

Article

# Morpholino-Substituted BODIPY Species: Synthesis, Structure and Electrochemical Studies

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**Abstract:** Functionalization of 4,4-difluoro-4-bora-3a,4a-diaza-s-indacene (BODIPY) chromophores at the 2,6-positions with iodo substituents and morpholino-substituted  $\alpha$ -methyl groups affords molecules with strong absorbance in the visible spectrum. The effect of such substitution on the solid-state arrangements, absorption, fluorescence and electronic properties of these dye molecules is reported. The spectroscopic and spectroelectrochemical measurements display intense absorptions in the UV-visible spectrum with bathochromic shifts, in comparison to unfunctionalized BODIPY, and a positive shift in redox potentials due to functionalisation of the BODIPY core. Halogen bonds are observed in the solid-state structures of both halogenated BODIPY species, which in one case leads to the formation of an unusual halogen bonded framework.

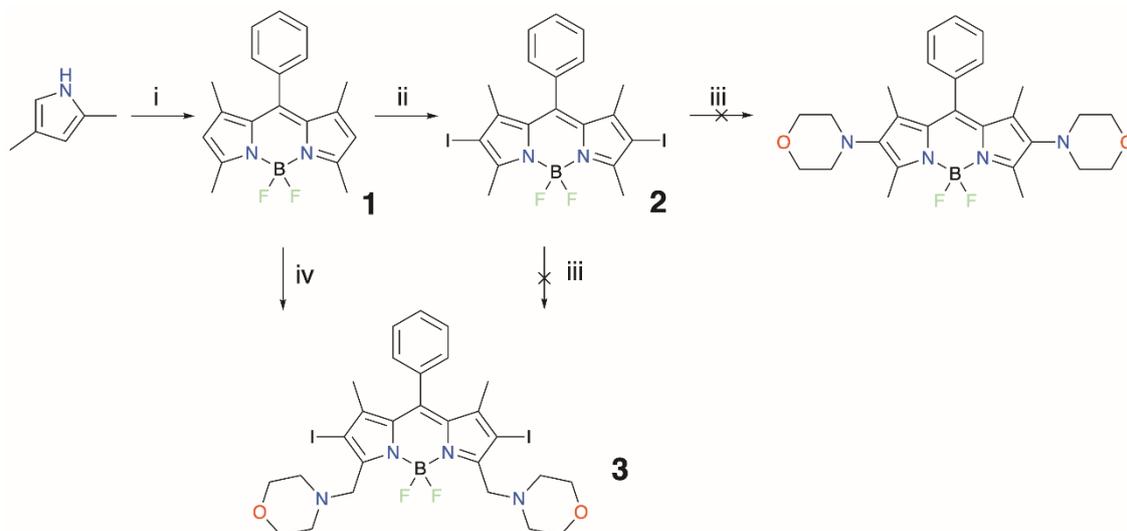
**Keywords:** boron-dipyrromethene dye; halogen bonding; polymorphism; Hirshfeld surface analysis; spectroelectrochemistry

## 1. Introduction

BODIPY (4,4-difluoro-4-bora-3a,4a-diaza-s-indacene) compounds are a class of dye molecules that have been investigated extensively because of their applications in biology, materials and supramolecular chemistry [1–4]. The robust nature of the compounds and their attractive photochemical properties have led to their use for various applications [1,2,4] including light-harvesting devices [4,5], solar cell technology [4,6,7] and chemical sensors [8]. In order to develop specifically targeted BODIPY compounds modification of the dipyrin core is required. Introduction of different substituents can significantly affect the optical behaviour of BODIPY species, for example heavy atoms, such as iodine, can favour the generation of triplet states [9,10], whereas other substituents, such as amines [11–15] can significantly alter the orbital energy manifold and quench fluorescence even when not conjugated with the dipyrin core [11].

Perhaps the most widely known strategy for BODIPY post-functionalization is halogenation. Halogenation offers both the possibility of directly modifying the BODIPY properties, due to the introduction of halogen substituents, but also raises the possibility of subsequent reactions to introduce further functionality [16–22]. In this study we demonstrate the extension of these strategies using BODIPY analogues to functionalise  $\alpha$ -methyl groups whilst also halogenating unsubstituted  $\beta$ -positions using direct electrophilic halogenation [4]. By using this simple approach, we were able to combine different functionalities in a single BODIPY molecule. In particular, our aim was to investigate the effect of combining halo-functionalisation at  $\beta$ -positions with non-conjugated tertiary amine substitution of the  $\alpha$ -methyl groups on the absorption and electronic properties of BODIPY species. Thus, we have identified three target BODIPY molecules, 1,3,5,7-tetramethyl-8-phenyl-4,4-difluoroboradiazaindacene

(1), 2,6-diiodo-1,3,5,7-tetramethyl-8-phenyl-4,4-difluoroboradiazaindacene (2) and 2,6-diiodo-5,7-dimethyl-1,3-bis(methylene) dimorpholino-8-phenyl-4,4-difluoroboradiazaindacene (3) (Scheme 1), which are studied herein.



**Scheme 1.** Synthesis of compounds 1–3. (i) benzoyl chloride, CH<sub>2</sub>Cl<sub>2</sub>, rt, 16 h; Et<sub>3</sub>N, BF<sub>3</sub>·Et<sub>2</sub>O, rt, 1h; (ii) NIS, CH<sub>2</sub>Cl<sub>2</sub>, rt, 30 min; (iii) morpholine, CH<sub>2</sub>Cl<sub>2</sub>, reflux, 3 h or morpholine, Et<sub>3</sub>N, reflux, 3 d; (iv) NIS, CH<sub>2</sub>Cl<sub>2</sub>, rt, 30 min; morpholine, reflux, 1 h.

## 2. Experimental Part

**General Experimental:** All reagents and solvents were purchased from Aldrich Chemicals, Alfa Aesar or Fisher Scientific UK, and used without further purification unless otherwise stated. Thin layer chromatography (TLC) was used to monitor the reactions. Column chromatography was used for purification (unless otherwise stated) using silica gel (Merck silica gel 60, 0.2–0.5 mm, 50–130 mesh). <sup>1</sup>H and <sup>13</sup>C{<sup>1</sup>H} NMR spectra were recorded at room temperature with either a Bruker AV400 or a Bruker AV3400, operating at 400 and 100 MHz, respectively. <sup>11</sup>B and <sup>19</sup>F NMR spectra were recorded at room temperature with Bruker AV3400, operating at 128 and 376 MHz, respectively. Chemical shifts are quoted in ppm relative to TMS (<sup>1</sup>H and <sup>13</sup>C{<sup>1</sup>H}). BF<sub>3</sub>·OEt<sub>2</sub> was used as an external reference for <sup>11</sup>B NMR measurements and CFCl<sub>3</sub> as an external reference for <sup>19</sup>F NMR. Elemental microanalyses were performed at the Analytical Services, School of Chemistry, University of Nottingham.

**Electrochemical Measurements:** Cyclic Voltammetry (CV) experiments were performed using an Autolab PGSTAT20 potentiostat with a three-electrode arrangement in a single-compartment cell under an atmosphere of argon. A glassy carbon working electrode, a Pt wire secondary electrode and a saturated calomel reference electrode were chemically isolated from the test solution by a bridge tube containing electrolyte solution and fitted with a porous vycor frit and were used in the cell. The solutions were 10<sup>−3</sup> M in test compounds and 0.4 M in [<sup>n</sup>Bu<sub>4</sub>N][BF<sub>4</sub>] as supporting electrolyte. The redox potentials are quoted versus the ferrocenium–ferrocene couple (Fc<sup>+</sup>/Fc) which was used as an internal reference [23]. Compensation for internal resistance was not applied. Dichloromethane was dried over molecular sieves under dinitrogen gas, ferrocene (Aldrich) was used as received and [<sup>n</sup>Bu<sub>4</sub>N][BF<sub>4</sub>] was prepared by a literature method [24].

**Photophysical Measurement:** Absorption UV-Vis and fluorescence measurements were carried out at the University of Nottingham. Absorption UV-Vis spectra of solutions were measured with a Perkin Elmer Lambda 25 UV-Vis spectrometer in 1 cm length quartz cuvettes. Excitation and emission photoluminescence spectra of solution samples were performed using an Edinburgh Instruments FLS920 Fluorometer. The fluorescence samples were run in standard quartz fluorescence cuvettes, l = 1 × 1 cm. All spectroscopy experiments were performed using spectrophotometric grade solvent and

solution concentrations never exceeded  $A = 0.10$  at the excitation wavelength or any longer wavelength, to minimise inner filter effects.

**Synthetic Procedures:** Compounds **1** and **2** [9] were synthesised according to literature procedures. *2,6-diiodo-5,7-dimethyl-1,3-bis(methylene)dimorpholine,8-phenyl-4,4-difluoroboradiazaindacene (3)*. To a solution of **1** (170 mg, 0.51 mmol) in  $\text{CH}_2\text{Cl}_2$  (25 mL) was added 4 equivalents of N-iodosuccinimide (459 mg, 2.04 mmol). The resulting mixture was stirred at room temperature in the absence of light for 1 h. Morpholine (2 cm<sup>3</sup>, 2.04 mmol) was added and the reaction mixture was heated at reflux for 1 h. The crude product was concentrated under vacuum and purified by silica gel column chromatography (*n*-hexane/ $\text{CH}_2\text{Cl}_2$ , 1:2). The deep red-coloured fraction was collected, and the solvent removed under reduced pressure to yield the desired compound as a red solid. Yield 90 mg (26 %). <sup>1</sup>H NMR (400 MHz,  $\text{CDCl}_3$ )  $\delta$  7.57–7.55 (m, 3H), 7.29–7.30 (m, 2H), 3.84 (s, 4H) 3.36–3.76 (t,  $J = 4.6$  Hz, 8H), 2.65 (t,  $J = 4.6$  Hz, 8H), 1.43 (s, 6H); <sup>13</sup>C NMR (500 MHz,  $\text{CDCl}_3$ )  $\delta$  155.96, 146.46, 143.37, 134.58, 131.20, 129.74, 127.60, 67.11, 53.63, 29.72 16.98; <sup>11</sup>B NMR (128 MHz,  $\text{CDCl}_3$ )  $\delta$  0.6 (t,  $J = 33.1$  Hz); <sup>19</sup>F NMR: (376 MHz,  $\text{CDCl}_3$ )  $\delta$  –134.1 (q,  $J = 33.1$  Hz). (MALDI-TOF, DCTB/MeCN):  $m/z$  calculated 746.06, observed 746.1([M]<sup>+</sup>, 100%). IR (ATR)/cm<sup>–1</sup>: 2954 (w), 2849 (w), 2757 (w), 1772 (w), 1704 (s), 1536 (s), 1447 (m), 1411 (m), 1346 (s), 1175 (s), 1149 (m), 1102 (s), 1004 (m), 978 (m), 913 (m), 858 (m), 724 (s), 630 (w), 537 (w), 485 (w). Elemental Analysis: Calcd for  $\text{C}_{32}\text{H}_{43}\text{BF}_2\text{I}_2\text{N}_4\text{O}_2$ : C, 46.97; H, 5.30; N, 6.85. Found: C, 46.45; H, 5.11; N, 6.35.

**Single Crystal X-ray Diffraction Studies:** Single crystal diffraction data were collected for **2** and **3**, at 120(2) K on a Rigaku Oxford Diffraction SuperNova, Atlas GV1000 diffractometer using mirror monochromated Cu  $K_\alpha$  radiation. Using Olex2 [25] the structure was solved with the ShelXT [26] structure solution program using intrinsic phasing and refined with the ShelXL [27] refinement package using least squares minimisation. All non-hydrogen atoms were refined with anisotropic displacement parameters. Hydrogen atoms were placed in calculated positions and thereafter refined using a riding model with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$  or  $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C})$  for methyl H atoms. Visualizing interactions in molecular crystals using Hirshfeld surface tools and 2D fingerprint plot (FPs) were analysed using Crystal Explorer (Version 3.1, and 17.1) [28,29].

Crystal data for **2**:  $\text{C}_{19}\text{H}_{17}\text{BF}_2\text{I}_2\text{N}_2$ : space group  $P2_1/n$  (no. 14),  $a = 11.4551(3)$  Å,  $b = 12.9407(4)$  Å,  $c = 13.2532(4)$  Å,  $\beta = 90.788(2)^\circ$ ,  $V = 1964.43(10)$ ,  $Z = 4$ ,  $D_{\text{calc}} = 1.947$  g cm<sup>–3</sup>,  $\mu = 25.351$  mm<sup>–1</sup>,  $F(000) = 1096.0$ . A total of 7967 reflections were collected, of which 3866 were unique, with  $R_{\text{int}} = 0.0275$ . Final  $R_1$  ( $wR_2$ ) = 0.0273 (0.0683) with GOF = 1.074.

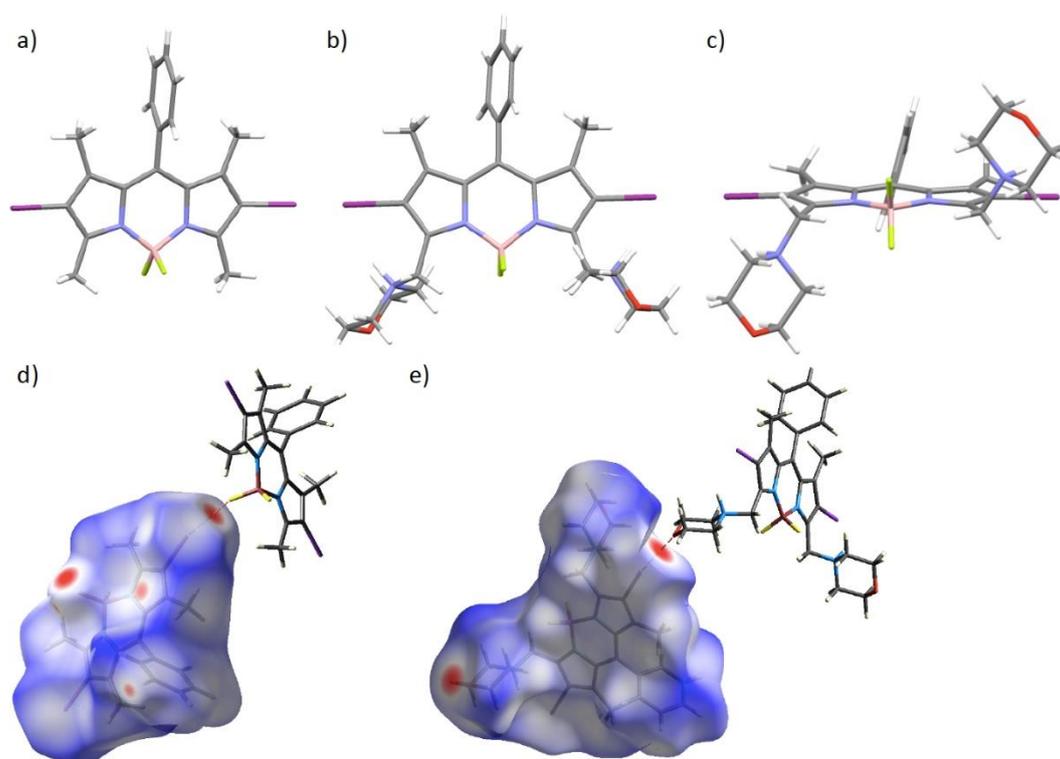
Crystal data for **3**:  $\text{C}_{27}\text{H}_{31}\text{BF}_2\text{I}_2\text{N}_4\text{O}$ . Orthorhombic, space group  $P2_12_12_1$  (no. 19)  $a = 6.5790(10)$  Å,  $b = 18.7712(4)$  Å,  $c = 23.3004(4)$  Å,  $V = 2877.50(9)$  Å<sup>3</sup>,  $Z = 4$ ,  $D_{\text{calc}} = 1.722$  g cm<sup>–3</sup>,  $\mu = 17.541$  mm<sup>–1</sup>,  $F(000) = 1464.0$ . A total of 27,321 reflections were collected, of which 5705 were unique, with  $R_{\text{int}} = 0.0921$ . Final  $R_1$  ( $wR_2$ ) = 0.0348 (0.0710) with GOF = 1.003. Flack parameter = –0.022(7).

### 3. Results and Discussion

**2** was synthesised by direct electrophilic halogenation of **1** [9] using N-iodosuccinimide (NIS) [9,30]. Interestingly, the direct reaction of the  $\beta$ -halogenated BODIPY (**2**) with morpholine did not result in nucleophilic substitution of either halogenated  $\beta$ -positions or the  $\alpha$ -methyl groups. However, reaction of **1** with NIS in  $\text{CH}_2\text{Cl}_2$  for 30 minutes followed by reflux with morpholine for a further hour gave **3**, in 26% yield, which combines both iodo and amino substituents. The intermediate species with iodinated  $\alpha$ -methyl groups could not be isolated [11]. Instead, the direct reaction of **1** with morpholine, in the presence of NIS, resulted in amine substitution of this position.

The packing motif observed for BODIPY compounds can significantly affect their absorption and emission behaviour [30,31]. Although in this case neither **2** nor **3** exhibited solid-state emission ( $\lambda_{\text{excitation}} = 365$  nm), we investigated the packing arrangements exhibited by **2** and **3**. Single crystals of **2** and **3** were grown by vapour diffusion of diethyl ether into *n*-hexane solutions of the target compound (Figure 1). Structural determination by single-crystal X-ray diffraction revealed the BODIPY species to crystallise in the monoclinic space group  $P2_1/n$  for **2** and in the orthorhombic space group  $P2_12_12_1$  for

3. The X-ray crystal structures confirmed the chemical composition of both compounds as determined by NMR spectroscopy.



**Figure 1.** Crystal structure of (a) **2** and (b,c) **3**, in the latter case illustrating the relative orientation of the morpholino substituents in relation to the dipyrroin core; (d) Hirshfeld surface representation of (d) C–I···F interaction in **2** and (e) C–I···O interaction in **3**. C = grey, N = blue, B = pink, F = green, I = purple, O = red, H = white.

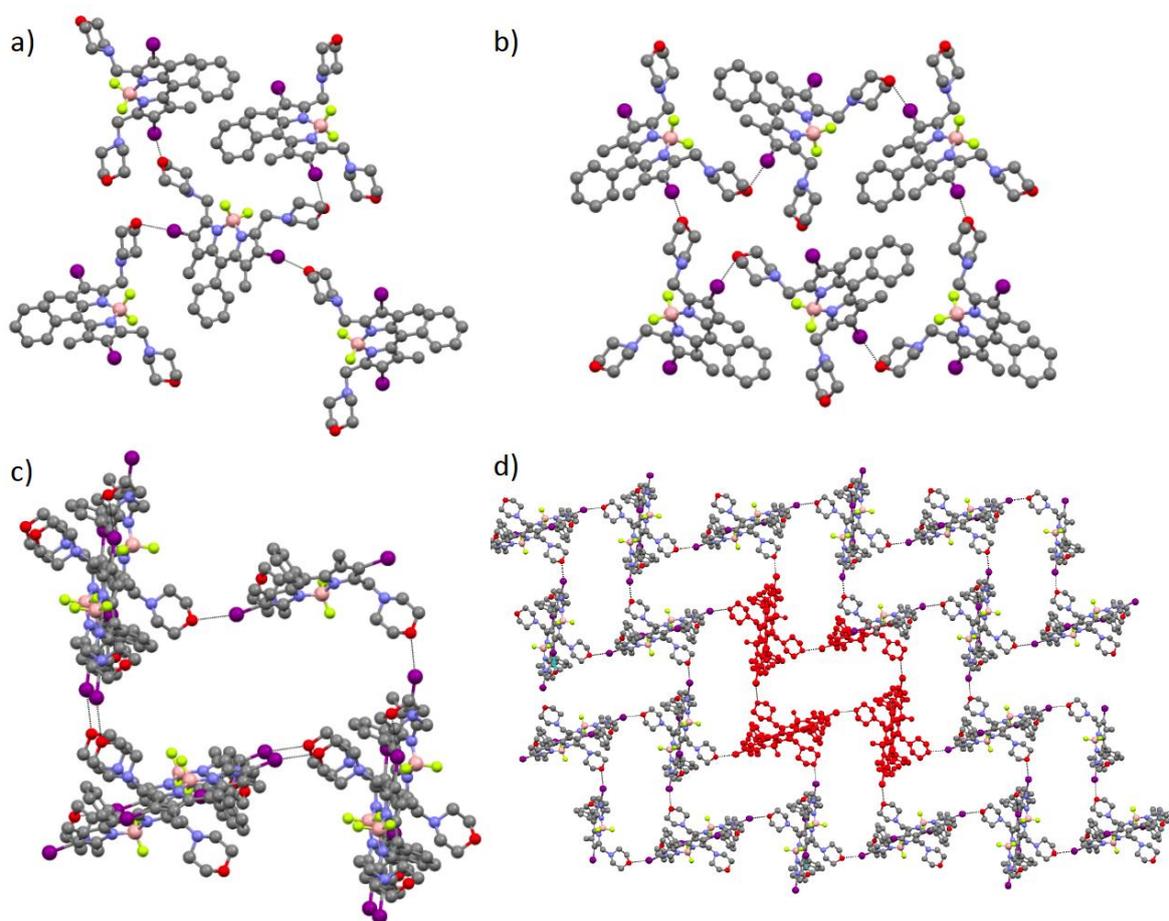
In comparison to previously reported boron-dipyrromethene compounds, the X-ray crystal structures of **2** and **3** revealed the expected planarity of the BODIPY core that extended through the 2,6-diiodo groups (Figure 1a–c), with the *meso* phenyl ring approximately orthogonal to the indacene plane (angle between least-squares planes of the *meso* phenyl ring and indacene group: **2** = 88.2°; **3** = 75.9°). In **3**, the morpholino appendages are found to be positioned above and below the indacene plane as a result of the orientation of the methylene group linking the two groups [ $C_{\alpha\text{pyrrole}}-C_{\alpha\text{carbon}}-N_{\text{Morph}}$  = 109°] (Figure 1c).

Three polymorphs of 2,6-diiodo-1,3,5,7-tetramethyl-8-phenyl-4,4-difluoroboradiazaindacene, named **2** herein, have been reported [32–37]. Inspection of the packing of **2** reveals that this structure is polymorph I [32–35]. In this polymorph, a zig-zag linear pattern is observed due to the formation of C–I···F interactions between an iodine-atom on one BODIPY and the fluorine atoms from the neighbouring BODIPY (Figure 1d). Bond lengths and angles are consistent with type I [38] halogen interactions [39–41] [for **2**: I···F = 3.1165(19) Å; C–I···F = 163.89(11)°, B–F···I = 152.64(19)°] (see Supplementary Materials for view of zig-zag chain). In order to confirm the occurrence of these C–I···F interactions in **2**, Hirshfeld surface analysis was undertaken.

Inspection of the Hirshfeld surfaces of **2** (Figure 1d) indicates that the most significant intermolecular interactions involve the C–I···F interaction discussed above. The iodine atom not involved in this C–I···F interaction does not engage in halogen bonding but rather acts as an acceptor in a short (C–H···I) contact involving a methyl group of an  $\alpha$ -carbon from the neighbouring BODIPY. Polymorph II [36] exhibits a distinct packing arrangement to polymorph I. Although a zig-zag linear pattern is also observed in this polymorph, the packing arises due to intermolecular I···I interactions

[I...I = 3.78(8) Å; C-I...I = 114.30(2)°, 163.30(2)°; 112,50(2)°, 160.20(3)°] (see Supplementary Materials for view of zig-zag chain). Polymorph III [37], like polymorph I, adopts C-I...F interactions between adjacent molecules and forms a zig-zag arrangement but of a more compact nature with a less-linear C-I...F angle than that observed in polymorph I [I...F = 3.24(8) Å; C-I...F = 108.18(6)°, B-F...I = 173.25(3)°] (see Supplementary Materials for view of polymorph III).

Hirshfeld surface analysis was also performed for **3** revealing the presence of C-I...O interactions (Figure 1e) [42–44]. The unusual C-I...O halogen interaction, [C-I...O = 2.972(6) Å, C-I...O = 166.70(2)°, C-O...I = 105.00(5)°] exhibits a short I...O distance, over 20% shorter than the sum of mean van der Waals radii of the two atoms (3.49 Å) [44], suggesting a strong interaction. In contrast to **2**, no close C-I...F interactions are observed in **3**, possibly due to presence of the morpholino moieties, resulting in a more favourable I...O interaction being adopted. The packing diagram of **3** confirms the presence of halogen bonds leading to the formation of a framework structure (Figure 2). Each molecule of **3** forms C-I...O halogen bonds to four nearest neighbour molecules, ultimately forming a three-dimensional halogen-bonded framework with pseudo-diamondoid topology. No I...I or F...F interactions are observed in the structure of **3**. It is notable that the presence of additional halogen bond acceptors in **3**, the morpholino oxygen atoms, results in a significantly different crystalline framework structure in comparison to the zig-zag chains observed in all polymorphs of **2**.



**Figure 2.** Views of the crystal structure of **3** illustrating the overall halogen bonded framework and illustrating (a) each molecule connected to four nearest neighbours, (b) the formation of a six-membered ring and (c) an adamantane-like unit that is a component of (d) the overall pseudo-diamondoid network. In (d) the adamantane-like unit, shown in (c), is highlighted in red. C = grey, N = blue, B = pink, F = green, I = purple, O = red, H = white. Black dashed lines represent C-I...O halogen bonds.

The optical properties of **2** and **3** give UV-visible spectral profiles characteristic of BODIPY compounds (Table 1) with bands corresponding to the ( $S_0 \rightarrow S_1$ ,  $\pi-\pi^*$ ) transition of the dipyrroin ligands. The band maxima of **2** and **3** in  $\text{CH}_2\text{Cl}_2$  were observed at 534 and 539 nm, respectively, showing that iodination at the 2- and 6-positions of the dipyrroin results in bathochromic shifts of approximately 30 nm in the absorption maximum of **2** in comparison with **1** and a 19 nm bathochromic shift in comparison to the previously reported mono-iodinated derivatives [32]. The morpholino substituted species, **3**, displays intense absorptions of approximately three times the intensity of those observed for **2**, which also show a further small bathochromic shift. This results in substantial colour changes for solutions of the three compounds (Figure 3). It is also noticeable that for **3**, fluorescence is essentially quenched with very low quantum yield,  $\Phi_F = 0.0006$  in  $\text{CH}_2\text{Cl}_2$  in comparison to values of 0.63 [33] or 0.05 [33] for **1** and **2**, respectively. This quenching can be attributed to the presence of the tertiary amine of the morpholino-substituent.

**Table 1.** Summary of cyclic voltammetry data, absorption wavelengths ( $\lambda_{\text{abs}}$ ) and extinction coefficients ( $\epsilon$ ) for neutral and reduced species and emission data for compounds 1–3. <sup>a</sup>

Compound	$E_{1/2}/\text{V}$ ( $\Delta E$ )	$\lambda_{\text{abs}}/\text{nm}$ ( $\epsilon \times 10^{-4}/\text{dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$ ) <sup>b</sup>		$\lambda_{\text{em}}/\text{nm}$ <sup>e</sup>	$\Phi_F$ Quantum Yield <sup>e</sup>
		Neutral Species	Radical Anion		
1	−1.60 (0.08) <sup>c</sup>	501 (1.1)	287 (0.2), 340 (0.2), 473 (0.1), 528 (0.1), 571 (0.1)	511 <sup>f</sup>	0.63 <sup>f</sup>
2	−1.37 (0.08) <sup>c</sup>	384 (0.7), 534 (6.4)	304 (1.0), 343 (1.1), 493 (0.2), 528 (0.4), 569 (1.0)	550 <sup>f</sup>	0.05 <sup>f</sup>
3	−1.30 (0.07) <sup>d</sup>	389 (2.2), 538 (20.2)	303 (2.3), 357 (2.0), 529 (0.6), 568 (1.7)	556	0.0006

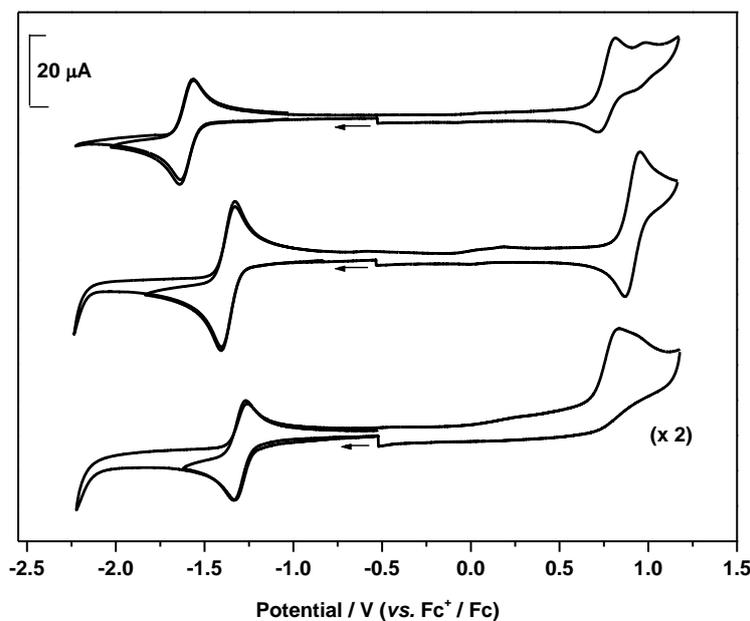
<sup>a</sup> In  $\text{CH}_2\text{Cl}_2$  containing  $[\text{nBu}_4\text{N}][\text{BF}_4]$  (0.4 M); <sup>b</sup> at 243 K; <sup>c</sup> oxidation process at  $E_{1/2} + 0.77$  V; <sup>d</sup> oxidation process at  $E_p^a + 0.83$  V; <sup>e</sup> recorded in  $\text{CH}_2\text{Cl}_2$ ; <sup>f</sup> from reference [32].



**Figure 3.** Photograph showing the different in colour of  $\text{CH}_2\text{Cl}_2$  solutions of (left to right) **1**, **2**, and **3**.

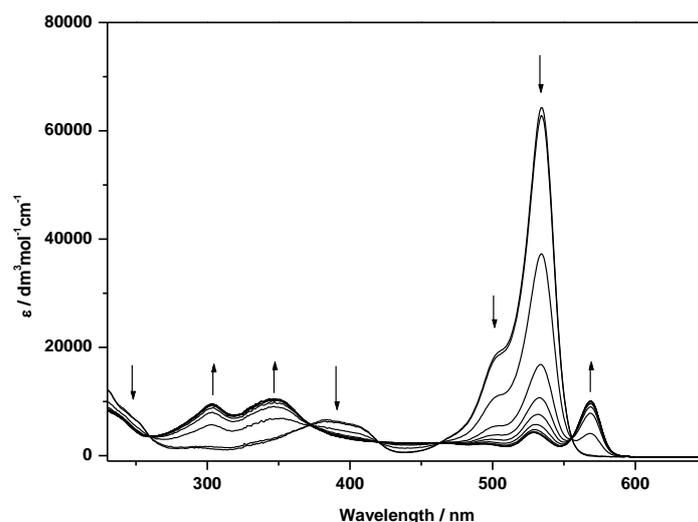
The redox behaviour of **1**, **2** and **3** in  $\text{CH}_2\text{Cl}_2$  solution was investigated by cyclic voltammetry and UV-Vis spectroelectrochemistry (Table 1). In line with previous observations all oxidations and reductions are based on the dipyrroin component of the molecule [45]. An electrochemically reversible one electron reduction process is observed for each compound (Figure 4) with the presence of the iodo substituents shifting this redox process to more positive values by more than 0.2 V, i.e., making the compound easier to reduce. The introduction of the morpholino substituents in **3** shifts the potential to a more positive potential in comparison to **2** suggesting a further decrease of electron density in the BODIPY manifold. In all cases oxidation processes were observed at positive potentials (Figure 4 and Table 1). For **1** and **2**, the oxidation appears reversible at faster scan rates, but as the scan rate

is decreased, the magnitude of the current in the second half cycle of the CV is reduced suggesting instability of the oxidised species in the diffusion layer. Inclusion of morpholine in **3** appears to change the nature of the oxidation process compared to **1** and **2**, with no return wave observed in the second half-cycle of the CV experiment. As would be expected, the presence of electron-withdrawing iodo substituents in **2** and **3** makes the BODIPY framework less electron rich and therefore the corresponding oxidation process more difficult compared to that in **1**. The oxidation processes were not studied further.

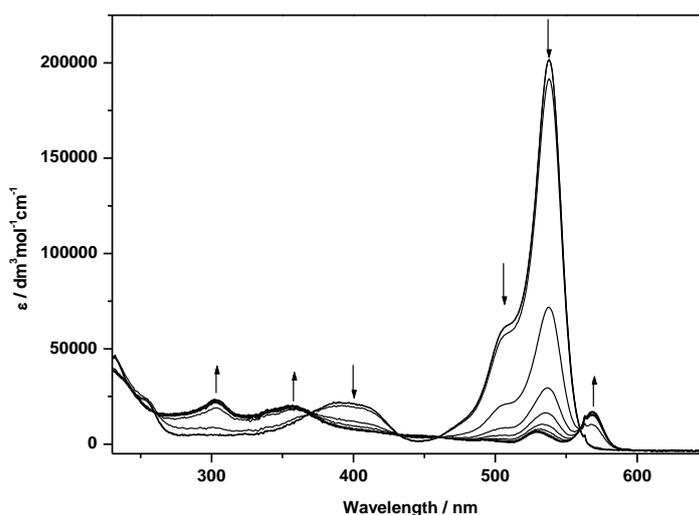


**Figure 4.** Cyclic voltammograms of **1** (upper traces), **2** (middle traces) and **3** (lower traces) at RT in  $\text{CH}_2\text{Cl}_2$  with  $[\text{nBu}_4\text{N}][\text{BF}_4]$  (0.4 M) as supporting electrolyte.

In situ spectroelectrochemical methods were employed to probe the UV-Vis spectroscopic changes that accompany the one-electron reduction processes of **2** and **3** (Table 1; Figures 5 and 6; see Supplementary Materials for further figures). The in situ one-electron reduction of each compound was followed by UV-vis spectroelectrochemistry at an optically transparent electrode. Generation of the one-electron reduced radical anion shows depletion of the major bands for the neutral species (ca. 534 nm for **2**, and 538 nm for **3**) and the corresponding development of a series of bands in the range 510–570 nm for **2** and **3**, all with significantly smaller absorption coefficients. In the range 290–450 nm, reduction leads to the development of new bands at ca. 300 and 350 nm for **2** and ca. 303 and 353 nm for **3**. Applying a potential sufficient to reverse the reduction process did not regenerate the spectral profile of **2** or **3** fully, as would be seen if the process was chemically reversible. This is indicative of a chemical instability of the generated radical anion resulting in decomposition to produce an unknown species. These results are largely consistent with unsubstituted BODIPY species [45] despite the substitution of the dipyrin core and shift of redox potentials to more positive potentials.



**Figure 5.** UV-vis spectroelectrochemistry of **2** in  $\text{CH}_2\text{Cl}_2$  with  $[\text{tBu}_4\text{N}][\text{BF}_4]$  as electrolyte at 243 K. Arrows indicate progress of reduction.



**Figure 6.** UV-vis spectroelectrochemistry of **3** in  $\text{CH}_2\text{Cl}_2$  with  $[\text{tBu}_4\text{N}][\text{BF}_4]$  as electrolyte at 243 K. Arrows indicate progress of reduction.

#### 4. Conclusions

In summary, we have demonstrated that it is possible to functionalize BODIPY with non-conjugated tertiary amine substitution in combination with iodo-substitution on the dipyrroin core giving rise to highly coloured dye molecules with intense absorptions. The shift in absorption wavelength is paralleled by a shift in reduction potentials to more positive values upon introduction of iodo and tertiary amine groups to the BODIPY scaffold. The introduction of iodo-substituents also leads to the adoption of halogen bonding interactions in the solid-state structures of **2** and **3**. Whereas **2** exhibits C–I···F interactions, **3** forms C–I···O interactions, which lead to each molecule of **3** acting as a fourfold-connected tecton and thus creating a three-dimensional framework of pseudo-diamondoid topology. In summary our study demonstrates that the introduction of different chemical functionalities can be used in an additive fashion to shift reduction potentials of the BODIPY core and can be used to create frameworks through the adoption of halogen bonding interactions.

**Supplementary Materials:** The following are available online at <http://www.mdpi.com/2073-4352/10/1/36/s1>. Additional Figures. CCDC 1963595 and 1963596 contain the supplementary crystallographic data for this

paper. These data can be obtained free of charge from the Cambridge Crystallographic Data Centre via [www.ccdc.cam.ac.uk/data\\_request/cif](http://www.ccdc.cam.ac.uk/data_request/cif).

**Author Contributions:** Conceptualization, N.R.C. and D.L.K.; Methodology, H.H., E.S.D. and N.R.C.; Validation, H.H., E.S.D. and W.L.; Investigation, H.H. and E.S.D.; supervision, N.R.C. and D.L.K. All authors have read and agreed to the published version of the manuscript.

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**Conflicts of Interest:** There are no conflicts to declare.

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