

Article

# Raman Spectroscopic Study of Coal Samples during Heating

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**Featured Application:** Temperature dependent in-situ Raman spectroscopy was applied to study the structure of coals at high temperature.

**Abstract:** Raman spectroscopy can be used to record the characteristic spectra of carbonaceous materials. The D and G bands are the most popular and most important spectral characteristics when discussing carbonaceous materials. In this paper, a Raman spectroscopic study of different coals was first carried out using a 355 nm wavelength laser beam as an excitation source. The spectral parameters of the resultant spectra were evaluated and analyzed. Raman spectral characteristics of different kinds of coals were explored. The high temperature-dependent Raman spectra of the coals were further collected in a temperature range from 298 to 1473 K in order to investigate the transformations of the internal structure of the coals during the pyrolysis process. An abnormal blue shift of the G band occurred at moderate temperature (600–900 K), and the intensity of the G band became weaker at high temperatures, indicating pyrolysis and graphitization of the sample at moderate and high temperature, respectively.

**Keywords:** coal; in-situ Raman spectroscopy; pyrolysis; characteristic Raman bands

## 1. Introduction

Raman spectroscopy has been used to study carbonaceous materials since the 1970s [1]. As a unique type of spectroscopy, Raman spectroscopy has many advantages in coal structure analysis over other techniques [2]. Raman spectroscopy can provide fast, simple, repeatable, and lossless qualitative and quantitative analysis of defects, and has been widely used in the coal chemical industry. Raman spectroscopy can be used to analyze the structure of coal and the quality of organic matter in order to characterize the catalysts in the coal chemical process and to detect heavy fraction leaching [3].

Green et al. [4] have studied the identification of minerals in coals, one of the principal areas of the field, and identified calcite, pyrite, and dolomite by using a micro-Raman spectrometer. Tuinstra and Koenig [5] and Friedel and Carlson [6] have published the first Raman spectrum of coal and reported G (1575–1620  $\text{cm}^{-1}$ ) and D-bands (1355–1380  $\text{cm}^{-1}$ ). Tuinstra and Koenig have suggested that the 1575  $\text{cm}^{-1}$  band can be assigned to the graphite  $E_{2g}$  mode with  $D^4_{6h}$  crystal symmetry, that the band at 1370  $\text{cm}^{-1}$  can be assigned to the  $A_{1g}$  mode, and that the ratio  $I_D/I_G$  relates to the average graphite domain dimension. La. Friedel and Carlson have debated the origin of the band at 1580  $\text{cm}^{-1}$  by investigating IR absorption and the Raman scattering of very finely ground graphite (C–C bonds broken), coal, and carbon black. They have suggested that the band at 1350  $\text{cm}^{-1}$  most likely arises from graphitic structures, and not from aromatic or conjugated carbonyls, as had been suggested

earlier. Johnson and Thomas [7] have found that the width of the D-band for chars originates from various ranks of coals and decreases with increasing heat treatment temperature (from  $240\text{ cm}^{-1}$  at  $500\text{ }^{\circ}\text{C}$ , with  $L_a = 2\text{ nm}$ , to  $40\text{ cm}^{-1}$  at  $1900\text{ }^{\circ}\text{C}$ , with  $L_a = 7\text{ nm}$ ). On the other hand, Benny-Bassez and Rouzaud [8] have found that the G+D' band of different coals decreases with increasing coal rank, although they did not observe a trend (the D' peak located around  $1620\text{ cm}^{-1}$  is due to irregular carbon and often appears with the G peaks). Angoni [9] has postulated a simple qualitative analysis in order to distinguish carbonaceous materials with low, medium, and high structural organization by analyzing 15 samples that included graphite, coke, coal and anthracite. He has found that the width of the D band relates to the degree of its disorder, but that neither the D nor the G bands can explain the degree of this disorder.

After analyzing chars generated from the pyrolysis and gasification of Australian brown coal by Raman spectroscopy, Chunzhu Li et al. [10] have identified 10 characteristic bands that they believe are attributable to the highly disordered carbon structure of coal [11]. Sheng [12] has studied the relationship between the microstructure and the combustion reactivity of coal char by Raman spectroscopy [13].

Recently, much research has been carried out on the application of Raman spectroscopy to coal. Raman spectroscopy has been applied to rapidly determine organic matter maturity and value the chemical structure and thermal maturity of vitrinite [14,15]. The relationship between coal structures and combustion characteristics has been investigated by micro-Raman spectroscopy, which can provide a new approach to rapidly predicting the coal properties by observing the reasonable correlations between the Raman spectral parameters and  $V^{\text{daf}}$  [2]. In addition, Raman spectroscopy has been used to distinguish components of graphitized coals and quantitatively evaluate vitrinite reflectance in shale [16,17].

An in-situ study at high temperature [18] can provide valuable information concerning structural evolution for fundamental research on coal pyrolysis. The changes in chemical bonds identified by in-situ infrared spectroscopy studies of coal pyrolysis [19,20] provided useful information on chemical bond dissociation. An in-situ high temperature Raman technique can also offer a unique tool for observing different characteristics of coal. In-situ Raman spectroscopy [21] has been widely used in many fields. However, research on temperature-dependent in-situ Raman spectroscopy of carbon materials is lacking. This paper describes an in-situ high temperature Raman spectroscopy study on selected coals and the evolution of the internal structural of coal during heating.

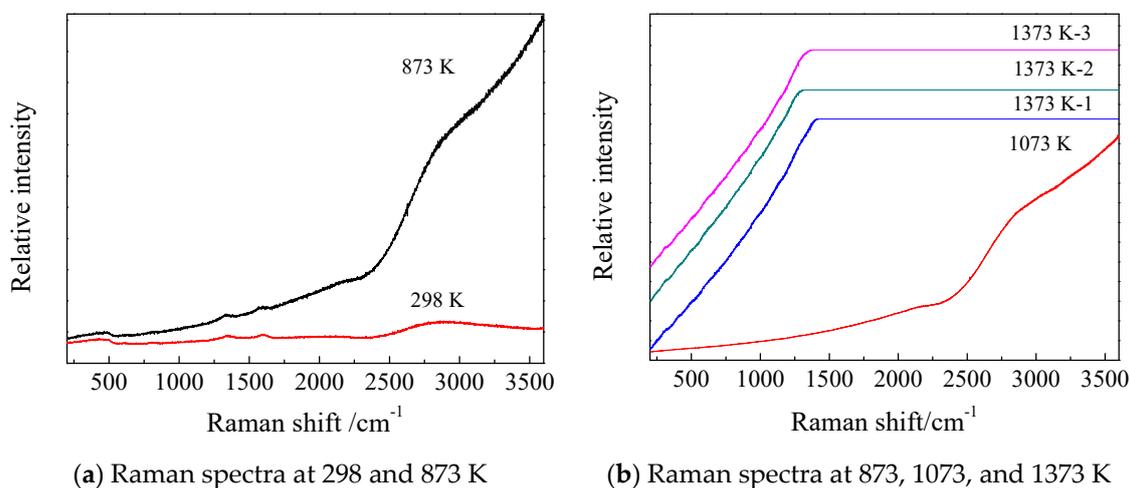
## 2. Experimental

### 2.1. Raman Spectrometer

Temperature-dependent Raman spectra were collected using a Horiba Jobin Yvon LabRam 800 HR Raman spectrometer. In comparison to inorganic materials such as aluminosilicates, carbonaceous materials or other organic components behave less sensitively when subjected to a laser beam of a short wavenumber and therefore have less cross-sectional penetration. The test sample was heated in an Ar atmosphere by a Linkam 1500 heating stage with a temperature deviation of less than  $\pm 1\text{ K}$  in order to investigate the microstructure evolution of the samples. A charge-coupled-device (CCD) detection system was used to record the Raman scattering light at an accumulated mode of  $20 \times 20$  (20 times, each of 20 s). The samples were held at the targeted temperature for 60 min before collecting Raman spectra in order to ensure they reached the targeted.

At first, 532 nm and 355 nm were selected as the excitation laser wavelengths for temperature-dependent in-situ Raman spectra tests. Figure 1 shows the in-situ Raman spectra of the coal samples using a 532 nm excitation laser. The spectra in the higher wavenumber range experienced interference from the overwhelming fluorescence, which became more severe over time. The fluorescence appeared dramatically when the temperature was higher than 1073 K and remained for a period of time. All coal samples had similar fluorescence interference. Under the circumstances, D and G bands would be

difficult to detect. As a result, 355 nm was selected as the excitation laser for the temperature-dependent in-situ Raman spectra test. A 355 nm ultraviolet laser beam of about 20 mW was focused on the sample as an excitation source through a 4× objective lens. A 355 nm laser excitation wavelength was used to ensure a smaller penetration depth in order to collect relative structure information. It was estimated that the penetration depth was several hundred nanometers within the coal samples under the 355 nm laser radiation.



**Figure 1.** In-situ Raman spectra of coal samples in Ar atmosphere by using a 532 nm excitation laser.

## 2.2. Coal Samples and Preparation

Six coal samples of varying ranks and with different carbon contents were selected. These samples are shown in Table 1. The elemental content of the coal samples was measured by Elementar vario ELIII (Germany) and listed in Table 1. The coal samples were first ground to the size of 0.2 mm and then dried at 40 °C for 12 h. The dry coal powders were stored for further characterization by high temperature Raman spectrometry.

**Table 1.** Elemental composition of coal samples.

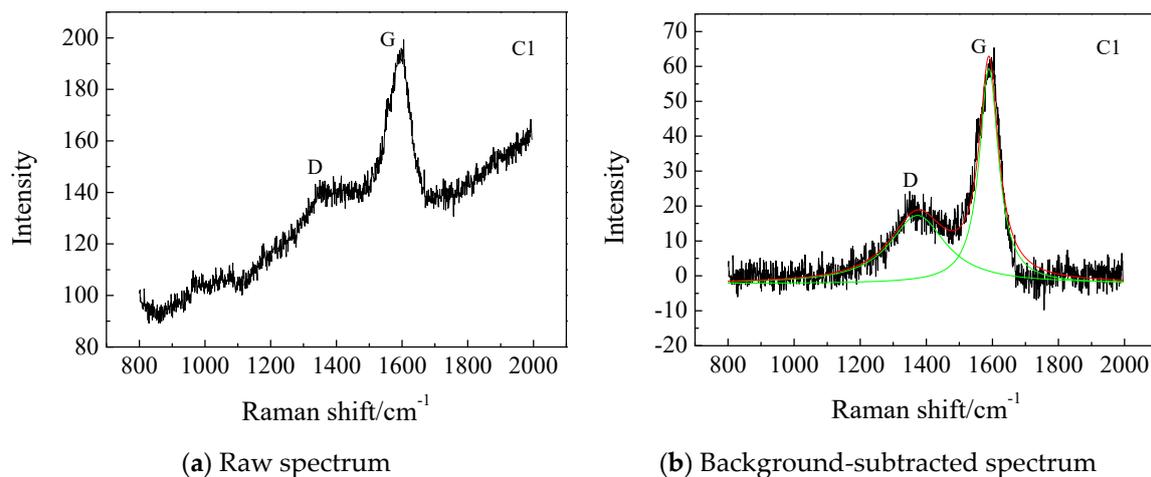
Sample ID	N (%)	C (%)	H (%)	S (%)	O (%)
C1	1.84	79.69	4.67	0.27	11.50
C2	1.92	77.84	4.72	0.41	7.23
C3	1.81	78.56	4.50	0.58	11.00
C4	1.88	78.54	4.26	0.75	10.70
C5	2.12	76.63	5.36	0.64	9.76
C6	1.84	77.02	4.62	0.44	11.80

The spectral parameters (e.g.,  $P_G$ ,  $P_D$ ,  $I_D/I_G$ ,  $W_D$ , and  $W_G$ ) were calculated from the deconvoluted D and G bands and are discussed as below. All Raman spectra were fitted three times, and the data shown above indicate the average value. The curve-fitting procedure was deconvoluted by Origin 8.0.

## 3. Results and Discussion

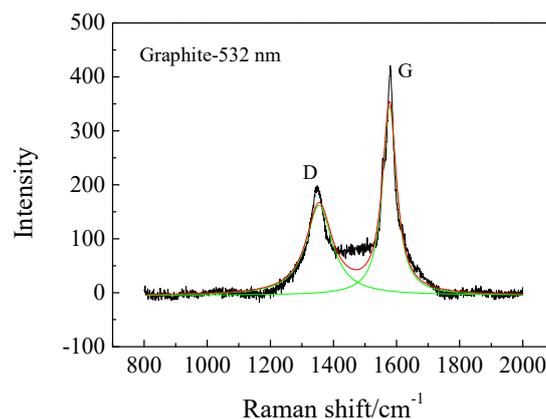
### 3.1. Room Temperature Raman Spectra of a Typical Coal Sample

Figure 2a shows a typical room temperature Raman spectrum obtained from one of the tested coals (Coal C1). Due to the low transmission efficiency of the ultraviolet light used in the spectrometer at 355 nm, a strong fluorescence interference, as evidenced by the sloping background with the Raman shift, was observed. To avoid the interference from the background, the background was subtracted using the baseline method. The background-subtracted spectrum is shown in Figure 2b.



**Figure 2.** A typical room-temperature Raman spectrum of one of the tested coals, C1.

Two bands are discernible at approximately 1360 and 1590  $\text{cm}^{-1}$ . These bands correspond to the D and G bands for graphite in Figure 2. Due to a disordered and defective diamond band or an  $\text{sp}^3$  hybridization carbon, the D band was assigned to the  $\text{A}_{1g}$  mode, which relates to the average graphite domain dimension, while the G band was assigned to the graphite  $\text{E}_{2g}$  mode with  $\text{D}^4_{6h}$  crystal symmetry and resulted from a graphitic, tangential band or  $\text{sp}^2$  hybridization [22]. However, unlike what observed for graphite, the bands were very diffuse and therefore overlapped, most likely due to the fluorescent background. In order to determine the wavenumber position, the relative peak intensity, and the width or full width at half maximum (FWHM), the overlapped Raman bands were further deconvoluted into two well-defined peaks using origin 8.0. (Lorentz function), as shown in Figure 2b. The deconvoluted peaks were then discussed to reveal the structural difference and structural evolution of the coals during pyrolysis. A typical room temperature Raman spectrum of the graphite is shown in Figure 3. In comparison to Figure 3, the coal in Figure 2b showed that both D and G bands had shifted slightly towards the red spectral range.



**Figure 3.** A typical room temperature Raman spectrum of the tested graphite.

Furthermore, the difference in the band position of  $P_G$ ,  $P_D$ ,  $W_D$ , and  $W_G$  and the area ratio ( $I_D/I_G$ ) of the G and D bands were calculated.

### 3.2. Microstructure of Coals of Varying Ranks

Figure 4 shows the background subtracted Raman spectra of the tested coal samples. The coal samples clearly showed similar spectral profiles, having two distinct diffuse and overlapped D and G bands.

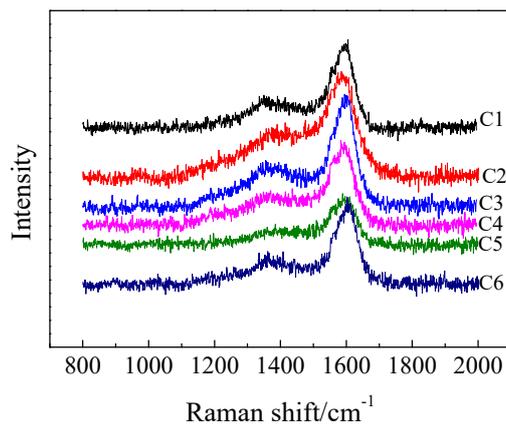


Figure 4. Background subtracted spectra of coals of different ranks.

As shown in Figure 5a, as the carbon content of the coals increased, the deconvoluted D and G bands shifted little. Within a certain range, the D and G bands of the coals seemed independent of the carbon content. However, compared with the G band, the D band shifted more noticeably, suggesting that the carbon content had a greater effect on the D band than on the G band. The D band represents the defects in the lattice of the C atom, while the G band represents the in-plane stretching vibration of the  $sp^2$  hybridization of the C atom. With the increase in carbon content, the C=C in the coal samples increases, but the number of disordered substances decreases [23].

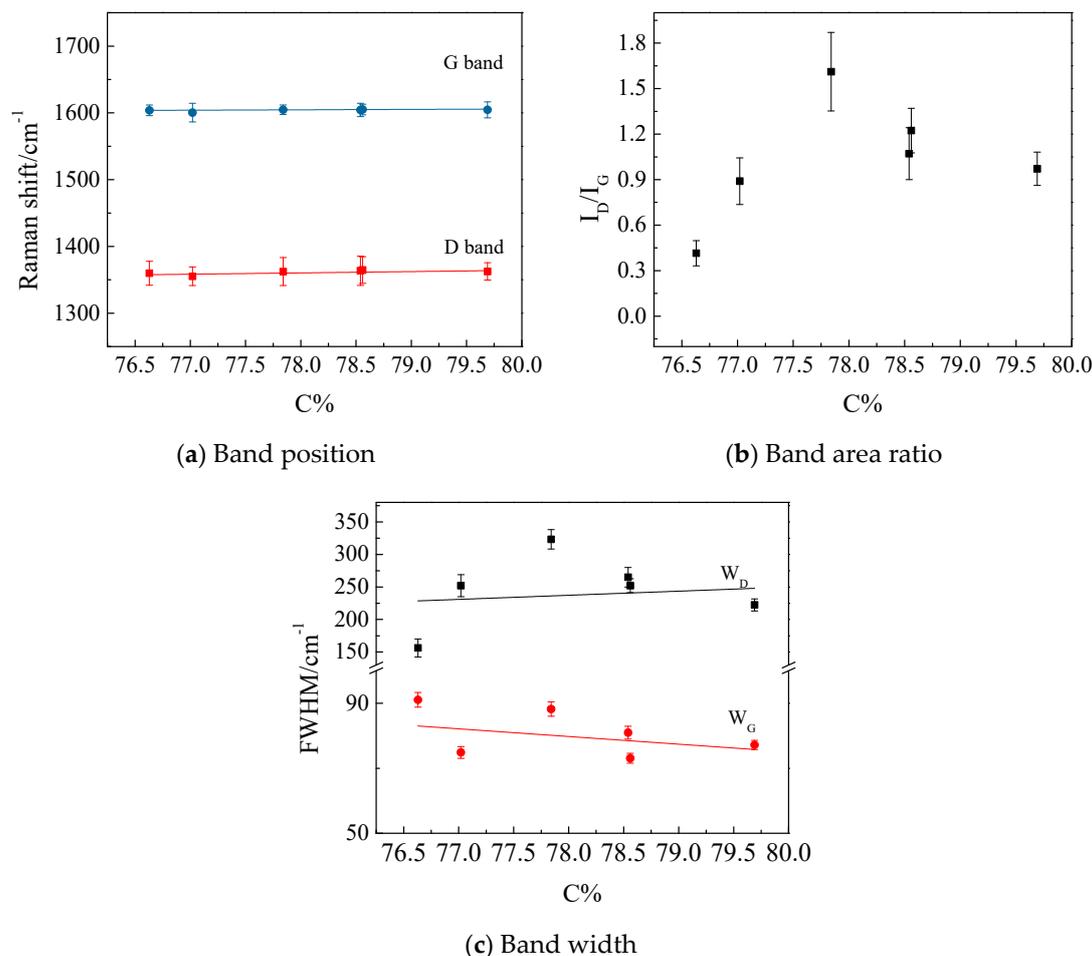


Figure 5. Effect of coal rank on (a) the band position ( $P_G$  and  $P_D$ ), (b) area ratio ( $I_D/I_G$ ), and (c) width of the coal samples.

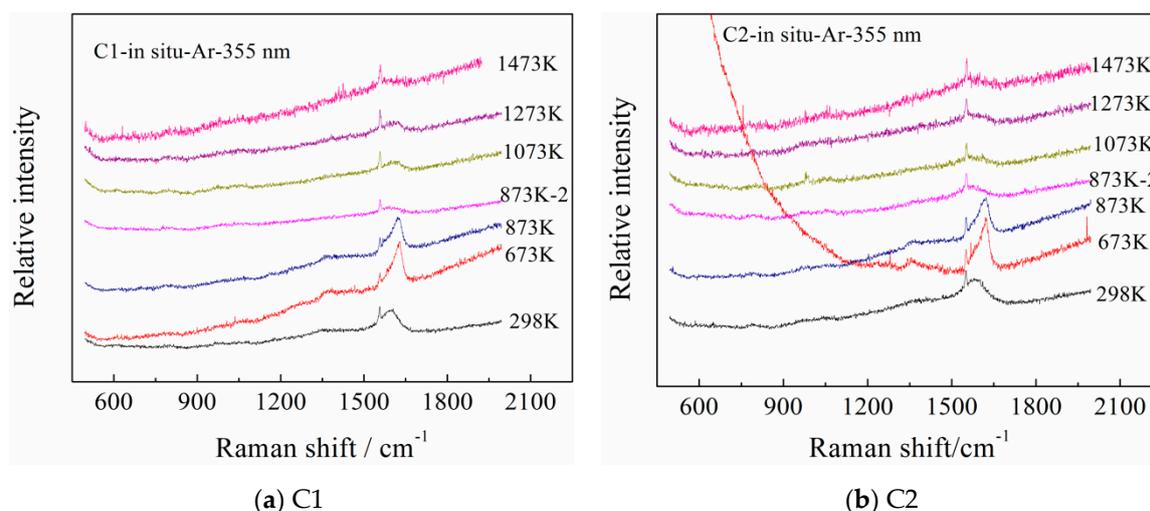
Figure 5b shows the effect of carbon content on the area ratio  $I_D/I_G$  of the deconvoluted D and G bands. As shown in Figure 5b,  $I_D/I_G$  increased linearly with the carbon content in the coals. The higher the value of  $I_D/I_G$ , the greater the number of defects that appeared in the C atom. As a result, the defects of the C atom increased with an increase of the carbon content [24].

The effect of the carbon content on the width (FWHM) of the deconvoluted G and D bands is presented in Figure 5c. The D band became broader as the carbon content of the coals increased. By contrast, the G band showed a negligible change.  $W_G$  can reflect the degree of ordering of carbon materials.

The parameters affected each other.  $I_D/I_G$  is used to quantitatively evaluate the degree of disorder of carbon materials. All parameters of the peaks G and D were affected by the first coalification jump.

### 3.3. In Situ High-Temperature Study of Different Coals by RAMAN

Figure 6a shows the in-situ temperature-dependent Raman spectra of coal C1 under an Ar atmosphere. The gas flow was monitored and controlled by the amount of bubbles from water—approximately 160 bubbles per minute. When the sample temperature was lower than 473 K, the sample was heated at a rate of 30 K per min. The heating rate was increased to 50 K per min when the sample temperature was higher than 473 K. The integral time and frequency were  $20 \times 20$  and were increased to  $30 \times 30$  at a temperature of 873 K. A yellow substance was observed at a temperature of approximately 673 K. This substance was believed to be tar oil [25]. The spectrum labeled as 873 K-2 appeared during the second recording, when the sample was reheated to 873 K after it was held at 873 K for 40 min. The sample was then cooled down at 50 K/min to room temperature. This cooling allowed for the cleaning of the silica glass prior to measurement.



**Figure 6.** Temperature dependence of in-situ Raman spectra of C1 and C2 under an Ar atmosphere using a 355 nm excitation laser.

Figure 6b shows the in-situ temperature-dependent Raman spectra of coal C2 under an Ar atmosphere. The integral time and frequency were set at  $30 \times 30$  s in the entire heating process. From Figure 6, it is clear that the peak position of the D band related to the low molecular weight carbon organics, while the intensity of the D band changed negligibly with the increasing temperature [26,27]. This implies an increase in the graphitization degree along with the increase in temperature. The low molecular weight carbon organics condensed considerably, starting to devolatilize at 673 K. Some of these organics stayed on the coal samples in a liquid state for some time. Most heavy molecular weight carbon organics seemed to condense and devolatilize starting at a temperature of 873 K. This suggests that there could be more heavy molecular weight carbon organics condensing and devolatilizing or moving from the coal samples in the range of 873 K to 1273 K, disappearing rapidly at 1473 K.

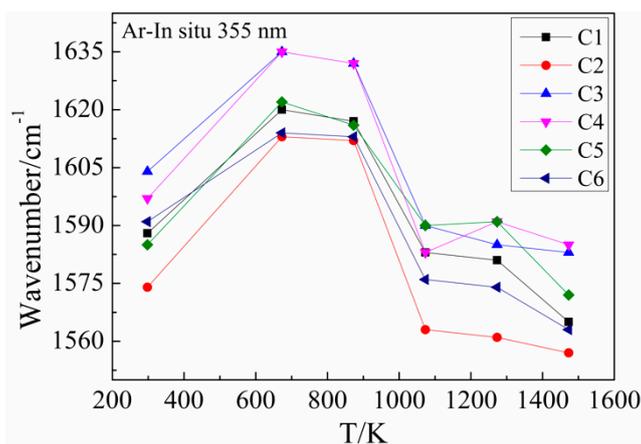
By contrast, the G peak position showed a blue shift with an increase in temperature, and its intensity became weaker due to the decreasing devolatilized content in the samples. This demonstrated that the G peak position was affected by the devolatilization of the coal samples.

Generally, the Raman band would shift to a lower wavenumber and become smaller in intensity with an increase in temperature if no phase transformation takes place.

From Figure 6b, it is clear that the spectrum at 673 K seemed quite different from the corresponding spectrum of C1 in Figure 6a. This difference was the result of the appearance and extension of the Rayleigh line. During pyrolysis, small liquid molecules on the surface where the sampling point is focused can occasionally enhance Rayleigh scattering, but this phenomenon occurs differently each time.

This finding may imply the presence of a liquid substance on the coal samples. A gas, liquid, and solid three-phase coexistence was reported during coal pyrolysis from 673 to 873 K [15,28]. When the temperature exceeded the curing temperature of 873 K, an adhesion phenomenon seemed to occur, resulting in the formation of tar oil or char [29,30], which covered the silica glass. The samples then needed to be cooled down at 50 K per min, from 873 K to room temperature, which allowed the silica glass to be cleaned. Finally, the samples were reheated to 873 K for further observation.

The effect of temperature on the peak position variations of the G band of the samples pyrolyzed under an Ar atmosphere is shown in Figure 7. Generally, the spectral lines show a linear decline trend as the temperature increases in the in-situ temperature-dependent Raman spectra [31,32]. As a result of the condensation of the low molecular weight carbon organics that was discovered to be coal tar [33], this variation in the peak position of the G band was complex and consisted of the coalification processes of condensation, crystallization, coking, and devolatilization during the entire pyrolysis process. As shown in Figure 7, from room temperature to 673 K, a blue shift was observed for the G band as more coal tar that was generated and adhered to the coal samples evaporated. A critical point was also observed at a temperature of 873 K. A red shift was observed, which was most likely due to the normal displacement and the devolatilization of coal tar generated in large quantities on the surface of the coal samples. This was followed by the coking stage from 873 to 1073 K and the devolatilization stage from 1073 to 1473 K. At the same time, the peak positions of the G band at 1073 and 1273 K were quite different from the peak position of the second value at 873 K. This finding confirms the results of the previous discussion [34].



**Figure 7.** In-situ temperature dependence of the peak position of the G band for the samples pyrolyzed under an Ar atmosphere.

#### 4. Conclusions

The carbon content has little effect on the positions of the D band and G band. The parameters ( $P_G$ ,  $P_D$ ,  $I_D/I_G$ ,  $W_D$ , and  $W_G$ ) of the D and G bands were discussed. The parameters affected each other. Among them,  $I_D/I_G$  is used to quantitatively evaluate the degree of disorder of carbon materials.

During an in-situ high temperature-dependent pyrolysis process, the peak position of the D band seems related to the low molecular weight carbon organics. Since the low molecular weight carbon organics condensed and devolatilized, the intensity of the D band changed negligibly with an increase in temperature. Meanwhile, the position of the G band was shifted to blue, and its intensity became weaker with an increase in temperature, which was the result of a decrease in devolatilized contents in the samples. The G peak position variation was affected by condensation and devolatilization of the low molecular weight carbon organics and the heavy molecular weight carbon organics inside the coal samples. This variation in the peak position of the G band reflects the coalification processes of condensation, crystallization, coking, and devolatilization during the entire pyrolysis process.

Raman spectroscopy has clear potential for the diagnostic identification of coal and for coking optimization.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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