

Supporting Information

Enhanced photo-electrochemical water oxidation using TiO₂-Co₃O₄ p-n heterostructures derived from in situ-loaded ZIF-67

S1. General procedures

S1.1 Materials:

All the chemicals were used directly without further purification. Hydrochloric acid (HCl, 35%) was purchased from OCI Company Ltd.; titanium butoxide (TBOT, 98%) was purchased from Sigma-Aldrich; 2-methylimidazole (2-Melm, 99%) was purchased from Acros Organics; cobalt (II) nitrate hexahydrate (Co(NO₃)₂·6H₂O, 97%) was purchased from Shinyo Pure Chemical Company Ltd.; and sodium sulfate anhydrous (Na₂SO₄, 99%) was purchased from Duksan. Milli-Q water (MΩ 18) was used in all the experiments.

S1.2. Characterization:

The morphology and composition of the TiO₂, TiO₂-ZIF-67, and TiO₂ nanostructures with Co₃O₄ were respectively studied using a field-emission scanning electron microscope (FESEM, S4800 Hitachi, Japan) and a field-emission transmission electron microscope (FETEM, Tecnai G2 F20, USA) attached to an energy dispersive spectrometer (EDS). The crystalline structures of the samples were acquired via X-ray diffraction (XRD; X'pert PRO MPD diffractometer, Netherlands) with a PANalytical X' Celerator detector and a Cu Kα (λ = 0.15406 nm) source. The absorption spectra of the newly developed materials were recorded on a Cary 5000 UV-Vis-NIR spectrophotometer (Agilent, USA). The surface elemental compositions and chemical states

were examined via X-ray photoelectron spectroscopy (XPS) using a K-Alpha spectrometer (Thermo Scientific, USA) equipped with a monochromatic Al-K α X-ray source.

S1.3. PEC tests:

PEC tests were performed using an electrochemical workstation (Zive SP1, WonATech Co., Ltd., Korea) in an electrochemical cell with a three-electrode system: the as-prepared photoanode was used as the working electrode, a platinum mesh was used as the counter electrode, and an Ag/AgCl electrode (in 3 M KCl) placed in a Luggin capillary was used as the reference electrode. The PEC test and other electrochemical measurements were performed using electrodes placed at appropriate positions in a quartz electrochemical cell measuring 50 mm \times 50 mm \times 20 mm (Hellma, HE.700.000. OG20.) with 0.3 M Na₂SO₄ (pH = 7) as the electrolyte. Linear sweep voltammetry (LSV) curves were recorded in a potential window ranging from -0.5 V to 1.3 V vs Ag/AgCl at a scan rate of 20 mV s⁻¹. Electrochemical impedance spectroscopy (EIS) was performed in the frequency range from 10⁻² Hz to 10⁵ Hz with an amplitude of 5 mV at a bias potential of 0 V vs. Ag/AgCl under solar simulator (AM 1.5G, 100 mW/cm²) irradiation. The incident photon-to-current conversion efficiency (IPCE) was measured at 1.23 V vs Ag/AgCl using a xenon arc lamp (75 W, Ushio UXL-75XE) filtered using a dual-grating monochromator and individual filters onto the test devices.

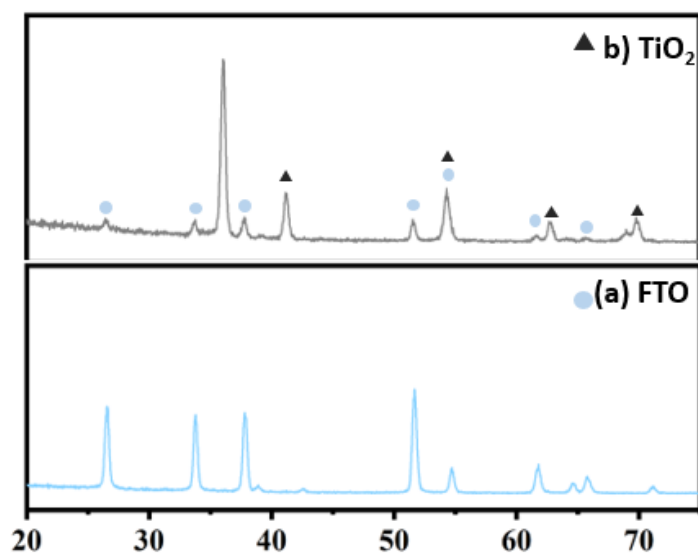


Figure S1. XRD diffraction patterns of (a) FTO(b) TiO_2

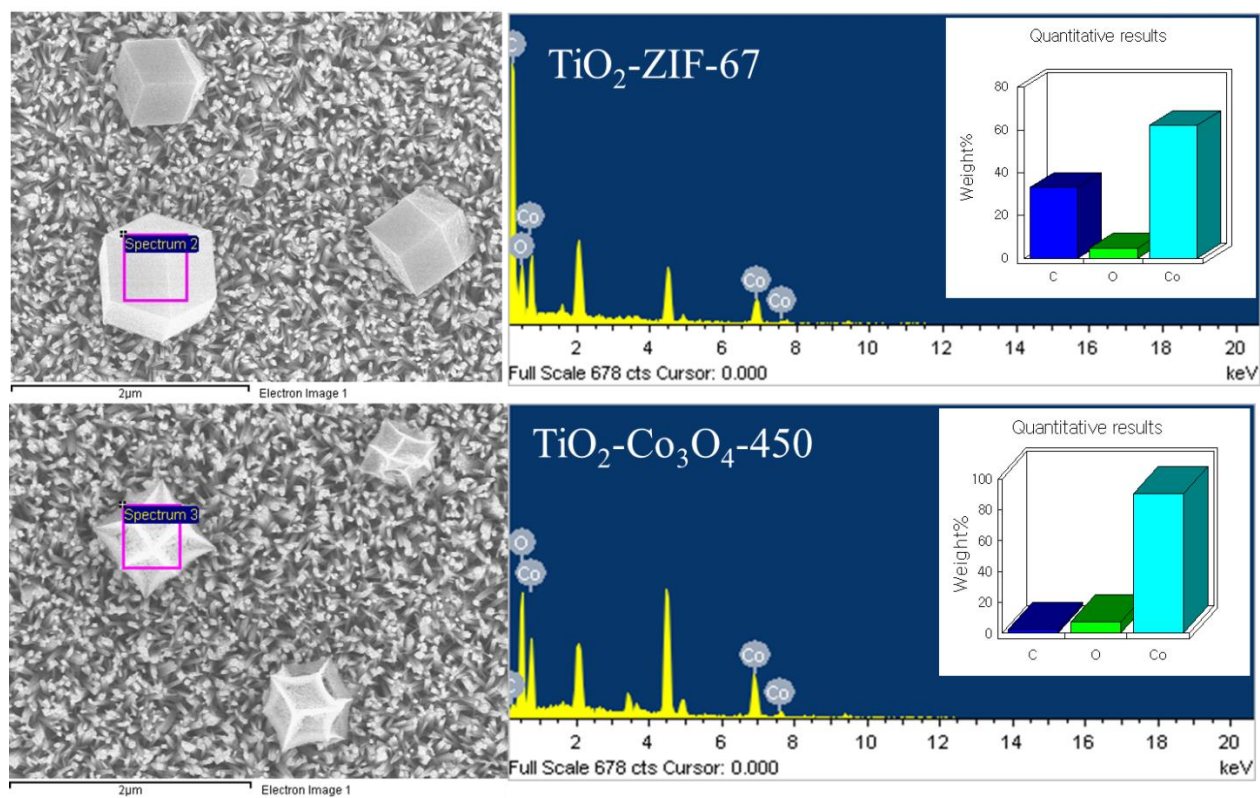


Figure S2. SEM-EDS analysis of ZIF-67 and TiO_2 - Co_3O_4 -450.

Table S1.EDS-elemental ratio of ZIF-67 and TiO₂-Co₃O₄-450

Elements	ZIF-67		Co ₃ O ₄ -450	
	Weight%	Atomic%	Weight%	Atomic%
C	32.98	66.94	2.39	9.09
O	4.80	7.32	7.31	20.89
Co	62.21	25.74	90.30	70.02

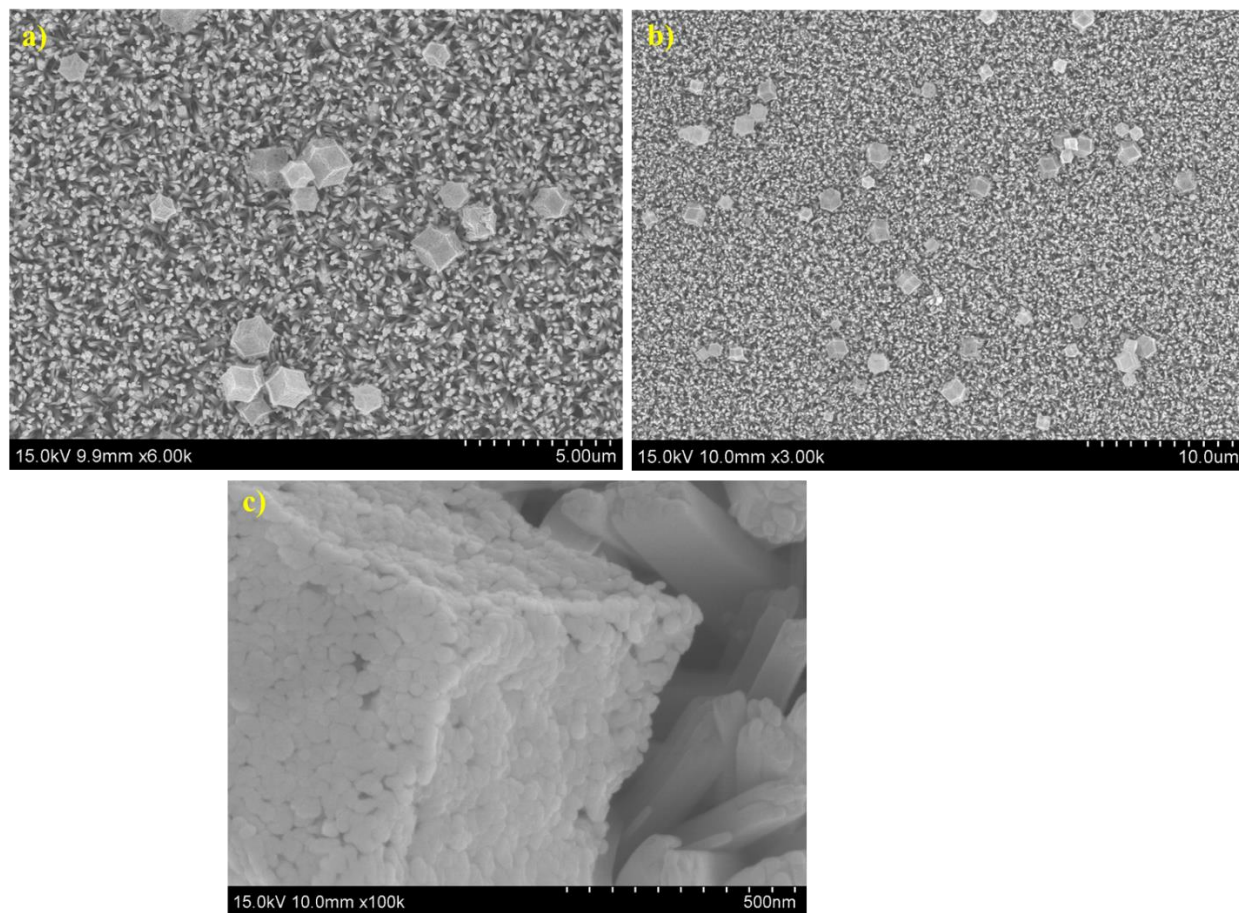


Figure. S3. SEM images of TiO₂-Co₃O₄-450 sample after 1h stability test

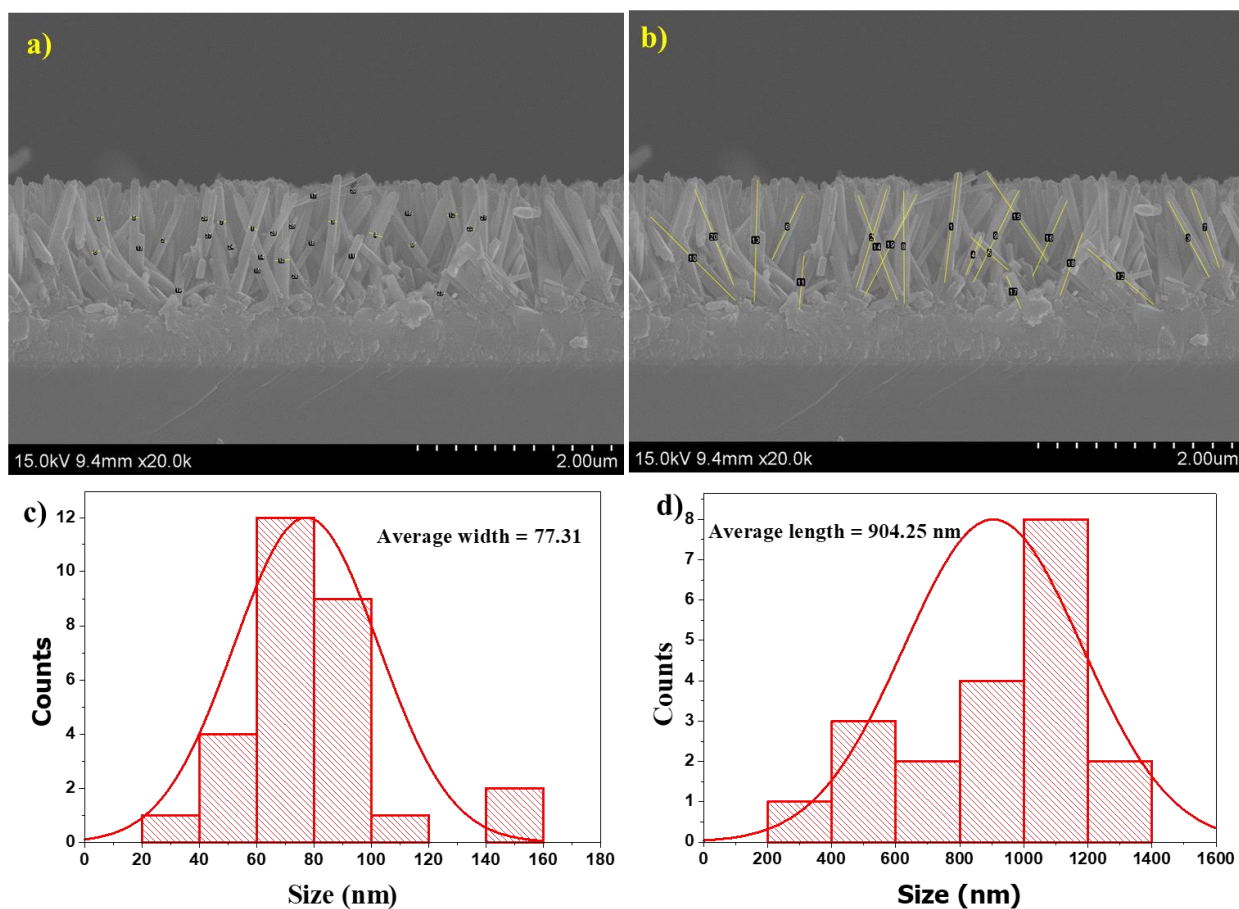


Figure. S4. The Sem and size distribution histogram of TiO₂ nanorod a,c) width of the TiO₂ b,d) Length of the TiO₂ nanorod.

Table S2. Comparative photoelectrochemical performance of selected TiO₂ or Co₃O₄ nanocomposites

Photoanode	Electrolyte/illumination	Photocurrent	Applied bias (V vs RHE)	Ref
Ag/NH ₂ -MIL-	0.5 M Na ₂ SO ₄	1.06 mA cm ⁻²	1.23 V vs	[1]

125/TiO ₂	Electrolyte/300 W Xe lamp with 100 mWcm ⁻²		RHE	
Co ₃ O ₄ /TiO ₂	1 M NaOH (pH = 14) 300 W xenon lamp (100 mWcm ⁻²)	1.04 mA cm ⁻²	1.23 V vs RHE	[2]
Co-MOF@TiFe.	1 M NaOH (pH=13.6)/ 100 mWcm ⁻²	1.78 mA cm ⁻²	1.23 V vs RHE	[3]
Co ₃ O ₄ /BiVO ₄	0.1M KPi buffer (pH7) 300 W Xe lamp with 100 mWcm ⁻²	2.71 mA cm ⁻²	1.23 V vs RHE	[4]
Co ₃ O ₄ /P-C ₃ N ₄ /TiO ₂	0.5 M Na ₂ SO ₄ / 150 W xenon lamp with 100 mWcm ⁻²	1.58 mA cm ⁻²	1.23 V vs RHE	[5]
CoNi-MOFs/BiVO ₄	0.5 M Na ₂ SO ₄ /100 mWcm ⁻²	3.2 mA cm ⁻²	1.23 V vs. RHE	[6]
NiFe-MOF/TiO ₂	0.5 M Na ₂ SO ₄ /100 mWcm ⁻²	0.77 mA cm ⁻² at	1.23 V (vs. RHE)	[7]
TiO ₂ -ZIF-67	0.5 M Na ₂ SO ₄ /100 mWcm ⁻²	1.8 mA cm ⁻²	1.85 V Vs RHE	This work
TiO ₂ -Co ₃ O ₄ -350	0.5 M Na ₂ SO ₄ /100 mWcm ⁻²	2.1 mA cm ⁻²	1.85 V Vs RHE	This work
TiO ₂ -Co ₃ O ₄ -450	0.5 M Na ₂ SO ₄ /100 mWcm ⁻²	2.4 mA cm ⁻²	1.85 V Vs RHE	This work
TiO ₂ -Co ₃ O ₄ -550	0.5 M Na ₂ SO ₄ /100 mWcm ⁻²	2.2 mA cm ⁻²	1.85 V Vs RHE	This work

References

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