XAS measurements were simultaneously conducted in both transmission and fluorescence modes at ambient atmosphere. The Fe foil was simultaneously measured together with the sample solutions and was further used as a standard reference for an in-line calibration of the energy shift during the synchrotron-operating period. The XAS data of the 100 mM solutions of  $Fe^{2+}$  and  $Fe^{3+}$  in acetonitrile-water (1:9 *v/v*) were also collected for comparing with the mixture of BODIPY-PH ligand and the  $Fe^{3+}$  ions. The XANES data were proceeded by using ATHENA software and were analyzed to examine the oxidation state and chemical speciation of the mixture between BODIPY-PH ligand and  $Fe^{3+}$  ions with respect to the standards. The EXAFS fitting was carried out by ARTEMIS software using the DFT-calculated structural model to clarify the structural information of the binding and complexation between BODIPY-PH ligand and  $Fe^{3+}$  ions.

### 2.9. Cell Viability and Imaging Experiments

Cell viability assay: Human embryonic kidney 293 (HEK-293) cells were cultured on a 75 cm<sup>2</sup> culture flask in Dulbecco's modified Eagle medium/high glucose (DMEM/HIGH GLUCOSE, GE healthcare Life Sciences HyClone Laboratories) supplemented with 10% fetal bovine serum (FBS, Gibco) and 1% penicillin streptomycin solution 100X (CORNING). All cells were cultured in a humidified incubator at 37 °C with 5% CO<sub>2</sub>. For cell viability assay, HEK-293 cells were seeded into 96-well cell culture plates at density  $7 \times 10^3$  cells/well and incubated for 24 h at 37 °C under humidified air containing 5% CO<sub>2</sub>. After that, cell culture media were replaced with BODIPY-PH in DMEM at various concentrations (0, 5, 10, 20, 30, 40, and 50 μM). After 24 h incubation, the cell viabilities were measured using MTT assay. The cells were washed with PBS (three times) before being treated with methylthiazolyldiphenyl-tetrazolium bromide (20 μL, 0.5 mg/mL, Sigma-Aldrich) for 3 h. Then, culture media were replaced with DMSO and the cell viabilities were determined through UV-vis absorption of the resulting formazan at wavelength 560 nm using microplate reader (BMG Labtech/SPECTROstar Nano).

Cell imaging experiments: For confocal imaging, HEK-293 cells were seeded at a density of  $7 \times 10^3$  cells per well in 8-well chambered coverglass with non-removable wells (Nunc Lab-Tek II) and incubated for 24 h at 37 °C under 5% CO<sub>2</sub>. After that, cell culture media were removed and the cells were pre-treated with a solution of  $Fe(NO_3)_3 \cdot 9H_2O$

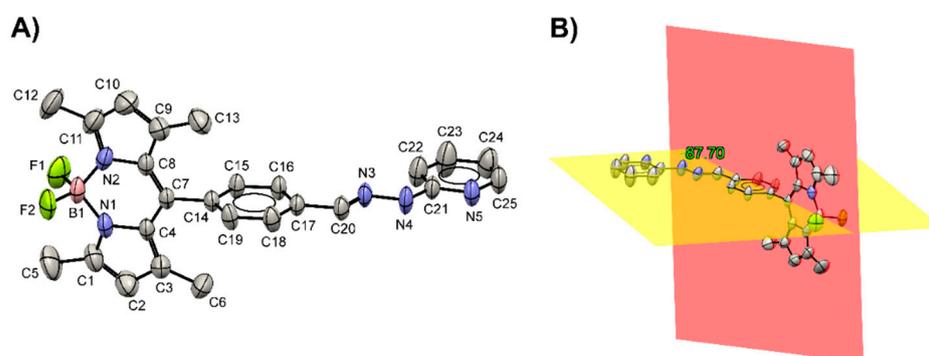
in DMEM with 5% FBS at final concentrations of 10 or 20  $\mu\text{M}$  for 30 min. Afterwards, the cells were washed three times with PBS before adding solutions of BODIPY-PH in DMEM with 5% FBS at final concentrations of 5 or 10  $\mu\text{M}$ . After 30 min incubation, the cells were washed three times with PBS to remove the non-uptake dye before being stained with 1.0  $\mu\text{M}$  Hoechst 33,342 (Thermo Fisher Scientific) for 10 min. All cells were brought to image under laser scanning confocal microscope (LSCM, Nikon A1Rsi). BODIPY-PH was excited by 488 nm laser and Hoechst 33,342 was excited by 405 nm laser. A 60 $\times$  oil immersion objective lens was used. Image analysis was performed using ImageJ software, version 1.52v (developed by Wayne Rasband, National Institutes of Health, Bethesda, MD, USA) and Java 1.8.0\_112 (64-bit) engine.

### 3. Results and Discussion

#### 3.1. Synthesis, Characterizations, and Crystal Structures of BODIPY-PH

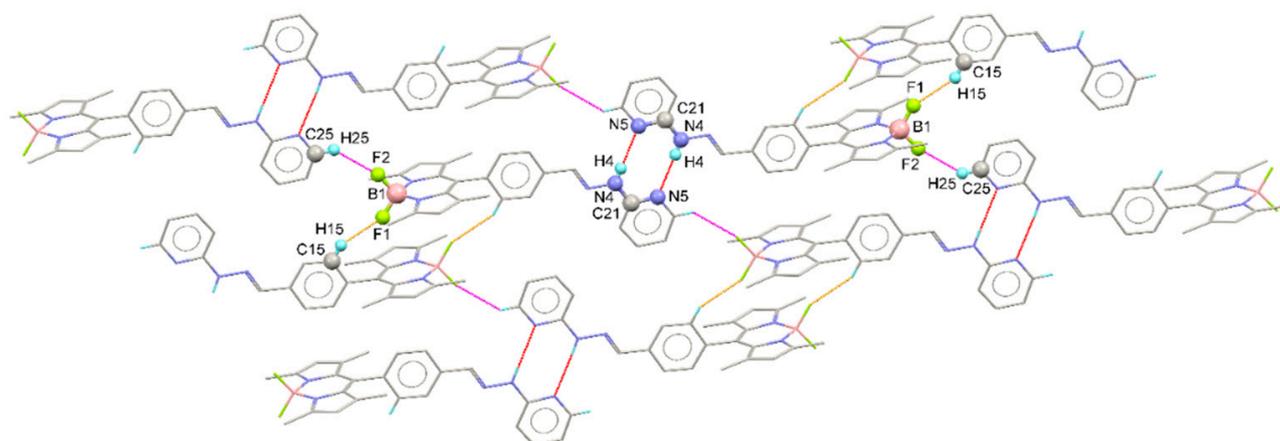
A new pyridylhydrazone-linked BODIPY probe (BODIPY-PH) has been prepared by simple condensation reaction between the known aldehyde-functionalized BODIPY (BODIPY-CHO) and 2-hydrazinopyridine as demonstrated in Scheme 1. Its  $^1\text{H}$  NMR spectra manifested the anticipated phenyl and pyridyl resonances at the aromatic region (6–8 ppm) as well as two methyl resonances of the BODIPY backbone (Figure S1, Supplementary Materials). The NH signal of this derivative was distinctly shifted downfield to around 11 ppm as seen in other pyridylhydrazone derivatives [18]. The observations of the positive molecular ions by electrospray mass spectrometry at  $m/z = 444.2165$  served as an additional confirmation for the identity of BODIPY-PH (Figure S3).

Additionally, the structure of BODIPY-PH was confirmed by single crystal X-ray crystallographic technique. The ORTEP plot showing the molecular structure of BODIPY-PH is illustrated in Figure 2A, while the corresponding crystal data and the refinement details are tabulated in Table S1. The compound BODIPY-PH crystallized in  $P\bar{1}$  space group of triclinic crystal structure with  $Z = 2$  possessing the asymmetric unit of a single BODIPY-PH molecule. From the crystal structure, the non-hydrogen atoms in the BODIPY moiety, excluding fluorine (C1–C4/C7–C11/N1–N2/B1) are arranged planarly as observed from the r.m.s. deviation of the C1–C4/C7–C11/N1–N2/B1 mean plane of 0.0377 Å. The boron atom is the most deviated atom within the BODIPY plane with the deviation of 0.074 (3) Å as it forms  $\text{sp}^3$  hybridization (other atoms possess  $\text{sp}^2$  hybridization). In addition, the non-hydrogen atoms in the benzene ring adjacent to the BODIPY moiety (C14–C19) and the benzylidenehydrazinylpyridine (C14–C25/N3–N5) group are located planarly. The corresponding r.m.s. deviations are 0.0056 Å and 0.0247 Å for the benzene and benzylidenehydrazinylpyridine groups, respectively. The mean planes of the BODIPY and benzene rings are roughly perpendicular to each other exhibiting the dihedral angle of 87.70° (9)° (Figure 2B).



**Figure 2.** (A) The ORTEP plot showing the molecular structure of BODIPY-PH together with a non-IUPAC atom-labeling scheme. Displacement ellipsoids are drawn at the 50% probability level (CCDC No. 2040756). (B) The mean planes constructed from 12 atoms (C1–C4/C7–C11/N1–N2/B1; red plane) and 6 atoms (C14–C19; yellow plane) are roughly perpendicular to each other.

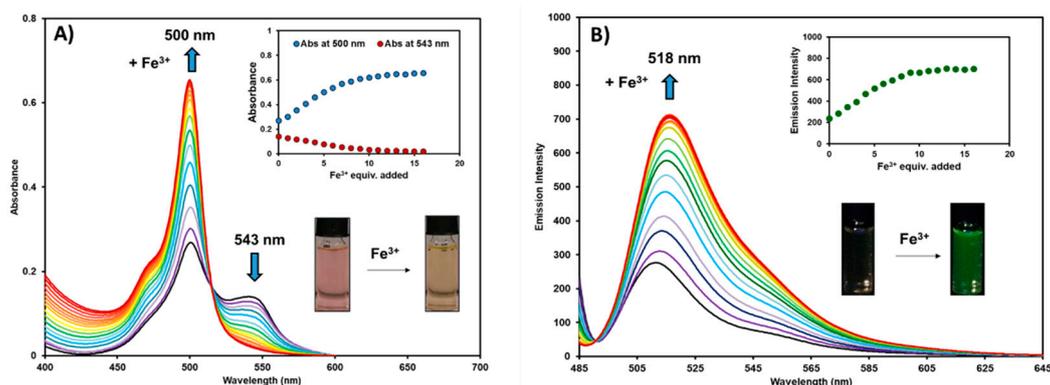
The intermolecular interactions within the crystals are displayed in Figure 3 and listed in Table S2. The N—H···N hydrogen bond is observed in the crystal structure of BODIPY-PH. The freely refined N—H bond distance is 0.89(2) Å, while the N4—H4···N5 bond angle is 177(2)°. The N—H···N hydrogen bonds (red dash lines in Figure 3) linked two molecules related by an inversion center to a  $R_2^2(8)$ -loop motif constructed from a H4—N4—C21—N5—H4—N4—C21—N5 synthon. Moreover, two different C—H···F interactions are observed. The first one (C25—H25···F2; magenta dash lines) joins the BODIPY-PH molecules related by a translational symmetry, while the second one (C15—H15···F1; orange dash lines) links the inversion-related BODIPY-PH molecules. According to the steric hindrance of the two perpendicular bulky groups, the centroids of aromatic rings are pushed off. Consequently, the distance between the ring centroids is larger than 4 Å. Hence, no  $\pi$ – $\pi$  interactions are observed in the crystal structure [44].



**Figure 3.** Views of the crystal packing which illustrate the  $R_2^2(8)$  N—H···N hydrogen bonding loop (red dash lines) as well as the C—H···F interactions (orange and magenta dash lines) in the extended structure of BODIPY-PH.

### 3.2. Photophysical Properties toward Ferric Ion Sensing

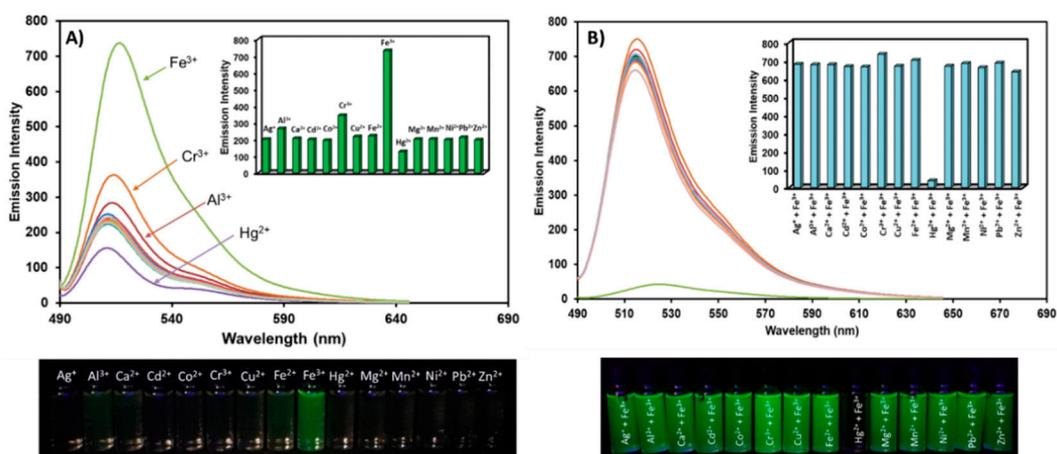
The photophysical properties of BODIPY-PH (10  $\mu$ M) were investigated toward ferric ions in an aqueous solution (acetonitrile-water = 1:9 *v/v*). Upon constant addition of 1 molar equivalent (10  $\mu$ M) to 16 molar equivalents (160  $\mu$ M) of  $\text{Fe}^{3+}$  (10  $\mu$ M increment) to the 1 molar equivalent (10  $\mu$ M) of BODIPY-PH, the absorption intensities of the BODIPY derivative at 500 nm gradually increased while the intensities at 543 nm concurrently decreased resulting in a color change from pink to orange (Figure 4A). As shown in Figure 4B, the emission spectra of BODIPY-PH exhibited around a 2.5-fold increase upon adding ferric ions and the emission maxima at the concentrations of  $\text{Fe}^{3+}$  from low to high displayed a slightly bathochromic shift from 512 nm to 518 nm. Both absorption and emission intensities remained constant at 10 equivalents of  $\text{Fe}^{3+}$  added indicating ferric ion saturation for BODIPY-PH. In addition, the fluorescence quantum yield of BODIPY-PH was enhanced from 1.3 ( $\pm 0.6$ ,  $n = 3$ ) % to 3.3 ( $\pm 0.6$ ,  $n = 3$ ) % after additions of 10 equivalents of  $\text{Fe}^{3+}$ . These results confirmed the potential of this pyridylhydrazone-tethered BODIPY as a fluorescence turn-on chemosensor for  $\text{Fe}^{3+}$  in aqueous media.



**Figure 4.** Changes in the absorption spectra of BODIPY-PH (A), and emission spectra of BODIPY-PH (B) (excitation wavelength = 475 nm) with respect to the increasing of  $\text{Fe}^{3+}$  equivalents (1–16 equiv.) in acetonitrile-water (1:9 *v/v*) solutions (Note that, the final concentration of BODIPY-PH was fixed to be 10  $\mu\text{M}$ ).

### 3.3. Sensing Properties of BODIPY-PH toward Ferric Ions

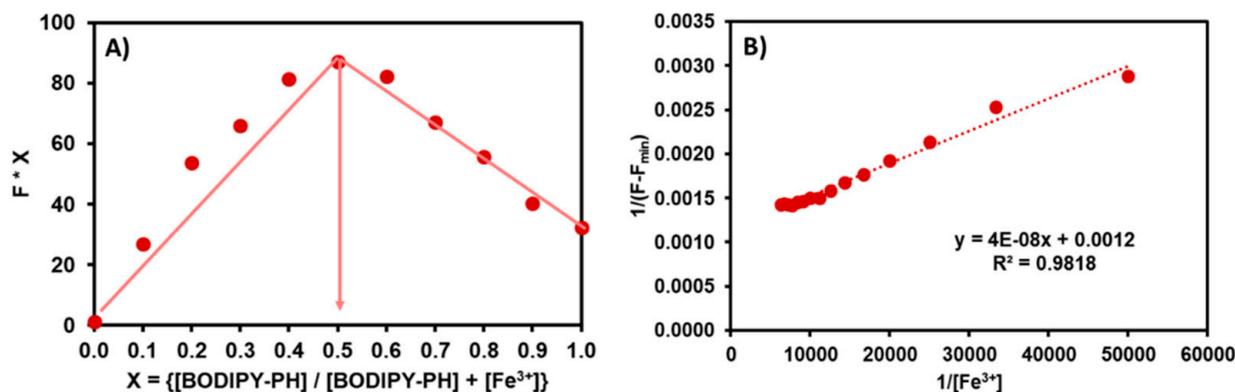
The selectivity of BODIPY-PH toward various metal cations was investigated by mixing them with 10 equivalence of  $\text{Ag}^+$ ,  $\text{Al}^{3+}$ ,  $\text{Ca}^{2+}$ ,  $\text{Cd}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Cr}^{3+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Fe}^{3+}$ ,  $\text{Hg}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Ni}^{2+}$ ,  $\text{Pb}^{2+}$ , and  $\text{Zn}^{2+}$ , respectively. As demonstrated in Figure 5A, only the presence of  $\text{Fe}^{3+}$  led to a strong green emission for BODIPY-PH suggesting its high selectivity toward ferric ions.  $\text{Al}^{3+}$  and  $\text{Cr}^{3+}$  induced a weak green emission due to their trivalent nature which has been observed in previous works [45–48], while  $\text{Hg}^{2+}$  caused a small fluorescence quenching. The interfering effect of the above-mentioned cations toward  $\text{Fe}^{3+}$  sensing was also investigated. As seen in Figure 5B, the co-existence of  $\text{Fe}^{3+}$  and  $\text{Hg}^{2+}$  could interfere with the detection of ferric ions in real applications as the fluorescent quenching was noticeably observed. Based on previous literature,  $\text{Hg}^{2+}$  could potentially bind to pyridyl unit in optical sensor systems [49–51], therefore,  $\text{Hg}^{2+}$  and  $\text{Fe}^{3+}$  could competitively bind to the pyridylhydrazone unit in BODIPY-PH.



**Figure 5.** (A) The fluorescence spectra of BODIPY-PH in the presence of various metal cations (10 equiv.) and the images BODIPY-PH (10  $\mu\text{M}$ ) upon addition of those metal ions. The inset displays bar charts of fluorescent intensities at 518 nm of BODIPY-PH after additions of metal ions. (B) The fluorescence spectra of BODIPY-PH in the presence of  $\text{Fe}^{3+}$  co-existing with other metal ions and the images BODIPY-PH (10  $\mu\text{M}$ ) in the presence of  $\text{Fe}^{3+}$  co-presenting with those metal ions. The inset shows bar charts of fluorescent intensities at 518 nm of BODIPY-PH in the existence of  $\text{Fe}^{3+}$  co-presenting with other metal ions.

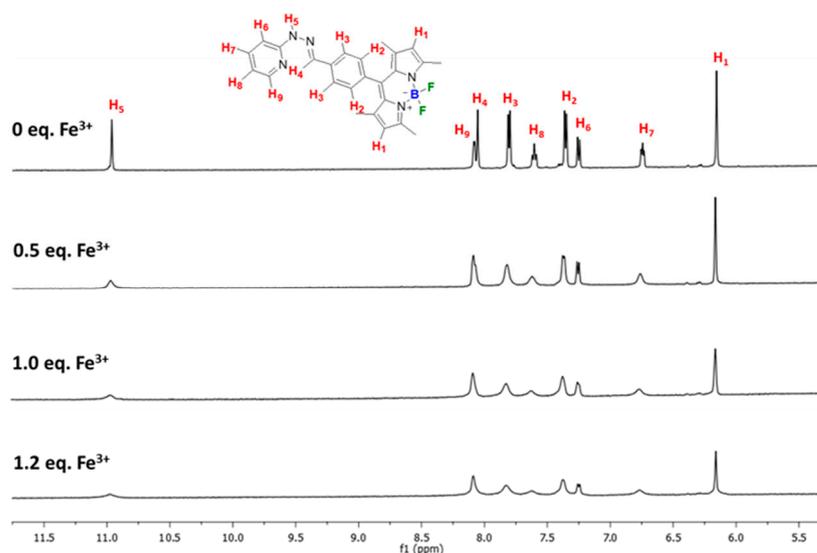
The binding stoichiometry of BODIPY-PH- $\text{Fe}^{3+}$  complex was proved by Job's plot analysis. As displayed in Figure 6A, the emission intensities at 518 nm were plotted against molar fractions of BODIPY-PH ( $[\text{BODIPY-PH}]/([\text{BODIPY-PH}] + [\text{Fe}^{3+}])$ ) at the

constant total concentrations of BODIPY-PH and  $\text{Fe}^{3+}$  ( $[\text{BODIPY-PH}] + [\text{Fe}^{3+}]$ ) of  $50 \mu\text{M}$ . The maximum emission intensity of this BODIPY derivative was observed at the mole fractions of 0.5, suggesting 1:1 binding stoichiometry of BODIPY-PH- $\text{Fe}^{3+}$  complex [38]. Next, the association constant ( $K$ ) for this metal-ligand formation was determined using Benesi-Hildebrand Equation (2). The linearly fitted plot between  $1/(F-F_{\min})$  and  $1/[\text{Fe}^{3+}]$  further confirmed the 1:1 binding ratio between BODIPY-PH and  $\text{Fe}^{3+}$  ions (Figure 6B). Derived from the slope of this plot, the  $K$  value was calculated as  $5.35 \times 10^4 \text{ M}^{-1}$  which is comparable with those of BODIPY-based sensor in the literature (Table 1).



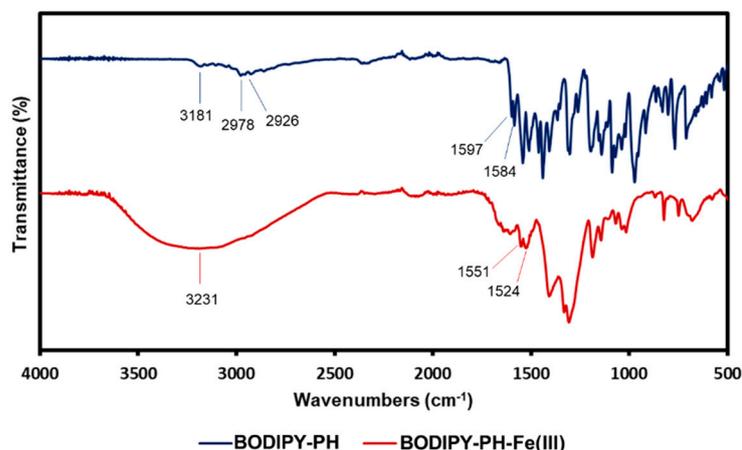
**Figure 6.** (A) Job's plot for BODIPY-PH in acetonitrile-water (1:9 *v/v*). (B) Benesi-Hildebrand binding plot generated from the analysis of fluorescence enhancements of BODIPY-PH upon addition of  $\text{Fe}^{3+}$ , and fitted in accordance with Equation (2).

Next, the  $\text{Fe}^{3+}$  binding at the pyridylhydrazone (PH) unit was investigated by  $^1\text{H}$  NMR spectroscopy. As  $\text{Fe}^{3+}$  is paramagnetic ions that can affect relaxation times of proton-NMR spectra, the NMR resonance frequency of protons close to  $\text{Fe}^{3+}$  binding unit could be altered upon binding [52]. As displayed in Figure 7, the hydrazone proton signal ( $\text{H}_4$ ) at 8.06 ppm completely disappeared and the amine proton resonance ( $\text{H}_5$ ) at 10.98 ppm almost completely disappeared upon the addition of  $\text{Fe}^{3+}$ . The rest of the proton peaks of pyridylhydrazone moiety became broader indicating that  $\text{Fe}^{3+}$  bound to the PH unit which is consistent with the previous findings [25,53,54]. In addition, the detections of molecular ions at  $m/z = 639.0923$   $\{[(\text{BODIPY-PH-H})\text{Fe}(\text{NO}_3)(\text{DMSO})]^+\}$  and  $684.0762$   $\{[(\text{BODIPY-PH-H})\text{Fe}(\text{NO}_3)(\text{DMSO})(\text{EtOH})]^+\}$  from electrospray ionization mass spectra (ESI-MS) confirmed the existence of iron-BODIPY complex (Figure S4).



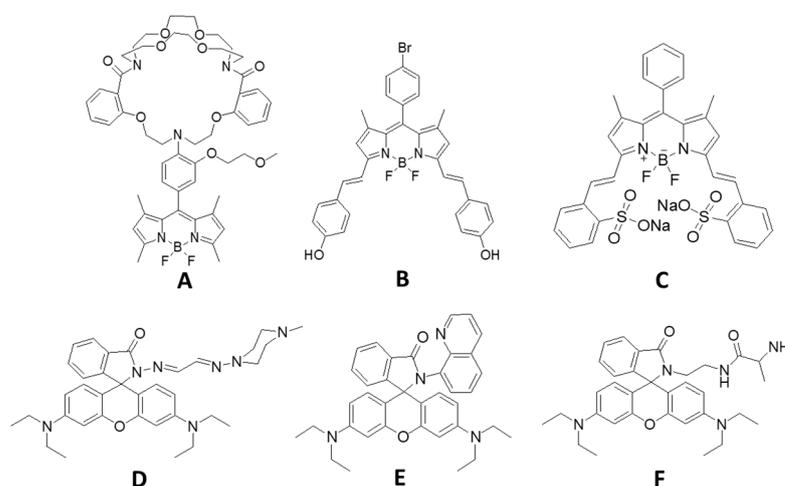
**Figure 7.**  $^1\text{H}$  NMR (500 MHz) spectra of BODIPY-PH in the presence of  $\text{Fe}^{3+}$  in  $\text{DMSO-d}_6$ .

FT-IR spectroscopy was also carried out to confirm the complex formation between  $\text{Fe}^{3+}$  and BODIPY-PH (Figure 8). For the BODIPY-PH spectrum, the characteristic peaks at 3181, 2978, 2926, 1597, and 1584  $\text{cm}^{-1}$  were assigned to N—H,  $\text{C}_{\text{methyl}}\text{—H}$ ,  $\text{C}_{\text{methyl}}\text{—H}$ , C=N, and C=N stretching vibrations, respectively. It was observed that the wavenumbers of the N—H and C=N peaks were found at the lower frequencies than those values reported in the literature [55,56], due to the larger conjugation length. After complex formation, the broad and intense O—H peak at 3231  $\text{cm}^{-1}$  regarding the  $\text{H}_2\text{O}$  ligands that coordinated to the  $\text{Fe}^{3+}$  was detected. Moreover, the C=N peaks of the BODIPY-PH- $\text{Fe}^{3+}$  complex were enlarged and shifted toward the lower wavenumbers (1551, 1524  $\text{cm}^{-1}$ ) compared to those of BODIPY-PH. This suggested that the Schiff base nitrogen was likely to coordinate with the  $\text{Fe}^{3+}$  ions [57] which is agreeable with the result from NMR spectroscopic technique.



**Figure 8.** Fourier transform infrared spectroscopy (FTIR) spectra of BODIPY-PH (blue line) and BODIPY-PH- $\text{Fe}^{3+}$  (red line).

In addition, the limit of detection (LOD) of BODIPY-PH was calculated from the plots of the increased fluorescence intensities and the  $\text{Fe}^{3+}$  concentrations, as demonstrated in Figure S5 by using  $3\sigma/m$  formula, where  $\sigma$  is the standard deviation of fluorescence signals of the chemosensor without analyte and  $m$  is the slopes of the plots. The LOD of BODIPY-PH was estimated as 0.58  $\mu\text{M}$ , which was slightly higher than that of other reported BODIPY-based  $\text{Fe}^{3+}$  sensors A and B and considerably higher than that of C (Table 1 and the structures of sensors show in Figure 9). However, compared with the reported rhodamine-based sensors tethered nitrogen-chelating moiety (sensor D, E, F), the LOD of BODIPY-PH was about 4 to 6-time lower, indicating superior sensitivity toward  $\text{Fe}^{3+}$ .



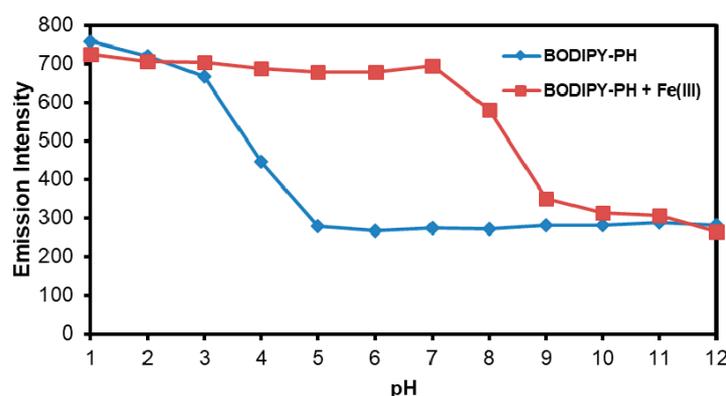
**Figure 9.** Structures of BODIPY- and rhodamine-based  $\text{Fe}^{3+}$  sensors in the literature.

**Table 1.** Comparison of BODIPY-PH with other reported BODIPY- and rhodamine-based sensors for ferric ion detection.

Sensor	Solvent System	LOD ( $\mu\text{M}$ )	Binding Constant (K) (Binding Mode: $\text{Fe}^{3+}$ : Ligand)	Sensing Mechanism	Reference
A	Acetonitrile-water (1:9 v/v)	0.13	$1.0 \times 10^4 \text{ M}^{-1}$ (1:1)	Turn-on	[36]
B	Ethanol	0.14	$3.1 \times 10^4 \text{ M}^{-1}$ (1:1)	Turn-on	[48]
C	water	0.014	$K_1 = 2.1 \times 10^4 \text{ M}^{-1}$ $K_2 = 8.8 \times 10^4 \text{ M}^{-1}$ (stepwise 1:1 and 2:1)	Turn-off	[58]
D	Methanol-water (1:9 v/v)	2.2	$6.9 \times 10^3 \text{ M}^{-1}$ (1:1)	Turn-on	[59]
E	Acetonitrile-water (1:1 v/v)	3.5	N.R.	Turn-on	[60]
F	Methanol-water (4:6 v/v)	3	N.R.	Turn-on	[61]
BODIPY-PH	Acetonitrile-water (1:9 v/v)	0.58	$5.4 \times 10^4 \text{ M}^{-1}$ (1:1)	Turn-on	This work

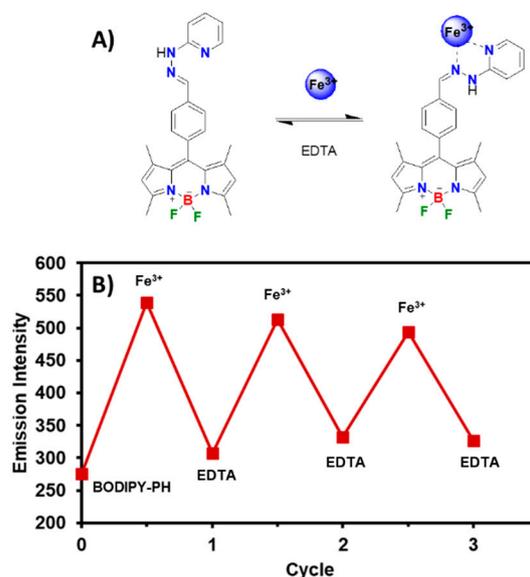
N.R. = Not Report.

The effect of pH toward photophysical properties of BODIPY-PH (10  $\mu\text{M}$ ) with and without the addition of  $\text{Fe}^{3+}$  was explored in acetonitrile-standard buffer solution (1:9 v/v) in the pH range of 1–12. As seen in Figure 10, BODIPY-PH exhibited low fluorescence signals at a pH range of 5–12, while showed the increased emission intensity with the decreased pH from 5 to 1. Upon addition of  $\text{Fe}^{3+}$  (10 molar equivalents or 100  $\mu\text{M}$ ) to the solution of BODIPY-PH (10  $\mu\text{M}$ ) in various pH, the highest fluorescent signal changes were observed in the pH range from 5 to 7 which could be a suitable pH range for  $\text{Fe}^{3+}$  detection in real water samples. The emission intensities of BODIPY-PH in the presence of  $\text{Fe}^{3+}$  decreased at pH > 7 indicating poor stability of BODIPY-PH- $\text{Fe}^{3+}$  complex at high pH values.

**Figure 10.** Fluorescence intensity variation at maximum emission of 518 nm of BODIPY-PH (10  $\mu\text{M}$ ) at media alternation of pH from 1 to 12.

Then, the reversibility and reusability of BODIPY-PH as proposed in Figure 11A was systematically investigated by the fluorescence titration of the dye with consecutive additions of  $\text{Fe}^{3+}$  (100  $\mu\text{M}$ ) and the common metal ion sequester, ethylenediaminetetraacetic acid (EDTA) (100  $\mu\text{M}$ ). As displayed in Figure 11B, the fluorescence signals reached the maximum point when BODIPY-PH (10  $\mu\text{M}$ ) was saturated by the addition of 10 equiv. of  $\text{Fe}^{3+}$  (100  $\mu\text{M}$ ) and quenched upon the further addition of EDTA (100  $\mu\text{M}$ ). Nevertheless,

when another 100  $\mu\text{M}$  of  $\text{Fe}^{3+}$  was added to the resulting mixture, the fluorescent intensity was almost recovered. This off-on-off cycle could be effectively repeated up to three times confirming the practical regeneration of the BODIPY-PH after ferric ion detection.



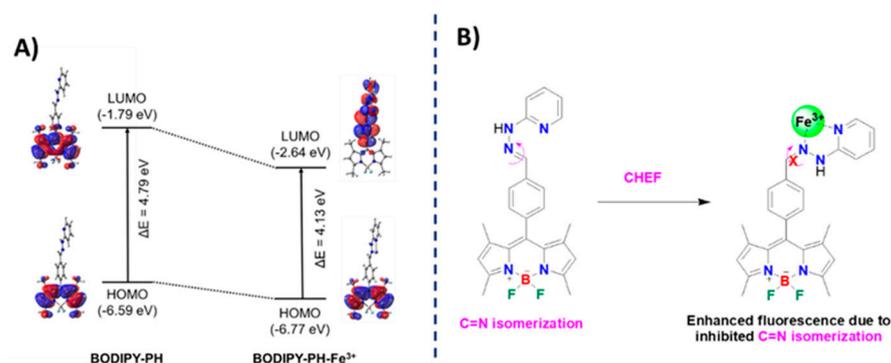
**Figure 11.** (A) The proposed binding mode between BODIPY-PH and  $\text{Fe}^{3+}$  and its reversible reaction through EDTA complexation. (B) Reversible changes in fluorescence intensity of BODIPY-PH (10  $\mu\text{M}$ ) at 518 nm upon alternate addition of  $\text{Fe}^{3+}$  (100  $\mu\text{M}$ ) and EDTA (100  $\mu\text{M}$ ).

### 3.4. Theoretical Calculations

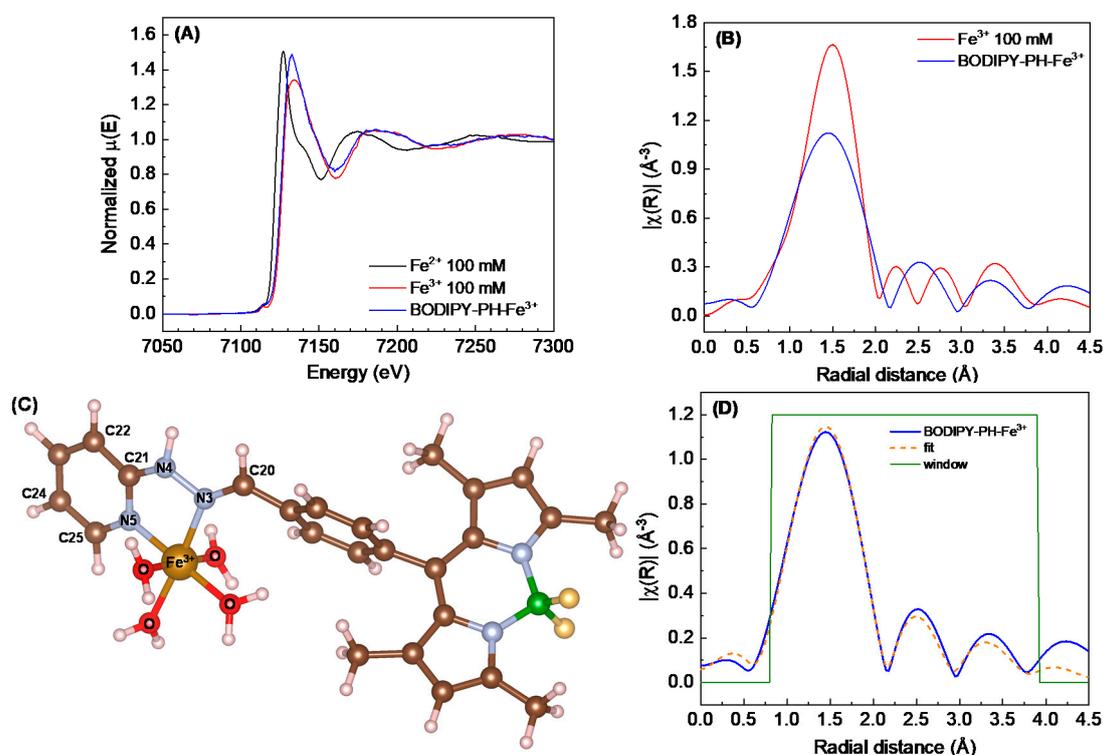
To identify the sensing mechanism of the BODIPY-PH toward  $\text{Fe}^{3+}$  complexation, we performed the DFT calculations using BHandHLYP/6-311G\* level of theory in the implicit water solvent and the LANL2DZ basis set for  $\text{Fe}^{3+}$ . The highest occupied molecular orbital (HOMO) and the lowest unoccupied molecular orbital (LUMO) of BODIPY-PH and BODIPY-PH- $\text{Fe}^{3+}$  complex with the isovalue of 0.02 and their corresponding energy gaps are illustrated in Figure 12A. The density distributions of HOMO- and LUMO-electrons on the BODIPY-PH are mainly located on the BODIPY backbone as also observed in other BODIPY derivatives according to the literature [62–64]. In addition, there are no  $\pi$ -electrons covered on the pyridylhydrazone moiety which indicates the separation of the  $\pi$ -conjugated system on the BODIPY core from the one on pyridylhydrazone unit. On the other hand, the electron density distributions of BODIPY-PH- $\text{Fe}^{3+}$  complex are shifted from the BODIPY core (in the HOMO) to the phenylene moiety and the  $\text{Fe}^{3+}$  binding site (in the LUMO). This result indicates the increased  $\pi$ -conjugation in the BODIPY-PH- $\text{Fe}^{3+}$  complex and the inhibition of C=N isomerization leading to the fluorescence enhancement (Figure 12B). These theoretical results confirm the chelation-enhanced fluorescence (CHEF) effect in BODIPY-PH upon  $\text{Fe}^{3+}$  binding as shown in Figure 12B.

### 3.5. Investigation of the Binding and Complexation between BODIPY-PH Ligand and $\text{Fe}^{3+}$ Ions by X-ray Absorption Spectroscopy (XAS)

XAS spectra at Fe K-edge were collected in both XANES and EXAFS regions to reveal the structural information upon binding of  $\text{Fe}^{3+}$  ions to the BODIPY-PH ligand (named as BODIPY-PH- $\text{Fe}^{3+}$ ). As shown in Figure 13A, the XANES spectrum of BODIPY-PH- $\text{Fe}^{3+}$  illustrates the rather similar absorption edge and fingerprint feature with that of the standard solutions of  $\text{Fe}^{3+}$  in acetonitrile-water (1:9 v/v) solution. In other words, the oxidation state of Fe in the BODIPY-PH- $\text{Fe}^{3+}$  remains +3 and the complexation of BODIPY-PH- $\text{Fe}^{3+}$  may resemble the octahedral geometry as observed in the  $\text{Fe}^{3+}$  aqua complex.



**Figure 12.** (A) Frontier molecular orbitals (HOMO and LUMO) of BODIPY-PH and BODIPY-PH-Fe<sup>3+</sup> compounds and their corresponding energy gaps ( $\Delta E$  in eV) computed by BHandHLYP/6-311G\* level of theory. The HOMO and LUMO orbitals are plotted with isovalue of 0.02. (B) Proposed sensing mechanism of BODIPY-PH upon Fe<sup>3+</sup> binding.



**Figure 13.** (A) X-ray absorption near edge structure (XANES) spectra probed at Fe K-edge of the mixture between BODIPY-PH ligand and Fe<sup>3+</sup> (BODIPY-PH-Fe<sup>3+</sup>) compared to the 100 mM standard solutions of Fe<sup>2+</sup> and Fe<sup>3+</sup> in acetonitrile-water (1:9 v/v). (B) Comparison of the extended X-ray absorption fine structure (EXAFS) data of the BODIPY-PH-Fe<sup>3+</sup> with respect to the one of 100 mM standard solutions of Fe<sup>3+</sup>. (C) The DFT-calculated structural model of BODIPY-PH-Fe<sup>3+</sup> used for EXAFS fitting indicating the neighboring atoms around Fe<sup>3+</sup>. (D) EXAFS data fitting by employing ARTEMIS software of the BODIPY-PH-Fe<sup>3+</sup>.

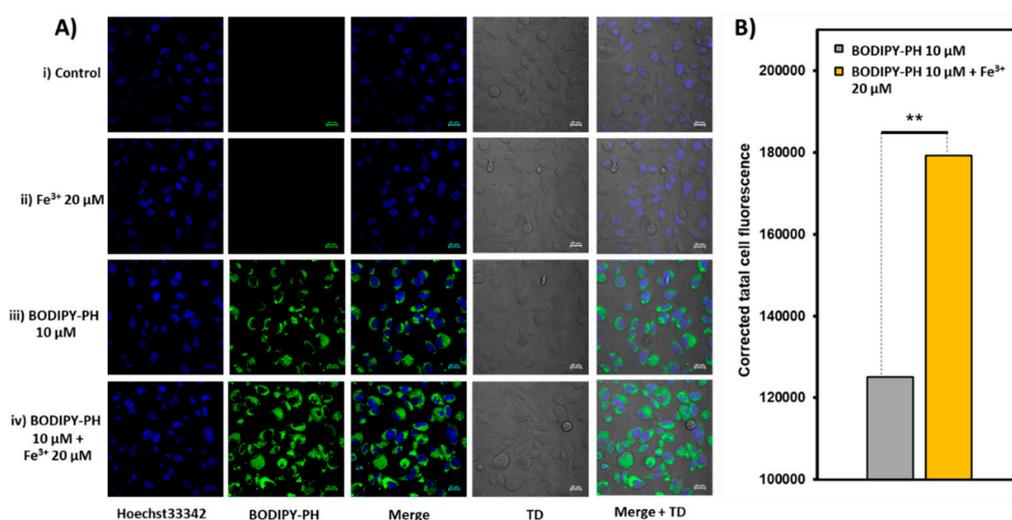
Closer inspection into the EXAFS data plotted in R-space (Figure 13B), the amplitude at the first shell of BODIPY-PH-Fe<sup>3+</sup> is relatively lower and the first-shell feature becomes broader than those of the standard Fe<sup>3+</sup> solution. This observation suggests the deviation from the symmetric octahedral coordination sphere in the Fe<sup>3+</sup> aqua complex by means of the binding of Fe<sup>3+</sup> with the BODIPY-PH ligand. Moreover, the radial distance distribution at the second and the third shell around the Fe<sup>3+</sup> in the BODIPY-PH-Fe<sup>3+</sup> is different from the feature of the Fe<sup>3+</sup> aqua complex. To identify the structural complexation of

BODIPY-PH-Fe<sup>3+</sup>, the structural optimization of the octahedral Fe<sup>3+</sup> complex coordinated with four H<sub>2</sub>O molecules and two nitrogen (N3 and N5) at the pyridylhydrazone unit of BODIPY-PH ligand (Figure 13C) was performed by DFT calculation. The EXAFS data fitting was operated by using this DFT-calculated structure as a model. The well-matching of the structural model with the experimental data is observed (Figure 13D) showing the corresponding EXAFS-fitting parameters in Table S3. Noteworthy, the EXAFS data fitting reveals the complexation of Fe<sup>3+</sup> ions with BODIPY-PH ligand in 1:1 ratio and forms the octahedral coordination geometry with the additional four H<sub>2</sub>O molecules.

### 3.6. Cell Imaging Study

To prove the capability of BODIPY-PH for Fe<sup>3+</sup> analysis in a cellular matrix, the cell imaging experiment was investigated in the human embryonic kidney (HEK293) cell line acting as a cellular model. As an excessive uptake of iron can cause acute kidney injury in human [65], this work could provide an alternative fluorescence tool to study oxidative damage of iron toward kidney cells. Based on cell viability assay (Figure S6), BODIPY-PH is non-toxic to HEK293 cells at concentrations ranging from 5 to 50 μM. So, this BODIPY concentration range was next employed for the cell culture experiments with and without Fe<sup>3+</sup> ions.

Live cell-imaging experiments incubated with 5 μM and 10 μM of BODIPY-PH were done in parallel to investigating the effect of dye concentrations toward Fe<sup>3+</sup> sensing and the confocal laser scanning microscopy (CLSM) (images of HEK293 cells of both experimental sets are demonstrated in Figure S7A and Figure 14A), respectively. After incubating with ferric ions for 30 min at the twice concentrations of the dye, the HEK293 cells exhibited increased fluorescence signals which was consistent with the corrected total cell fluorescence (CTCF) bar chart (Figure S7B) and Figure 14B). Both cell-imaging experimental sets manifested ~1.4-fold fluorescence enhancement (Figure S8) which is comparable to the fluorescence enhancement of compound A (Figure 9) in HCT-116 cells upon addition of Fe<sup>3+</sup> [36]. Additionally, the TD images confirmed a good viability of the cells throughout the experiments. These results strongly suggested the capability of BODIPY-PH for Fe<sup>3+</sup> analysis in live cells without any observed cell damages.



**Figure 14.** (A) Confocal laser scanning microscopy (CLSM) images of HEK-293 cells. Row i, control HEK-293 cells with no treatment. Row ii, HEK-293 cells incubated with Fe<sup>3+</sup> (20 μM) for 30 min. Row iii, HEK-293 cells incubated with BODIPY-PH (10 μM) for 30 min. Row iv, HEK-293 cells incubated with Fe<sup>3+</sup> (20 μM) for 30 min and subsequently incubated with BODIPY-PH (10 μM) for another 30 min (For BODIPY-PH: λ<sub>ex</sub> = 488 nm, λ<sub>em</sub> = 509 nm, scale bar = 20 μm, TD = transmitted light differential interference contrast image). All images were collected at the same microscopy settings. (B) Bar chart showing corrected total cell fluorescence (CTCF) of HEK-293 cells incubated with BODIPY-PH (10 μM) before and after the addition of Fe<sup>3+</sup> (20 μM). Statistical analysis is based on student's two tailed t-test (\* *p* < 0.05, \*\* *p* < 0.01, \*\*\* *p* < 0.001).

#### 4. Conclusions

In summary, we have synthesized a new BODIPY derivative (BODIPY-PH) containing pyridylhydrazone moiety acting as a  $\text{Fe}^{3+}$  ion-binding site. This compound displayed fluorescent enhancement response toward ferric ion binding in acetonitrile-water (1:9 *v/v*) solution due to the effect from the chelation-enhanced fluorescence (CHEF) process. This ligand exhibited high selectivity for  $\text{Fe}^{3+}$  ions over various metal ions, however, its fluorescent sensing ability toward  $\text{Fe}^{3+}$  could be interfered by the co-presence of  $\text{Hg}^{2+}$ . The Job's plot and the Benesi-Hildebrand analysis indicated a 1:1 binding mode between  $\text{Fe}^{3+}$  and BODIPY-PH with a binding constant of  $5.4 \times 10^4 \text{ M}^{-1}$ . The X-ray absorption spectroscopy (XAS) further confirmed the  $\text{Fe}^{3+}$  binding by BODIPY-PH ligand and suggested an asymmetric octahedral geometry of the resulted BODIPY-PH- $\text{Fe}^{3+}$  complex. The low limit of detection of 0.58  $\mu\text{M}$ , the  $\text{Fe}^{3+}$ -sensing capability at neutral pH, and the reversible nature toward  $\text{Fe}^{3+}$  binding suggested the potential of BODIPY-PH for  $\text{Fe}^{3+}$  detection in real water samples. Finally, the fluorescence imaging results in living HEK293 cells indicated that the BODIPY-PH possesses excellent cell permeability, good biocompatibility, and high stability under physiological conditions and can be applicable for detecting  $\text{Fe}^{3+}$  in living cells.

**Supplementary Materials:** The following are available online at <https://www.mdpi.com/article/10.3390/chemosensors9070165/s1>, Figure S1:  $^1\text{H}$  NMR spectra of BODIPY-PH, Figure S2:  $^{13}\text{C}$  NMR spectra of BODIPY-PH, Figure S3: Mass spectra of BODIPY-PH, Figure S4: Mass spectra of BODIPY-PH- $\text{Fe}^{3+}$  complexes, Figure S5: Standard calibration curve demonstrating the fluorescence intensity increments at 518 nm ( $n = 3$ ) as a function of  $\text{Fe}^{3+}$  concentrations incubated with BODIPY-PH (10  $\mu\text{M}$ ) in acetonitrile-water (1:9 *v/v*), Figure S6: Percent cell viability estimated by MTT proliferation test versus incubation concentrations of BODIPY-PH. HEK-293 cells were cultured in the presence of BODIPY-PH (0–50  $\mu\text{M}$ ) at 37 °C for 24 h, Figure S7: (A) Confocal laser scanning microscopy (CLSM) images of HEK-293 cells. Row i, control HEK-293 cells with no treatment. Row ii, HEK-293 cells incubated with  $\text{Fe}^{3+}$  (10  $\mu\text{M}$ ) for 30 min. Row iii, HEK-293 cells incubated with BODIPY-PH (5  $\mu\text{M}$ ) for 30 min. Row iv, HEK-293 cells incubated with  $\text{Fe}^{3+}$  (10  $\mu\text{M}$ ) for 30 min and subsequently incubated with BODIPY-PH (5  $\mu\text{M}$ ) for another 30 min (For BODIPY-PH:  $\lambda_{\text{ex}} = 488 \text{ nm}$ ,  $\lambda_{\text{em}} = 509 \text{ nm}$ , scale bar = 20  $\mu\text{m}$ , TD = transmitted light differential interference contrast image). All images were collected at the same microscopy settings. (B) Bar chart showing corrected total cell fluorescence (CTCF) of HEK-293 cells incubated with BODIPY-PH (5  $\mu\text{M}$ ) before and after the addition of  $\text{Fe}^{3+}$  (10  $\mu\text{M}$ ). Statistical analysis is based on student's two tailed t-test (\*  $p < 0.05$ , \*\*  $p < 0.01$ , \*\*\*  $p < 0.001$ ), Figure S8: Bar chart showing the relative corrected total cell fluorescence (CTCF) of HEK-293 cells incubated with BODIPY-PH (5 and 10  $\mu\text{M}$ ) comparing between before and after the addition of  $\text{Fe}^{3+}$  (10 and 20  $\mu\text{M}$ ), Table S1: Crystal data and structure refinement for BODIPY-PH, Table S2: Geometric parameters and symmetry operations for intermolecular H-bond interactions ( $\text{\AA}$ ,  $^\circ$ ), Table S3: EXAFS fitting data for investigating the structural information of the complexation between BODIPY-PH ligand and  $\text{Fe}^{3+}$  ion.

**Author Contributions:** J.N.: methodology, investigation, validation, and writing—original draft. C.S.: formal analysis and writing—review and editing. R.D.: investigation, validation, and writing—review and editing. A.K.: formal analysis and writing—review and editing. W.W.: formal analysis and writing—review and editing. S.W.: formal analysis and writing—review and editing. P.R.: resources and supervision. K.C.: conceptualization, methodology, resources, supervision, project administration, funding acquisition, writing—review and editing. All authors have read and agreed to the published version of the manuscript.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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