

Review

# Recent Advancements in Two-Dimensional Layered Molybdenum and Tungsten Carbide-Based Materials for Efficient Hydrogen Evolution Reactions

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**Abstract:** Green and renewable energy is the key to overcoming energy-related challenges such as fossil-fuel depletion and the worsening of environmental habituation. Among the different clean energy sources, hydrogen is considered the most impactful energy carrier and is touted as an alternate fuel for clean energy needs. Even though noble metal catalysts such as Pt, Pd, and Au exhibit excellent hydrogen evolution reaction (HER) activity in acid media, their earth abundance and capital costs are highly debatable. Hence, developing cost-effective, earth-abundant, and conductive electrocatalysts is crucial. In particular, various two-dimensional (2D) transition metal carbides and their compounds are gradually emerging as potential alternatives to noble metal-based catalysts. Owing to their improved hydrophilicity, good conductivity, and large surface areas, these 2D materials show superior stability and excellent catalytic performances during the HER process. This review article is a compilation of the different synthetic protocols, their impact, effects of doping on molybdenum and tungsten carbides and their derivatives, and their application in the HER process. The paper is more focused on the detailed strategies for improving the HER activity, highlights the limits of molybdenum and tungsten carbide-based electrocatalysts in electro-catalytic process, and elaborates on the future advancements expected in this field.

**Keywords:** two-dimensional nanostructures; tungsten carbide; energy conversion; hydrogen evolution reaction; Mo<sub>x</sub>C electrocatalysts



**Citation:** Karuppasamy, K.; Nicholson, A.; Vikraman, D.; Choi, J.-H.; Hussain, S.; Ambika, C.; Bose, R.; Alfantazi, A.; Kim, H.-S. Recent Advancements in Two-Dimensional Layered Molybdenum and Tungsten Carbide-Based Materials for Efficient Hydrogen Evolution Reactions. *Nanomaterials* **2022**, *12*, 3884. <https://doi.org/10.3390/nano12213884>

Academic Editor: Francesc Viñes Solana

Received: 12 October 2022

Accepted: 31 October 2022

Published: 3 November 2022

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## 1. Introduction

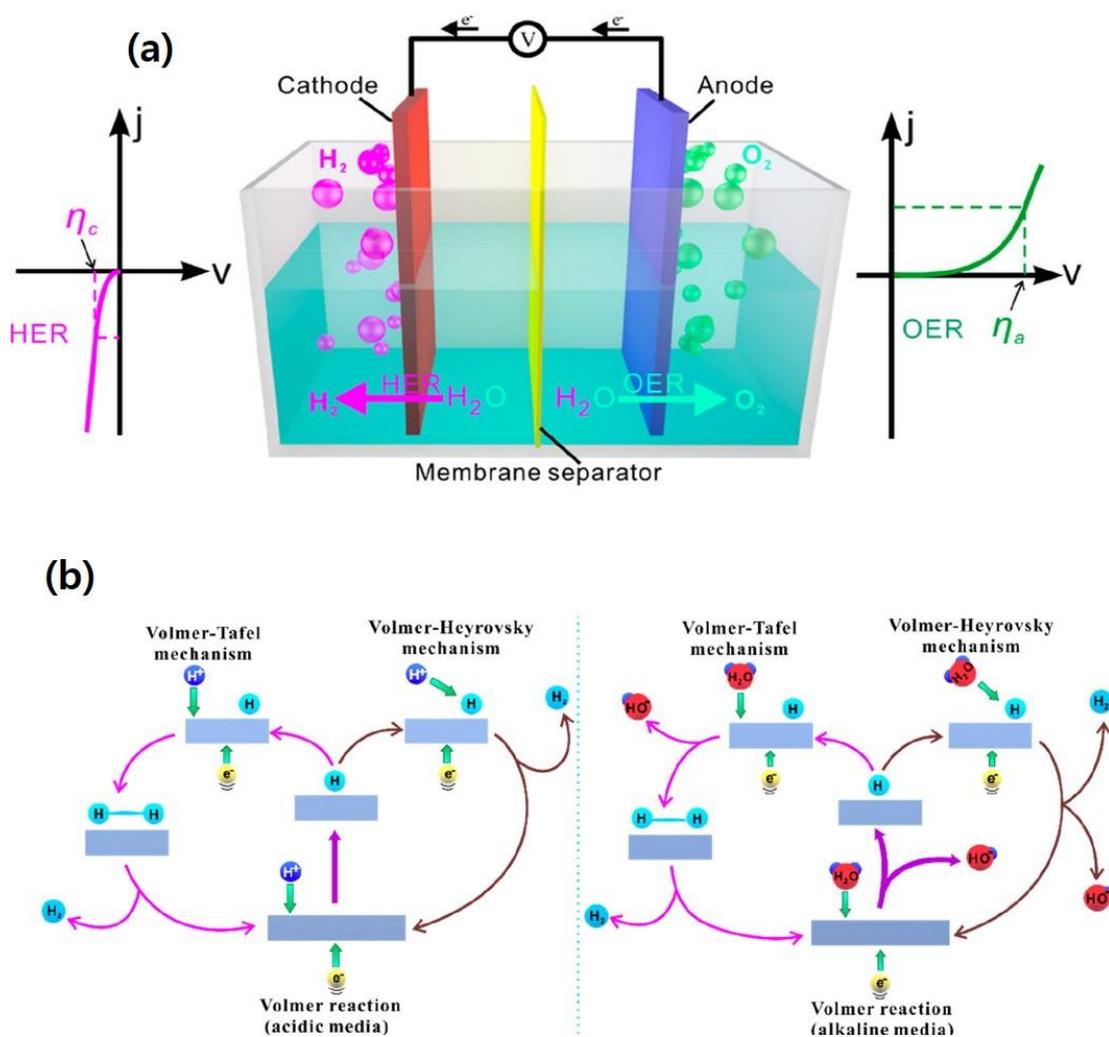
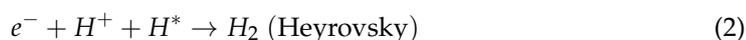
Developing clean, sustainable, and affordable energy systems with abundant, non-toxic, and inexpensive materials for the future is a global challenge [1]. Hydrogen, a light element with a high energy density, is the cleanest alternative to fossil fuels [2,3] and is broadly investigated for application in green-energy technologies. The formation of hydrogen gas through electrolysis of water is termed hydrogen evolution reaction (HER), which has gained considerable attention in recent years owing to their various catalytic applications, especially in photoelectrochemical and electrochemical water-splitting [4,5]. For HER, robust, efficient, cost-effective, and earth-abundant electrocatalysts that demonstrate superior electrochemical activity are required. Platinum is a promising electrocatalyst for HER despite being expensive and scarce. In general, electrocatalysis is defined as a kind of catalysis that results in the alteration of the rate of electrochemical reactions that occurs on the electrode surface. The electrochemical performance of HERs strongly depend on the hydrogen adsorption configurations, chemical composition, and surface functionalization [6]. Researchers around the world are determined to identify a suitable non-platinum

electrocatalyst that is non-precious, durable, earth-abundant, and stable. Earth-enriched materials, such as electrocatalysts, are not feasible for large-scale hydrogen production for industrial applications. Scientists have employed density functional theory calculations to identify superior electrocatalysts for HERs to produce hydrogen energy. An efficient electrocatalyst exhibits high electrocatalytic activity, low overvoltage, good electrochemical stability, high exchange current density, and decreased Tafel slope, and is cost effective for HERs as well [7]. The performance of HER electrode materials could be improved by optimizing the geometrical aspects (such as the specific surface area), forming core-shell nanostructures, nanoengineering electrodes, modifying the structure and chemical composition, and tuning the electronic structure of the electrode materials by doping cations and anions into the crystal lattice of the material. The intrinsic catalytic activity can be improved by doping and alloying the metal elements with non-metal elements. Over the past few decades, many efforts have been made to find highly active and long-lasting electrocatalyst materials with optimized characteristics for HERs in alkaline and acid media to realize an efficient hydrogen production [8–11]. To date, several precious and non-precious metal elements and their alloys, carbon-based materials, transition metal-based sulfides, phosphides, dichalcogenides, carbides, nitrides, and phosphosulfates have been investigated for HERs [1,12,13]. Recently, transition metal carbides (TMCs), transition metal nitrides (TMNs), and carbonitrides have been exploited for HERs because of the rich chemistry of these MXene compounds, including a high electrical conductivity, strong hydrophilicity, and surface functionalization. These catalysts have been proven highly efficient for HER. Among the known TMCs, molybdenum and tungsten carbides display substantial catalytic activities. The MXene-based compounds  $\text{Mo}_2\text{CT}_x$  and  $\text{Ti}_2\text{CT}_x$  have been investigated both computationally and experimentally as HER catalysts. The catalytic activity of  $\text{Mo}_2\text{CT}_x$  is higher than that of  $\text{Ti}_2\text{CT}_x$  [14–16]. Density functional theory calculations predicted that  $\text{Mo}_2\text{CT}_x$  can be categorized into three separate classes based on their hydrogen adsorption properties, and in all these categories, the hexagonal close-packed (hcp) site is the lowest free energy adsorption site for oxygen for both the nitride and carbide ( $M = \text{Mo}, \text{W}, \text{and Cr}$ ). For conventional HER electrolyzers, different non-noble metal catalysts have been widely studied in the past [17,18]. In majority of these studies, nickel- and cobalt-based compounds were used as the potential catalysts for HER. However, the stability of these compounds under neutral pH conditions is limited [19,20]. Tungsten carbide is a plausible earth-abundant catalyst that is cost effective and non-toxic, and shows a promising HER catalytic activity [17,19,21,22]. Further, this particular catalyst exhibits advantageous catalytic composition, activity, and stability under neutral pH conditions.

Nevertheless, the number of topical reviews on the catalytic activity of  $\text{Mo}_x\text{C}$  and  $\text{WC}/\text{W}_2\text{C}$  is limited, and thus, a detailed and systematic review of the current advancements in molybdenum- and tungsten-based electrocatalysts for HER activity is necessary at this juncture. This review is focused on the current progresses made in two-dimensional (2D) layered TMCs, especially molybdenum and tungsten carbides, for HER in acidic and alkaline solutions. The basic principles, mechanism, and various electrochemical parameters of HERs are elucidated here. State-of-the-art catalyst design and fabrication strategies for  $\text{Mo}_2\text{C}$  and tungsten carbide ( $\text{WC}, \text{W}_2\text{C}$ )-based catalysts for HER are presented. Finally, we discuss the next generation of  $\text{Mo}_2\text{C}$  and tungsten carbide ( $\text{WC}, \text{W}_2\text{C}$ )-based HER catalysts and their role in green hydrogen production and realizing a sustainable energy future.

## 2. Basic Principles of Electrocatalytic HER

Electrocatalytic water-splitting (EWS) is a widely used process that transforms electrical energy into hydrogen fuel, with natural water as the prime source and oxygen and hydrogen as the byproducts. During the EWS process, two types of half-cell reactions occur, viz. evolution of oxygen at the anode, namely oxygen evolution reaction (OER) and production of hydrogen at the cathode, namely HER. The overall EWS process is demonstrated in Figure 1a. Several experimental and theoretical studies have been conducted on the HER mechanism, in alkaline and acidic media through the Volmer–Tafel or Volmer–Heyrovsky routes as displayed in Figure 1b [23]. When HER occurs in an acidic solution, first, hydrogen is electrochemically adsorbed over the electrode surface, followed by the generation of molecular hydrogen due to desorption [24]. In the final step, the adsorbed hydrogen is transformed into  $H_2$  as follows, Equations (1)–(3):



**Figure 1.** (a) Schematic mechanism of OER and HER processes during overall water-splitting process, (b) fundamental principles of HER over the surface of electrocatalysts in acidic and alkaline environments [23]. Copyright 2020 The American Chemical Society.

When HER occurs in an alkaline solution, the water molecules dissociate to form protons, which involves the Volmer and Heyrovsky steps (Figure 1b). Next, unlike the acidic solution, no change occurs in the Tafel step as demonstrated in Equations (4)–(6) [25],



### 2.1. Parameters Governing the Electrocatalytic HER Process

In order to understand the electrocatalytic process of HERs, several significant parameters need to be determined, including the onset potential or overpotential ( $\eta$ ), Gibb's free energy, turnover frequency (TOF), faradaic efficiency, Tafel slope, exchange current density, and stability. These parameters provide insights into the performance of the catalysts.

#### 2.1.1. Overpotential $\eta$

During water electrolysis, the accumulation of  $H_2$  over the surface of the electrode hinders the electrode reactions, resulting in a sluggish reaction rate. In order to overcome this issue, an additional potential called the overpotential ( $\eta$ ) could be applied during the electrolysis process. Overpotential is basically the difference between the thermodynamic potential for a given HER and the applied potential at which the catalyst operates at a specific current under specific conditions. In general, the water electrolysis equation has three major components (Equation (7)), i.e., the overpotential  $\Delta E_{irreversible}$ , the theoretical disintegration potential  $E_{reversible}$ , and the dropped voltage  $IR$  [26].

$$E_{electrolysis} \rightarrow E_{reversible} + \Delta E_{irreversible} + IR \quad (7)$$

Here,  $IR$  is the product of resistance ( $R$ ) (arising from contact wires, points, and electrolytes) and electric current ( $I$ ). Based on earlier reports [26,27], the catalyst's performance is evaluated by the overpotential at a current density of  $10 \text{ mA cm}^{-2}$ , which is considered as the standard. The electrocatalysts that can yield a very low overpotential at high current densities are called as ideal electrocatalysts [26].

#### 2.1.2. Faradic Efficiency

It indicates the consumption efficiency of transferable electrons through an external circuit during the course of an electrochemical reaction [28]. Further, it can be determined by the ratio between the experimentally determined and theoretically calculated amount of hydrogen produced as expressed below:

$$\text{Faradaic Efficiency} = H_{2\text{produced}}^{\text{experimental}} / H_{2\text{produced}}^{\text{theoretical}} \quad (8)$$

Usually, the quantity of  $H_2$  produced is experimentally deduced by gas chromatography or water–gas displacement techniques, whereas the theoretical values are obtained by galvanostatic or potentiostatic electrolysis.

#### 2.1.3. Tafel Plot

Tafel slope implies the steady-state current density dependency of the change in overpotential, which can be derived by linear-sweep voltammetry (LSV). Using the Tafel plot for the HER process, the reaction kinetics and reaction mechanism can be revealed using the following expression [29]:

$$\eta = b \log j + a \quad (9)$$

In the above equation,  $b$  indicates the Tafel slope,  $j$  represents the current density, and  $a$  is the logarithmic of the exchange current. The Tafel slope is used to determine the

reaction kinetics of an electrochemical HER process, and the reaction kinetics reveal the rate determining step, plausible reaction pathway, and fundamental catalytic behavior of the electrocatalysts under equilibrium conditions. Usually, if the value of “*b*” is small, then faster reaction kinetics prevail and vice versa. The Tafel values are always very low with larger current densities for ideal electrocatalysts.

#### 2.1.4. Turnover Frequency

TOF is a facile and direct fundamental activity marker to identifying the efficiency of the catalyst for the electrochemical HER and OER. Furthermore, it is a measure of the quantity of reactant-consumed or product-formed per unit time for the specified catalyst [30]. In other words, it signifies the ratio between the volumetric rate of reaction and the number of active centers per unit volume. Usually, the TOF of a catalyst is reported with the function of overpotential differentiation. A large number of methods are reliable for loading mass with plenty of variations; hence, to rectify such shortcomings, TOF is the most important parameter for determining HER catalytic performance. The exceptional performance of the electrocatalysts can be unveiled by their high TOF values during the electrochemical process in HER. The commonly used equation for calculating the TOF in electrocatalysis is given as follows,

$$TOF = jxN_A / (Fnx\tau) \quad (10)$$

where *n* indicates the number of electrons necessary to transfer for the production of one molecule of the product, usually for HER and OER, *n* = 2 and 4, respectively. The other parameters *j*, *F*, *N<sub>A</sub>*, and *τ* are the current density, Faraday constant, Avogadro constant, and accurate number of active sites of the catalysts during the reaction, respectively.

#### 2.1.5. Gibbs Free Energy ( $\Delta G_H$ )

According to the Sabatier principle, the electrical charges could be transferred facilely on the electrocatalyst's surface, and the bonding of the adsorbed hydrogen yields molecular hydrogen gas as the product. Based on this principle, we can easily determine the Gibbs free energy and reaction kinetics of HER. For instance, if the bonding between the surface of the electrocatalyst and the adsorbed hydrogen is very strong, the desorption reaction is restricted; hence, the Heyrovsky–Tafel step can occur, whereas a weak bonding between the surface of the electrocatalyst and the adsorbed hydrogen adversely affects the Volmer step. The adsorption of hydrogen over the surface of the electrocatalyst is evaluated by the Gibbs free energy function  $\Delta G_H$ . The reaction kinetics of the electrochemical process in the HER could be determined from the  $\Delta G_H$  value for hydrogen adsorption. Therefore,  $\Delta G_H$  is one of the important parameters for identifying the electrochemical process in HER.

#### 2.1.6. Electrochemical Active Surface Area

The electrochemical active surface area (ECSA) of an electrocatalyst, especially for a porous substrate, determines the exposure of active sites. This approach is used to identify the surface area with double layer capacitance (EDLC) for the electrochemical process. Therefore, an EDLC within a narrow potential range and at different sweep rates is usually employed. If there is no Faradaic reaction in the measured range, then a substantial differential capacitance equivalent to the EDLC (*C<sub>dl</sub>*) is obtained. Based on the obtained specific capacitance (*C<sub>s</sub>*), the ECSA is calculated using the following expression [26,31],

$$ECSA = C_{dl}/C_s \quad (11)$$

According to McCrory et al., the average value of *C<sub>s</sub>* in an alkaline solution is in the range of 40 μF/cm<sup>2</sup> [32]. If the proposed catalysts have a large number of active exposed sites and highly unsaturated atom co-ordinations, then the ECSA is large, indicating the excellent performance of the catalysts.

## 2.2. Necessity of Noble-Metal Free Catalysts for HER

For hydrogen fuel production, it is crucial to develop efficient catalysts and to choose suitable electrolytes to speed up the HER rate. For H<sub>2</sub> generation, electrocatalytic systems usually integrate expensive noble metals, such as platinum catalysts, because of their unique features, such as low over potentials and ultrafast kinetics, for driving the HERs [33]. For developing electrocatalysts that are highly active, stable, durable, earth-abundant, non-toxic, exceedingly conductive, inexpensive, robust, environment-friendly, and efficacious for next-generation HER non-noble metals are being replaced with commercial high-cost Pt/C because of the scarcity of noble metals [34,35]. Moreover, significant theoretical and experimental research progress has been made by scientists in recent years in this direction. Nevertheless, there is an urgent need to replace the noble metal catalysts with promising highly conductive electrocatalysts that are efficient, cost-effective, and abundant and to achieve the goal of exceptional HER catalytic performance, which is similar to those of the existing electrocatalysts [36].

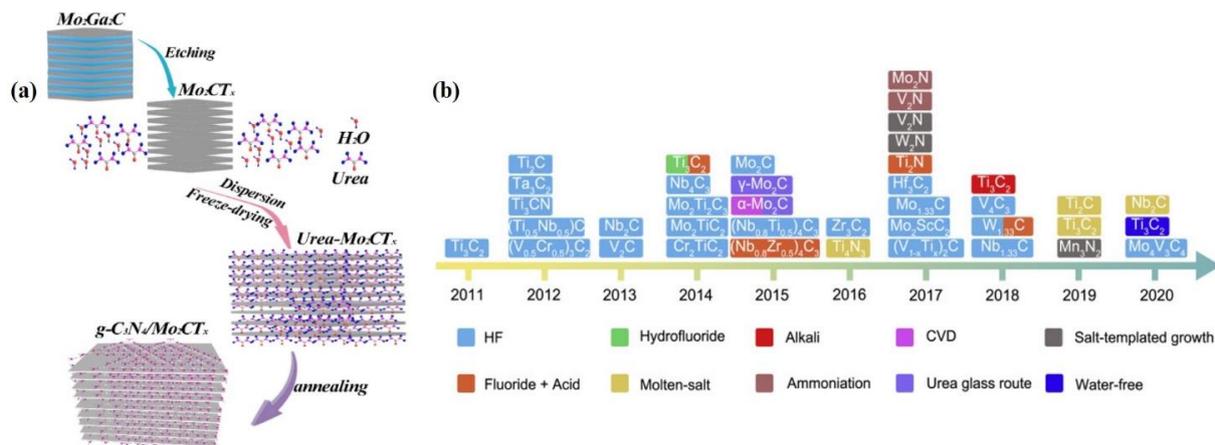
Designing suitable electrocatalysts and selecting the corresponding electrolytes are two primary influencing factors that accelerate HERs [25]. By employing efficient catalysts for HER, the reaction kinetics can be greatly enhanced by lowering the activation energy, and thus, reducing the overpotential. Many factors should be taken into consideration for developing the ideal electrocatalysts, including the cost of the material, catalytic activity, synthesis method, robustness, and earth-abundant materials with long-term utilization. Electrocatalysts with outstanding kinetics, which show negligible overpotential, are the prerequisite for replacing the expensive and limited Pt-based catalysts in HER. In order to discover the ideal electrocatalyst for HER, significant efforts have been made by researchers all over the world. Exchange current density and Tafel slope are the two key parameters that govern the catalytic efficiency. In particular, inexpensive and abundant electrocatalyst materials are required to replace non-abundant materials like platinum. The HER performance of platinum is superior to any other material, and it performs outstandingly both as an H<sub>2</sub> evolution and oxidation catalyst with an excellent catalytic activity and is thus the benchmark for HER reactions. Based on previous studies on the development of HER electrocatalysts, it can be suggested that significant progress is being made both experimentally and theoretically in realizing noble-metal-free catalysts [37]. The recent development of TMCs, especially molybdenum and tungsten carbide-based electrocatalysts, for realizing an improved HER activity and their salient catalytic features are summarized in the following sections.

## 2.3. Synthesis of TMCs

Usually, there are two different categories of 2D TMCs/TMNs, including non-layered and layered materials. Various synthetic techniques are used to determine the structure of the materials. In order to fabricate 2D TMCs/TMNs, two widely accepted approaches are employed:

(a) The first approach is the top-down approach, which is utilized to construct 2D TMCs/TMNs from their bulk counterparts using liquid exfoliation, mechanical exfoliation, and scotch-tape methods.

(b) In the second bottom-up approach, a 2D nanostructure is formed by chemical vapor deposition (CVD) as well as by solution-related chemical syntheses processes such as hydrothermal and solvothermal and physical vapor deposition processes. Numerous materials synthesized through these processes have their own advantages and disadvantages that depend on the raw materials, processes, cost effectiveness, solvent, and precursors. To date, the top-down process is considered the most effective process to construct MXenes as shown in Figure 2 [15]. Unlike the top-down process, the bottom-up process requires less precursor chemicals and can construct atomic ingredients into a few layers of nanostructures. Further, the bottom-up process comprises several beneficial properties, including morphology control, precise size control, and surface termination of MXenes in a specific ordered manner.



**Figure 2.** (a) Illustration for the preparation of the  $\text{Mo}_2\text{CT}_x/\text{g-C}_3\text{N}_4$  hybrid by mechanical exfoliation using HF as etchant, (b) schematic of the synthetic exfoliation protocols of molybdenum and other MXenes reported over the last two decades [14,15]. Copyright 2021 Elsevier.

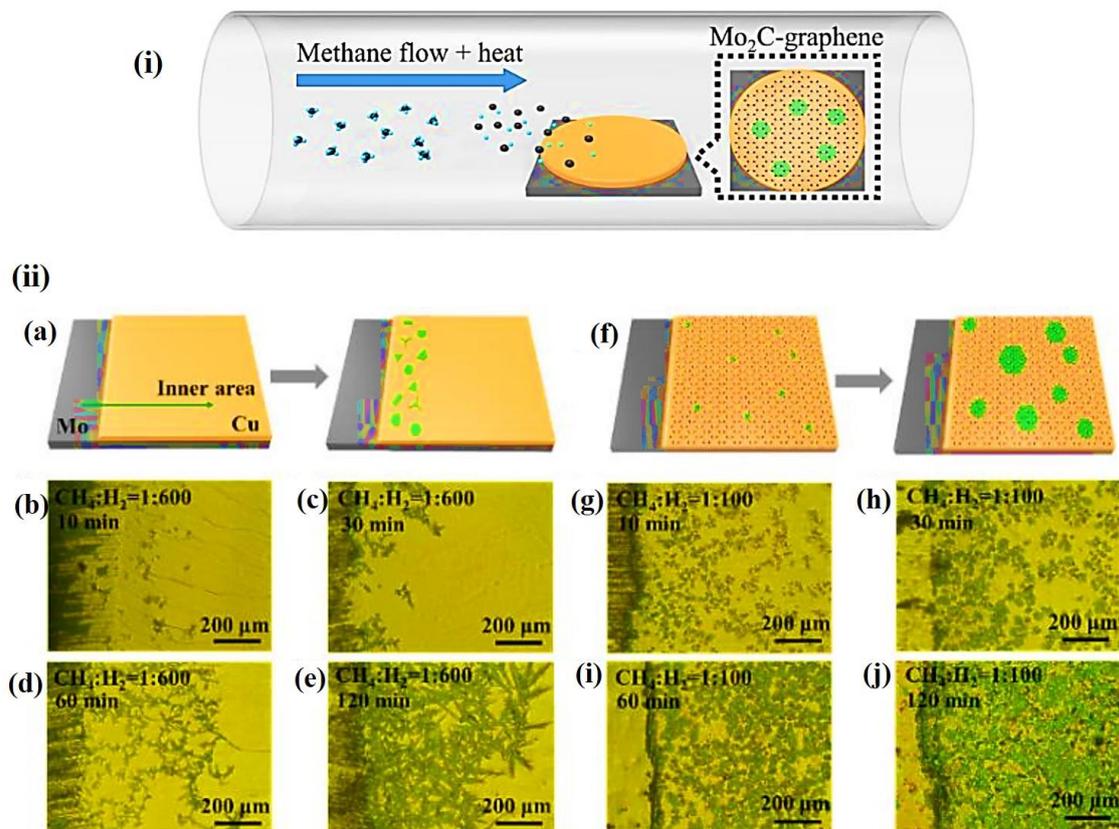
### 2.3.1. Exfoliation Processes

The exfoliation process is widely used to construct few-layered or single-layered structures of 2D metal carbides from their bulk counterparts. During exfoliation, the type of etching influences the performance of the electrocatalysts. The electrocatalyst materials can be mechanically exfoliated or synthesized by liquid exfoliation via fluoride etching, alkaline etching, electrochemical etching, and water-free etching. For instance,  $\text{Mo}_2\text{CT}_x$  can be prepared liquid exfoliation using different etching solutions like LiF, NaF, KF, and  $\text{NH}_4\text{F}$  mixed with HCl [38]. In particular, ultrapure  $\text{Mo}_2\text{CT}_x$  can be obtained with a solution mixture of  $\text{NH}_4$  and HCl at the very low temperature of  $140^\circ\text{C}$ . Wan et al. recently constructed a  $\text{g-C}_3\text{N}_4/\text{Mo}_2\text{CT}_x$  hybrid via selective etching with HF, followed by calcination [14]. The schematic illustration of the preparation of  $\text{Mo}_2\text{CT}_x$  from  $\text{Mo}_2\text{Ga}_2\text{C}$  using HF etchant is shown in Figure 2a. In this case, the ratio of HF and HCl were maintained at 3:2, and the reaction occurred for 72 h. In the second step, urea and the derived  $\text{Mo}_2\text{CT}_x$  were mixed in a ratio of 1:20 and calcined at  $550^\circ\text{C}$  to obtain the  $\text{Mo}_2\text{CT}_x/\text{g-C}_3\text{N}_4$  hybrid. In the case of the scotch-tape process, a piece of tape is adhered to the surface of the bulk 2D material, so that few layers of the top 2D flakes become attached to the tape. Subsequently, this tape is pressed on the preferred substrate to transfer the few-layered 2D flakes. Initially, the scotch tape process (mechanical exfoliation) was employed to synthesize few-layered graphene from bulk graphite. In the recent decade, it has been used to exfoliate 2D MXene materials as well owing to its cost effectiveness and simplicity. Zhang et al. deposited  $\text{Ti}_3\text{C}_2$  thin films over an  $\text{SiO}_2$  substrate by three different processing techniques, including mechanical exfoliation, wafer dipping, and spin coating [39].  $\text{Mo}_4\text{Ce}_4\text{Al}_7\text{C}_3$  flakes were synthesized through mechanical exfoliation process by Gkountras et al. very recently [40]. The schematic synthetic exfoliation protocols of various molybdenum and other carbide MXenes over last two decades are shown in Figure 2b [15].

### 2.3.2. CVD

The CVD process is a type of bottom-up process to construct highly pure and high-performance solid MXene materials. Usually, a thin layer or a few layers of the required materials are constructed by heating the gaseous reactants in a CVD furnace, between the substrate and gaseous reactants or after the in situ chemical reactions between gaseous precursors. Chuan et al. established the production methodology of 2D ultrathin  $\alpha\text{-Mo}_2\text{C}$  crystals using CVD to produce a stable few-nanometers-thick layer with a large area [41]. The group used methane as the carbon source and a combination of Cu and Mo foils as the substrate for the growth at a temperature of  $1000\text{--}1100^\circ\text{C}$ . Similarly, Geng et al. fabricated  $\text{Mo}_2\text{C}$  and  $\text{Mo}_2\text{C}$ -graphene hybrid films of various thickness by CVD using a

liquid Cu surface instead of foil substrates [42]. Usage of liquid Cu completely avoids the grain boundary formation as it grows on the supported Mo foil, which in turn acts as a catalyst during the growth process. The same group used methane as the precursor with  $H_2$ , which acts as a reducing as well as carrier gas during the formation at a temperature of  $1100\text{ }^\circ\text{C}$ . Likewise, Kang et al. recently developed a  $Mo_2C$ -graphene structure over an Mo/Cu substrate by CVD using a different ratio of methane to hydrogen with different time intervals, such as 10, 30, 60, and 120 min, as shown in Figure 3 [16]. Interestingly, the optical images of bare  $Mo_2C$  showed the growth of uneven  $Mo_2C$  crystals at the edges of the substrate (Figure 3b–e). Once graphene is introduced (Figure 3f), it becomes even, and the  $Mo_2C$ -graphene is distributed all over the surface of the substrate as depicted in Figure 3g–j.



**Figure 3.** (a) Typical representation for the formation of  $Mo_2C$ -graphene over a Cu/Mo substrate under various flow rates (low and high) of  $CH_4$  to  $H_2$ . (ii) (b–e) Optical microscope image of  $Mo_2C$  over a Cu/Mo substrate under various flow rates (low and high) of  $CH_4$  to  $H_2$  at different time intervals of (b) 10 min, (c) 30 min, (d) 60 min, and (e) 120 min, (f) formation of  $Mo_2C$ -graphene over a Cu/Mo substrate at high flow rate, (g–j) Optical microscope image of  $Mo_2C$ -graphene over a Cu/Mo substrate under various flow rates (low and high) of  $CH_4$  to  $H_2$  at different time intervals of (g) 10 min, (h) 30 min, (i) 60 min, and (j) 120 min [16].

Hot filament CVD (HF-CVD) has been used to grow vertically aligned graphene nanoribbons (VA-GNR) from vertically aligned carbon nanotubes (VA-CNTs) [43]. After the growth process, a thin layer of Mo ( $\sim 75\text{ nm}$ ) is deposited over the top of VA-GNR and proceeded to atomic H treated for carburization of metallic Mo. For a competitive performance study,  $Mo_2C$  without a VA-GNR support was prepared. Jia et al. synthesized nanosheets (NSs) of nitrogen-incorporated  $Mo_2C$  (N/ $Mo_2C$  NSs) in a two-step process. First, ultrathin  $MoO_2$  NSs were prepared by chemical vapor reduction of commercial  $MoO_3$  powders under an  $Ar/H_2$  environment at  $900\text{ }^\circ\text{C}$  for 60 min. In the second step,

dicyandiamide ( $C_2H_4N_4$ ) powder was mixed well with 100 mg of  $MoO_2$  NSs obtained from the first step and placed inside a tube furnace under an  $Ar/H_2$  atmosphere. Next, the mixture was subjected to calcination twice simultaneously at  $450\text{ }^\circ\text{C}$  and  $700\text{ }^\circ\text{C}$  for 2 h, resulting in the formation of  $N-Mo_2C$  NSs. These synthesized  $N-Mo_2C$  NSs exhibited an excellent HER activity with overpotentials of 48.3 and 99 mV vs. RHE at the cathodic current densities of 1 and  $10\text{ mA cm}^{-2}$ , respectively. A very low Tafel slope of  $44.5\text{ mV dec}^{-1}$ , a large exchange current density of  $0.1\text{ mA cm}^{-2}$ , and excellent long-term stability (stable for more than 12 h) were observed for the resulting catalyst [44].

### 2.3.3. Hydrothermal/Solvothermal Process

Compared to the CVD process, the hydrothermal and solvothermal processing of 2D TMCs/TMNs facilitates an inexpensive, facile, and large-scale production of these electrocatalysts. Usually, powdery target products can be obtained through the hydrothermal/solvothermal process by adding suitable starting reagents in water/other solvents at a high pressure and temperature. Compared to the hydrothermal approach, other solvents are used to replace water in the solvothermal method. Notably, the solvothermal reaction temperature should be higher than the boiling temperature of the solvent when all the starting reagents are kept in a closed vessel. Some interesting reports on hydrothermally processed 2D TMCs/TMNs are discussed herein to gain more insights into these materials and their potential as electrocatalysts. Glucose hydrolyzed  $Mo_2C$ -RGO (RGO denoted reduced graphene oxide) hybrids were developed by Pan et al. by hydrothermal carbonization using a graphene oxide (GO) suspension and ammonium heptamolybdate at  $180\text{ }^\circ\text{C}$  [45]. Interestingly, amorphous  $Mo_2C$  changed into 3D particles with an average size of 50 nm in the absence of GO sheets. Cheng et al. synthesized  $\beta$ - $Mo_2C$  with a Mo:amine ratio of 1:4 (heptamolybdate  $((NH_4)_6Mo_7O_{24}\cdot 4H_2O)$  or sodium molybdate  $(Na_2MoO_4)$  in 4-Cl-phenylenediamine in 60 mL of deionized water) in an acidic (pH 3) solution to ensure the formation of proper amine-oxide hybrids [46]. In the case of Fe-doped molybdenum carbide, 0.02–0.1 equimolar ratio of  $FeCl_3\cdot 6H_2O$  was added to a wet precipitate, and a usual purification criterion was followed. The synthesized  $\beta$ - $Mo_2C$  was further subjected to a long-term stability analysis at both low ( $-140\text{ mV}$ ) and high ( $-240\text{ mV}$ ) overpotentials for an appropriate period of 9 h. In contrast, two-step protocols were used by Meihong et al. to synthesize  $Mo_2C/CC$  (CC represents carbon cloth) [47]. In the first step, a hybrid between the molybdate anions and the cetyltrimethyl ammonium cations was formed over the CC by hydrothermal processing. Next, a simple in situ thermal treatment was performed in an  $Ar$  atmosphere to form  $Mo_2C@C$  hybrids at  $900\text{ }^\circ\text{C}$ . The resulting  $Mo_2C@C$  exhibited a record high HER activity of  $140\text{ mV @ }10\text{ mA cm}^{-2}$  and  $124\text{ mV per dec}^{-1}$  (overpotentials and Tafel slope, respectively).

## 3. HER Activity of 2D TMCs

### 3.1. Molybdenum Carbide ( $Mo_2C$ ) and Its Composite

Molybdenum carbide  $Mo_xC$  is known to exhibit high electronic conductivity and has a Pt-like electronic configuration. The excellent HER activity of  $Mo_2C$  in both alkaline and acidic media was first demonstrated in 2012 [48]. Nevertheless, the application of  $Mo_2C$  as a substitute of noble metal Pt is hindered by two notable obstructions [19,22]. First, the temperature-controlled reaction or commercial carburization in gas phase restricts the nanostructuring. Second, the reduction of metal starting reagents needs a high elevated temperature and highly flammable hydrocarbon gases such as  $CO$ ,  $CH_4$ , and  $C_2H_6$ , which damage the nanostructure, reduce the number of active sites, and deteriorate the reaction surface. Hence, engineering the gas phase constituents, surface area is essential to forming various phases of metal carbides with various catalytic activities. The ultra-prime progress on synthesis protocols has been made to overcome these problems. To avoid particle coking and sintering, various process like incorporation of carbon supports or metal-organic frameworks (MOFs) into the carbide matrix are adopted. Further, the inclusion of

heterometal atoms also improves the overall catalytic activity. These issues are discussed in detail in the following sections with examples.

### 3.1.1. Molybdenum Carbide ( $\text{Mo}_2\text{C}$ )

Initially, molybdenum carbide ( $\text{Mo}_2\text{C}$ ) was prepared by Qamar and his colleague by a two-step fabrication process that involved the incorporation of an Mo source into the MOF matrix MIL-53(Al). Next, carburization was conducted to facilitate the growth and nucleation of  $\text{Mo}_2\text{C}$  nanocrystals into a well-defined porous texture. Moreover, they obtained an overpotential of 165 @ 10 mA, a low Tafel slope of 63.6 mV dec<sup>-1</sup>, and observed a prolonged durability (20 h) in 1.0 M KOH medium [49].

By contrast, Chen et al. prepared  $\beta$ - $\text{Mo}_2\text{C}$  nanoparticles (NPs) by an in situ carburization using CNTs and ammonium molybdate. These NPs exhibited much smaller onset potential values of 63 @ 1 mA cm<sup>-2</sup> and 105 @ 10 mA cm<sup>-2</sup>, a small Tafel slope of 55.2 V dec<sup>-1</sup>, and an exchange current density of  $1.4 \times 10^{-2}$  mA cm<sup>-2</sup> [50]. The  $\beta$ - $\text{Mo}_2\text{C}$  solution was prepared by mixing ammonium heptamolybdate (( $\text{NH}_4$ )<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub>·4H<sub>2</sub>O) or sodium molybdate ( $\text{Na}_2\text{MoO}_4$ ) in 4-Cl-o-phenylenediamine with an Mo:amine ratio = 1:4. Acidic pH was maintained by adding hydrochloric acid until the formation of amine-oxide hybrid precipitates. Pure  $\beta$ - $\text{Mo}_2\text{C}$  was obtained as a result of centrifugation at 5500 rpm for 10 min; the solution was heated to 50 °C with stirring for 2.5 h. In contrast, for Fe-doped molybdenum carbide synthesis, 0.02–0.1 molar equivalents of  $\text{FeCl}_3/6\text{H}_2\text{O}$  were added in 60 mL of water. Thus, the prepared  $\beta$ - $\text{Mo}_2\text{C}$  exhibited long-term durability at both lower and higher overpotentials of –140 and –240 mV, respectively, for 9 h. Later, Jia et al. reported nitrogen-doped  $\text{Mo}_2\text{C}$  NSs prepared using a two-step process. These NSs exhibited an excellent HER activity with an overpotential of 48.3 and 99 mV vs. RHE at cathodic current densities of 1 and 10 mA cm<sup>-2</sup>, respectively. A very low Tafel slope of 44.5 mV dec<sup>-1</sup>, a large exchange current density of 0.1 mA cm<sup>-2</sup>, and excellent long-term stability of more than 12 h were observed [44]. Chaoyun et al. developed  $\text{Mo}_2\text{C}$  by a combined hydrothermal and calcination process. The product obtained from the hydrothermal treatment was heated in a tube furnace at 800 °C for about 3 h in an Ar atmosphere. Finally,  $\beta$ -black molybdenum carbide ( $\beta$ - $\text{Mo}_2\text{C}$ ) was obtained by H<sub>2</sub> reduction for 20 min. The as-prepared catalyst exhibited the best HER activity with the smallest onset potential of 80 mV @ 1 mA cm<sup>-2</sup>, a small Tafel slope of 55 mV dec<sup>-1</sup>, an overpotential of –165 mV @ 10 mA cm<sup>-2</sup>, and a large current density of 60 mA cm<sup>-2</sup> @ 200 mV [51].

Chen et al. prepared multiple phases of molybdenum carbide ( $\text{Mo}_2\text{C}$ ) by solution-based processes. The amine-metal oxide hybrid was precipitated from an aqueous solution of ammonium molybdate and 4-Cl-o-phenylenediamine (4Cl-o-PDA) by adjusting the pH to below 3. The amine-metal oxide hybrid was heated at 750 °C for 12 h to form  $\alpha$ - $\text{MoC}_{1-x}$ , at 850 °C for 12 h to produce  $\beta$ - $\text{Mo}_2\text{C}$ , and at 850 °C for 24 h to form  $\gamma$ - $\text{MoC}$ . p-Phenylenediamine (pPDA) was used as the precursor to produce  $\eta$ - $\text{MoC}$  at 1050 °C with a dwell time of zero. The  $\beta$ - $\text{Mo}_2\text{C}$  electrode exhibited a Tafel slope of 120 mV/decade and an exchange current of 17.29  $\mu\text{A cm}^{-2}$  ( $\eta = 180$ –220 mV). Comparatively, the Tafel analysis of  $\gamma$ - $\text{MoC}$  yielded an exchange current of 3.2  $\mu\text{A cm}^{-2}$  and a Tafel slope of 121.6 mV dec<sup>-1</sup> ( $\eta = 270$ –310 mV) [46].

The  $\text{MoC}$ - $\text{Mo}_2\text{C}$  hetero-nanowires (NWs) were prepared by immersing ( $\text{NH}_4$ )<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub>·4H<sub>2</sub>O in aniline solution by Huanlei et al. During the preparation process, it is keen to note that the solution should be kept at 50 °C for 4 h in an oil bath for its slower reaction process; the final products were filtered and thoroughly washed with ethanol, and then dried at 50 °C overnight. HER parameters such as overpotential, Tafel slope, and onset potentials in 0.5 M H<sub>2</sub>SO<sub>4</sub> and 1.0 M KOH were 126 and 120 mV, 43 and 42 mV dec<sup>-1</sup>, and 38 and 33 mV, respectively. The exchange current density ( $j_0$ ) of  $\text{MoC}$  was  $1.1 \times 10^{-2}$  mA cm<sup>-2</sup> [52].

Pan et al. successfully synthesized novel molybdenum carbide ( $\text{Mo}_x\text{C}$ ) using a simple one-pot pyrolysis method. One-dimensional  $\text{MoO}_3$ -EDA prepared by the already reported procedure was subjected to a pyrolysis treatment at 600 °C under an Ar gas atmosphere, and molybdenum carbides  $\text{Mo}_x\text{C}$  with high-valence state Mo nanobelts were obtained. The as-prepared products displayed tremendous HER electrocatalytic activity with a low onset

overpotential of 50 mV and a small Tafel slope of 49.6 mV dec<sup>-1</sup> in an acidic medium (0.5 M H<sub>2</sub>SO<sub>4</sub>). In addition, the catalysts necessitate merely overpotentials of 143 and 234 mV to achieve current densities of 10 and 220 mA cm<sup>-2</sup>, respectively with good durability after 2000 cycles [53].

Bin et al. presented a unique synthesis strategy to produce porous MoC<sub>x</sub> nano-octahedrons via an in situ wet chemical process and confined carburization reactions. The group used metal–organic frameworks (MOFs) and guest polyoxometalates (POMs) that resides inside the pores of MOFs and acts as receiver of the guest metal species into the MOFs. The co-precursor—namely transition metals (Mo, W, and V)—carbides enable easy synthesis in the presence of metal species that would be impossible if single MOF sources were used. MoC<sub>x</sub> nano-octahedrons possess HER onset potentials of 87 and 92 @ 1 mV, 142 and 157 @ 10 mV vs. RHE, low Tafel slopes of 53 and 59 mV dec<sup>-1</sup>, and exchange-current densities of 0.023 and 0.029 mA cm<sup>-2</sup> in acidic and alkaline media with excellent long-term stability (more than 10 h) [54]. A meso-microporous structured 3DHP-Mo<sub>2</sub>C catalyst fabricated by Meng et al. involves a four-step process. The constructed 3DHP-Mo<sub>2</sub>C apparently exhibited a comparatively larger current density—1 mA cm<sup>-2</sup> with a low overpotential of 75 mV—than 3DHPC and Com-Mo<sub>2</sub>C (235 and 202 mV at the same current density). Further, 3DHP-Mo<sub>2</sub>C exhibits a relatively low Tafel slope value with excellent stability (10 h) in 1M KOH solution [55].

### 3.1.2. Mo<sub>2</sub>C and Carbon Composites

For electronic modification and nanostructuring, different carbon-based supports including GO, CNTs, and carbon fiber are employed along with metal carbides. The inclusion of well-defined carbon supports enhances the particle dispersity and so the number of active sites, thereby leading to improvement of the electronic conductivity and overall catalytic performance towards HER activity by means of altering the electrochemical structure. A various carbon-supported metal carbides for improved HER performance are elaborated in this section.

Through a solid-state process, Cui et al. decorated Mo<sub>2</sub>C NPs over graphitic carbon sheets (Mo<sub>2</sub>C/GCSs) using (NH<sub>4</sub>)<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub>·4H<sub>2</sub>O and sodium alginate (ALG) as the precursors at 900 °C under an Ar atmosphere. The resulting Mo<sub>2</sub>C/GCSs catalyst was tested in 0.5 M H<sub>2</sub>SO<sub>4</sub> for HER, and a low overpotential (120 mV @ 1 mV) and Tafel slope (62.6 mV dec<sup>-1</sup>) were obtained in the measured range. In addition, the catalyst showed good electrical conductivity with a large exchange current density of 0.125 mA cm<sup>-2</sup> and long durability over 3000 electrochemical cycles.

Using a similar solid-state strategy, Pu et al. developed Mo<sub>2</sub>C quantum dot-based nitrogen-incorporated graphitic carbon layer (Mo<sub>2</sub>C-QDs/NGCLs). Firstly, through the sonication method, the precursors such as chitosan biopolymer and ammonium molybdate ((NH<sub>4</sub>)<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub>·4H<sub>2</sub>O) were dissolved in distilled water. During this stage, a chitosan–Mo complex could be formed by means of MoO<sub>4</sub><sup>2-</sup> conjugate with OH and NH<sub>2</sub> of chitosan. Later, the formed complex was subjected into an annealing process at 900 °C under Ar environment for 2 h, producing a hybrid Mo<sub>2</sub>C-QDs/NGCLs material. For HER analysis, Mo<sub>2</sub>C-QDs/NGCLs delivers the appropriate overpotential of 150 mV @ 10 mA cm<sup>-2</sup> and very mere Tafel slopes values of 81.9 mV dec<sup>-1</sup> in 0.5 M H<sub>2</sub>SO<sub>4</sub> suggesting its good stability. The same catalyst offers the low overpotential and Tafel slope values of 111 mV and 57.8 mV dec<sup>-1</sup> @ 10 mA cm<sup>-2</sup> even in the exceptional alkaline 1.0 M KOH electrolyte [56].

Huang Yang's group developed Mo<sub>2</sub>C NPs using a widely used two-step strategy using ammonium molybdate tetrahydrate ((NH<sub>4</sub>)<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub>·4H<sub>2</sub>O) and dopamine hydrochloride (C<sub>8</sub>H<sub>11</sub>NO<sub>2</sub>·HCl) as the precursors. The uniform dispersion of Mo<sub>2</sub>C over carbon micro flowers (Mo<sub>2</sub>C/NCF) was realized through the self-polymerization of dopamine. For the HER properties, the overpotentials of 85 and 144 mV @ current densities of 1 and 10 mA cm<sup>-2</sup>, respectively, for Mo<sub>2</sub>C/NCF were obtained. Tafel study divulged that the linear region exhibited a slope of 55 mV decade<sup>-1</sup> in a 0.5M H<sub>2</sub>SO<sub>4</sub> solution. HER activities

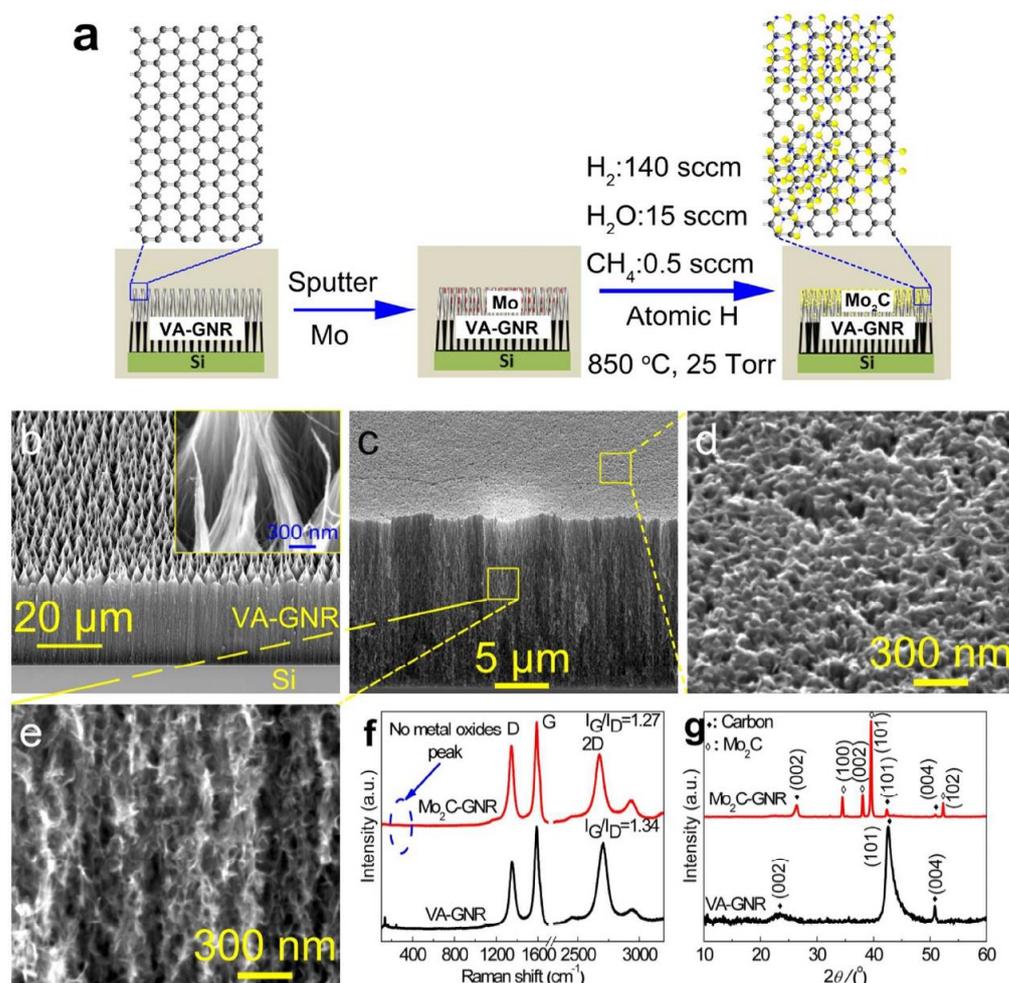
of Mo<sub>2</sub>C/NCF in 1 M KOH were found to be enhanced with an overpotential of 38 and 100 mV @ 1 and 10 mA cm<sup>-2</sup>, respectively, and an excellent stability (8 h) [57].

Dezhi et al. constructed Mo<sub>2</sub>C NPs embedded in Ketjenblack carbon (KB) by a simple in situ carbonization process. The augmented Mo<sub>2</sub>C/KB catalyst reveals a distinctive HER activity in both acidic and alkaline conditions with the least Tafel slopes of 49 and 48 mV dec<sup>-1</sup>, and overpotential of 180 mV @ 10 mA/cm<sup>2</sup>, an exchange current density of  $3 \times 10^{-3}$  mA cm<sup>-2</sup>, and an excellent stability (10 h) [58]. Can et al. prepared hierarchical Mo<sub>2</sub>C/C NS hybrids by employing a novel preparation process involving the carburizing of ammonium molybdate and citric-acid-coated NaCl cube crystals under an Ar gas atmosphere. HER performance of Mo<sub>2</sub>C/C hybrids is studied at both alkaline and acid media with the very low overpotential of 125 mV @ 10 mA cm<sup>-2</sup> in 1 M KOH and 180 mV @ 10 mA cm<sup>-2</sup> in 0.5M H<sub>2</sub>SO<sub>4</sub>, while the observed Tafel values are 72 mV dec<sup>-1</sup> in 1.0 M KOH and 71 mV dec<sup>-1</sup> in 0.5M H<sub>2</sub>SO<sub>4</sub> [59].

Cuncai et al. developed MoC/C by a simple one-step ultrasonic spray pyrolysis (USP) by dissolving MoCl<sub>5</sub> in ethanol. During the USP process, the starting material passes via the heating zone inside the series outlets to safeguard the indistinguishable heating. The resultant MoC/C from the pyrolysis has offered the least Tafel slope of 63.6 mV dec<sup>-1</sup> @ 20 mA cm<sup>-2</sup> in 0.5 M H<sub>2</sub>SO<sub>4</sub>, and exists a long term stability during the electrolysis [60]. Chunyong and coworkers synthesized nanostructured molybdenum carbide of smaller size ranging from 2.5 nm and 5 nm for MoC and Mo<sub>2</sub>C on a graphene support, respectively, through a facile in situ approach. For the production of Mo<sub>2</sub>C-G and MoC-G through this approach, GO powders and (NH<sub>4</sub>)<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub>·4H<sub>2</sub>O were employed as the starting reagents. The HER performances of Mo<sub>2</sub>C-G and MoC-G with the least overpotential ( $\eta_{10}$ ) of 150 and 221 mV @ 10 mA cm<sup>-2</sup>, values of Tafel slopes of 57, and 88 mV dec<sup>-1</sup> respectively. The exchange current density ( $j_0$ ) of the MoC-G and Mo<sub>2</sub>C-G are observed to be 0.00255 mA cm<sup>-2</sup> and 0.00258 mA cm<sup>-2</sup> respectively and demonstrates the exceptional stability over 20 h [61].

Huo Lili and his team established an inexpensive, noble-metal-free mesoporous Mo<sub>2</sub>C/graphene (m-Mo<sub>2</sub>C/G) catalyst with two-dimensional layered structure by a nanocasting method using glucose as a carbon source for in situ assembled mesoporous KIT-6/graphene (KIT-6/G) as a template. The m-Mo<sub>2</sub>C/G catalyst outrages large catalytic behavior with outstanding long-term stability for HER over a wide pH range at small onset potential of 8 mV. Overpotential ( $\eta_{10}$ ) for a cathodic current density of 10 mA cm<sup>-2</sup> of 135 mV, a Tafel slope of 58 mV dec<sup>-1</sup>, and an exchange current density of  $6.31 \times 10^{-2}$  mA cm<sup>-2</sup> in acidic media have been observed. Similarly for alkaline medium, an onset potential of 41 mV,  $\eta_{10}$  of 128 mV, an exchange current density of 0.00409 mA cm<sup>-2</sup>, and Tafel slope of 56 mV dec<sup>-1</sup> have been observed [8]. By using the HF-CVD process, vertically aligned graphene nanoribbons (VA-GNR) were grown from vertically aligned carbon nanotubes (VA-CNTs) [43]. After the growth process, a thin layer of Mo (~75 nm) was deposited on the top of VA-GNR and atomic H treatment was conducted for carburization of metallic Mo as displayed in Figure 4. For a comparison, the performance study of Mo<sub>2</sub>C without the VA-GNR support was also prepared.

The corresponding scanning electron microscopy (SEM) images reveal the vertically aligned nanoribbons over Mo<sub>2</sub>C as shown in Figure 4b–e. Its respective spectral analyses such as Raman and X-ray diffraction (XRD) imply the purity of the prepared catalyst and confirms the absence of metal oxide in the prepared materials as shown in Figure 4f,g.



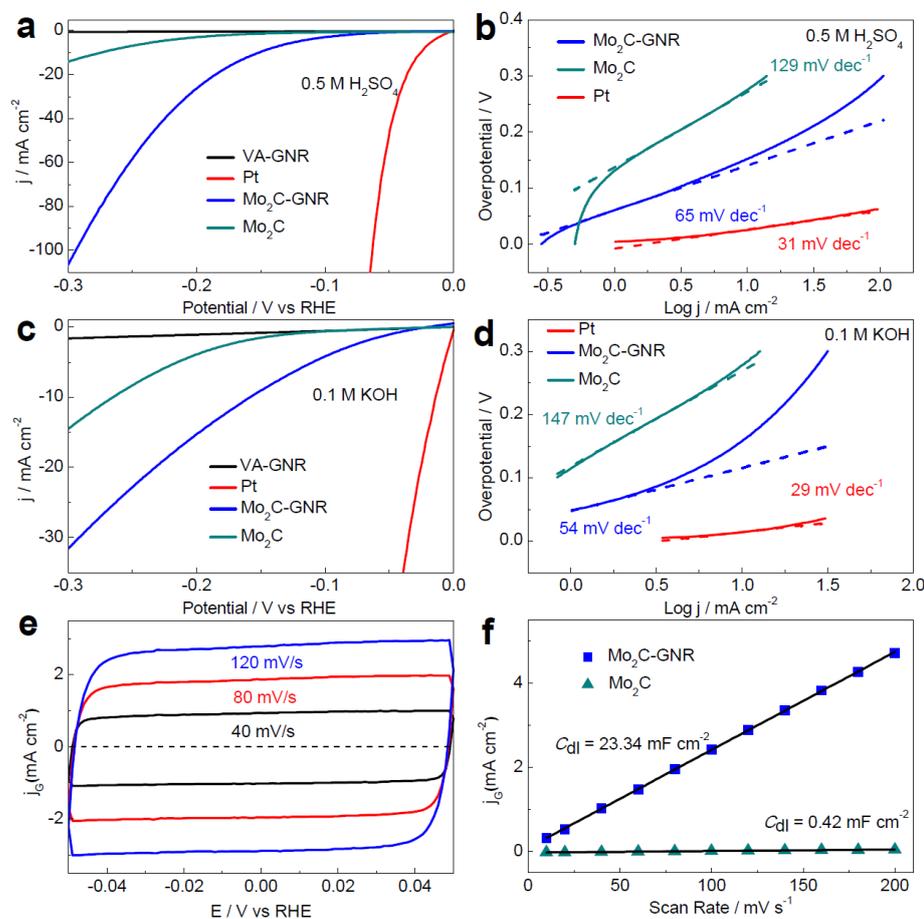
**Figure 4.** (a) Schematic showing the formation of Mo<sub>2</sub>C over VA-GNR, (b–e) field-emission SEM image of VA-GNR and Mo<sub>2</sub>C-GNR at high and low magnifications, (f,g) their corresponding Raman spectra in the range of 300–3100 cm<sup>-1</sup> and XRD patterns in the 2θ range of 10–60°. These spectral analyses confirmed the formation of Mo<sub>2</sub>C over the GNR matrix [43]. Copyright 2017 The American Chemical Society.

As depicted in Figure 5a–d, the HER performance of Mo<sub>2</sub>C-GNR in acidic (0.5 M H<sub>2</sub>SO<sub>4</sub>) as well as alkaline (1 M KOH) solutions confirmed the excellent HER activity of Mo<sub>2</sub>C-GNR. Compared to the current density values (29 and 31 mA cm<sup>-2</sup>) of Pt/C and those of the stand-alone Mo<sub>2</sub>C (~275 and ~266 mA cm<sup>-2</sup>), the above catalyst Mo<sub>2</sub>C-GNR exhibited a better current density of 10 mA cm<sup>-2</sup> at a low (η) of ~152 and ~121 mV. Similarly, the Tafel slope values of the Mo<sub>2</sub>C-GNR and Mo<sub>2</sub>C catalysts in acidic and alkaline solutions were 65 and 54 mV dec<sup>-1</sup> and 129 and 147 mV dec<sup>-1</sup>, respectively.

As shown in Figure 5e, a systematic cyclic voltammetry (CV) test was conducted to determine the EDLC behavior (C<sub>dl</sub>) of the resultant catalysts. As evident from Figure 5f, the capacitance values of Mo<sub>2</sub>C-GNR and Mo<sub>2</sub>C electrodes are 23.34 and 0.42 mF cm<sup>-2</sup>, along with the roughness factors of 1060 and 19, respectively.

Kasinath et al. synthesized RGO-based composites with Mo<sub>2</sub>C rods using a two-step process that involved aniline complexation followed by annealing. The RGO-Mo<sub>2</sub>C catalyst exhibited an excellent electrochemical activity toward HER in an acidic medium and achieved the highest current density of 125 mA/cm<sup>2</sup> at 400 mV and the lowest Tafel slope of 67 mV dec<sup>-1</sup> [62]. The presence of GO in the Mo<sub>2</sub>C-RGO composite reduced the particle size (~10 nm) and chiefly improved the carbonization reaction. Kai et al. fabricated Mo<sub>2</sub>C-NCNTs by a two-step process. The Mo<sub>2</sub>C-NCNTs offers the lowest overpotentials of 72 and

147 mV @ 1 and 10 mA cm<sup>-2</sup> current densities, respectively. The Mo<sub>2</sub>C-NCNTs catalyst has further shown the least Tafel slope of 71 mV dec<sup>-1</sup>, exchange current density of 114.6 μA cm<sup>-2</sup> [63]. Youn et al. fabricated Mo<sub>2</sub>C/CNT-graphene using CNTs and graphene oxide (GO) in ethanol solution followed by the addition of metal precursor (MoCl<sub>5</sub>) resulting in the Mo orthoester formation. On addition of urea (1 molar ratio (R)) yielded a metal-urea complex on CNT-GO, which upon heat treatment under N<sub>2</sub> atmosphere at 750 °C results in Mo<sub>2</sub>N/CNT-GR hybrids. Upon increasing, the R-value by a factor of eight, Mo<sub>2</sub>C/CNT-GR been generated instead of Mo<sub>2</sub>N. The unreacted nitrogen and oxygen atoms were removed from the precursor solutions as NO<sub>x</sub> during the heat treatment. The Mo<sub>2</sub>C/CNT-graphene displays the maximum activity for HER with a least overpotential and Tafel slope of 62 mV and 58 mV dec<sup>-1</sup> respectively @ 1 mA cm<sup>-2</sup>. The calculated exchange current density is in the range of 6.20 × 10<sup>-2</sup> mA cm<sup>-2</sup> along with very good stability in H<sub>2</sub>SO<sub>4</sub> solution [64].



**Figure 5.** (a,b) LSV and Tafel slope plots of Mo<sub>2</sub>C nanocrystals over VA-GNR, bare VA-GNR and Mo<sub>2</sub>C in an acidic medium, (c,d) LSV and Tafel slope plots of Mo<sub>2</sub>C nanocrystals over VA-GNR, bare VA-GNR, and Mo<sub>2</sub>C in an alkaline medium, (e,f) corresponding CV and current density vs. scan rate plots at various scan rates to determine  $C_{dl}$  [43]. Copyright 2017 The American Chemical Society.

Jian et al. prepared Mo oxyanions-loaded cotton T-shirts using molybdenum carbide NP (BCF/Mo<sub>2</sub>C) electrodes via a direct annealing treatment method. The BCF/Mo<sub>2</sub>C displayed an enhanced electrocatalytic performance with extremely low overpotentials of 88 and 115 mV to drive a current density of 20 mA cm<sup>-2</sup>, and Tafel slopes of 84.8 mV dec<sup>-1</sup> and 52.4 mV dec<sup>-1</sup> in alkaline and acidic media. The hybrid's 50 h stability test showed a decreased cathodic current density in both the alkaline and acidic solutions [65]. Mo<sub>2</sub>C-based crosslinked carbon networks (Mo<sub>2</sub>C/CLCN) were prepared by Jia et al. Equal volumes of CuCl<sub>2</sub>·2H<sub>2</sub>O (1.4 mmol) and (NH<sub>4</sub>)<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub>·4H<sub>2</sub>O (1 mmol) solutions were

mixed subsequently and heated at 450 °C for 4 h under an Ar-H<sub>2</sub> (10%) atmosphere. After the calcination, the black Cu-MoO<sub>2</sub> rods were obtained. Furthermore, Cu-MoO<sub>2</sub> rods has been heated up to 1000 °C under Ar (10 sccm) atmosphere to form Mo<sub>2</sub>C based crosslinked carbon network (Mo<sub>2</sub>C/CLCN). Mo<sub>2</sub>C/CLCN display a low onset potential of −85 mV @ 10 mA cm<sup>−2</sup> an operating overpotential of 145 mV @ 10 mA cm<sup>−2</sup> with a Tafel slope of 48.2 mV dec<sup>−1</sup>, with a large exchange current density (0.062 mA cm<sup>−2</sup>), and outstanding long-term cycling stability in acidic electrolyte [66].

Fan et al. prepared mesoporous Mo<sub>2</sub>C/nanocrystals by dispersing ammonium molybdate (0.225 g) in ethylenediamine (1.35 g) by ultrasonication and stirring for at least 20 min. Then, carbon tetrachloride (2.31 g) was added, and the resultant mixture was heated, refluxed, and stirred at 90 °C for 6 h. Next, heat-treatment of the polymer composites was performed under nitrogen flow at 800 °C with a heating rate of 3.08 °C min<sup>−1</sup> and kept under these conditions for 6 h to carbonize the polymer. The optimized electrocatalyst Mo<sub>2</sub>C/nanocrystals requires relatively low overpotentials of 140 mV to produce a current density of 10 mA cm<sup>−2</sup> for the electrocatalytic HER, a Tafel slope of 116 mV dec<sup>−1</sup>, a large exchange current density of 0.0361 mA cm<sup>−2</sup> and to remain stable for 120 h in acidic medium [67].

### 3.1.3. Heterometal Atom Doped Mo<sub>2</sub>C

The doping of hetero-metal atoms with metal carbides especially molybdenum carbide could effectively alter the inherent electrocatalytic activity of Mo by changing its electronic structure. Since, the d-band electron structure of Mo<sub>x</sub>C is same as Pt, the Mo has the restricted activity due to their empty d-orbitals which in turns to cause strong hydrogen-binding and so lower the rate of hydrogen desorption. The doping of hetero-metal atoms in Mo<sub>x</sub>C could downshift the d-band electron center regarding to EF which enhances the kinetic of HER activity. For instance, Ji-Sen et al. designed and fabricated a 2D hybrid consisting of Mo<sub>2</sub>C incorporated by N, P-codoped carbon shells and RGO (Mo<sub>2</sub>C@NPC/NPRGO) by following a three-step protocol. Mo<sub>2</sub>C@NPC/NPRGO has established outstanding HER performance in both alkaline and acidic media with a very low overpotential (34 mV for 10 mA cm<sup>−2</sup>), exchange current density 1.09 mA cm<sup>−2</sup>, Tafel slope (33 mV dec<sup>−1</sup>), and long stability for 10 h in 0.5 M H<sub>2</sub>SO<sub>4</sub> [68]. Jun et al. constructed Mo<sub>2</sub>C porous nanostructures in a controlled way, including 2D NSs and 1D NWs. First, the MoO<sub>3</sub> NSs were prepared by hydrothermal processing. In this process, the precursor solutions, containing a mixture of 2-methylimidazole and cobalt nitrate hexahydrate, were immersed into methanol. This led to the formation of MoO<sub>3</sub>@ZIF-67 hybrid NSs as the ZIF-67 layer was coated on top of the MoO NSs. The coated thin layers were subjected to carburization process under ultrapure Ar atmosphere at 800 °C for 3 h to form nitrogen- and cobalt-doped Mo<sub>2</sub>C porous NSs (2D-N, Co-Mo<sub>2</sub>C), which largely contract the morphology of the MoO<sub>3</sub> NSs. 2D-N, Co-Mo<sub>2</sub>C shows an onset potential as low as 25 mV @ 1 mA cm<sup>−2</sup>, 71 mV @ 10 mA cm<sup>−2</sup>. Tafel slope of 2D-N, Co-Mo<sub>2</sub>C is 40 mV dec<sup>−1</sup> in 0.1 M HClO<sub>4</sub>. 2D-N, Co-Mo<sub>2</sub>C shows an onset potential as low as 36 mV @ 1 mA cm<sup>−2</sup>, 92 mV @ 10 mA cm<sup>−2</sup>, Tafel slope of 2D-N, Co-Mo<sub>2</sub>C is 47 mV dec<sup>−1</sup> in 0.1 M KOH [69].

Co-doped Co-Mo<sub>2</sub>C NWs were prepared by Huanlei et al. using Mo<sub>3</sub>O<sub>10</sub>(C<sub>6</sub>H<sub>5</sub>NH<sub>3</sub>)<sub>2</sub>·2H<sub>2</sub>O and a varied CoCl<sub>2</sub>·6H<sub>2</sub>O precursors in aqueous aniline solution. 1 M HCl aqueous solution was added to the aniline solution to adjust the pH level to 4.5 for thermodynamic reaction. Finally, the pink products obtained were filtered using ethanol, and then dried at 50 °C overnight and then calcined at 750 °C for 5 h resulting in the formation of Co-doped Co-Mo<sub>2</sub>C NWs. Before testing for HER activity, the products were treated through 1.0 M HCl to remove residual Co species. For preparing Mo<sub>2</sub>C NWs, the similar procedures were employed without the addition of CoCl<sub>2</sub>·6H<sub>2</sub>O. Co-Mo<sub>2</sub>C NWs (Co/Mo ratio—0.02) has exhibited a low overpotential ( $\eta_{10}$  = 140 and 118 mV @ 10 mA cm<sup>−2</sup>;  $\eta_{100}$  = 200 and 195 mV @ 100 mA cm<sup>−2</sup>), least Tafel slope values (39 and 44 mV dec<sup>−1</sup>), and merely lower onset overpotential (40 and 25 mV) in 0.5 m H<sub>2</sub>SO<sub>4</sub> and 1.0 m KOH, respectively [70]. Through a joined hydrothermal and carburization process, Ni-doped Mo<sub>2</sub>C over nickel

foam (Ni-Mo<sub>2</sub>C/NF) has prepared by Xiong et al. [71]. The resultant Ni-Mo<sub>2</sub>C/NF is observed to be extremely active toward HER and needs a low onset potential of 21 mV @ 1 mA/cm<sup>2</sup>, 150 mV @ 100 mA/cm<sup>2</sup> which are noticeably smaller than the onset potential of bare Mo<sub>2</sub>C/NF (87 mV) @ 1 mA/cm<sup>2</sup>. The calculated Tafel slopes of 54.2 and 36.8 mV dec<sup>-1</sup> are found for bare Mo<sub>2</sub>C/NF and Ni-Mo<sub>2</sub>C/NF respectively. Further Ni-Mo<sub>2</sub>C/NF delivers the exchange current density of 0.51 mA cm<sup>-2</sup> which is 12 folds greater than that of Mo<sub>2</sub>C/NF. Likewise, the Mo<sub>2</sub>C@nanocrystals catalyst has fabricated by Yipu et al. provides a mere low Tafel slope value of 60 mV dec<sup>-1</sup> and a maximum exchange current density of 0.096 mA cm<sup>-2</sup> with excellent stability in HER even after 80 h-long over a wide pH range (pH 0–14) [72].

Shiping et al. employed solid-state thermolysis reactions to prepare molybdenum-carbide-modified N-doped carbon-vesicle encapsulated Ni NPs (Mo<sub>x</sub>C-Ni@ NCV) Mo<sub>x</sub>C-Ni@ NCV which revealed high catalytic HER activity in acidic electrolyte, indicating a synergistic effect between  $\gamma$ -MoC and Ni. The HER catalyzed by sample Mo<sub>x</sub>C-Ni@ NCV has reached the stable state with the least overpotential and Tafel slope of 75 mV and 45 mV dec<sup>-1</sup> respectively @ 10 mA cm<sup>-2</sup>. The same catalyst delivers the large exchange current density of 0.95 mA cm<sup>-2</sup> in the measured range [73]. A hexamethylenetetramine (HMT) based one-step thermal decomposition method was employed by Yuan et al. to develop Co<sub>3</sub>Mo<sub>3</sub>C decorated CNT (Co<sub>3</sub>Mo<sub>3</sub>C/CNT). The purpose of adding HMT favors the formation of molybdate ion ligand, besides acting as self-reducing agent. Co<sub>3</sub>Mo<sub>3</sub>C/CNT electrocatalyst, points the low onset and overpotentials of 42 mV, and 124 mV, exchange current density of 0.415 mA cm<sup>-2</sup> and Tafel slope of 249 mV dec<sup>-1</sup> @ 10 mA cm<sup>-2</sup> in seawater [74].

Cuicui et al. used in situ confining carburization to prepare well-regulated structures of Mo<sub>2</sub>C NSs in N-doped carbon (MCNS/NC) using Mo-based inorganic–organic lamellar mesostructures at 900 °C. They demonstrated the HER efficiency at the least overpotential of 19 mV and an onset-potential of ~0 mV @ 10 mA cm<sup>-2</sup>. The determined Tafel slope was 28.9 mV dec<sup>-1</sup> in acid media [75]. Zhao et al. performed a two-step process to fabricate MoS<sub>2</sub>/Mo<sub>2</sub>C hybrid NSs. Firstly, they vertically aligned MoS<sub>2</sub> has grown over carbon fiber paper (CFP) by hydrothermal method (200 °C using Na<sub>2</sub>MoO<sub>4</sub> and thiourea). Then, the formed MoS<sub>2</sub> NSs are undergone carburization in Ar/H<sub>2</sub>/CH<sub>4</sub> atmospheres at 750 °C to form MoS<sub>2</sub>/Mo<sub>2</sub>C hybrid NSs. The as-prepared hybrid MoS<sub>2</sub>/Mo<sub>2</sub>C NSs shows the least Tafel slope of 48 mV dec<sup>-1</sup> with an outstanding HER performance (63 mV @ 10 mA/cm<sup>2</sup>) as well as excellent stability [76]. The electrochemical characteristics of various Mo<sub>2</sub>C-based electrocatalysts for improved HER activity are listed in Table 1.

It is well known that TMCs have no limits to control the carbon content in TMCs' phase and so no forfeits for the position of the active sites. In specific the carbon atoms in the metal carbides like Mo<sub>2</sub>C and WC are occupied at the interstitial site of metal ion (Mo and W) mostly stimulates the mixing and re-hybridization processes between the d-orbitals of metal atoms and s, p orbital of C-atom [77]. Henceforth, the electronic density of d-band at the Fermi level matches almost that of Pt-atom and showed Pt-like catalytic behavior. Additionally, as stated above, to obtain the thorough insight into the metal carbides catalysts such as Mo<sub>2</sub>C and WC active sites towards HER, the ECSA is one of the dominant factors to determine the activity of the catalyst which is in linear-dependent with the C<sub>dI</sub> of the resultant catalysts. In addition, to identify the real accurate active sites of the catalysts, the Gibbs free energy  $|\Delta G_{H^*}|$  could be calculated for all sites (including Mo and C in Mo<sub>2</sub>C; W and C in W<sub>2</sub>C or WC) [77].

**Table 1.** Electrochemical characteristics of various Mo<sub>2</sub>C-based electrocatalysts for improved HER activity.

Catalyst	Synthesis	Morphology	$\eta_j$ (mV)	$j$ (mA cm <sup>-2</sup> )	Tafel Slope (mV dec <sup>-1</sup> )	Ref
N-Mo <sub>2</sub> C NSs	CVD	nanosheets	99	10	48.3	[44]
Mo <sub>2</sub> C-RGO	Hydrothermal carbonization	nanoparticles	70	10	57.3	[45]
$\beta$ -Mo <sub>2</sub> C	Hydrothermal/calcination	nanospheres	240	10	120	[46]
Mo <sub>2</sub> C/CC	Hydrothermal/calcination	Nano-island	140	10	124	[47]
Mo <sub>2</sub> C	Chemical activation process	nanoparticle	240	20	56	[48]
$\beta$ -Mo <sub>2</sub> C/C	hydrothermal	Irregular ill-defined particles	330	20	72	[49]
Mo <sub>2</sub> C/CNT	In situ-carburization	nanoparticle	64	1	63	[50]
$\beta$ -Mo <sub>2</sub> C	Hydrothermal/calcination	nanoparticle	165	10	55	[51]
MoC-Mo <sub>2</sub> C	Controlled carbonization	nanoparticle	33	10	42	[52]
Mo <sub>x</sub> C	One-pot pyrolysis	nanobelts	50	10	49.6	[53]
3DHP-Mo <sub>2</sub> C	Scalable salt-template process	Highly interconnected 3D porous network	75	1	75	[55]
Mo <sub>2</sub> C/NCF	High temperature calcination	nanoflowers	144	10	65	[57]
Mo <sub>2</sub> C/C	pyrolysis	nanosheets	180	10	72	[59]
Mo <sub>2</sub> C-G	In situ-carburization	nanoparticle	150	10	57	[61]
Mo <sub>2</sub> C/CLCN	Hydrothermal/calcination	nanorods	155	10	48.2	[66]
Mo <sub>2</sub> C/NC	Polymerization/carbonization	sprout	140	10	114.4	[67]
1D Mo <sub>2</sub> C	In situ-carburization	nanosheets	36	10	47	[69]

RGO—reduced graphene oxide; G—graphene; CNT—carbon nanotube; NSs—nanosheets; NC—nitrogen doped carbon; 1D—one dimensional; 3DHP—three dimensional highly porous network; NCF—nitrogen doped carbon framework;  $\eta_j$ —overpotential;  $j$ —current density.

According to various reported literatures, the  $|\Delta G_{H^*}|$  values at various adsorption sites of Mo is lower than the carbon sites suggesting the actual active site is Mo in Mo<sub>2</sub>C catalyst where the hydrogen is favorably profoundly adsorbed or desorbed [77,78]. Similarly, the real active site W in WC catalyst could be identified by its low  $|\Delta G_{H^*}|$  at various W-adsorption sites and so prefers hydrogen adsorption/desorption. Mostly active edge sites have contributed for the efficient catalytic activity of metal carbides including Mo<sub>2</sub>C and W<sub>2</sub>C/WC. In the specialized case, few research groups purposefully activated the basal or the planar edge sites to enhance the overall catalytic activity of metal carbides [2,79,80]. Generally, to confirm the existence of active sites and their roles towards HER process, various techniques including X-ray absorption near edge structure analysis, DFT calculations, and Fourier transform of the Mo or W extended X-ray absorption structure analysis can be used [81].

### 3.2. Tungsten Carbide (WC/W<sub>2</sub>C)

Tungsten carbide was first invented by Levy and Boudart in the year 1973. This material displays Pt-like performance in various catalytic reactions [82]. In addition, tungsten carbide shows a Pt-like performance with high immunity to catalytic poisonous gases such as CO and H<sub>2</sub>S; such a feature is not observed even in Pt. Based on the typical metal-carbon bonding, tungsten carbide is analogous to molybdenum carbide and has progressed in a parallel technique. There are two different phases of tungsten carbide including WC and W<sub>2</sub>C, which are commonly investigated as the electrocatalysts for HER process [82]. As compared to WC, W<sub>2</sub>C possesses higher catalytic activity [83]. Since, the ratio of tungsten to carbon starting reagents in W<sub>2</sub>C is uncontrollable, commercial carbide synthesis protocols using hydrocarbon gases like CO, C<sub>2</sub>H<sub>6</sub> and CH<sub>4</sub> could not work. This may be attributed to the extremely fast carbon diffusion at the solid–gas interface. However, this phase of tungsten carbide could be achieved by using crystalline non-volatile solid carbon precursors such as multi-walled carbon nanotube (MWCNTs), RGO, graphene, etc. At that time, various attempts had been made to determine the viability of employing

tungsten carbide and its compounds as electrocatalysts for water splitting reactions. The nanocrystalline thin films based on tungsten carbide has showed excellent catalytic activity towards HER and reported elsewhere by Zheng et al. [84].

### 3.2.1. Various Nanostructures of Tungsten Carbide Catalyst

The 1D tungsten carbide (WC) NWs with high porosity are initially established by Ren Bowen and his co-worker by two-step synthetic protocols involving solution treatment followed by carbonization process. The as-prepared catalyst displays a very good HER catalytic activity with the least onset potential of 39 mV. In acidic medium, the calculated overpotential  $\eta_{10}$  and Tafel slope of 118 mV and 55 mV dec<sup>-1</sup> @ 10 mA cm<sup>-2</sup>. Whereas in alkaline medium, same catalyst delivers the onset potential,  $\eta_{10}$ , and Tafel slopes of 56 mV, 122 mV, and 56 mV dec<sup>-1</sup> respectively indicating its excellent catalytic activity [85]. Likewise, another 1D form of nanocrystalline WC nanowalls were prepared by Jin and his group through a two-step CVD processes. Initially the nanocrystalline diamond layer was fabricated over silicon substrate by HF-CVD process. Later, the WC nanowall was grown over the NCD layer through the direct current plasma CVD process. The prepared catalysts reveals excellent HER performance with 248mV of overpotential, and 67 mV dec<sup>-1</sup> of Tafel slope @ 10mA cm<sup>-2</sup>, in acidic solution [86]. Interestingly, the cage-confinement pyrolysis protocol has been adopted to fabricate ultra-small WC nanoclusters/NPs as reported by Xu Yan-Tong and his co-workers. An RHO-[Zn(eim)<sub>2</sub>] type zeolitic metal azolate with hydrophobic nanocavities (d = 1.84 nm) framework MAF-6, Heim = 2-ethylimidazolate, possessing large nanocages and small apertures (average diameter = 0.76 nm), is chosen to restrain the metal source W(CO)<sub>6</sub>. WC nanocluster exhibits the least overpotential and Tafel slope of 51 mV and 49 mV dec<sup>-1</sup> @ 10 mA cm<sup>-2</sup> respectively. In addition the large exchange current density of 2.4 mA cm<sup>-2</sup> has recorded which is quite high among all the molybdenum/tungsten oriented catalysts in acidic solution [87]. A facile one-step protocol to prepare metastable WC NPs from phosphotungstic acid (H<sub>3</sub>PW<sub>12</sub>O<sub>40</sub>) was established by Takafumi et al. The NPs with an average diameter of 1 to 50 nm dispersed over the C-substrate were prepared through a simple operation series that involved the incorporation of carbon black with H<sub>3</sub>PW<sub>12</sub>O<sub>40</sub> followed by heat treatment at 1000 °C. The produced WC NPs exhibited an outstanding catalytic activity for HER, an onset potential of 147 @ 1 mAcm<sup>-2</sup>, with the least Tafel slope of ~50 mV dec<sup>-1</sup> [88]. Tang Chaoyun and his team investigated the WC-supported Pt electrocatalyst using the microwave-assisted ethylene glycol process. The tungsten carbide-supported Pt electrocatalyst featured a small onset overpotential of ~100 mV @ 1 mA cm<sup>-2</sup>, Tafel slope of 73 mVdec<sup>-1</sup>, and an exchange current density of 0.241 mA cm<sup>-2</sup> [89].

### 3.2.2. Tungsten Carbide and Its Composites

Zheng et al. prepared WC<sub>x</sub>/C composite by a combined solution combustion synthesis (SCS) method integrated with in-situ carbothermal reduction reaction. The WC<sub>x</sub>/C composite exhibited a good HER catalytic activity, with  $\eta_{10}$  of -264 mV and Tafel slope of 85 mV dec<sup>-1</sup> @ 10 mA cm<sup>-2</sup> in 0.5 M H<sub>2</sub>SO<sub>4</sub> [90]. A 2D W<sub>2</sub>C single crystal incorporated graphene (2D WC-graphene) has been prepared through three steps by Mengqi et al. Initially, a Ga droplet is placed on W substrates at the center region of the quartz tube. Secondly, Ga-W substrates was heated to 980–1020 °C at a rate of 30 °C/min under the combined flow of Ar and H<sub>2</sub>; in the last step the substrates were exposed to carbon source (methane (CH<sub>4</sub>)) for 30 min under Ar/H<sub>2</sub> resulting in the growth of 2D WC-graphene in-plane heterostructures. The overpotential of the catalyst is very small (120 mV @ 10 mA cm<sup>-2</sup>) and the determined Tafel slope is 38 mV dec<sup>-1</sup> in an acidic solution [91]. Han Lei and coworkers studied the formation of nitrogen-doped carbon spheres encapsulated by W<sub>2</sub>C nanocrystallites using a combined in situ polymerization and carburization process. They recorded the overpotential of 290 mV in acidic solution and bit higher of 300 mV in alkaline solution @ 10 mA cm<sup>-2</sup> for the resultant catalyst. In addition, the catalyst has exhibited the Tafel slopes of 110 mV dec<sup>-1</sup> and 133 mV dec<sup>-1</sup> in acidic and alkaline media,

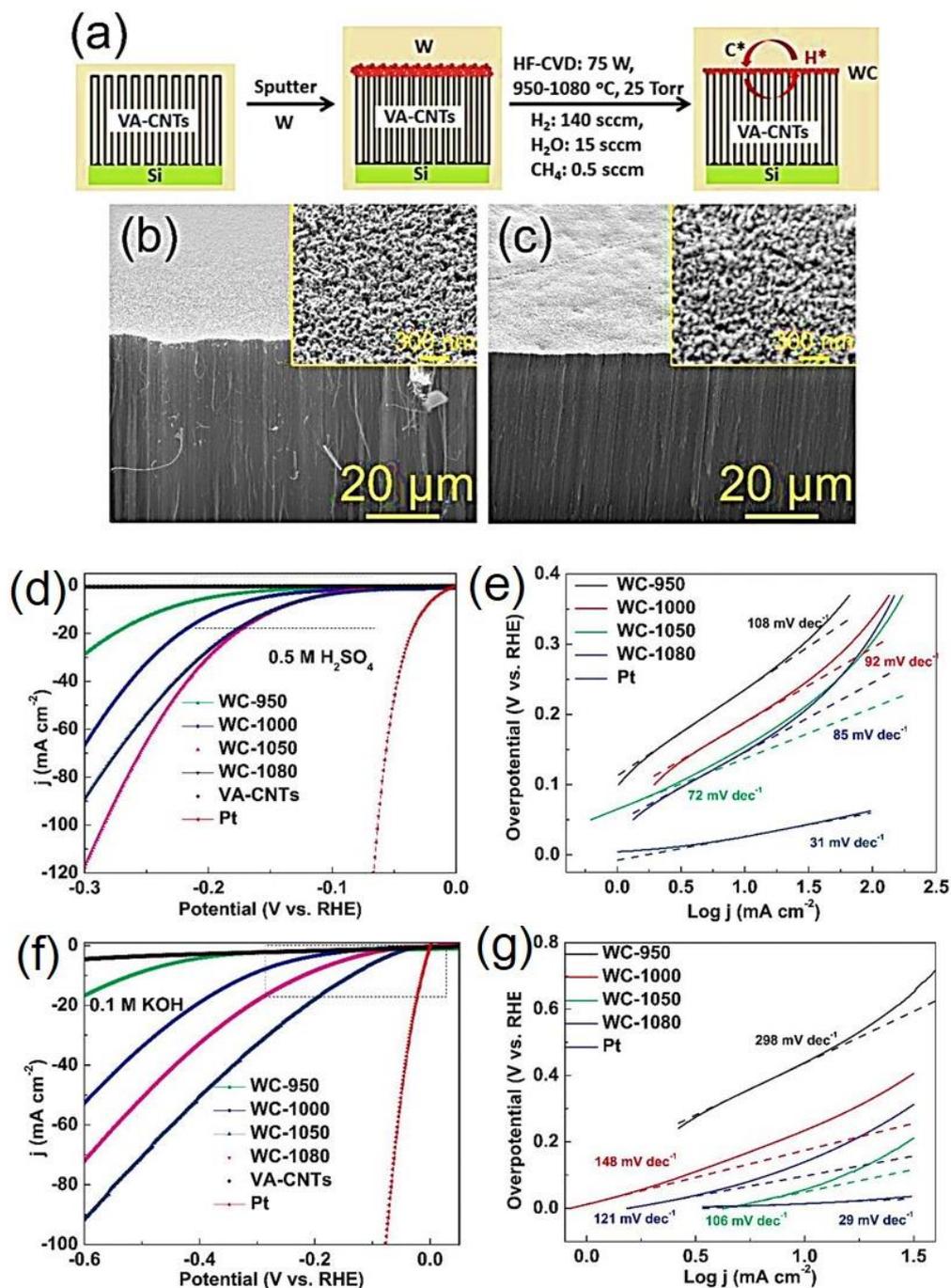
signifying the supportable HER kinetics in acidic medium [92].  $W_2C$  and tungsten disulfide ( $WS_2$ ) have been constructed on vertically aligned MWCNT forests in recent years. Initially an ultra-thin layer of tungsten been placed on vertically aligned MWCNT forests followed by heating at  $770\text{ }^\circ\text{C}$  under inert atmosphere or under sulfur-vapor for getting  $W_2C$  or  $W_2S$ , respectively. The WC-CNT and  $WS_2$ -CNT composites have the overpotentials as low as  $489\text{ mV @ }10\text{ mA cm}^{-2}$ ,  $689\text{ @ }10\text{mA cm}^{-2}$  and the Tafel slope value  $122, 182\text{ mV dec}^{-1}$  in  $0.5\text{ MH}_2\text{SO}_4$  [93]. Qiufang et al. reported a facile two-step methodology to fabricate ultra-small and phase-pure  $W_2C$  particles supported on MWCNTs ( $W_2C$ /MWCNT).  $W_2C$ /MWCNT exhibited impressive electrocatalytic HER performance, featuring a small onset overpotential of  $\sim 50\text{ mV @ }1\text{mA cm}^{-2}$ ,  $123\text{ mV @ }10\text{mA cm}^{-2}$ , Tafel slope of  $45\text{ mV dec}^{-1}$ , and long-term cycling stability [83].

Recently, Xiujun et al. developed WC nanocrystals via sputtering as well as HF-CVD. These NCs were grown over vertically aligned CNTs by a two-step process as displayed in Figure 6 ( $W_2C$ -CNTs). The corresponding SEM image confirming the formation of WC nanocrystals over the VA-CNTs (Figure 6b,c) and it exhibit an excellent catalytic activity for HER in both alkaline and acidic solutions as shown in Figure 6d–f [94]. The overpotentials ( $\eta_{10}$ ) in acid solution drives a current of  $10\text{ mA cm}^{-2}$  of  $145\text{ mV}$ , onset potential of  $15\text{ mV}$ , exchange current density @  $300\text{ mV}$  of  $117.6\text{ mV}$  and Tafel slope values of  $72\text{ mV dec}^{-1}$ . Further, for alkaline media ( $\eta_{10}$ ) of  $137\text{ mV}$ , onset potential of  $16\text{ mV}$ , exchange current density @  $300\text{ mV}$  of  $33.1\text{ mV}$ , and Tafel slope of  $106\text{ mV dec}^{-1}$  were observed [94]. The electrochemical characteristics of various  $W_2C$ /WC-based electrocatalysts for improved HER activity are listed in Table 2.

**Table 2.** Electrochemical characteristics of various  $W_2C$ /WC-based electrocatalysts for improved HER activity.

Catalyst	Synthesis	Morphology	$\eta_j(\text{mV})$	$j(\text{mA cm}^{-2})$	Tafel Slope ( $\text{mV dec}^{-1}$ )	Ref
$W_2C$ /MWCNT	carburization	nanoparticles	123	10	485	[83]
$W_2C$ -thin film	CVD	nanograins	263	10	42.2	[84]
$W_2C$	Plasma assisted carburization	nanowires	118	10	55	[85]
WC	Plasma assisted deposition	nanowall	160	10	67	[86]
WC	Cage confinement pyrolysis	nanoparticle	51	10	49	[87]
$W_2C$	High temperature calcination	nanoparticles	368	20	50	[88]
Pt/WC	High temperature calcination	Spherical particles	22	10	28.8	[89]
$WC_x/C$	Combustion reaction	nanoparticle	264	10	85	[90]
2D WC-G	Liquid metal solvent-based co-segregation strategy	Single crystals	120	10	38	[91]
N-doped WC	In-situ polymerization/carburization	nanospheres	290	10	110	[92]
WC/CNT	CVD	nanoflakes	435	10	103	[93]
$W_2C$	carburization	nanoparticles	50	10	45	[83]
$W_2C$	HF-CVD	nanocrystal	117.6	10	72	[94]

G—graphene; CNT—carbon nanotube; NSs—nanosheets; NC—nitrogen doped carbon; CVD—chemical vapor deposition; HF-CVD—high filament chemical vapor deposition.



**Figure 6.** (a) Schematic of the formation of WC nanocrystals over VA-CNTs, (b,c) Field-emission SEM image of bare VA-CNT and WC nanocrystals over VA-CNT that formed at 1050 °C for 6 h, (d,e) LSV and Tafel slope plots of WC nanocrystals over VA-CNT at different temperatures in an acidic medium, (f,g) LSV and Tafel slope plots of WC nanocrystals over VA-CNT at different temperatures in an alkaline medium [94]. Copyright 2015 The American Chemical Society.

Due to their ideal chemical reactivity and tunability, 2D-TMCs are largely used as the potential supports for nanostructured particle catalysts and that the existence of noble metal nanoparticles favors the reduction of M constituent of the MXene ( $M_{n+1}X_nT_x$ ) and eliminate the surface functional groups effectively [95]. The further reduction of these supports could be led to form metal components in the supports to nanoparticle which in turns to contact with the surface of the supports thereby leading to form a well-ordered intermetallic compound material via reactive metals support interactions (RMSIs) in few cases. RMSI

denotes to the chemical reaction of the support and metal that stimulates the bimetallic nanostructures' formation, which is determined through the superior thermodynamic stability of the subsequent intermetallic compounds. MXenes could easily favor this process by means of its 2D structures with M-C bonds. These M-C bonds are usually weaker than that of M-O bonds in a characteristic oxide support. Owing to this enhanced chemical reactivity, it permits RMSIs to ensue at very low temperature and, so, facilitate the control of size of the particle in contrast to high temperature reduction that is necessary for early bulk carbides and transition metal oxides [96,97]. In few cases, over MXenes intermetallic compounds might be produced that is impossible in conventional carbide or oxide supports and their complete physicochemical properties could be controlled through in-situ chemical reduction at adequate temperature. Forming these types of supports with ideal reactivity and chemical stability is still a challenging task. Recently Li et al. have investigated the formation of Pt<sub>3</sub>Ti and Pt<sub>3</sub>Nb intermetallic compounds through the RMSIs on Pt/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> and Pt/Nb<sub>2</sub>CT<sub>x</sub> catalysts and reported elsewhere [81].

#### 4. Summary and Future Outlook

TMCs especially Mo<sub>2</sub>C and WC/W<sub>2</sub>C are considered as the perfect alternatives to noble metal catalysts, such as Pt, Pd, and Au, for HER, although their HER catalytic activity is restricted. Different synthesis protocols and catalyst design strategies limit the variation in catalytic activities between the noble metal and the Mo<sub>2</sub>C- and WC/W<sub>2</sub>C-based catalysts. Although Mo<sub>x</sub>C and WC/W<sub>2</sub>C-based compounds are well known catalysts in chemical engineering, the research on their function as electrocatalysts is still in its primary stage. This review is focused on the recent development of Mo<sub>x</sub>C and WC/W<sub>2</sub>C-based electrocatalysts for HER activity. One of the major concerns in designing and fabricating the Mo<sub>x</sub>C and WC/W<sub>2</sub>C-based electrocatalysts is the low-density of active sites. Moreover, structural collapse and particle agglomeration occur at high carburization temperatures, leading to the formation of bulk materials. To overcome these drawbacks, diverse nanostructure engineering methods have been reported. Doping and nanostructuring have emerged as vital tools for altering the properties and behaviors of these catalysts. However, issues, such as mass production of catalysts, understanding the mechanism of HER activity, and controlling the nanostructure, still need to be addressed.

Primarily, the excellent HER catalytic activity of the recent metal carbide catalysts alone is not suitable for their practicability. Various synthetic protocols have been developed to control the composition and nanostructure of these electrocatalysts, and further enhancements in the catalytic activity of these catalysts are required. Doping heterometal atoms or carbon supports, such as CNT, RGO, and graphene, could efficiently enhance the performance of the Mo<sub>2</sub>C and WC/W<sub>2</sub>C-based catalysts. Recently, MOF-derived catalyst designing strategies have been in the limelight and could be prolonged to fabricate Mo<sub>2</sub>C and WC/W<sub>2</sub>C-based electrocatalysts with refined stereo-structures and components. Emerging innovative or new-fangled templates can also support the construction of hierarchical nanostructures, including hollow structured or core-shell WC/W<sub>2</sub>C and Mo<sub>2</sub>C-based composites. In addition, these metal carbides can be integrated with other potential efficient compounds such as metal sulfides, phosphides, and selenides to make multicomponent catalyst systems to improve the overall catalytic activity.

Second, extensive and widespread efforts are required to merge the experimental and theoretical evaluations in order to understand the correlation between the physicochemical structures and the electrochemical behaviors of the Mo<sub>2</sub>C and WC/W<sub>2</sub>C-electrocatalysts. Advancements in catalyst structures and examination of the active site origins can aid in the atomic-scale study of the growth and nucleation processes of metal carbide catalysts. More systematic and detailed theoretical studies on the electrochemical HER activities of the Mo<sub>2</sub>C and WC/W<sub>2</sub>C-based electrocatalysts are necessary. These attempts and efforts would guide the scientific community to better tune the chemical and physical properties of the Mo<sub>2</sub>C and WC/W<sub>2</sub>C-electrocatalysts.

In a nutshell, the mass production of hydrogen from water electrolysis is known to be a highly challenging task, and we still need to go a long way to achieve this goal. It majorly depends on the accessibility of cost effective electrocatalysts. However, the reported well-designed, morphology controlling synthesis methods may not be appropriate for the large-scale production of such catalysts. Hence, the growth of cost-effective and scalable processes for large-scale fabrication of Mo<sub>2</sub>C and WC/W<sub>2</sub>C-based electrocatalysts are essential.

**Author Contributions:** Conceptualization, methodology and writing—original draft preparation, K.K.; resources and writing original draft, A.N.; methodology, writing—original draft preparation and resources, D.V.; software, J.-H.C.; resources, S.H.; writing—original draft preparation, C.A.; writing—review and editing, R.B.; writing—review and editing, A.A.; conceptualization, visualization, supervision, project administration and funding acquisition H.-S.K. All authors have read and agreed to the published version of the manuscript.

**Funding:** This work was supported by the research program of Dongguk University in 2022 (No. S-2022-G0001-00016).

**Data Availability Statement:** Not applicable.

**Conflicts of Interest:** The authors declare no competing financial interests.

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