



## Article

# Examining the Impact of Dimethyl Sulfide Emissions on Atmospheric Sulfate over the Continental U.S.

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**Abstract:** We examined the impact of dimethylsulfide (DMS) emissions on sulfate concentrations over the continental U.S. by using the Community Multiscale Air Quality (CMAQ) model version 5.4 and performing annual simulations without and with DMS emissions for 2018. DMS emissions enhance sulfate not only over seawater but also over land, although to a lesser extent. On an annual basis, the inclusion of DMS emissions increase sulfate concentrations by 36% over seawater and 9% over land. The largest impacts over land occur in California, Oregon, Washington, and Florida, where the annual mean sulfate concentrations increase by ~25%. The increase in sulfate causes a decrease in nitrate concentration due to limited ammonia concentration, especially over seawater, and an increase in ammonium concentration with a net effect of increased inorganic particles. The largest sulfate enhancement occurs near the surface (over seawater), and the enhancement decreases with altitude, diminishing to 10–20% at an altitude of ~5 km. Seasonally, the largest enhancement in sulfate over seawater occurs in summer, and the lowest in winter. In contrast, the largest enhancements over land occur in spring and fall due to higher wind speeds that can transport more sulfate from seawater into land.

**Keywords:** DMS; seawater; SO<sub>2</sub>; sulfate; CMAQ



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## 1. Introduction

Sulfate is an important component of atmospheric fine particles and can also affect the cloud condensation nuclei and the aerosol radiative forcing [1–4]. Sulfate is formed in the atmosphere mainly from the oxidation of sulfur dioxide (SO<sub>2</sub>) via gas- and aqueous-phase chemistry. Sulfur can be emitted from natural as well as anthropogenic sources. Significant reductions in anthropogenic SO<sub>2</sub> emissions from fossil fuel combustion have occurred in the U.S. during the past decades due to stringent control measures in response to tightened air quality standards. For example, the U.S. Environmental Protection Agency (EPA) reported a 93% reduction in anthropogenic SO<sub>2</sub> emissions during 1980–2021 despite a 187% increase in gross domestic product, 111% increase in vehicle miles travelled, and 25% increase in energy consumption during the same period (<https://www.epa.gov/air-trends/air-quality-national-summary#air-quality-trends> (accessed on 24 August 2022)). Similar changes have occurred in many other parts of the world [5]. The contribution of sulfate to fine particles varies with time and location. Tanner et al. [6] analyzed trends in annual pollutant concentrations in eastern Tennessee and reported that ammonium sulfate contributed 48% in 1999 and 41% in 2013 to fine particle mass concentrations in Look Rock, Tennessee. Malm et al. [7] performed an analysis of measured sulfate concentrations at the IMPROVE (Interagency Monitoring of Protected Visual Environments) sites for 2001–2015 and reported that, on average, ammonium sulfate contributed 30% to the summertime reconstructed fine particle mass. There are some specific regions in the U.S. where other

species contribute a substantial fraction to fine particle mass. For example, aerosol nitrate and organic aerosol contribute ~85% of wintertime fine particle mass in San Joaquin Valley in California [8]. Atmospheric sulfate concentrations in the U.S. have decreased over the years; however, they are still an important constituent of fine particles.

Dimethyl sulfide (DMS;  $\text{CH}_3\text{SCH}_3$ ) can be emitted from natural as well as anthropogenic sources. Natural sources of DMS include emissions from the ocean, wetlands, and plants and soil, with the ocean being the largest source [9]. Phytoplankton in seawater [9] contain dimethylsulfoniopropionate (DMSP), which produces DMS upon breakdown. DMS emissions can also be emitted from anthropogenic sources. For example, paper and pulp industries and rayon/cellulosics manufacturing facilities can emit DMS; however, their emissions are an order of magnitude lower than natural emissions [10]. Estimates of oceanic DMS emissions range between 17.6 and 34.4 Tg S yr<sup>-1</sup> [3] and are ~10 times larger than the anthropogenic DMS emission estimates of 2.20 Tg S yr<sup>-1</sup> [10]. Once emitted into the atmosphere, the oxidation of DMS forms sulfur-containing products such as SO<sub>2</sub> and sulfuric acid [11,12]. These products then lead to the formation of sulfate aerosols. While the anthropogenic sulfur emissions are decreasing, natural sulfur emissions remain mostly unchanged. Thus, natural oceanic DMS emissions play an increasingly important role in the atmosphere, and their impact on sulfate needs a careful examination. Several previous studies have been conducted to examine the impact of natural sulfur emissions on sulfate [3,13–20]. However, most of these studies employed global or hemispheric models and were conducted with older modeling systems for historic conditions which did not reflect recent decreases in anthropogenic sulfur emissions and/or covered only limited areas over the U.S. For example, Park and coworkers [13] used the GEOS-Chem version 5.03 model ([geos-chem.seas.harvard.edu](http://geos-chem.seas.harvard.edu)) with a horizontal resolution of 2° × 2.5° and conducted simulations for 2001. They included natural sulfur emissions from oceans, volcanoes, and biomass burning sources and reported that these emissions can enhance ammonium sulfate concentrations in the U.S. Zhao et al. [20] implemented oceanic DMS emissions in the CMAQv5.3 modeling system [21]. Using the hemispheric Community Multiscale Air Quality (CMAQ) model ([www.epa.gov/cmaq](http://www.epa.gov/cmaq), accessed on 15 December 2022) version 5.3 with 108 km horizontal grids, they reported that DMS emissions can enhance surface sulfate concentration over many areas of the northern hemisphere, including the U.S., but they did not conduct any regional-scale CMAQ simulations. In one of the two previous regional-scale modeling studies we are aware of, Mueller et al. [18] applied CMAQv4.6 with a horizontal resolution of 36 × 36 km for 2002 atmospheric conditions to examine the impact of natural emissions on sulfate. They included natural sulfur emissions from the ocean, coastal wetlands, freshwater lakes, the Great Salt Lake, soil, and geothermal sources; performed simulations over the continental U.S., Canada, Mexico, and surrounding oceanic areas; and reported that these emissions can enhance ammonium sulfate over the U.S. Perraud et al. [19] examined the impacts of a hypothetical scenario with zero anthropogenic SO<sub>2</sub> emissions in a large coastal urban area in California. They applied a three-dimensional air quality model (The University of California, Irvine—California Institute of Technology regional airshed model) with 5 km horizontal grids over California and reported that particle formation would be reduced by two orders of magnitude in the case of zero anthropogenic SO<sub>2</sub> emissions; however, particles would continue to form from natural and anthropogenic sources of organosulfur compounds. In this study, we add to this body of knowledge generated in earlier studies by examining the impact of oceanic DMS emissions on sulfate over the continental U.S. for recent atmospheric conditions using the latest version of CMAQ (version 5.4) with horizontal grid spacings of 12 km and constrained by lateral boundary conditions from corresponding large-scale CMAQv5.4 simulations performed over the northern hemisphere.

## 2. Materials and Methods

To quantify the impact of oceanic DMS, we used a simulation without DMS emissions (NODMS) and a simulation with DMS emission (WDMS). They both shared the base

model configuration and emissions (except DMS). The WDMS simulation used the on-line implementation of DMS emissions, while the NODMS simulation did not use the on-line DMS emissions. Both simulations included additional chemistry implemented for DMS. The base model configuration, the additional chemistry, and DMS emissions implementations are described below.

### 2.1. Base Model Configuration

CMAQ is a state-of-the-science three-dimensional chemical transport model containing detailed treatments of important atmospheric processes. It has been widely used in the U.S. and many other countries for air quality simulations involving research as well as regulatory activities [21–25]. Here, we used the recently released CMAQv5.4 [26] for simulating air quality for the entire year of 2018. The modeling domain covered the entire continental U.S., parts of Canada and Mexico, and surrounding oceanic areas with 12 km horizontal grid spacings and 35 vertical layers of varying thickness with a surface layer height of 20 m. The modeling domain contained a total of  $459 \times 299 \times 35$  (4,803,435) grid cells. Meteorological fields were generated using the Weather Research and Forecasting (WRFv4.3.3, [27]) model with the RRTMG (Rapid Radiative Transfer Model for GCMs) for long- and short-wave radiation [28], Kain–Fritsch scheme [29] with lightning assimilation [30,31] for convective parameterization, Morrison microphysics scheme [32], Pleim–Xiu land surface model [33,34], and the Asymmetric Convective Model version 2 (ACM2; [35,36]) for the planetary boundary layer (PBL) scheme. Four-dimensional data assimilation (FDDA) using “analysis nudging” was applied to continuously nudge the wind, temperature, and water vapor mixing ratio above the PBL [37] toward the NAM 12 km analysis every three hours. The Meteorology Chemistry Interface Processor (MCIPv5.3.3) [38] was used to process the WRF results into CMAQ-ready input files. Torres-Vazquez et al. [39] compared WRFv4.1.1 predictions with observed data over the U.S. (same modeling domain used in this study) and reported model performance for wind speed, wind direction, temperature, and mixing ratios. The model configuration between the two studies is same, except that we used WRFv4.3.3 and they used WRFv4.1.1. The WRF model performance in this study is similar to that reported by Torres-Vazquez et al. [39]. We employed the standard cloud and aero7 modules, and the Carbon Bond chemical mechanism (version 6, release 5 (CB6r5)) [40] with chlorine chemistry [41].

The model included emissions from various natural and anthropogenic sources. Anthropogenic source emissions include contributions from mobile, point, and non-point sources and were obtained from the 2017 National Emissions Inventory and projected for 2018. Additional information on different types of emissions can be found in EPA [42]. Mobile source emissions include contributions from on- and off-road vehicles, commercial marine vessels, and railroads. Point source emissions include contributions from electric generating units, oil and gas facilities, industrial sources not included in electric generating units and oil/gas facilities, agricultural field burning activities, wild/prescribed burning activities, and airports. Non-point source emissions include contributions from agriculture, area fugitive dust, oil and gas facilities, residential wood combustion, and others. In addition, emissions from Mexico and Canada were also included. The details of these emissions can be found at the EPA website ([www.epa.gov/cmaq/equates#emissions\\_modeling](http://www.epa.gov/cmaq/equates#emissions_modeling), accessed on 10 December 2022). Biogenic emissions were calculated using the in-line Biogenic Emission Inventory System (BEIS4). In-line sea-salt spray emissions were calculated following Gantt et al. [43], while in-line lightning emissions were calculated following Kang et al. [44]. The simulation in this study did not include any wind-blown dust emissions.

### 2.2. DMS Emission Implementation

In addition, we included DMS flux (F) from the ocean, estimated using the total gas transfer velocity ( $k_T$ ) and the DMS concentration in seawater ( $C_w$ ) (F) as follows [3]:

$$F = k_T \times C_w \quad (1)$$

DMS concentrations in seawater can be obtained from several databases [3,45,46]. Here, we used the monthly climatological seawater concentration from the Surface Ocean and Lower Atmosphere project ([www.bodc.ac.uk/solas\\_integration/implementation\\_products/group1/dms](http://www.bodc.ac.uk/solas_integration/implementation_products/group1/dms), accessed on 26 April 2022) [3] for calculating the DMS flux. The  $k_T$  in Equation (1) is calculated using the water side gas transfer velocity ( $k_w$ ) and the atmospheric gradient fraction ( $\gamma$ ) as follows:

$$k_T = k_w \times (1 - \gamma) \quad (2)$$

$k_w$  can be estimated using several parameterizations [47–49]. Lana et al. [3] estimated global DMS emissions using these parameterizations and reported that the Wanninkhof [49] parameterization produces the highest estimates, and the Liss and Merlivat [47] parameterization produces the minimum estimates. Consistent with the study of Zhao et al. [20], for the northern hemisphere, we also used the Liss and Merlivat [47] parameterization (Equation (3)) to calculate DMS flux.

$$k_w = \begin{cases} \frac{0.17U_{10}}{(S_{CDMS}/600)^{2/3}} & U_{10} \leq 3.6 \text{ m/s} \\ \frac{2.85U_{10}-9.65}{(S_{CDMS}/600)^{1/2}} & 3.6 < U_{10} \leq 13 \text{ m/s} \\ \frac{5.9U_{10}-49.3}{(S_{CDMS}/600)^{1/2}} & U_{10} > 13 \text{ m/s} \end{cases} \quad (3)$$

The Schmidt number ( $S_{CDMS}$ ) describes the diffusion of DMS in seawater and is calculated using sea surface temperature ( $T_{Sw}$ , °C) [50]:

$$S_{CDMS} = 2674.0 - 147.12 \times T_{Sw} + 3.726 \times T_{Sw}^2 - 0.038 \times T_{Sw}^3 \quad (4)$$

The atmospheric gradient fraction,  $\gamma$ , can be calculated using Equation (5):

$$\gamma = 1.0 / (1.0 + k_a / Hk_w) \quad (5)$$

$H$  is dimensionless Henry's Law Coefficient (liquid/gas) and can be calculated as follows:

$$H = h_{Dacey} \times R \times T_{Sw} \quad (6)$$

where  $R$  is the universal gas constant (0.08205746 atm L/K mol),  $T_{Sw}$  is the sea surface temperature, and  $h_{Dacey}$  is the solubility of DMS:

$$h_{Dacey} = \frac{1.0}{e^{(-\frac{3547.0}{273.15+T_{Sw}})+12.64}} \quad (7)$$

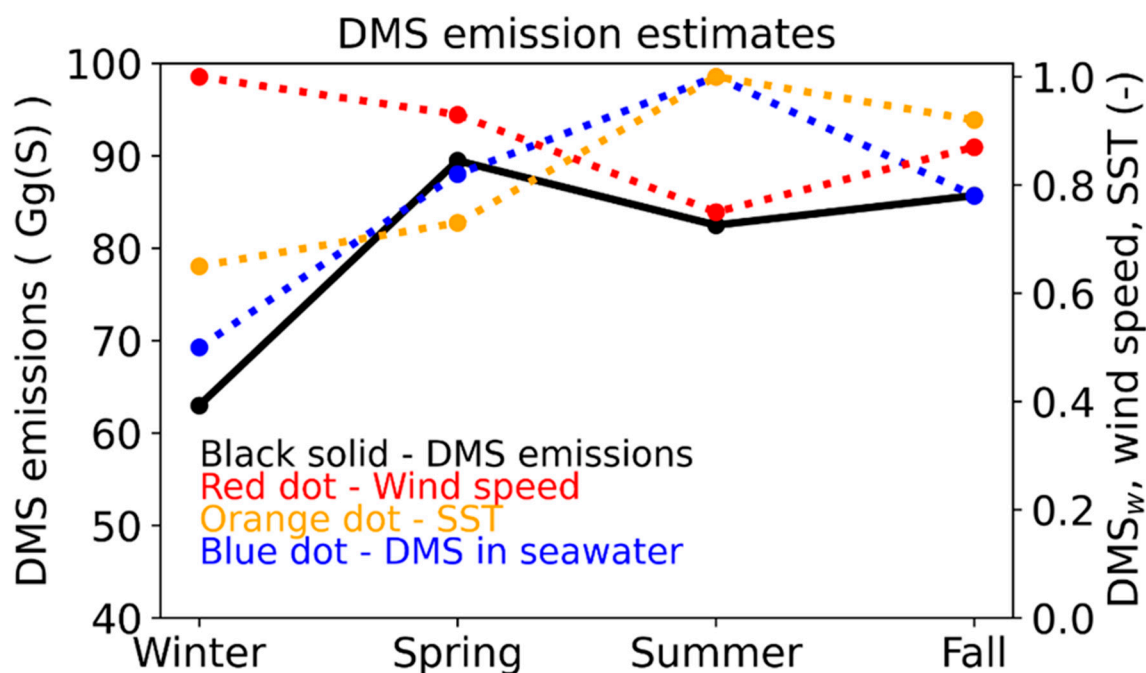
The wind-speed-dependent airside transfer coefficient,  $k_a$  is defined by

$$k_a = 659 \times U_{10} \sqrt{\left(\frac{M_{DMS}}{M_{H_2O}}\right)} \quad (8)$$

where  $M_{DMS}$  is the molecular weight of DMS and  $M_{H_2O}$  is the molecular weight of water [51].

Seasonal DMS emissions (estimated using the scheme described above) over the modeling domain (only grid cells emitting DMS emissions are used for the calculation) are shown in Figure 1. The highest estimate (89.5 Gg(S)—gigagram of sulfur) occurred in spring (March–May) and the lowest (62.9 Gg(S)) in winter. DMS emissions estimates in summer (June–August) and fall (September–November) were lower than those in spring, but higher than those in winter (December–February). Mean DMS concentration in seawater, wind speed, and sea surface temperature (SST) are also shown in Figure 1. DMS emissions increased with higher DMS concentrations in seawater, higher wind speed, and higher SST. Both the DMS concentration in seawater and SST were the highest in summer and

the lowest in winter. However, wind speed was the highest in winter and the lowest in summer. Summertime DMS emissions were lower than those in spring due to lower wind speed. The lowest DMS emissions in winter were driven by the lowest DMS concentration in seawater and SST. The annual DMS emissions were 320 Gg(S) with ~21 Gg(S) in January and ~29 Gg(S) in July. Smith and Mueller [52] used a similar domain extent and reported 18 Gg(S) in January and 45 Gg(S) in July. January emissions estimates between the two studies vary by <20%, while July emissions estimates vary by ~50%, likely due to the differences in the temporal allocation of annual emissions between the two studies.



**Figure 1.** Seasonal DMS emissions over the modeling domain calculated using the Liss and Merlivat [47] parameterization for the water side gas transfer velocity ( $k_w$ ). DMS<sub>w</sub>, wind speed, and SST values shown on the right y-axis are dimensionless. Normalized values were calculated by dividing the actual values by their maximum values. Winter represents December–February, spring represents March–May, summer represents June–August, and fall represents September–November.

### 2.3. DMS Chemistry Implementation

DMS chemistry was previously implemented in the hemispheric CMAQ model [20]. Seven chemical reactions involving the oxidation of DMS by hydroxyl radical (OH), nitrate radical (NO<sub>3</sub>), chlorine radical (Cl), chlorine oxide (ClO), bromine oxide (BrO), and iodine oxide (IO) were used in the hemispheric CMAQ model. The CB6r5 chemical mechanism does not include any bromine or iodine chemistry; thus, the reactions of DMS with BrO and IO were not included in this study. In addition, since the reaction of ClO only contributes ~0.1% of the total DMS oxidation [20], it was also excluded in this study. As shown in Table 1, four chemical reactions for DMS oxidation were included. DMS chemistry produces SO<sub>2</sub>, which is then further oxidized by gas-phase reaction with OH to produce sulfate aerosol. The in-cloud oxidation of SO<sub>2</sub> by hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>), ozone (O<sub>3</sub>), methyl hydroperoxide (MEPX), peroxyacetic acid (PACD), and oxygen catalyzed by iron (Fe[III]) and manganese (Mn[II]) further produces aerosol sulfate.

### 2.4. Boundary Conditions

Boundary conditions were generated from the hemispheric CMAQ model [56] simulations using CMAQv5.4. Two different model simulations were performed using the hemispheric CMAQv5.4: one simulation without DMS emissions and the other with DMS emissions. Results obtained with these hemispheric simulations were used to prepare



two different sets of boundary conditions for the 12 km modeling domain for this study. The first 12 km simulation contained no DMS emissions and used the hourly boundary condition derived from the hemispheric model without DMS emissions (NODMS); the second 12 km simulation contained DMS emissions and used the hourly boundary condition derived from the hemispheric model with DMS emissions (WDMS). The differences in the model results are thus solely attributable to the DMS emissions that occur in both the 12 km and hemispheric domain. Initial conditions for the model were prepared using results from a previously generated model simulation for the EPA's Air Quality Time Series (EQUATES) ([www.epa.gov/cmaq/equates](http://www.epa.gov/cmaq/equates), accessed on 12 October 2022). The simulation was started on 22 December 2017 and continued until 31 December 2018. The simulations for the first 10 days were used for the spin-up purpose. The analyses of the model results were performed for the entire year of 2018.

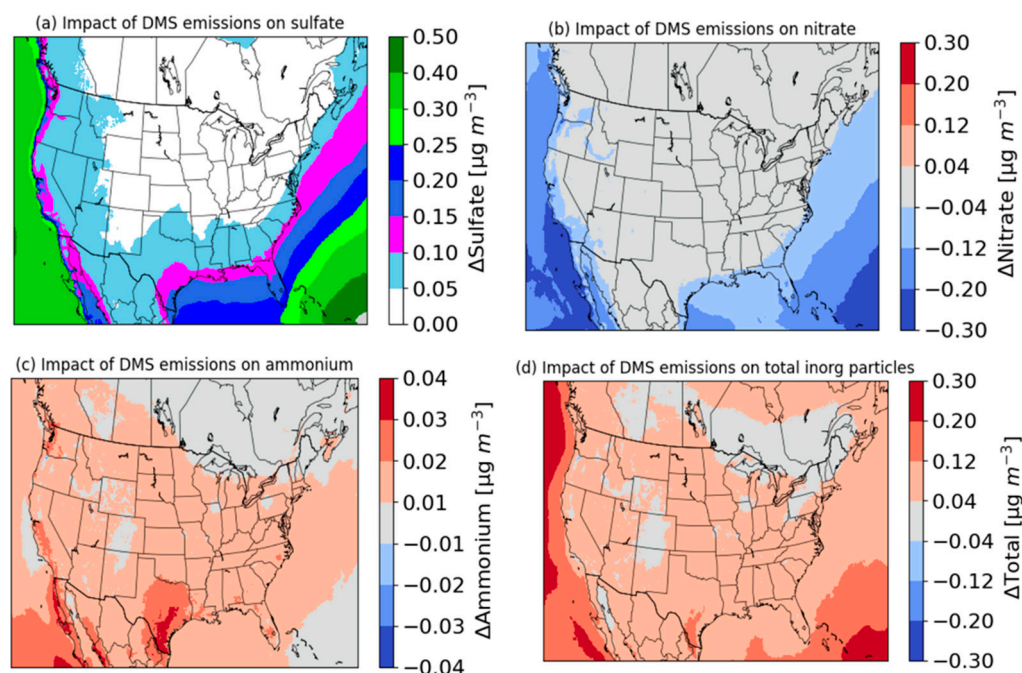
**Table 1.** Gas-phase chemical reactions for DMS oxidation in CMAQ.

No.	Reaction	Rate Expression ( $\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ )	References
1	$\text{DMS} + \text{OH} = \text{SO}_2 + \dots$ (abstraction channel)	$k = 1.12 \times 10^{-11} e^{-250/T}$ $T = \text{temperature in Kelvin}$	[53]
2	$\text{DMS} + \text{OH} = 0.75 \times \text{SO}_2 + \dots$ (addition channel)	$k_0 = 1.99 \times 10^{-39} e^{-5270/T}$ $k_\infty = 1.26 \times 10^{-10} e^{+340/T}$ $k = \{k_0[M]/(1 + k_0[M]/k_\infty)\} F^Z$ $Z = \{(1/N) + \log_{10}[k_0[M]/k_\infty]^2\}^{-1}$ $F = 1.0$ and $N = 1.0$ $[M] = \text{total pressure, molecules/cm}^3$	[53]
3	$\text{DMS} + \text{NO}_3 = \text{SO}_2 + \dots$	$k = 1.93 \times 10^{-13} e^{+520/T}$	[53]
4	$\text{DMS} + \text{Cl} = 0.86 \times \text{SO}_2 + \dots$	$k = 3.4 \times 10^{-13} e^{+2081/T}$	[54,55]

### 3. Results and Discussion

#### 3.1. Annual Impact of DMS Emissions on Surface Inorganic Particles

Predicted annual mean changes (WDMS—NODMS) in surface sulfate, nitrate, ammonium, and total inorganic (sulfate + nitrate + ammonium) particle concentrations due to DMS emissions are shown in Figure 2. DMS emissions enhance sulfate concentrations not only over seawater but also over land area. The largest enhancements occur over seawater and then gradually decrease over the interior portion of the modeling domain. DMS is emitted from the ocean and has a relatively short lifetime (~0.5 day; [57]); thus, the highest enhancements occur over seawater. The impact is particularly high over the Pacific Ocean due in part to the prevailing wind direction carrying the impact of western boundary conditions into the modeling domain. DMS emissions enhance sulfate concentrations by 0.10–0.50  $\mu\text{g}/\text{m}^3$  (12–67%) over a large portion of seawater, with a mean over-ocean enhancement of 0.24  $\mu\text{g}/\text{m}^3$  (36%). DMS emissions also enhance sulfate concentrations by 0.05–0.25  $\mu\text{g}/\text{m}^3$  over a large area of land, with a mean enhancement of 0.055  $\mu\text{g}/\text{m}^3$  (9%). The impacts near coastal areas are greater than those over the interior portion of land. The largest impacts over land occur in coastal areas of Washington, Oregon, California, southern Texas, and Florida, in which they generally increase sulfate in some areas by more than 30–40%. They also increase sulfate by up to ~50% in some grid cells along the coastal areas. Considering the entire state, DMS emissions increase annual mean sulfate by 0.15  $\mu\text{g}/\text{m}^3$  (23%) over Florida, 0.11  $\mu\text{g}/\text{m}^3$  (24%) over California, 0.09  $\mu\text{g}/\text{m}^3$  (25%) over Oregon, 0.09  $\mu\text{g}/\text{m}^3$  (25%) over Washington, and 0.09  $\mu\text{g}/\text{m}^3$  (9%) over Texas. While the impacts over some coastal states are relatively large, the impacts over many interior states covering large areas of the modeling domain are relatively small (<0.05  $\mu\text{g}/\text{m}^3$ ; ~4–6%).



**Figure 2.** Predicted annual mean changes (WDMS—NODMS) in surface (a) sulfate, (b) nitrate, (c) ammonium, and (d) total inorganic particle concentrations by DMS emissions. Changes in inorganic particle concentrations are calculated as the net effect due to increases in sulfate and ammonium and decreases in nitrate concentrations.

DMS emissions decrease nitrate concentrations by  $0.04\text{--}0.28\text{ }\mu\text{g}/\text{m}^3$  (6–67%) over seawater and  $0\text{--}0.28\text{ }\mu\text{g}/\text{m}^3$  (0–69%) over coastal areas and increase ammonium concentrations by up to  $0.04\text{ }\mu\text{g}/\text{m}^3$  (~14% at the location of the maximum increase). Overall, they decrease mean nitrate by  $0.13\text{ }\mu\text{g}/\text{m}^3$  (39%) over the entire seawater and  $0.016\text{ }\mu\text{g}/\text{m}^3$  (5%) over the entire land area and increase annual mean ammonium by  $0.013\text{ }\mu\text{g}/\text{m}^3$  (35%) over the entire seawater and  $0.012\text{ }\mu\text{g}/\text{m}^3$  (7%) over the entire land area. Ammonia preferentially reacts with sulfuric acid, forming ammonium sulfate or bisulfate [9]. Any remaining ammonia then reacts with nitric acid, forming ammonium nitrate. Ammonia concentrations, especially over seawater, tend to be limited and are consumed by the additional sulfuric acid produced by DMS emissions, increasing both the ammonium and sulfate concentrations and decreasing nitrate concentrations. Ammonium nitrate is replaced with ammonium sulfate or bisulfate; thus, the impact on ammonium tends to be smaller than the impacts on sulfate and nitrate. The impact of DMS emissions on total inorganic particles is calculated as the net effect due to increases in sulfate and ammonium and decreases in nitrate concentrations. The impacts on total inorganic particles are lower than those on sulfate due to the decreases in nitrate concentrations. However, DMS emissions still increase annual mean total inorganic particles by  $0.04\text{--}0.30\text{ }\mu\text{g}/\text{m}^3$  (5–33%) over seawater and  $0.04\text{--}0.20\text{ }\mu\text{g}/\text{m}^3$  (2–20%) over land.

We can compare our results to estimates from previous global- and regional-scale modeling studies (Table 2). Using annual results of the GEOS-Chem model, Park et al. [13] reported that natural sulfur emissions can enhance ammonium sulfate by  $0.11\text{ }\mu\text{g}/\text{m}^3$  over the western and eastern U.S. In our study, DMS emissions increased ammonium sulfate by  $0.07\text{ }\mu\text{g}/\text{m}^3$  over land areas of the modeling domain. Thus, the impacts in our study are lower than the values presented in Park et al. [13] due to model-to-model and emissions differences (they used natural sulfur emissions from the ocean as well as other sources). Mueller et al. [18] used the CMAQv4.6 model and reported that natural emissions can enhance ammonium sulfate by  $0.12\text{ }\mu\text{g}/\text{m}^3$  in winter and  $0.27\text{ }\mu\text{g}/\text{m}^3$  in summer over the modeling domain. In our study, DMS emissions increased ammonium

sulfate by  $0.08 \mu\text{g}/\text{m}^3$  in winter and  $0.18 \mu\text{g}/\text{m}^3$  in summer over the modeling domain. Thus, the impacts shown in our study are lower than those presented in Mueller et al. [18], due in part to the fact that they used natural sulfur emissions from the ocean as well as other sources. Using CMAQv5.3.3 results over the northern hemisphere, Zhao et al. [20] reported that DMS emissions enhance annual mean sulfate by  $0.08 \mu\text{g}/\text{m}^3$  across the entire U.S. In this study, DMS emissions increased annual mean sulfate by  $0.055 \mu\text{g}/\text{m}^3$  over the land area of the modeling domain. The impacts are also less than those in the findings by Zhao et al. [20] due to differences in meteorological conditions between the simulation years, grid size, and model-to-model differences (two different versions of the model were used in these studies).

**Table 2.** A comparison of sulfate enhancement in different studies.

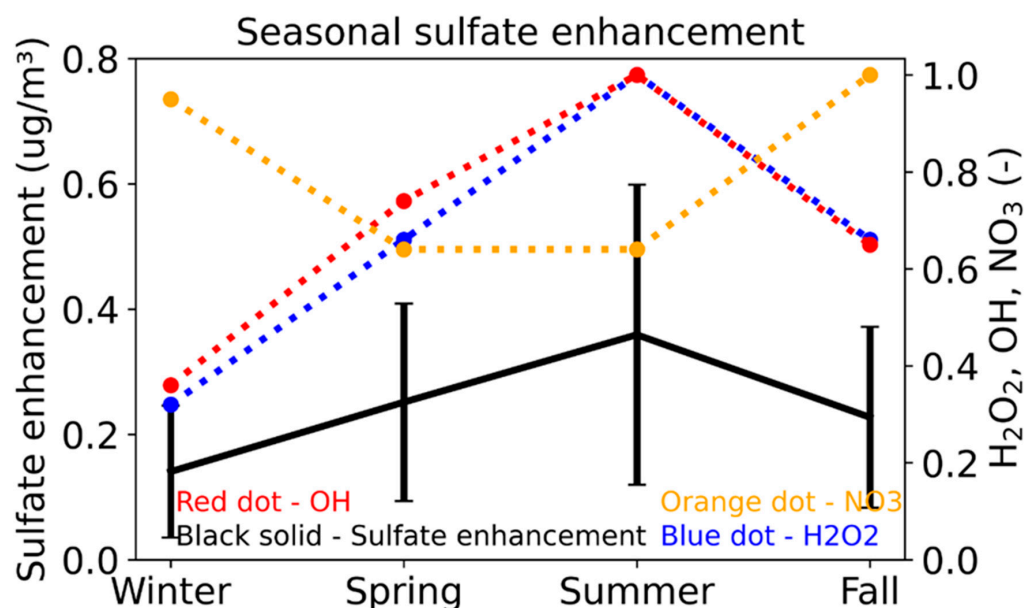
Studies	Enhancements $\mu\text{g}/\text{m}^3$	Sources Considered	Geographic Region	Season
Park et al. [13]	Ammonium sulfate 0.11	DMS from seawater, volcanoes, and biomass burning activities	Eastern and western US	Annual
This study	Ammonium sulfate 0.07	DMS from seawater	Entire land area of the modeling domain	Annual
Mueller et al. [18]	Ammonium sulfate 0.12	DMS and hydrogen sulfide from seawater, coastal wetlands, freshwater, Great Salt Lake, soils, volcanoes, and fumaroles	Entire modeling domain	Winter (December–February)
This study	Ammonium sulfate 0.08	DMS from seawater	Entire modeling domain	Winter (December–February)
Mueller et al. [18]	Ammonium sulfate 0.27	DMS and hydrogen sulfide from seawater, coastal wetlands, freshwater, Great Salt Lake, soils, volcanoes, and fumaroles	Entire modeling domain	Summer (June–August)
This study	Ammonium sulfate 0.18	DMS from seawater	Entire modeling domain	Summer (June–August)
Zhao et al. [20]	Sulfate 0.08	DMS from seawater	Entire US	Annual
This study	Sulfate 0.055	DMS from seawater	Entire US	Annual

### 3.2. Seasonal Impacts of DMS Emissions on Surface Sulfate

The mean seasonal impacts of DMS emissions on sulfate along with OH,  $\text{NO}_3$ , and  $\text{H}_2\text{O}_2$  concentrations over seawater are shown in Figure 3. DMS enhances sulfate in each season, with the greatest enhancement occurring in summer and the smallest in winter. The enhancement in summer is  $\sim 2.5$  times greater than that in winter, which is consistent with increased emissions and increased oxidation. The formation of sulfate from DMS occurs in two steps. In the first step, DMS is oxidized into  $\text{SO}_2$  via the OH-,  $\text{NO}_3$ -, and Cl-initiated reactions. DMS oxidation via the Cl-initiated reaction is generally smaller than the other pathways, since Cl levels are lower than other oxidants. In the second step, the resulting  $\text{SO}_2$  is oxidized into sulfate via the gas-phase reaction with OH and aqueous-phase reactions with  $\text{H}_2\text{O}_2$  and other oxidants. Sulfur oxidation by  $\text{H}_2\text{O}_2$  in cloud and OH in the gas phase are important processes for producing sulfate in the atmosphere [9]. While  $\text{NO}_3$  levels are higher in winter and fall, both OH and  $\text{H}_2\text{O}_2$  concentrations peak in summer. Thus, both the DMS and  $\text{SO}_2$  oxidation can occur efficiently, producing the



highest sulfate enhancement in summer. In contrast, both OH and H<sub>2</sub>O<sub>2</sub> concentrations are the smallest in winter, producing the lowest enhancement in sulfate.



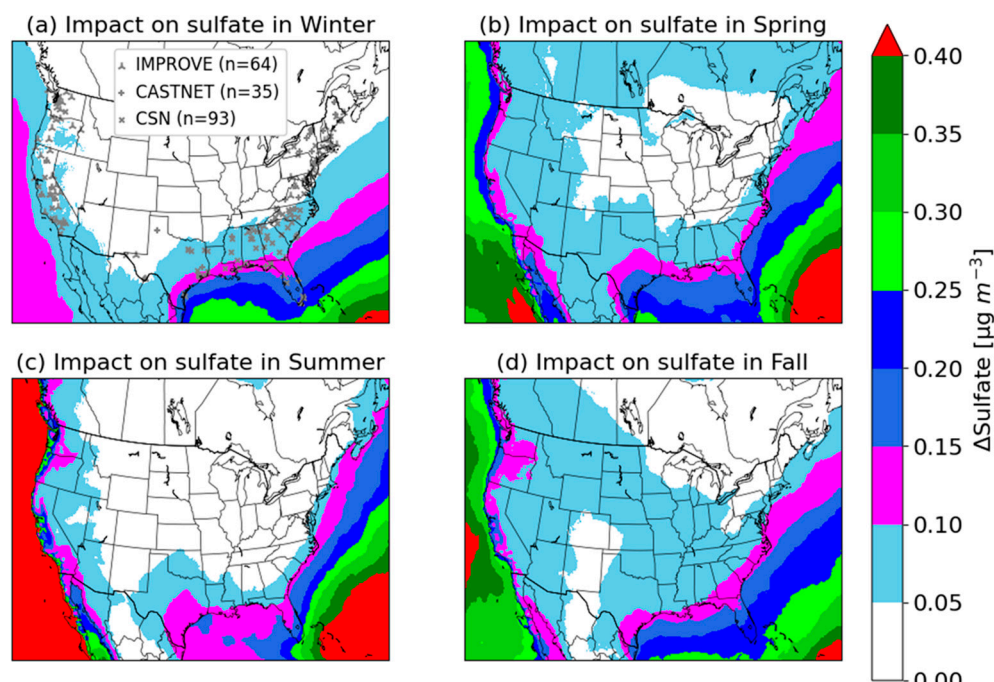
**Figure 3.** Seasonal impacts of DMS emissions on sulfate over seawater. H<sub>2</sub>O<sub>2</sub>, OH, and NO<sub>3</sub> concentrations shown on the right y-axis are dimensionless. Normalized values were calculated by dividing the actual concentrations by their maximum concentrations and are shown on y-axis. Winter represents December–February, spring represents March–May, summer represents June–August, and fall represents September–November. Solid vertical lines represent standard deviation of sulfate enhancement.

The spatial distribution of the seasonal impacts is shown in Figure 4. Consistent with the annual results shown in Figure 2a, larger impacts occur over seawater than those over land in all seasons. DMS-induced sulfate enhancements are lower in winter over most areas than those in other seasons. In winter, however, DMS emissions increase sulfate over the Gulf of Mexico and part of the Atlantic Ocean by larger margins than those over the Pacific Ocean, leading to higher sulfate concentrations in the winter. In other seasons, DMS increases sulfate over the Pacific Ocean by larger margins due to the contribution from the western boundary condition, higher DMS emissions, and higher oxidant levels. The impacts of DMS emissions over seawater are the highest in summer; however, the impacts tend to be localized over coastal areas. The spatial distributions of the impacts on sulfate are different in spring and fall than those in summer. The mean enhancements over seawater in spring and fall are lower than those in summer. However, DMS emissions produce greater impacts over a large interior portion of land in spring and fall due to higher wind speed, which can effectively transport the resulting sulfate to land. Wind speed is the highest in winter; however, sulfate enhancements are the lowest; thus, the impacts over land are not as widespread as predicted in spring and fall. The average DMS-induced sulfate enhancements over land are 0.04 µg/m<sup>3</sup> (6%), 0.07 µg/m<sup>3</sup> (8%), 0.05 µg/m<sup>3</sup> (11%), and 0.06 µg/m<sup>3</sup> (13%) in winter, spring, summer, and fall, respectively.

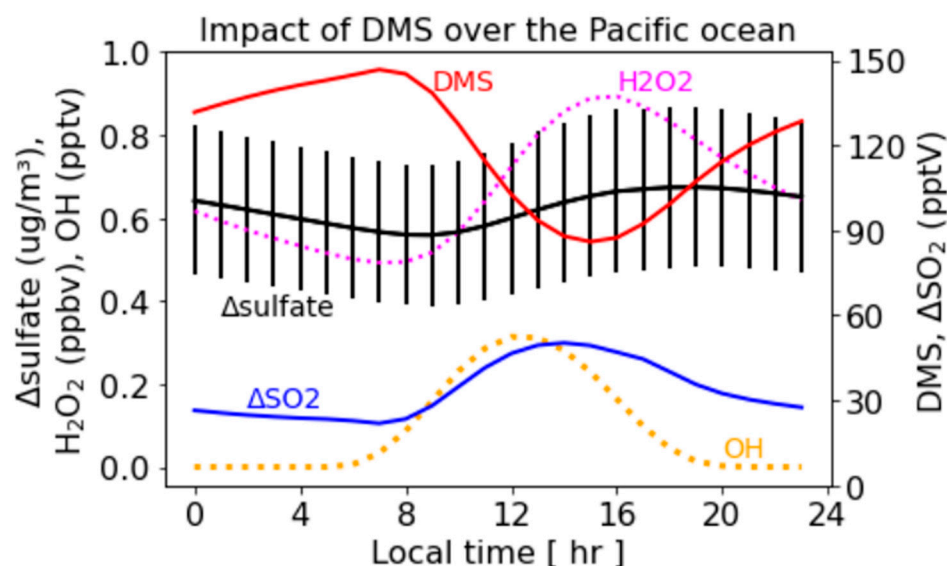
### 3.3. Diurnal Variation in the Sulfate Enhancement

To examine the diurnal variation in DMS-induced sulfate, we selected all grid cells over the Pacific Ocean (including the Gulf of California) in July. The diurnal variation in H<sub>2</sub>O<sub>2</sub>, OH, DMS, and SO<sub>2</sub> and the sulfate produced by DMS emissions over the Pacific Ocean are shown in Figure 5. OH concentrations remain low at night, start increasing during the day, and peak around noon. H<sub>2</sub>O<sub>2</sub> concentrations are also lower at night, increase during the day, and peak during the mid-afternoon. DMS concentrations slowly increase from midnight to morning, as DMS is continuously emitted but is only oxidized by NO<sub>3</sub> at night, which is lower in summer than in winter. The DMS concentrations decrease

sharply during the day, reach their lowest level around mid-afternoon, and then increase again. The decrease in DMS concentrations during the day occurs due to higher OH concentrations.  $\text{SO}_2$  levels remain relatively flat during the night despite  $\text{SO}_2$  production from the oxidation of DMS by  $\text{NO}_3$ .  $\text{SO}_2$  concentrations start increasing during the day as more DMS is oxidized by higher daytime OH levels, peak at mid-afternoon, and then decrease at night. Sulfate from the DMS emissions tends to decrease slowly during the night, as little sulfate is produced since OH and  $\text{H}_2\text{O}_2$  are lower at night. Sulfate levels increase during the day and peak in the late afternoon due to increased oxidation by OH (in gas phase) and  $\text{H}_2\text{O}_2$  in clouds. Thus, DMS-induced sulfate tends to be higher during the day and lower at night.



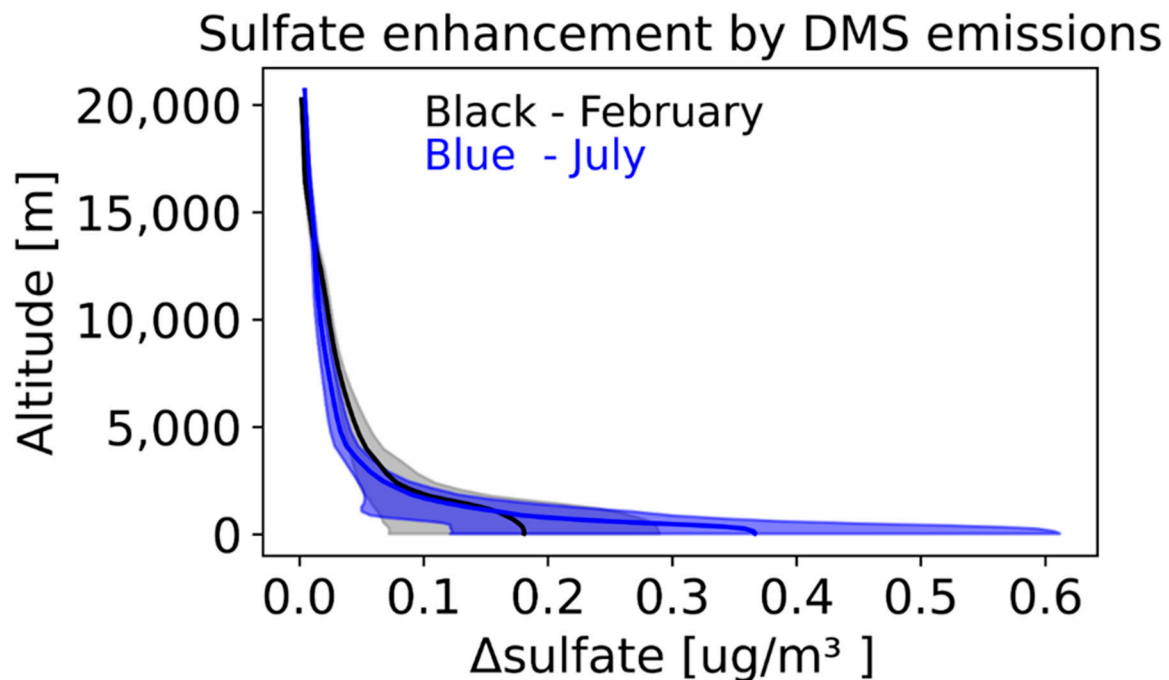
**Figure 4.** Spatial distribution of the seasonal impacts of DMS emissions on sulfate in (a) winter, (b) spring, (c) summer, and (d) fall. (a) shows monitoring locations for observed data used in Figure 8.



**Figure 5.** Diurnal variation in the impacts of DMS emissions on sulfate over the Pacific Ocean in July. All grid cells over the Pacific Ocean (including the Gulf of California) were used for the calculation. Solid vertical lines represent standard deviation of sulfate enhancement.

### 3.4. Impact of DMS Emissions on Sulfate Aloft

To examine the impact of DMS emissions aloft, we focused on a winter (February) and a summer (July) month. DMS-induced sulfate enhancements over the entire oceanic area with altitude are shown in Figure 6. DMS emissions enhance surface-layer sulfate by  $\sim 0.2 \mu\text{g}/\text{m}^3$  in February and  $\sim 0.4 \mu\text{g}/\text{m}^3$  in July. These emissions consistently produce more sulfate in summer than in winter. DMS-induced sulfate enhancements are highest at the surface and decrease with altitude, diminishing to  $\sim 10\text{--}20\%$  at an altitude of  $\sim 5 \text{ km}$ . Thus, the DMS emissions enhance sulfate not only at the surface layer but also aloft, though its impact tends to be limited in the lower troposphere.

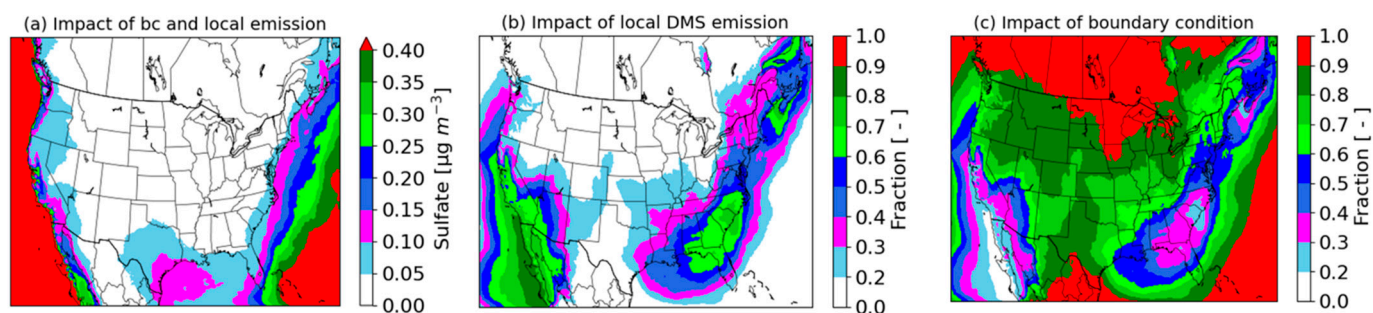


**Figure 6.** Impact of DMS emissions on sulfate aloft. Solid lines represent mean and shaded areas represent standard deviation of sulfate enhancement.

### 3.5. Impact of DMS Emissions and Boundary Conditions on Sulfate

To understand the relative importance of in-domain emissions and transport on sulfate, we completed an additional sensitivity simulation similar to WDMS, but without DMS emissions, in the 12 km domain. The additional sensitivity simulation, therefore, only has DMS and products from the boundary conditions, so we refer to it as BCDMS. The in-domain contribution is calculated as the difference between WDMS (both BC and local emissions) and BCDMS (only BC). The boundary contribution is calculated as the difference between BCDMS (only BC) and NODMS (no DMS). The combined impacts of in-domain DMS emissions and boundary conditions (WDMS—NODMS) on July mean sulfate are shown in Figure 7a. Consistent with Figures 2a and 4c, DMS emissions and boundary conditions produce higher impacts over seawater than over land, and they also enhance sulfate in many coastal areas. Figure 7b,c show the fractional contributions of in-domain DMS emissions (WDMS—BCDMS) and outside-domain DMS emissions (BCDMS—NODMS) to the total sulfate enhancement (WDMS—NODMS) shown in Figure 7a. In-domain DMS emissions increase sulfate over seawater by fractions of  $>0.6$  over the southern Pacific coast, the southern Atlantic coast, and the coast of Maine (Figure 7b). In contrast, boundary conditions enhance sulfate over seawater by fractions of  $>0.6$  over the northern Pacific coast, further offshore in the Atlantic/Gulf of Mexico, the tip of Florida, and over a large land area (Figure 7c). While the combined impacts of DMS emissions and boundary conditions over the interior portion of the modeling domain are relatively small ( $<0.05 \mu\text{g}/\text{m}^3$ ), the impact over this region is dominated by boundary conditions, i.e., the long-range transport of sulfate formed from DMS emissions over the

hemisphere. These results suggest that boundary conditions play an important role on sulfate enhancements not only near the boundaries but also over the interior portion of the modeling domain. Thus, it is important to employ appropriate boundary conditions in the model; otherwise, regional-scale model predictions will not reflect the full effects of DMS emissions on the relevant temporal and spatial scales. It should be noted that boundary conditions obtained with the DMS emissions from the hemispheric model include the impacts on DMS,  $\text{SO}_2$ , and sulfate. These results suggest that enhancements in sulfate from DMS on the Pacific Northwest of  $0.25\text{--}0.3\text{ }\mu\text{g}/\text{m}^3$  are primarily from emissions outside the domain. This is consistent with large fluxes between the Pacific Northwest and Alaska from May through August, as evident in Zhao et al. [20] and Hulswar et al. [58].



**Figure 7.** (a) Combined impact of DMS emissions and boundary conditions on sulfate, (b) fractional impact of within-domain DMS emissions on sulfate, and (c) fractional impact of outside-domain DMS emissions (specified through boundary conditions) on sulfate in July.

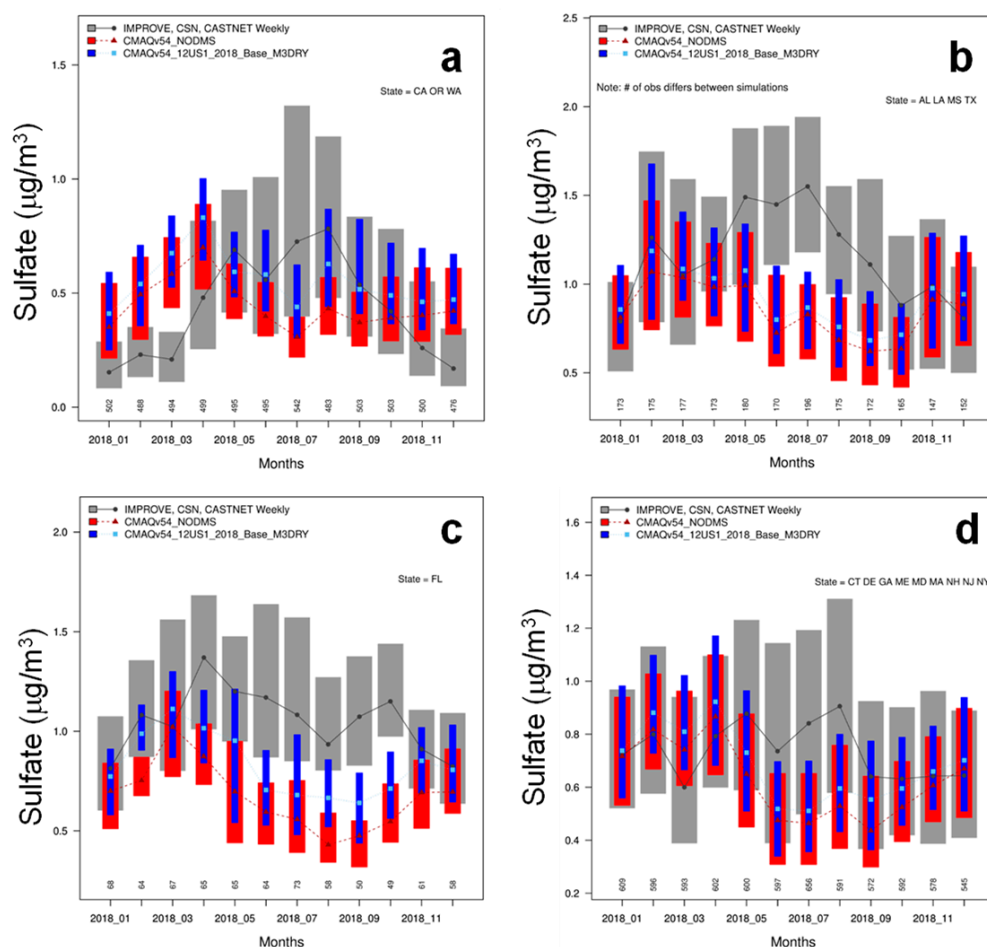
### 3.6. Comparison with Observed Sulfate

The largest impact of DMS emissions on land occurs in coastal areas. Thus, we focused on model performance in coastal states and compared model predictions with observed data in four separate groups: (1) the Pacific Coast states (Washington, Oregon, California), (2) the Gulf Coast states (Texas, Louisiana, Mississippi, Alabama), (3) Florida, and (4) the Atlantic Coast states (Georgia, South Carolina, North Carolina, Virginia, Maryland, Delaware, New Jersey, New York, Connecticut, Rhode Island, Massachusetts, New Hampshire, Maine). We calculated the monthly mean observed sulfate concentrations by combining data from IMPROVE (Interagency Monitoring of Protected Visual Environments), CASTNET (Clean Air Status and Trends Network), and CSN (Chemical Speciation Network) sites within each group. We then calculated the mean sulfate concentrations without and with DMS emissions by combining predictions at all sites within the group and compared them to the corresponding observed data.

Figure 8 shows that the effect on performance varies by region and season. In all regions, the NODMS underpredicts May through October observations. Adding DMS improves the model performance. For the cooler months, the performance is mixed because the base model overpredicts in some places and underpredicts in others. For cooler months (January–April, November, December), NODMS overpredicts compared to observations, so adding DMS deteriorates the model performance. In the Gulf Coast states, NODMS only overpredicts in January and December, and DMS emissions improve the comparison with observed data in all other months. A similar pattern is also noticeable in Florida and the Atlantic Coast states. Thus, additional sulfate from DMS emissions tends to deteriorate the model performance in cooler months while improving it in warmer months.

To further examine the impact on model performance, we isolated IMPROVE sites located within 10 km from the coastline in the Pacific Coast states and Florida and compared monthly mean sulfate concentrations without and with DMS emissions to observed data (Figure 9). Similar to Figure 8, DMS emissions deteriorate the model performance in cooler months. However, they improve the comparison in warmer months by larger margins than those shown in Figure 8. Thus, DMS emissions can improve the model performance by a larger margin near the coastal areas than the interior areas.





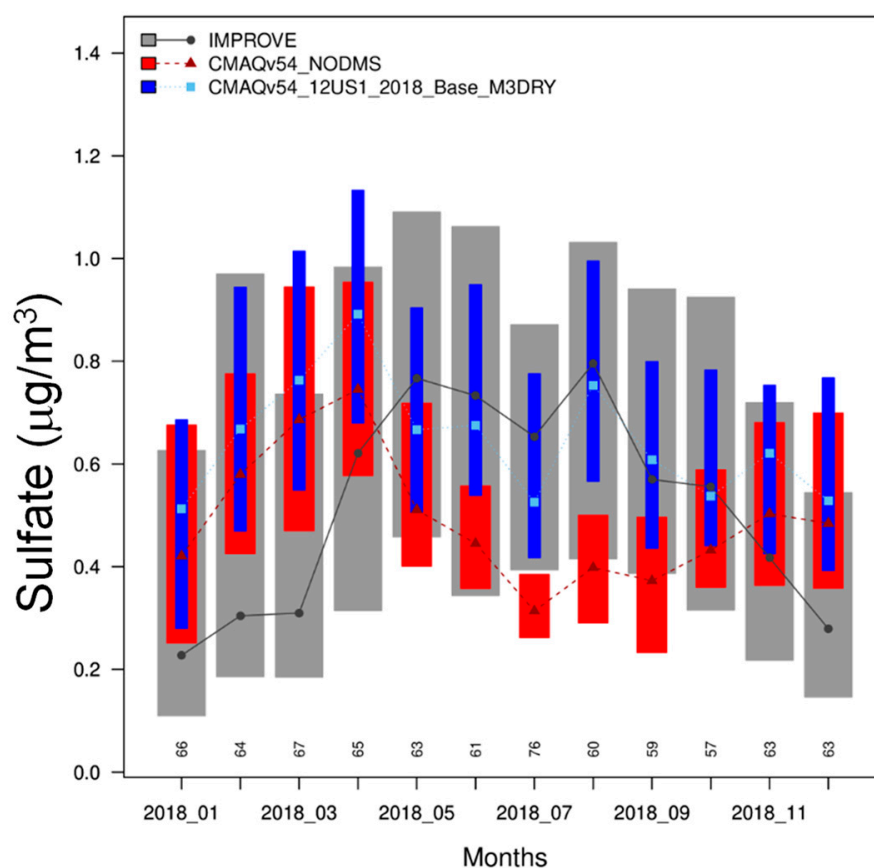
**Figure 8.** A comparison of predicted sulfate without and with DMS emissions to observed sulfate in (a) the Pacific Coast states, (b) the Gulf Coast states, (c) Florida, and (d) the Atlantic Coast states. Boxplot of observed sulfate (grey boxes and lines) from IMPROVE, CSN, and CASTNET; model without DMS emissions (red boxes and lines); and model with DMS emissions (blue boxes and lines) by month. The shaded box represents the interquartile range (25% to 75%) of the data and the point indicates the median value. The numbers below each box indicate the total number of observations available for that month from the three networks. See Figure 4a for the locations of monitoring sites.

### 3.7. Changes in DMS-Initiated Sulfate Enhancement with Horizontal Grid Resolution

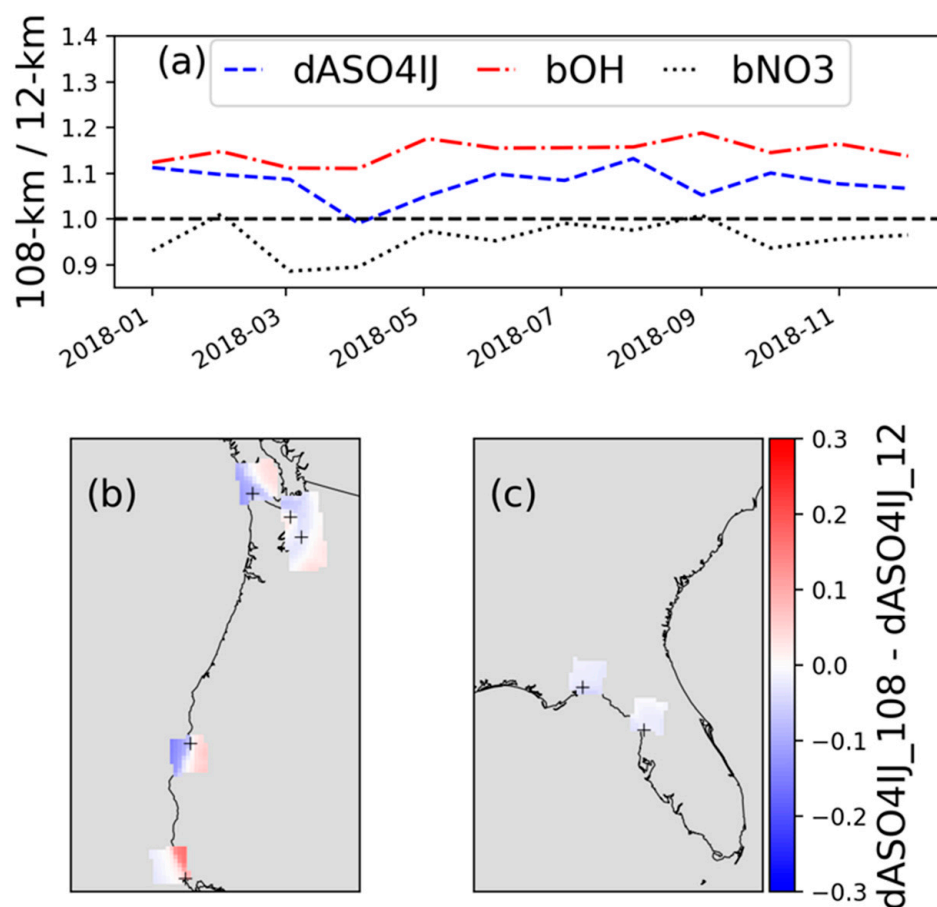
Most of the prior studies employed large-scale models using coarse horizontal grid resolution to examine the impact of oceanic DMS emissions on sulfate. In this study, we used a finer horizontal grid resolution (12 km) over the continental U.S. We also conducted CMAQ simulations without and with DMS emissions over the northern hemisphere using coarse (108 km) horizontal grids, from which we created boundary conditions for model simulations over the continental U.S. Here, we examined the impacts of DMS emissions on sulfate using results from simulations obtained with 12 km and 108 km horizontal grids. We calculated DMS-initiated monthly mean sulfate enhancements over the entire seawater of the modeling domain covering the continental U.S. using both the 108 km and 12 km horizontal grids. Over the entire seawater, ratios of sulfate enhancements with 108 km grids to 12 km horizontal grids are greater than 1.0 (Figure 10a), which indicates that model with 108 km grids produces slightly more DMS-initiated sulfate than the model with 12 km grids. Ratios of model OH and NO<sub>3</sub> concentrations over the seawater with 108 km and 12 km grids are also shown in Figure 10a. The OH concentrations over the entire seawater are consistently greater with 108 km grids than those with 12 km grids (ratio > 1.0). In contrast, the NO<sub>3</sub> concentrations over the entire seawater are generally lower with 108 km grids than those with 12 km grids



(ratio < 1.0). The grid resolution affects many parameters in the model which, in turn, affect DMS-initiated sulfate enhancements. However, OH appears to play an important role in affecting the DMS-initiated sulfate enhancement. The higher DMS-initiated sulfate with 108 km grids occurs primarily due to additional DMS oxidation via higher OH concentrations and additional SO<sub>2</sub> oxidation with OH concentrations. However, such enhancements with 108 km grids do not occur in all areas. For example, the model with 108 km grids generally produces more sulfate enhancements over the Pacific Ocean. In contrast, the model with 108 km grids produces less sulfate enhancements over portions of the Gulf of Mexico and the Atlantic Ocean, primarily due to reduced DMS oxidation by lower NO<sub>3</sub> concentrations and reduced SO<sub>2</sub> oxidation by lower H<sub>2</sub>O<sub>2</sub> concentrations. Thus, DMS-initiated sulfate enhancements are dependent on grid resolution and geographic location. To further illustrate the complexity of evaluating the impacts of grid resolution at individual sites, the annual mean difference in DMS-initiated sulfate at the seven coastal IMPROVE sites (same sites used in Figure 9) are shown in Figure 10b,c. Each 108 km grid cell contains 81 approximately 12 km grid cells. The model with 108 km grids produces lower sulfate enhancements consistently at the two sites in Florida than those with the model with 12 km grids. However, the impacts at the western U.S. sites are mixed within 108 km grid cells. The model with 108 km grids produces more sulfate enhancements at some sites while producing less sulfate enhancements at other sites. Such results show the importance of considering grid resolution in evaluating the impacts of DMS-initiated sulfate enhancements.



**Figure 9.** A comparison of model-predicted sulfate without and with DMS emissions to observed sulfate at seven coastal sites (IMPROVE sites located within 10 km of the coast) in California (PORE1, REDW1), Washington (MAKA2, OLYM1, PUSO1), and Florida (CHAS1, SAMA1). Boxplot of observed sulfate (grey boxes and lines); model without DMS emissions (red boxes and lines); and model with DMS emissions (blue boxes and lines) by month. The shaded box represents the interquartile range (25% to 75%) of the data and the point indicates the median value. The numbers below each box indicate the total number of observations available for that month.



**Figure 10.** (a) Ratios of 108 km to 12 km sulfate enhancement (dSO<sub>4</sub>, blue solid), DMS simulation OH concentration (bOH, red dashed), and DMS simulation NO<sub>3</sub> (bNO<sub>3</sub>, blue dotted) for ocean cells within the 12US1 domain. dSO<sub>4</sub>, bOH, and bNO<sub>3</sub> are shown on the y-axis and dates are shown on the x-axis; (b,c) annual mean difference in sulfate enhancements between 108 km and 12 km horizontal grid resolutions zoomed in on the Pacific Coast (b) and Florida (c). Seven IMPROVE sites were used in the calculation: PORE1 and REDW1 in California, MAK2, OLYM1, and PUSO1 in Washington, and CHAS1 and SAMA1 in Florida.

#### 4. Conclusions

We examined the impacts of DMS emissions on sulfate concentrations over the continental U.S. and surrounding areas by using CMAQv5.4 with a horizontal grid resolution of 12 km and boundary conditions derived from a comparable hemispheric-scale model. DMS emissions increase annual mean sulfate concentrations over seawater and land areas, with the largest impact over seawater. In-domain DMS emissions increase sulfate more over coastal areas than over interior portions of land, with the largest impacts occurring over Washington, Oregon, California, and Florida. Seasonally, the largest enhancements over seawater occur in summer and the lowest in winter. These enhancements help reduce the low bias in modeled sulfate at coastal sites in California, Washington, and Florida. The over-seawater enhancements in spring and fall are lower than those in summer but greater than those in winter. However, enhanced sulfate can occur deeper into land areas in spring and fall due to the prevailing higher wind speeds, causing slightly higher enhancements than those in summer. DMS emissions increase sulfate throughout the lower troposphere, though the impacts at higher altitudes decrease sharply compared to those near the surface. Predicted sulfate with DMS deteriorates the model performance in cooler months, while improving it in warmer months. DMS plays a larger role in affecting the model performance in coastal areas. Sulfate enhancements from DMS emissions on a hemispheric scale represented through boundary conditions for the 12 km domain play an important role

in affecting sulfate enhancements, not only near the boundaries, but also in the interior portion of the modeling domain.

As indicated by our results, the oxidation of DMS in the troposphere and subsequent conversion to SO<sub>2</sub> are key processes for the formation and growth in sulfur-containing aerosols in the marine boundary layer, which then can also undergo long-range transport and influence background aerosol sulfate levels in continental areas. Though our current implementation of DMS chemistry captures the dominant oxidation pathways via OH and NO<sub>3</sub>, future enhancements in the representation of halogen chemistry in the model enable a more complete representation of DMS oxidation by reactive halogens as well as multiphase pathways [16]. Additionally, Veres et al. [59] recently identified hydroperoxymethyl thioformate (HPMTF) as a new DMS oxidation product in measurements performed during the Atmospheric Tomography mission over marine environments. The addition of HPMTF and its oxidation pathways could potentially slow down DMS oxidation to SO<sub>2</sub> and thus reduce the abundance of SO<sub>2</sub> and sulfate in areas of DMS emissions but enhance their levels in downwind cloud-free regions where HPMTF can continue to oxidize to produce SO<sub>2</sub> and sulfate. In contrast, Novak et al. [60] suggest that in the cloudy marine boundary layer, the HPMTF lifetime is considerably reduced due to uptake in cloud droplets, limiting the production of SO<sub>2</sub>. However, the paucity of direct observations of the full suite of DMS oxidation products across the variable conditions of the marine boundary layer make it challenging to adequately constrain the reaction pathways and SO<sub>2</sub> yields from DMS oxidation. Future efforts will focus on enhancing the representation of these pathways, as better estimates of these multiphase chemical pathways evolve.

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