



## Article

# Effect of Covalent Organic Frameworks Containing Different Groups on Properties of Sulfonated Poly(ether ether ketone) Matrix Proton Exchange Membranes

Xiaoyu Meng<sup>1,2,\*</sup>, Yinan Lv<sup>1</sup>, Lei Ding<sup>3</sup>, Luman Peng<sup>1</sup>, Qiwang Peng<sup>1</sup>, Chuanbo Cong<sup>1,2</sup>, Haimu Ye<sup>1,2</sup>   
and Qiong Zhou<sup>1,2</sup>

<sup>1</sup> Department of Materials Science and Engineering, College of New Energy and Materials, China University of Petroleum, Beijing 102249, China

<sup>2</sup> Beijing Key Laboratory of Failure, Corrosion, and Protection of Oil/Gas Facilities, China University of Petroleum, Beijing 102249, China

<sup>3</sup> CSSC Systems Engineering Research Institute, Beijing 100036, China

\* Correspondence: mengxy@cup.edu.cn

**Abstract:** The rich  $-\text{SO}_3\text{H}$  groups enable sulfonated poly (ether ether ketone) (SPEEK) to possess excellent proton conductivities in proton exchange membrane (PEM), but cause excessive water absorption, resulting in the decline of dimensional stability. It is a challenge to resolve the conflict between conductivity and stability. Owing to its unique structural designability, covalent organic frameworks (COFs) have been used to regulate the performances of PEMs. The authors propose the use of COFs with acidic and basic groups for meeting the requirements of proton conductivity and dimensional stability. Herein, COFs containing different groups (sulfoacid, pyridine, and both) were uniformly dispersed into the SPEEK matrix by in situ synthesis, and the effects on the properties of SPEEK matrix PEMs were revealed. The sulfoacid group significantly improves proton conductivities. At 60 °C, under 95% RH, the conductivity of the SPEEK/TpPa- $\text{SO}_3\text{H}$ -20 composite membrane was  $443.6 \text{ mS}\cdot\text{cm}^{-1}$ , which was 3.3 times that of the pristine SPEEK membrane. The pyridine group reduced the swelling ratio at 50 °C from 220.7% to 2.4%, indicating an enhancement in dimensional stability. Combining the benefits of sulfoacid and pyridine groups, SPEEK/TpPa-( $\text{SO}_3\text{H}$ -Py) composite membrane has a conductivity of  $360.3 \text{ mS}\cdot\text{cm}^{-1}$  at 60 °C and 95% RH, which is 1.86 times that of SPEEK, and its swelling ratio is 11.8%, about 1/20 of that of SPEEK membrane. The method of in situ combination and regulation of groups open up a way for the development of SPEEK/COFs composite PEMs.

**Keywords:** PEMs; SPEEK; proton conductivities; dimensional stabilities; covalent organic frameworks



**Citation:** Meng, X.; Lv, Y.; Ding, L.; Peng, L.; Peng, Q.; Cong, C.; Ye, H.; Zhou, Q. Effect of Covalent Organic Frameworks Containing Different Groups on Properties of Sulfonated Poly(ether ether ketone) Matrix Proton Exchange Membranes. *Nanomaterials* **2022**, *12*, 3518. <https://doi.org/10.3390/nano12193518>

Academic Editor: Félix Zamora

Received: 15 September 2022

Accepted: 6 October 2022

Published: 8 October 2022

**Publisher's Note:** MDPI stays neutral with regard to jurisdictional claims in published maps and institutional affiliations.



**Copyright:** © 2022 by the authors. Licensee MDPI, Basel, Switzerland. This article is an open access article distributed under the terms and conditions of the Creative Commons Attribution (CC BY) license (<https://creativecommons.org/licenses/by/4.0/>).

## 1. Introduction

Proton exchange membrane fuel cells (PEMFCs) have the advantages of high energy density, low operating temperature, fast start-up speed, and simple structure [1–3]. Playing an important role in PEMFCs, proton exchange membranes (PEMs) need to have high proton conductivity, excellent thermal stability, and mechanical stability [4]. As the most widely used commercially available PEMs, Nafion is a fluoropolymer made of sulfonated polytetrafluoroethylene, which has high proton conductivities [5]. However, problems such as high cost, complex preparation process, and reduced dimensional stabilities, proton conductivities in high-temperature environments, and unstable relative humidity (RH) limit the further application of Nafion [5,6]. Therefore, it has become a research hotspot to find alternative materials for PEMs with low prices and great properties.

Sulfonated polyether ether ketone (SPEEK), as a potential PEM material, has excellent mechanical properties, and a chemical stability [7]. Similar to the Nafion microphase structure, SPEEK is composed of hydrophilic domains and hydrophobic domains. As

the hydrophobic domains, the PEEK skeleton improves mechanical properties [4]. The hydrophilic domains of SPEEK are composed of  $-\text{SO}_3\text{H}$  side groups, which play the role of proton conductors. The hydrophilic domains ( $-\text{SO}_3\text{H}$  and water) form spherical ion clusters. As the water content increases, the number of spherical ion clusters increases, and these spherical ion clusters connect to form connected proton transport channels [8–11]. SPEEK with a high degree of sulfonation (DS) is more likely to form ion clusters and has higher proton conductivities [12]. However, with it comes swelling caused by excessive water absorption, especially at higher service temperatures. This reduces the dimensional stability of the PEMs and also limits the operating temperature range of SPEEK-based PEMs. To overcome this shortcoming, some second phases are added to the SPEEK matrix, such as graphene oxide (GO), metal–organic frameworks (MOFs), and covalent organic frameworks (COFs). More active groups can be introduced into SPEEK ionomers as proton conductors to improve the connectivity of ion clusters, reduce the activation energy required for proton transport, and effectively improve proton conductivity. Therefore, a stable and effective carrier is needed to introduce active groups into the SPEEK matrix.

More second-phase particle-doped polymer PEMs have been reported. If there is no ideal compatibility between the doped particles and the polymer matrix, the phase separation phenomenon will occur, which will make work unsatisfactory. The excellent compatibility between covalent organic frameworks (COFs) and polymer matrices makes them have great potential for application in PEMs [13]. COFs are a type of nanoparticles that connect organic monomers through strong covalent bonds and are formed by orderly stacking by non-covalent interactions. Compared with the traditional crystalline porous structure, COFs can accurately pre-design the topological structure and design functional COFs to realize the structure and chemical control of specific functions [14–16]. In addition, diverse structural units, stable porosity, and thermal stability enable COFs to be used in gas storage and separation [17], catalysis [18], sensing [19], photoelectric [20], and proton conduction [21]. Using molecular design and structural optimization, the content, and location of acidic groups and basic groups in COFs materials can be precisely controlled. Peng et al. [22] successfully prepared fully sulfonated COFs (NUS–10) by reacting amino monomers containing sulfonate groups with Tp. It was found that the proton conductivity of TpPa–1 (which had the same structure but does not contain sulfonic acid groups) at room temperature and 97% RH was only  $2.40 \times 10^{-5} \text{ S}\cdot\text{cm}^{-1}$ , which was almost negligible. The proton conductivity of NUS–10 was as high as  $3.96 \times 10^{-2} \text{ S}\cdot\text{cm}^{-1}$  under the same conditions. Adding acidic groups in the confined nanospace can effectively increase the proton conductivity of COFs. The participation of basic groups will allow protons to be transported along the hydrogen bond network composed of acids and bases, thereby improving the efficiency of proton transfer [23]. TpPa–( $\text{SO}_3\text{H}$ -Py) containing both sulfonic acid groups and pyridine groups were successfully synthesized, in which the sulfonic acid groups were used as the acidic sites and the pyridyl groups were used as the basic sites, and ideal proton conductivities were obtained [24].

Based on the unique molecular structure of the porous material COFs, it provides more options for improving the performance of PEMs. COFs exist in the form of a filler in polymer, which can not only form acid–base interactions with the polymer to obtain a high proton conductive electrolyte material but also inhibit the swelling behavior of the polymer in solution. However, COFs are usually presented in the form of fluffy microcrystalline powder, with a hard texture and low solubility, making it difficult to post-process COFs powder. Therefore, how to uniformly disperse COFs powder into a polymer matrix is an urgent problem to be solved. To improve the compatibility of COFs and polymer matrix, Yin et al. [25] chose a Schiff base COFs network (SNW–1) with inherent amino groups and introduced it into the Nafion matrix to prepare composite membranes. The  $-\text{N}-\text{H}-$  bond in the SNW–1 structure can interact with  $-\text{SO}_3\text{H}$  in Nafion to improve the dispersion of SNW–1 in the matrix. Fan et al. [26] doped triazole functionalized COFs into SPEEK matrix and used triazole groups and  $-\text{SO}_3\text{H}$  groups to form acid–base pairs at the interface to obtain uniformly dispersed composite membranes. The proton conductivity of the

SPEEK/HPW@COF composite membrane at 65 °C and 40% RH was 35.5 times that of SPEEK. In general, these methods have obvious limitations on the choice of COFs.

Herein, we improved the dispersion of COFs in the SPEEK matrix by in situ synthesis, thereby improving the proton conductivity and mechanical properties of PEMs. At the same time, the performances of the PEMs were further improved by adjusting the COFs' frame structure. We prepared TpPa–SO<sub>3</sub>H and TpPa–Py by the in situ synthesis in the SPEEK matrix, respectively. TpPa–SO<sub>3</sub>H can provide additional proton conduction sites for proton conduction and improve the proton conductivity of PEMs. The pyridine group of TpPa–Py had a strong interaction with the –SO<sub>3</sub>H group of the SPEEK matrix and improved the mechanical stability of PEMs. The –SO<sub>3</sub>H and pyridine groups had excellent performance in improving the proton conductivity and mechanical stability of PEMs, respectively. Therefore, the –SO<sub>3</sub>H and pyridine groups were introduced into the COFs structure at the same time to improve the comprehensive performance of the PEMs.

## 2. Materials and Methods

### 2.1. Materials and Chemicals

Poly (ether ether ketone) (VICTREX<sup>®</sup> 450 PF) was produced by Victrex (Lancashire, England); concentrated H<sub>2</sub>SO<sub>4</sub> was used as a sulfonating agent to achieve SPEEK as described in previous work [27]. 1,3,5-triformylphloroglucinol (Tp), 2,5-diaminobenzene-sulfonic acid (Pa–SO<sub>3</sub>H), 2,5-diaminopyridine (Pa–Py) were purchased from Sigma-Aldrich Co., Ltd. (Shanghai, China). All the other reagents mentioned in this work were produced by Beijing chemical reagent factory (Beijing, China).

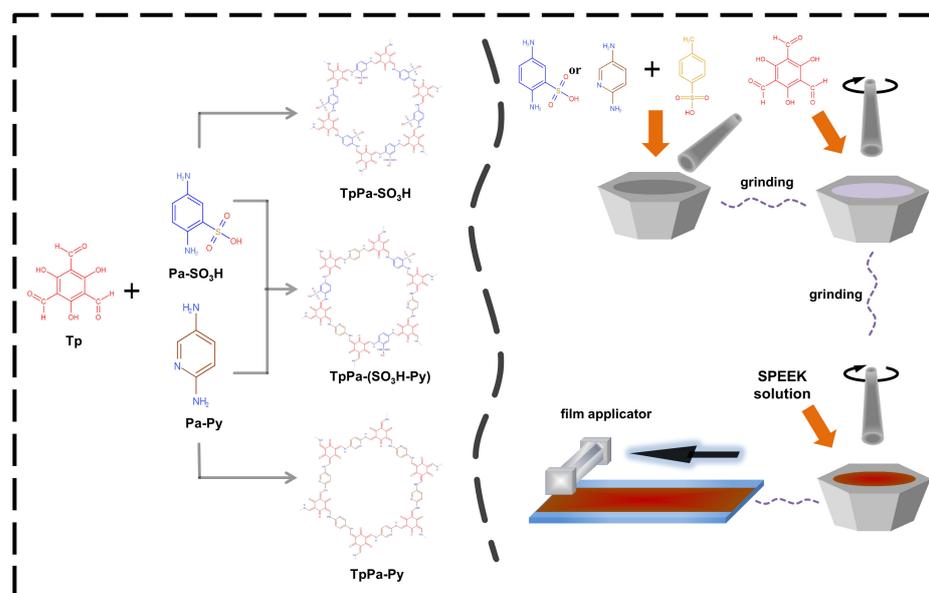
### 2.2. Synthesis of SPEEK

PEEK powder (25 g) was added into concentrated H<sub>2</sub>SO<sub>4</sub> (250 mL), which was stirred at 50 °C. After PEEK was completely dissolved into H<sub>2</sub>SO<sub>4</sub>, the solution was vigorously stirred for another 4 h 10 min. The reaction was terminated by pouring the sulfonated polymer solution into ice water under mechanical stirring. The precipitated SPEEK was repeatedly washed using deionized (DI) water until the pH of the rinse water was approximately 7. The SPEEK material was obtained after stirring again for 2 h and drying in a blast oven at 60 °C for 48–72 h.

### 2.3. Preparation of SPEEK/COF Composite Membranes

The SPEEK was added to DMSO, stirred at 60 °C until SPEEK was dissolved, and formed a SPEEK/DMSO solution with a mass concentration of 20%. A 20 wt% SPEEK/DMSO (4.66 mL) solution was cast in a glass mold and kept in an oven under 80 °C for 72 h to prepare the SPEEK membrane. Pa–SO<sub>3</sub>H and PTSA were placed in a mortar and ground for 5–10 min to form a uniform mixed powder. Tp was added to the above-mixed powder and grinding continued for 5 min. Then, deionized water (150 µL) was added, and ground for 20 min to make a uniform slurry. SPEEK/DMSO solution was added to the above slurry and ground for about 30 min to make it uniformly mixed to obtain a film-forming liquid. The proportion of each raw material is shown in the supporting information. The film-forming solution was coated on a glass plate and dried in a blast drying oven at 80 °C for 72 h to obtain an in situ synthesized SPEEK/TpPa–SO<sub>3</sub>H composite PEM.

The synthesis process of TpPa–Py and TpPa–(SO<sub>3</sub>H-Py) is the same as the synthesis process of TpPa–SO<sub>3</sub>H, and the usage ratio of each raw material is shown in Tables S1–S3. The monomers and process diagram used to prepare composite membranes are shown in Scheme 1.



**Scheme 1.** Synthetic route of SPEEK/COFs composite membranes.

#### 2.4. Characterizations

X-ray diffraction (XRD, Bruker D8Focus, Bremen, Germany) was utilized to analyze the physical topology crystallization of COF nanoparticles. The chemical composition of COF nanoparticles was characterized by Fourier transform infrared (FTIR, Bruker Tensor II, Bremen, Germany). The thermal stability of SPEEK and SPEEK/COF membranes were evaluated through thermal gravimetric analysis under the nitrogen atmosphere in the range of 40 °C to 800 °C (TGA, TA Q5000, New Castle, DE, USA) (Figure S1). The scanning electron microscope (SEM, HITACHI SU8010, Tokyo, Japan) and transmission electron microscopy (TEM, FEI F20, Oregon, USA) were utilized to characterize the COF feature and the cross-section morphologies of as-prepared membranes. Solid-state nuclear magnetic resonance ( $^{13}\text{C}$  SSNMR, JEOL JNM-ECZ600R, Japan) was used to characterize the structure of COF. The mechanical properties of the composite membranes were measured on a universal material testing machine (KQL, WDL-10, Yangzhou, China) at a strain rate of 10 mm/min.

#### 2.5. Water Uptaking and Swelling Ratio

The dry membranes were cut into rectangle samples (about 1 cm  $\times$  3 cm) and its mass ( $W_{dry}$ ) and area ( $S_{dry}$ ) were recorded. Then, the samples were immersed in deionized water at different temperatures for 12 h, and the moisture on the surface of the membrane was quickly removed, and then the wet weight ( $W_{wet}$ ) and wet area ( $S_{wet}$ ) of the samples were measured. Each sample was measured at least three times and the average value was calculated. The water absorption and swelling degree of the film is calculated by the following formula:

$$\text{Water Uptaking} = \frac{(W_{wet} - W_{dry})}{(W_{dry})} \times 100\% \quad (1)$$

$$\text{Area swelling ratio} = \frac{(S_{wet} - S_{dry})}{(S_{dry})} \times 100\% \quad (2)$$

### 2.6. Proton Conductivity

The proton conductivity was characterized by CHI660E. The test was carried out under different temperature and relative humidity (RH) conditions in the YSGTS-50 constant temperature and humidity test chamber. All membrane samples were cut into 1.5 cm × 3 cm sample strips before the test and placed in a constant temperature and humidity test chamber to be sufficiently stable under the tested conditions. The proton conductivity of the membrane was calculated according to Equation (3):

$$\sigma = \frac{L}{AR} \quad (3)$$

where  $\sigma$  represents the proton conductivity of the membrane ( $\text{mS}\cdot\text{cm}^{-1}$ ),  $L$  is the distance between the two platinum electrodes (cm),  $R$  ( $\Omega$ ) is the resistance value of the membrane, and  $A$  is the cross-sectional area of samples.

### 2.7. Ion Exchange Capacity (IEC) and Degree of Sulfonation (DS)

The ion exchange capacity (IEC) was determined by trans titration method. The membranes were soaked in 2 mol·L<sup>-1</sup> NaCl solution for 24 h, the purpose was to completely replace the H<sup>+</sup> in the -SO<sub>3</sub>H groups of membranes with Na<sup>+</sup> in the NaCl solution. Phenolphthalein solution was used as indicator, and the replaced acidic NaCl solutions were titrated with 0.01 mol·L<sup>-1</sup> NaOH solution. The calculation formula of IEC is according to Equation (4):

$$IEC = \frac{C_{NaOH} \times V_{NaOH}}{M_{dry}} \quad (4)$$

where  $V_{NaOH}$  is the volume of NaOH solution used for titration;  $C_{NaOH}$  is the concentration of NaOH solution;  $M_{dry}$  is the mass of dry samples of membranes used for measurement.

The degree of sulfonation (DS) of the SPEEK pristine membrane can be calculated from the theoretical value of IEC with the following formula:

$$DS = \frac{288 \cdot IEC}{1000 - 102 \cdot IEC} \times 100\% \quad (5)$$

## 3. Results and Discussion

### 3.1. SPEEK/TpPa-SO<sub>3</sub>H and SPEEK/TpPa-Py Composite Membranes

#### 3.1.1. Characterization of TpPa-SO<sub>3</sub>H and TpPa-Py

TpPa-SO<sub>3</sub>H and TpPa-Py were selected to be synthesized in situ in SPEEK matrix to prepare PEMs. To characterize the synthesis of TpPa-SO<sub>3</sub>H and TpPa-Py in the membranes, the composite membranes were dissolved several times in DMSO solvent at 80 °C and centrifuged, and then the COFs powders were separated from it. The morphology of TpPa-SO<sub>3</sub>H and TpPa-Py were obtained by SEM and TEM (Figure 1). Consistent with previously reported in the literature [24,28], the morphology of TpPa-SO<sub>3</sub>H was spherical, and the diameter of each spherical particle is about 100 nm. The morphology of TpPa-Py was fibrous, which is also nanoscale in size. Compared with the SPEEK matrix, the characteristic element of COF is N. It can be seen from the EDS results that the washed powder contained the N element, and the distribution of the N element in the powder was uniform.

The FTIR spectrum (Figure 2a,d) confirmed the successful synthesis of TpPa-SO<sub>3</sub>H and TpPa-Py in SPEEK matrix. The complete consumption of the raw materials was indicated by the absence of the characteristic N-H stretching band (3428–3338 cm<sup>-1</sup>) of the free diamine and aldehydic -CH=O (1643 cm<sup>-1</sup>) of the free Tp. The absorption band at 2893 cm<sup>-1</sup> corresponded to the C-H stretching peak of the Tp aldehyde group [29]. The nearby peaks at 1590 cm<sup>-1</sup> and 1230 cm<sup>-1</sup> were attributed to the stretching vibration peaks of C=C and C-N, which proved the enol-ketone mutual variation structured [30].

The peaks at  $1084\text{ cm}^{-1}$  and  $1027\text{ cm}^{-1}$  were attributed to the stretching vibration peaks of sulfonate  $\text{O}=\text{S}=\text{O}$  on  $\text{TpPa}-\text{SO}_3\text{H}$  [31].

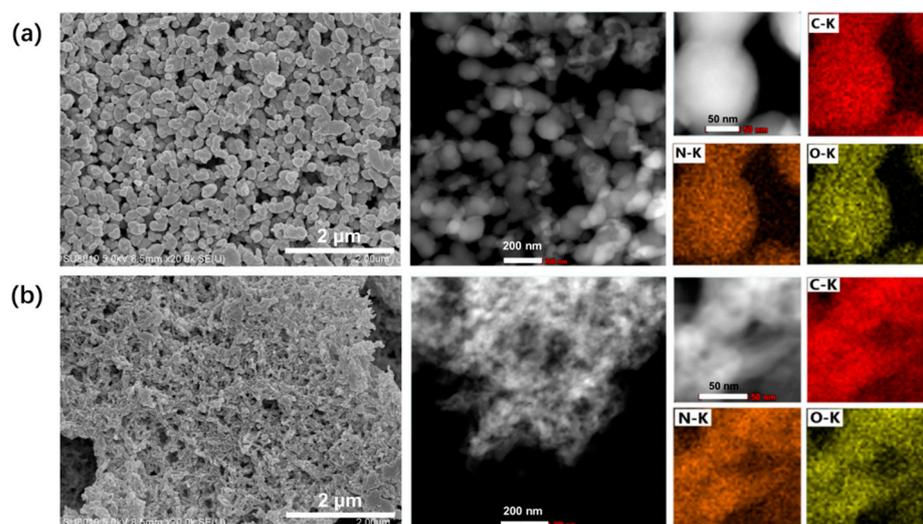


Figure 1. SEM, TEM and EDS-mapping of (a)  $\text{TpPa}-\text{SO}_3\text{H}$  and (b)  $\text{TpPa}-\text{Py}$ .

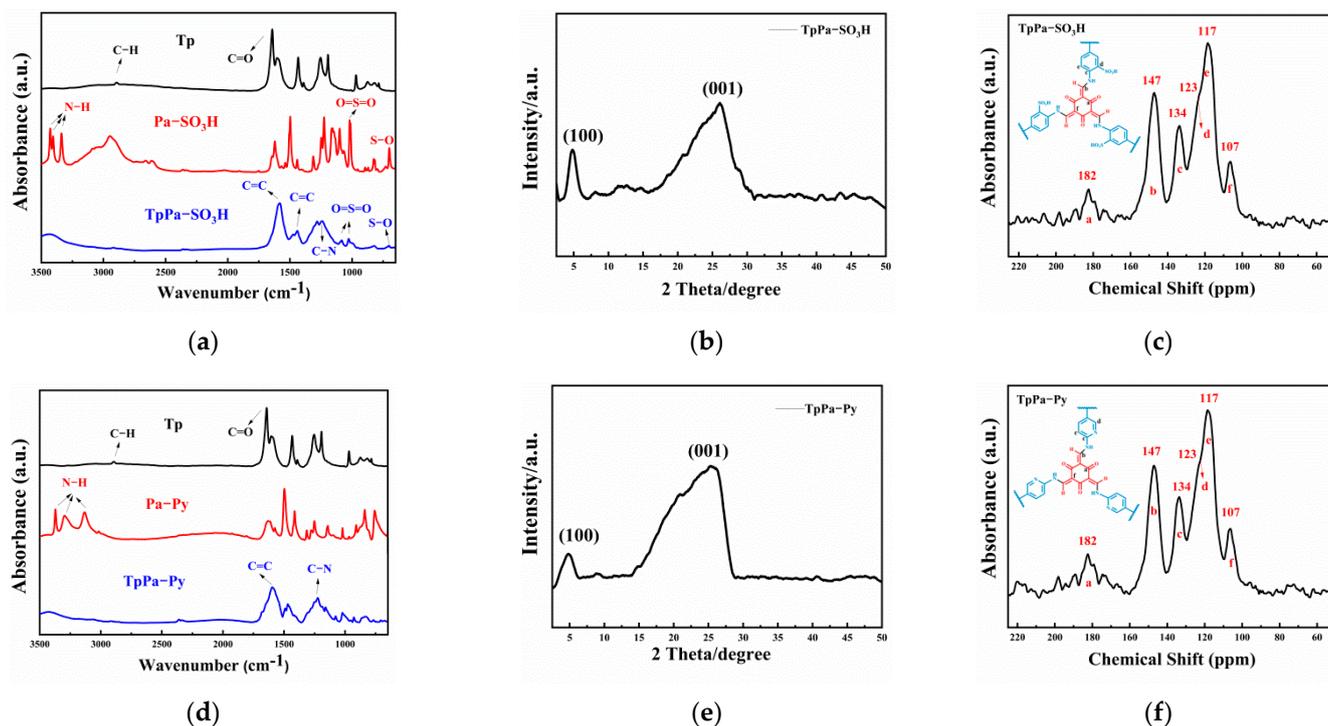


Figure 2. (a) FTIR, (b) XRD, (c) NMR spectra of  $\text{TpPa}-\text{SO}_3\text{H}$  and (d) FTIR, (e) XRD, (f) NMR spectra of  $\text{TpPa}-\text{Py}$ .

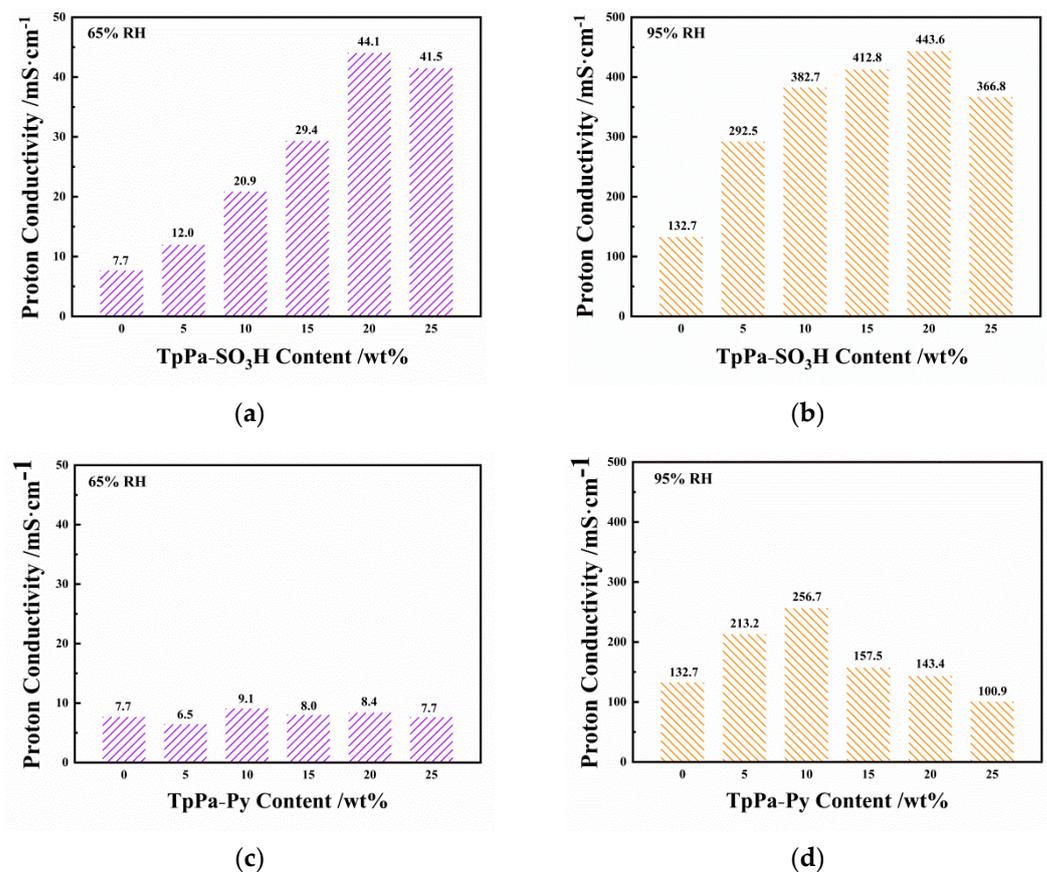
The ordered structure of  $\text{TpPa}-\text{SO}_3\text{H}$  and  $\text{TpPa}-\text{Py}$  was determined by XRD (Figure 2b,e). There was a diffraction peak at  $2\theta = 4.9^\circ$ , corresponding to the reflection of the (100) plane of the  $\text{TpPa}-\text{SO}_3\text{H}$  lattice [32]. Due to the reflection from the (001) plane, a broad peak appeared near  $26^\circ$ , indicating that there was a  $\pi$ - $\pi$  stacking between COF single molecules [33]. The diffraction peak at  $4.9^\circ$  corresponded to the reflection peak of the  $\text{TpPa}-\text{Py}$  (100) plane. The broad peak at  $26.0^\circ$  at the (001) plane confirmed that  $\text{TpPa}-\text{Py}$  was formed in the form of crystals and  $\pi$ - $\pi$  stacks [24].

To analyze the structure of the TpPa–SO<sub>3</sub>H and TpPa–Py samples separated from composite membranes, a <sup>13</sup>C SSNMR test was performed (Figure 2c,f). The carbonyl carbon peak of TpPa–SO<sub>3</sub>H and TpPa–Py at 182 ppm indicated the presence of typical keto-enol tautomers. The signals near 123 ppm and 117 ppm correspond to carbon atoms near the sulfonic acid group in TpPa–SO<sub>3</sub>H and the pyridine group in TpPa–Py. Moreover, the peak intensity at 117 ppm is significantly higher than other peaks, which indicated that the carbon in the residual SPEEK in the sample also participates in the resonance. All the results of SEM, TEM, FTIR, XRD, and <sup>13</sup>C SSNMR showed that TpPa–SO<sub>3</sub>H and TpPa–Py were successfully synthesized in the composite membranes.

### 3.1.2. Proton Conductivities of SPEEK/TpPa–SO<sub>3</sub>H and SPEEK/TpPa–Py Composite Membranes

Proton conductivity is an important factor in evaluating the performance of proton exchange membranes. The proton conductivity of SPEEK, SPEEK/TpPa–SO<sub>3</sub>H, and SPEEK/TpPa–Py composite membranes at 60 °C, under 65% RH and 95% RH was shown in Figure 3, respectively. After adding TpPa–SO<sub>3</sub>H and TpPa–Py nanoparticles, the proton conductivity of the SPEEK/COF membranes continued to increase with the addition of COF nanoparticles and reached the maximum value at 20 wt% and 10 wt%, respectively. Under the humidity of 65% RH and 95%, the satisfactory proton conductivity exhibited by the SPEEK/TpPa–SO<sub>3</sub>H-20 composite membrane which was 44.1 mS·cm<sup>−1</sup> and 443.6 mS·cm<sup>−1</sup>, 5.7 times and 3.3 times than SPEEK membranes, respectively. Meanwhile, the higher proton conductivity exhibited by the SPEEK/TpPa–Py-10 composite membrane was 9.1 mS·cm<sup>−1</sup> and 256.7 mS·cm<sup>−1</sup>, 1.2 times and 1.9 times of SPEEK membranes, respectively. The decline of proton conductivity appeared at the addition content of 25 wt% TpPa–SO<sub>3</sub>H and 15 wt% TpPa–Py. TpPa–SO<sub>3</sub>H and TpPa–Py agglomerated in the SPEEK matrix making it difficult to form a connected proton channel in the membrane. Protons are mainly transferred through water or acidic sites. The framework of TpPa–SO<sub>3</sub>H contains both the acidic group –SO<sub>3</sub>H and the basic group –C–N–, which can improve the water retention of the composite membrane under low humidity and accelerate the dissociation of the proton donor (–SO<sub>3</sub>H). Moreover, the good dispersion of TpPa–SO<sub>3</sub>H provided ordered sulfonic acid groups, which increased the efficiency of proton conduction, and promoted the rapid transmission of protons. Therefore, the SPEEK/TpPa–SO<sub>3</sub>H composite membranes exhibited excellent proton conductivity in both low humidity and high humidity. Compared with the acidic sulfonic acid group, the basic pyridine group in TpPa–Py has weaker hydrophilicity. However, the pyridine group and the sulfonic acid group can form acid–base pairs to improve the efficiency of proton transfer. The above reasons make SPEEK/TpPa–SO<sub>3</sub>H and SPEEK/TpPa–Py membranes had considerable proton conductivity in high- and low-humidity environments. Compared with TpPa–Py, TpPa–SO<sub>3</sub>H had a more obvious effect on improving the proton conductivity.

After testing, it can be seen that the DS of the SPEEK used in the experiment was about 47.3%. According to Table 1, it can be seen that the IEC change trend of the SPEEK/TpPa–SO<sub>3</sub>H composite membranes is consistent with the proton conductivities. The IEC increased with the increase in the TpPa–SO<sub>3</sub>H content. When the TpPa–SO<sub>3</sub>H content is 20 wt%, the IEC reached 2.26 mmol·g<sup>−1</sup>, and the IEC reached the highest of the SPEEK/TpPa–SO<sub>3</sub>H composite membranes. When the amount of TpPa–SO<sub>3</sub>H content increased to 25 wt%, the IEC decreased slightly. The IEC of the SPEEK/TpPa–Py composite membranes is generally lower than that of the SPEEK/TpPa–SO<sub>3</sub>H composite membranes, reaching 1.61 mmol·g<sup>−1</sup> when the TpPa–Py content was 10 wt%, which was the highest value in the SPEEK/TpPa–Py composite membranes.



**Figure 3.** Proton conductivity at 60 °C, (a,c) 65% RH and (b,d) 95% RH of SPEEK, SPEEK/TpPa–SO<sub>3</sub>H and SPEEK/TpPa–Py.

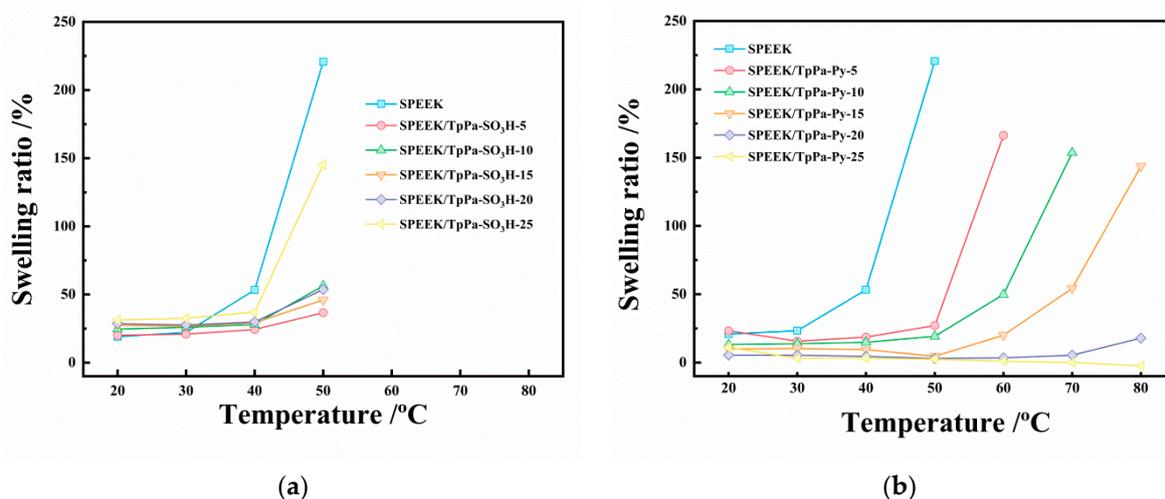
**Table 1.** IEC of the membranes and DS of the SPEEK.

| Membranes                         | IEC (mmol·g <sup>-1</sup> ) | DS (%) |
|-----------------------------------|-----------------------------|--------|
| SPEEK                             | 1.41                        | 47.3   |
| SPEEK/TpPa–SO <sub>3</sub> H-5    | 2.04                        | -      |
| SPEEK/TpPa–SO <sub>3</sub> H-10   | 2.11                        | -      |
| SPEEK/TpPa–SO <sub>3</sub> H-15   | 2.16                        | -      |
| SPEEK/TpPa–SO <sub>3</sub> H-20   | 2.26                        | -      |
| SPEEK/TpPa–SO <sub>3</sub> H-25   | 2.11                        | -      |
| SPEEK/TpPa–Py-5                   | 1.50                        | -      |
| SPEEK/TpPa–Py-10                  | 1.61                        | -      |
| SPEEK/TpPa–Py-15                  | 1.45                        | -      |
| SPEEK/TpPa–Py-20                  | 1.43                        | -      |
| SPEEK/TpPa–Py-25                  | 1.34                        | -      |
| SPEEK/TpPa–(SO <sub>3</sub> H-Py) | 2.09                        | -      |

### 3.1.3. Dimensional Stability

Swelling is an important parameter to characterize the performance of PEM. Water uptake is closely related to the free volume and hydrophilicity of the membrane and plays an important role in the transport of protons (Figure S3a,b). However, excessive swelling caused by excessive water uptake may destroy the chemical stability of the membrane. The swelling ratios of the composite membranes at 20 °C–50 °C were shown in Figure 4. The swelling ratio of the SPEEK membrane at 50 °C was as high as 220.7%. While the swelling ratio of SPEEK/TpPa–SO<sub>3</sub>H-20 at 50 °C was 53.7%, which was only about 1/4 of SPEEK. Owing to the hydrophilic groups –SO<sub>3</sub>H attached to the TpPa–SO<sub>3</sub>H skeleton, the SPEEK/TpPa–SO<sub>3</sub>H composite membrane can combine with more water molecules at

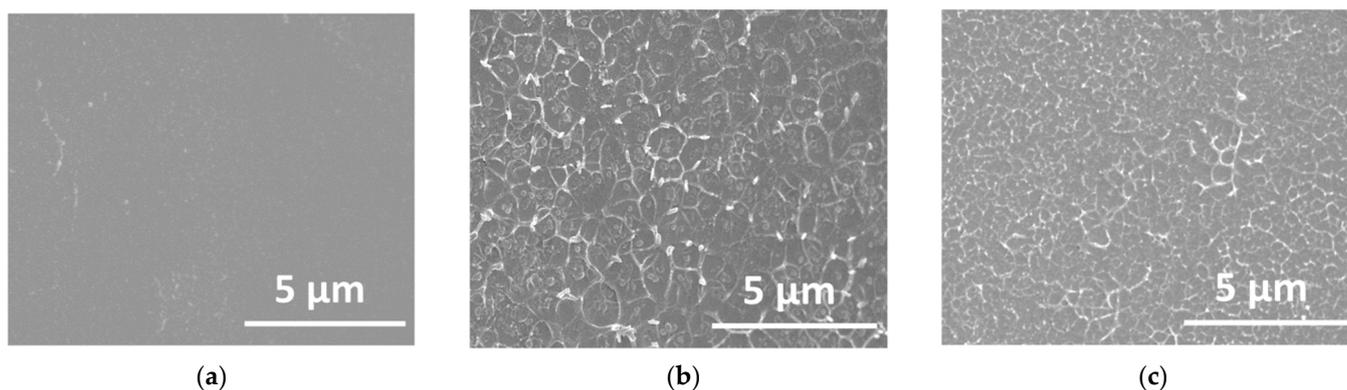
low temperatures, which improves the proton conductivity of the composite membrane, but the swelling ratio of the composite membrane will also increase with the increase in temperature. Due to the interaction such as the hydrogen bond between TpPa–SO<sub>3</sub>H and the polymer chain, which limits the migration of the polymer chain, the addition of TpPa–SO<sub>3</sub>H to the SPEEK matrix can significantly inhibit the swelling of the composite membranes, especially at high temperatures. Compared with TpPa–SO<sub>3</sub>H, TpPa–Py had a greater effect on improving the dimensional stability of PEM. The swelling ratio of SPEEK/TpPa–Py composite membranes was lower than that of SPEEK membranes. In the hydrated state at 50 °C, the area swelling ratio of SPEEK/TpPa–Py series membranes was in the range of 3–27.1%. With the content of TpPa–Py increased, the swelling ratio of the composite membranes decreased. When the temperature was higher than 50 °C, membranes absorbed water and swelled excessively, and the sizes of composite membranes cannot be measured. However, the composite membranes of SPEEK/TpPa–Py-20 and SPEEK/TpPa–Py-25 can still maintain excellent dimensional stability at 80 °C. The excellent inhibitory effect of TpPa–Py on the swelling ratio of the composite membranes is mainly due to the pyridine group in the TpPa–Py skeleton forming an ionic cross-linking network in the polymer matrix, which can effectively limit the excessive expansion of the SPEEK matrix.



**Figure 4.** Swelling ratio of (a) SPEEK and SPEEK/TpPa–SO<sub>3</sub>H, (b) SPEEK and SPEEK/TpPa–Py at different temperatures under 100% RH.

### 3.1.4. Structure and Morphology Characterization of SPEEK/TpPa–SO<sub>3</sub>H and SPEEK/TpPa–Py Composite Membranes

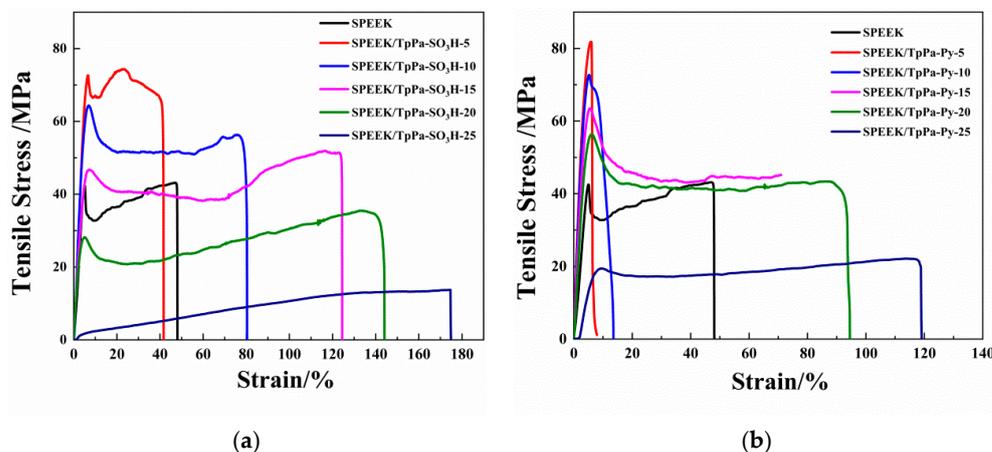
The cross-sectional scanning electron microscopy images of SPEEK, SPEEK/TpPa–SO<sub>3</sub>H-15, and SPEEK/TpPa–Py-15 composite membranes were shown in Figure 5. Compared with SPEEK, the cross sections of the composite membranes were rougher, but the textures were uniform, indicating that COFs were evenly dispersed in the matrix. Figure S2a–k showed the cross-sectional scanning electron microscopy images of all the SPEEK, SPEEK/TpPa–SO<sub>3</sub>H, and SPEEK/TpPa–Py composite membranes. With the increase in the amount of TpPa–SO<sub>3</sub>H and TpPa–Py nanoparticles, the composite membranes maintained a compact structure. Prepared by the in situ composite method, the polymer chains of SPEEK hindered the  $\pi$ – $\pi$  interaction between COF particles, [34] and the –NH– in the COF framework can interact with the –SO<sub>3</sub>H in the SPEEK matrix to improve compatibility between COF and SPEEK matrix. Moreover, the pyridine group and secondary amine group in the TpPa–Py structure had a strong interaction with the sulfonic acid group of SPEEK, which reduced the agglomeration of TpPa–Py in the matrix. The interaction between the COF and SPEEK matrix was beneficial to maintain the mechanical properties of composite membranes.



**Figure 5.** SEM cross-section images of (a) SPEEK, (b) SPEEK/TpPa–SO<sub>3</sub>H-15, (c) SPEEK/TpPa–Py-15.

### 3.1.5. Mechanical Properties

Figure 6 showed the stress–strain curves of SPEEK, SPEEK/TpPa–SO<sub>3</sub>H (Figure 6a), and SPEEK/TpPa–Py (Figure 6b) composite membranes at 25 °C. The hydrophobic backbone in SPEEK polymer chain makes membranes have good mechanical properties. The SPEEK membrane showed tensile stress of 42.5 MPa. When the loading of TpPa–SO<sub>3</sub>H was 5 wt%, the tensile strength of the composite membranes was as high as 47.6 MPa, indicating that the introduction of TpPa–SO<sub>3</sub>H improved the tensile strength of the composite membranes. There were two reasons for the improvement in mechanical properties. First, the ketoenamine (–C–N–) on the TpPa–SO<sub>3</sub>H ring can interact with the sulfonic acid and (–SO<sub>3</sub>H) acid base on the SPEEK aromatic ring. At the same time,  $\pi$ – $\pi$  interaction may also occur between the TpPa–SO<sub>3</sub>H framework and the SPEEK benzene ring [26]. The curve shows that the tensile strength of the composite film decreases with the increase in TpPa–SO<sub>3</sub>H filling amount because the content of PTSA in the TpPa–SO<sub>3</sub>H channel increases with the increase in the content of TpPa–SO<sub>3</sub>H. As a small molecule, PTSA increases the separation of polymer chains and reduces the intermolecular force of polymers [35]. When the filling amount of TpPa–SO<sub>3</sub>H is 25 wt%, it will inevitably agglomerate in the membrane, and plasticization of PTSA led to an increase in elongation. As shown in Figure 6b, the tensile strength of all SPEEK/TpPa–Py composite membranes was higher than that of the SPEEK membrane. When the filling amount of TpPa–Py was 5 wt%, the tensile strength of the composite membranes was as high as 81.8 MPa. The hydrogen bond interaction between the pyridine group in the TpPa–Py and the sulfonic acid group in SPEEK increased the yield strength of the composite membranes [36]. To a certain extent, this strong interaction can inhibit polymer chain migration.

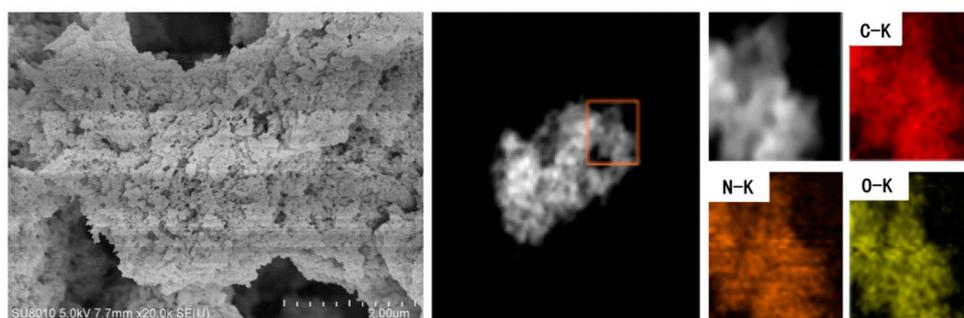


**Figure 6.** Stress–strain curves of (a) SPEEK and SPEEK/TpPa–SO<sub>3</sub>H, and (b) SPEEK and SPEEK/TpPa–Py.

### 3.2. SPEEK/TpPa-(SO<sub>3</sub>H-Py) Composite Membrane

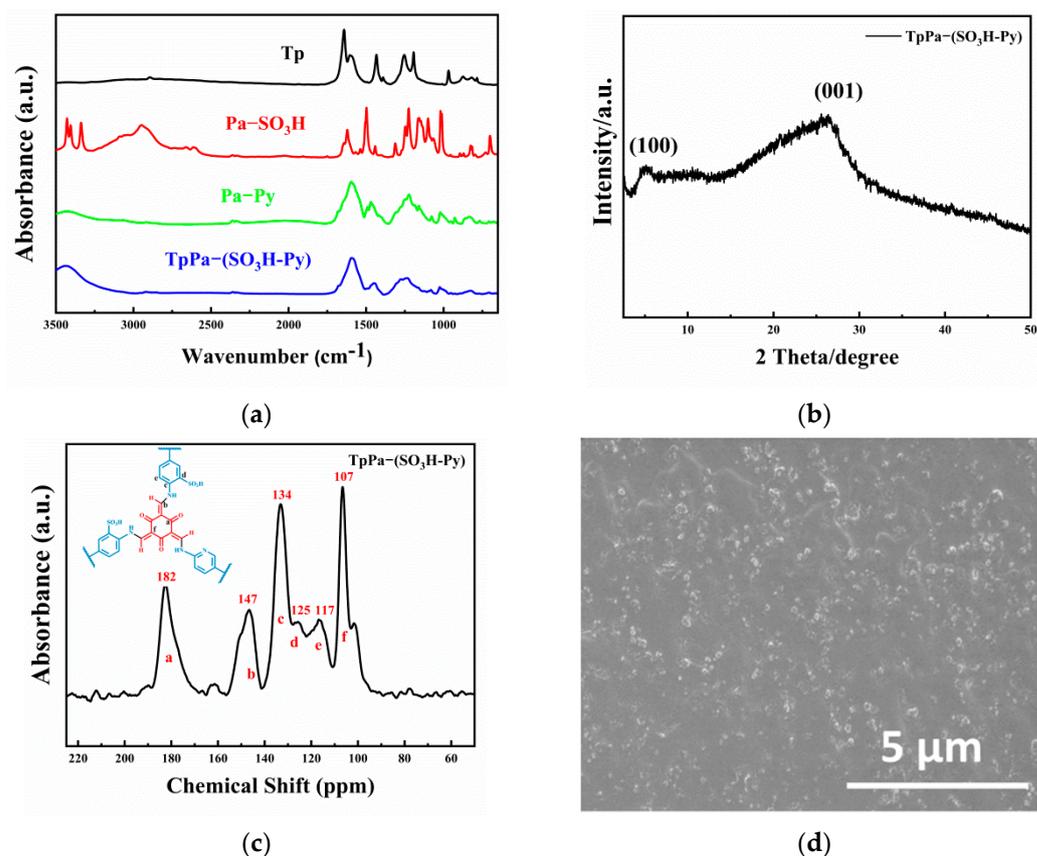
#### 3.2.1. Characterization of SPEEK/TpPa-(SO<sub>3</sub>H-Py)

From the above results, it is found that the -SO<sub>3</sub>H group has an excellent performance in promoting proton conduction and the pyridine group can effectively improve the dimensional stability and strength of PEMs. Therefore, the -SO<sub>3</sub>H group and the pyridine group were simultaneously introduced into the COF structure to improve the performances of PEMs. The content of TpPa-(SO<sub>3</sub>H-Py) was 20 wt% and the ratio of Pa-SO<sub>3</sub>H to Pa-Py was 1:1. Therefore, TpPa-(SO<sub>3</sub>H-Py) was synthesized and added to the SPEEK matrix by in situ synthesis to prepare composite membranes. To characterize the synthesis of TpPa-(SO<sub>3</sub>H-Py) in the membrane, the composite membranes were dissolved in DMSO at 80 °C, and then TpPa-(SO<sub>3</sub>H-Py) was separated from the solution. The morphology of TpPa-(SO<sub>3</sub>H-Py) was obtained by SEM and TEM (Figure 7). Consistent with the previously reported literature [24], the morphology of (TpPa-(SO<sub>3</sub>H-Py)) was not a simple mixture of spherical and fibrous, but formed by stacking layers on each other. Due to the centrifugal force, the TpPa-(SO<sub>3</sub>H-Py) is stacked together, but it can still be seen that the TpPa-(SO<sub>3</sub>H-Py) is also composed of nanoparticles (please see Figure S4 for other SEM images of TpPa-(SO<sub>3</sub>H-Py)). Compared with SPEEK matrix, the characteristic element of COF is N. It can be seen from the EDS of the washed powder that the washed powder contained the N element, and the distribution of the N element in the powder was uniform.



**Figure 7.** SEM, TEM, and EDS-mapping of TpPa-(SO<sub>3</sub>H-Py).

Figure 8a showed the FTIR spectra of TpPa-(SO<sub>3</sub>H-Py). The TpPa-(SO<sub>3</sub>H-Py) sample had characteristic peaks of C=C (1583 cm<sup>-1</sup>) and C-N (1232 cm<sup>-1</sup>), which proved the existence of enol-ketone mutual variation structure. The peaks at the wavenumbers of 1084 cm<sup>-1</sup> and 1027 cm<sup>-1</sup> were attributed to the O=S=O stretching vibration peaks of TpPa-(SO<sub>3</sub>H-Py). The ordered structures of TpPa-(SO<sub>3</sub>H-Py) were characterized by XRD (Figure 8b). Corresponding to the reflection of the (100) plane, there was a peak at 4.9° in the TpPa-(SO<sub>3</sub>H-Py) sample spectrum. The peak at 26° corresponded to the reflection peak of the (001) plane, which proved the existence of  $\pi$ - $\pi$  stacking in the synthesis process of TpPa-(SO<sub>3</sub>H-Py). Figure 8c showed the <sup>13</sup>C SSNMR spectrum of TpPa-(SO<sub>3</sub>H-Py), which further confirmed the synthesis of ketoenamine COFs. It was observed that the characteristic peak positions of the TpPa-(SO<sub>3</sub>H-Py) pattern were the same as the <sup>13</sup>C patterns of TpPa-SO<sub>3</sub>H and TpPa-Py. Keto-enol interconversion was also observed in TpPa-(SO<sub>3</sub>H-Py). The peak of C=O (ketone form) was 184 ppm, and the peak of C=C (enol form) was 147 ppm and 107 ppm. The resonance peak near 134 ppm corresponded to the characteristic peak of the carbon atom connected to nitrogen in the original reactive monomer aromatic diamine. The characteristic peaks of the carbon atoms near the sulfonic acid group and the pyridine group in the TpPa-(SO<sub>3</sub>H-Py) skeleton were located at 125 ppm and 117 ppm. This result matched well with the <sup>13</sup>C spectrum results of TpPa-(SO<sub>3</sub>H-Py) obtained in the previous report [24]. The above results indicated that TpPa-(SO<sub>3</sub>H-Py) was successfully synthesized in the SPEEK matrix.

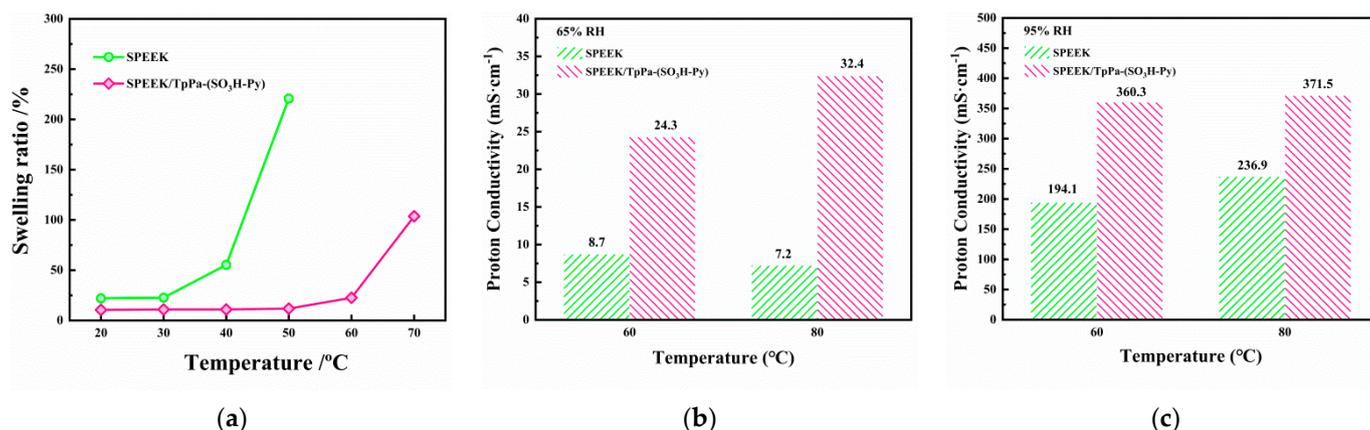


**Figure 8.** (a) FTIR, (b) XRD, (c) NMR spectra of TpPa-(SO<sub>3</sub>H-Py), (d) SEM cross-section images of SPEEK/TpPa-(SO<sub>3</sub>H-Py).

The cross-sectional morphology of the SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite film was observed by SEM, as shown in Figure 8d. The content of TpPa-(SO<sub>3</sub>H-Py) was 20 wt%, and the ratio of Pa-SO<sub>3</sub>H to Pa-Py was 1:1. There are no obvious cracks and defects on the cross-section of the composite membrane, and TpPa-(SO<sub>3</sub>H-Py) was evenly dispersed in the composite membrane.

### 3.2.2. Dimensional Stability and Proton Conductivity

The water uptake was shown in Figure S3c and the swelling ratio of SPEEK and SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membranes were shown in Figure 9a. Under the condition of 20 °C and 100% RH, the water uptake of the SPEEK and SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membranes were 39.3% and 44.3%, respectively; the area swelling ratios were 22.0% and 21.1%, respectively. Under 50 °C and 100% RH conditions, the water uptake of the SPEEK and SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membrane were 241.5% and 56.4%, respectively. At 50 °C, the swelling ratio of SPEEK was as high as 220.7%. In contrast, the swelling ratio of SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membrane was only 11.8% at 50 °C, and 103.6% at 70 °C. Under high-temperature conditions, the interaction between TpPa-(SO<sub>3</sub>H-Py) and SPEEK limited the movement of polymer chains and improve the dimensional stability of the composite membrane. In addition, the increase in the content of pyridine groups can further enhance the interaction with SPEEK, so that the swelling ratio of the composite membrane was at a much lower level. Generally, the addition of TpPa-(SO<sub>3</sub>H-Py) can effectively improve the dimensional stability of the composite membrane, and the swelling inhibition effect was best when the ratio of sulfonic acid groups to pyridine groups was 1:1 (for other sulfonic acid and pyridine ratio composite membranes, please read Figure S5).



**Figure 9.** (a) Swelling of SPEEK and SPEEK/TpPa-(SO<sub>3</sub>H-Py) from 20 °C to 70 °C at 100% RH; proton conductivity at (b) 65%RH and (c) 95%RH of TpPa-(SO<sub>3</sub>H-Py).

Figure 9b,c showed the conductivity of the composite membranes under different humidity environments. At 65% RH and 95% RH, the SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membranes showed higher proton conductivity than pure SPEEK membranes. At 60 °C and 65% RH, the proton conductivity of the SPEEK and SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membranes were 8.7 mS·cm<sup>-1</sup>, 24.3 mS·cm<sup>-1</sup>, respectively. When the relative humidity increased to 95%, the proton conductivity of the SPEEK and SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membranes were 194.1 mS·cm<sup>-1</sup> and 360.3 mS·cm<sup>-1</sup>, respectively. At 65% RH and 95% RH, the proton conductivity of the SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membrane is higher than that of SPEEK membrane and SPEEK/TpPa-Py composite membranes because the -SO<sub>3</sub>H group constructs a hydrogen bond network and improves the proton transfer efficiency. At 80 °C and 95% RH, the conductivity of the composite membrane is as high as 371.5 mS·cm<sup>-1</sup>, in contrast, the conductivity of the SPEEK pristine membrane is only 236.9 mS·cm<sup>-1</sup>.

Compared with SPEEK/TpPa-SO<sub>3</sub>H composite membranes, the conductivity of SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membrane slightly decreased due to the reduction of -SO<sub>3</sub>H but showed better dimensional stability and temperature resistance. At 50 °C, the swelling ratio of SPEEK/TpPa-SO<sub>3</sub>H-5 composite membrane was 36.7%, while that of SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membrane was only 11.8%. When the temperature reached 60 °C, the SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membrane still maintained good dimensional stability, and the swelling rate was only 22.6%. It shows that the pyridine group significantly improves the dimensional stability of SPEEK/TpPa-(SO<sub>3</sub>H-Py) composite membrane. In general, the construction of TpPa-(SO<sub>3</sub>H-Py) with -SO<sub>3</sub>H group and pyridine group in the composite membranes can improve proton conductivities, dimensional stability, and temperature resistance of the composite membranes.

#### 4. Conclusions

In summary, we propose a strategy for synthesizing COFs containing different groups in the SPEEK matrix to fabricate SPEEK/COFs composite proton exchange membranes. TpPa-SO<sub>3</sub>H containing -SO<sub>3</sub>H groups and TpPa-Py containing pyridine groups were synthesized in the SPEEK matrix by in situ synthesis, and evenly dispersed in the matrix. Meanwhile, the interactions between groups attached to COFs and -SO<sub>3</sub>H of SPEEK promote the deprotonation of -SO<sub>3</sub>H, constructing low-energy barrier channels for proton transport. Then the effects of different groups attached to COFs on the structure and properties of proton exchange membranes were investigated. The -SO<sub>3</sub>H group on the TpPa-SO<sub>3</sub>H framework provided additional sulfonate radicals, which enhanced the proton conductivity through the proton conduction mechanism and exhibits excellent proton conductivity. At 60 °C, under 65% RH and 95% RH, the conductivity of the SPEEK/TpPa-SO<sub>3</sub>H-20 composite membrane was as high as 44.1 mS·cm<sup>-1</sup> and

443.6 mS·cm<sup>-1</sup>, which were 5.7 times and 3.3 times that of SPEEK membrane, respectively. The pyridine group on the TpPa–Py skeleton provides higher stability for the composite membrane. Benefiting from the hydrogen-bonding interfaces between the pyridine groups of TpPa–Py and SPEEK, SPEEK/TpPa–Py had good dimensional stability. The swelling ratio of SPEEK was as high as 220.7% at 50 °C, while the swelling ratio of SPEEK/TpPa–Py-25 was only 2.4%. SPEEK/TpPa–Py-20 and SPEEK/TpPa–Py-25 still maintained good dimensional stability at 80 °C, and the swelling ratios are less than 20%. Combined with the high proton conductivity of –SO<sub>3</sub>H and the high stability of pyridine, these two groups were introduced into the COFs skeleton at the same time. The SPEEK/TpPa–(SO<sub>3</sub>H-Py) composite membranes had high proton conductivity and dimensional stability. At 60 °C and 95% humidity, the conductivity of the SPEEK/TpPa–(SO<sub>3</sub>H-Py) composite membrane was 360.3 mS·cm<sup>-1</sup>, which was 1.9 times that of SPEEK, and the swelling ratio was 11.8%, which was about 1/20 of that of SPEEK. This work reveals the outstanding performance of COFs in proton conduction and the great application potential of COFs nanomaterials containing different groups in PEMs.

**Supplementary Materials:** The following supporting information can be downloaded at: <https://www.mdpi.com/article/10.3390/nano12193518/s1>. Table S1: Amount of raw materials required for preparing SPEEK/TpPa–SO<sub>3</sub>H composite membranes; Table S2: Amount of raw materials required for preparing SPEEK/TpPa–Py composite membranes; Table S3: Amount of raw materials required for preparing SPEEK/TpPa–(SO<sub>3</sub>H-Py) composite membranes; Figure S1: TGA of membranes; Figure S2: SEM cross-section images of composite membranes; Figure S3: Water uptaking of membranes; Figure S4: Properties of composite membranes with different ratios of sulfonic acid groups to pyridine groups; Figure S5: Properties of composite membranes with different ratios of sulfonic acid groups to pyridine groups: (a) proton conductivity at 60 °C and 65% RH, (b) proton conductivity at 60 °C and 95% RH, (c) swelling ratio, and (d) water uptaking. Reference [37] is cited in the Supplementary Materials.

**Author Contributions:** Conceptualization, X.M. and Y.L.; methodology, X.M. and Y.L.; validation, X.M., Y.L., L.D., L.P. and Q.P.; formal analysis, X.M., Y.L. and L.P.; investigation, X.M., Y.L., L.D. and L.P.; writing—original draft preparation, Y.L.; writing—review and editing, X.M.; Y.L., L.P., C.C., H.Y. and Q.Z.; supervision, X.M. All authors have read and agreed to the published version of the manuscript.

**Funding:** This project was financially supported by Science Foundation of China University of Petroleum, Beijing (No. 2462020YXZZ019).

**Institutional Review Board Statement:** Not applicable.

**Informed Consent Statement:** Not applicable.

**Data Availability Statement:** The data presented in this study are available on request from the corresponding author.

**Conflicts of Interest:** The authors declare no conflict of interest.

## References

- Colicchio, I.; Wen, F.; Keul, H.; Simon, U.; Moeller, M. Sulfonated poly (ether ether ketone)–silica membranes doped with phosphotungstic acid. Morphology and proton conductivity. *J. Membr. Sci.* **2009**, *326*, 45–57. [[CrossRef](#)]
- Shin, D.W.; Guiver, M.D.; Lee, Y.M. Hydrocarbon-based polymer electrolyte membranes: Importance of morphology on ion transport and membrane stability. *Chem. Rev.* **2017**, *117*, 4759–4805. [[CrossRef](#)]
- Tan, A.R.; de Carvalho, L.M.; de Ramos Filho, F.G.; de Souza Gomes, A. Nanocomposite membranes based on sulfonated poly (etheretherketone) structured with modified silica for direct ethanol fuel cell. *Macromol. Symp.* **2006**, *245–246*, 470–475. [[CrossRef](#)]
- Byun, G.H.; Kim, J.A.; Kim, N.Y.; Cho, Y.S.; Park, C.R. Molecular engineering of hydrocarbon membrane to substitute perfluorinated sulfonic acid membrane for proton exchange membrane fuel cell operation. *Mater. Today Energy* **2020**, *17*, 100483. [[CrossRef](#)]
- Wong, C.; Wong, W.; Ramya, K.; Khalid, M.; Loh, K.; Daud, W.; Lim, K.; Walvekar, R.; Kadhum, A. Additives in proton exchange membranes for low-and high-temperature fuel cell applications: A review. *Int. J. Hydrogen Energy* **2019**, *44*, 6116–6135. [[CrossRef](#)]

6. Vinothkannan, M.; Ramakrishnan, S.; Kim, A.R.; Lee, H.K.; Yoo, D.J. Ceria Stabilized by Titanium Carbide as a Sustainable Filler in the Nafion Matrix Improves the Mechanical Integrity, Electrochemical Durability, and Hydrogen Impermeability of Proton-Exchange Membrane Fuel Cells: Effects of the Filler Content. *ACS Appl. Mater. Interfaces* **2020**, *12*, 5704–5716. [[CrossRef](#)] [[PubMed](#)]
7. Raja Sulaiman, R.R.; Rashmi, W.; Khalid, M.; Wong, W.; Priyanka, J. Recent progress in the development of aromatic polymer-based proton exchange membranes for fuel cell applications. *Polymers* **2020**, *12*, 1061. [[CrossRef](#)]
8. Maier, G.; Meier-Haack, J. Sulfonated aromatic polymers for fuel cell membranes. In *Fuel Cells II*; Scherer, G.G., Ed.; Springer: Berlin/Heidelberg, Germany, 2008; pp. 1–62. [[CrossRef](#)]
9. Wu, X.; Wang, X.; He, G.; Benziger, J. Differences in water sorption and proton conductivity between Nafion and SPEEK. *J. Polym. Sci. Part B Polym. Phys.* **2011**, *49*, 1437–1445. [[CrossRef](#)]
10. Xie, J.; Wood, D.L., III; Wayne, D.M.; Zawodzinski, T.A.; Atanassov, P.; Borup, R.L. Durability of PEFCs at high humidity conditions. *J. Electrochem. Soc.* **2004**, *152*, A104. [[CrossRef](#)]
11. Kreuer, K. On the development of proton conducting polymer membranes for hydrogen and methanol fuel cells. *J. Membr. Sci.* **2001**, *185*, 29–39. [[CrossRef](#)]
12. Xing, P.; Robertson, G.P.; Guiver, M.D.; Mikhailenko, S.D.; Wang, K.; Kaliaguine, S. Synthesis and characterization of sulfonated poly (ether ether ketone) for proton exchange membranes. *J. Membr. Sci.* **2004**, *229*, 95–106. [[CrossRef](#)]
13. Yuan, S.; Li, X.; Zhu, J.; Zhang, G.; Van Puyvelde, P.; Van der Bruggen, B. Covalent organic frameworks for membrane separation. *Chem. Soc. Rev.* **2019**, *48*, 2665–2681. [[CrossRef](#)] [[PubMed](#)]
14. Sun, Q.; Tang, Y.; Aguila, B.; Wang, S.; Xiao, F.S.; Thallapally, P.K.; Al-Enizi, A.M.; Nafady, A.; Ma, S. Reaction environment modification in covalent organic frameworks for catalytic performance enhancement. *Angew. Chem. Int. Ed.* **2019**, *58*, 8670–8675. [[CrossRef](#)] [[PubMed](#)]
15. Ji, W.; Xiao, L.; Ling, Y.; Ching, C.; Matsumoto, M.; Bisbey, R.P.; Helbling, D.E.; Dichtel, W.R. Removal of GenX and perfluorinated alkyl substances from water by amine-functionalized covalent organic frameworks. *J. Am. Chem. Soc.* **2018**, *140*, 12677–12681. [[CrossRef](#)]
16. Cui, D.; Peregichka, D.F.; MacLeod, J.M.; Rosei, F. Surface-confined single-layer covalent organic frameworks: Design, synthesis and application. *Chem. Soc. Rev.* **2020**, *49*, 2020–2038. [[CrossRef](#)] [[PubMed](#)]
17. Meri-Bofi, L.; Royuela, S.; Zamora, F.; Ruiz-González, M.L.; Segura, J.L.; Muñoz-Olivas, R.; Mancheño, M.J. Thiol grafted imine-based covalent organic frameworks for water remediation through selective removal of Hg (II). *J. Mater. Chem. A* **2017**, *5*, 17973–17981. [[CrossRef](#)]
18. Lin, S.; Diercks, C.S.; Zhang, Y.-B.; Kornienko, N.; Nichols, E.M.; Zhao, Y.; Paris, A.R.; Kim, D.; Yang, P.; Yaghi, O.M. Covalent organic frameworks comprising cobalt porphyrins for catalytic CO<sub>2</sub> reduction in water. *Science* **2015**, *349*, 1208–1213. [[CrossRef](#)] [[PubMed](#)]
19. Dalapati, S.; Jin, S.; Gao, J.; Xu, Y.; Nagai, A.; Jiang, D. An azine-linked covalent organic framework. *J. Am. Chem. Soc.* **2013**, *135*, 17310–17313. [[CrossRef](#)]
20. Bhunia, S.; Das, S.K.; Jana, R.; Peter, S.C.; Bhattacharya, S.; Addicoat, M.; Bhaumik, A.; Pradhan, A. Electrochemical stimuli-driven facile metal-free hydrogen evolution from pyrene-porphyrin-based crystalline covalent organic framework. *ACS Appl. Mater. Interfaces* **2017**, *9*, 23843–23851. [[CrossRef](#)]
21. Chandra, S.; Kundu, T.; Kandambeth, S.; BabaRao, R.; Marathe, Y.; Kunjir, S.M.; Banerjee, R. Phosphoric Acid Loaded Azo (–N=N–) Based Covalent Organic Framework for Proton Conduction. *J. Am. Chem. Soc.* **2014**, *136*, 6570–6573. [[CrossRef](#)]
22. Peng, Y.; Xu, G.; Hu, Z.; Cheng, Y.; Chi, C.; Yuan, D.; Cheng, H.; Zhao, D. Mechanoassisted synthesis of sulfonated covalent organic frameworks with high intrinsic proton conductivity. *ACS Appl. Mater. Interfaces* **2016**, *8*, 18505–18512. [[CrossRef](#)]
23. Mohammed, O.F.; Pines, D.; Nibbering, E.T.J.; Pines, E. Base-induced solvent switches in acid-base reactions. *Angew. Chem. Int. Ed.* **2007**, *46*, 1458–1461. [[CrossRef](#)]
24. Chandra, S.; Kundu, T.; Dey, K.; Addicoat, M.; Heine, T.; Banerjee, R. Interplaying intrinsic and extrinsic proton conductivities in covalent organic frameworks. *Chem. Mater.* **2016**, *28*, 1489–1494. [[CrossRef](#)]
25. Yin, Y.; Li, Z.; Yang, X.; Cao, L.; Wang, C.; Zhang, B.; Wu, H.; Jiang, Z. Enhanced proton conductivity of Nafion composite membrane by incorporating phosphoric acid-loaded covalent organic framework. *J. Power Sources* **2016**, *332*, 265–273. [[CrossRef](#)]
26. Fan, C.; Wu, H.; Li, Y.; Shi, B.; He, X.; Qiu, M.; Mao, X.; Jiang, Z. Incorporating self-anchored phosphotungstic acid@triazole-functionalized covalent organic framework into sulfonated poly (ether ether ketone) for enhanced proton conductivity. *Solid State Ion.* **2020**, *349*, 115316. [[CrossRef](#)]
27. Meng, X.Y.; Song, K.; Lv, Y.A.; Cong, C.A.B.; Ye, H.M.; Dong, Y.H.; Zhou, Q. SPEEK proton exchange membrane with enhanced proton conductivity stability from phosphotungstic acid-encapsulated silica nanorods. *Mater. Chem. Phys.* **2021**, *272*, 125045. [[CrossRef](#)]
28. Zheng, Y.; Shen, J.L.; Yuan, J.Q.; Khan, N.A.; You, X.D.; Yang, C.; Zhang, S.Y.; El-Gendi, A.; Wu, H.; Zhang, R.N.; et al. 2D nanosheets seeding layer modulated covalent organic framework membranes for efficient desalination. *Desalination* **2022**, *532*, 115753. [[CrossRef](#)]
29. Dey, K.; Pal, M.; Rout, K.C.; Kunjattu H, S.; Das, A.; Mukherjee, R.; Kharul, U.K.; Banerjee, R. Selective molecular separation by interfacially crystallized covalent organic framework thin films. *J. Am. Chem. Soc.* **2017**, *139*, 13083–13091. [[CrossRef](#)] [[PubMed](#)]

30. Kandambeth, S.; Mallick, A.; Lukose, B.; Mane, M.V.; Heine, T.; Banerjee, R. Construction of crystalline 2D covalent organic frameworks with remarkable chemical (acid/base) stability via a combined reversible and irreversible route. *J. Am. Chem. Soc.* **2012**, *134*, 19524–19527. [[CrossRef](#)]
31. Yao, J.; Xu, G.; Zhao, Z.; Guo, J.; Li, S.; Cai, W.; Zhang, S. An enhanced proton conductivity and reduced methanol permeability composite membrane prepared by sulfonated covalent organic nanosheets/Nafion. *Int. J. Hydrogen Energy* **2019**, *44*, 24985–24996. [[CrossRef](#)]
32. Jiang, W.; Cui, W.-R.; Liang, R.-P.; Qiu, J.-D. Difunctional covalent organic framework hybrid material for synergistic adsorption and selective removal of fluoroquinolone antibiotics. *J. Hazard. Mater.* **2021**, *413*, 125302. [[CrossRef](#)]
33. Shi, X.; Ma, D.; Xu, F.; Zhang, Z.; Wang, Y. Table-salt enabled interface-confined synthesis of covalent organic framework (COF) nanosheets. *Chem. Sci.* **2020**, *11*, 989–996. [[CrossRef](#)]
34. Kandambeth, S.; Biswal, B.P.; Chaudhari, H.D.; Rout, K.C.; Kunjattu H, S.; Mitra, S.; Karak, S.; Das, A.; Mukherjee, R.; Kharul, U.K.; et al. Selective molecular sieving in self-standing porous covalent-organic-framework membranes. *Adv. Mater.* **2017**, *29*, 1603945. [[CrossRef](#)] [[PubMed](#)]
35. Du, M.; Yang, L.; Luo, X.; Wang, K.; Chang, G. Novel phosphoric acid (PA)-poly (ether ketone sulfone) with flexible benzotriazole side chains for high-temperature proton exchange membranes. *Polym. J.* **2019**, *51*, 69–75. [[CrossRef](#)]
36. Yang, J.; Xu, Y.; Zhou, L.; Che, Q.; He, R.; Li, Q. Hydroxyl pyridine containing polybenzimidazole membranes for proton exchange membrane fuel cells. *J. Membr. Sci.* **2013**, *446*, 318–325. [[CrossRef](#)]
37. Lakshmi, R.M.; Choudhary, V.; Varma, I. Sulphonated poly (ether ether ketone): Synthesis and characterisation. *J. Mater. Sci.* **2005**, *40*, 629–636. [[CrossRef](#)]